



CHEMISTRY

BOOKS - JEEVITH PUBLICATIONS CHEMISTRY (KANNADA ENGLISH)

HYDROGEN

One Marks Questions And Answers

1. Define isotopes.



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2. Name the three isotopes of hydrogen.



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3. Which isotope of hydrogen is radioactive ?



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4. Name the isotope of hydrogen which used in nuclear reactor.



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5. Name the isotope of hydrogen which have atomic number one mass number two.



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6. What is the atomic number and mass number of proton and tritium.



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7. What is the importance of heavy water with regard to nuclear power generation ?



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8. Why does water has a high boiling point and a high melting point as compared to H_2S is not.



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9. Define hydrides.



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10. What are the classification of hydrides.



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11. How is heavy water produced from ordinary water ?



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12. What are the ways in which water molecules are bonded to an anhydrous salt to form a hydrate?



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13. Why is H_2 more reactive than D_2 ?



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14. Which of the substances present in water cause permanent hardness of water ?



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15. Give one example of ionic hydrides.



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16. Give one example of covalent hydrides.



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17. Give one example of complex hydrides.



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18. Give one example of metallic hydrides.



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19. Define D_2O .



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20. Which isotope of hydrogen does not have neutron ?



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21. Which out of nascent hydrogen and dihydrogen is more reactive ?



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1. How to prepare dihydrogen in laboratory by zinc.



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2. H_2O_2 is a better oxidizing agent than water.
Explain.



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3. How to prepare hydrogen by the electrolysis of water ?



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4. Explain the preparation of hydrogen gas from methane.



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5. How to prepare hydrogen gas from coal (water gas) ?



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6. How does dihydrogen react with halogens ?



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7. What happens when dihydrogen reacts with dioxygen ?



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8. How does dihydrogen react with dinitrogen ?



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9. Define ionic hydrides.



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10. Define covalent hydrides or molecular hydrides.



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11. Define metallic hydride or interstitial hydrides.



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12. Water is amphoteric substance. Why ?



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13. What happens when water reacts with sodium metal ?



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14. How to prepare hydrogen peroxide from BaO_2 ?



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15. Dilute solution of hydrogen peroxide cannot be heated strongly for its concentration. Explain.



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16. Can sodium bicarbonate make water hard ?



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17. Which of the hydrogen or deuterium undergoes reactions more rapidly and why ?



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18. Can we remove completely temporary hardness due to $Mg(HCO_3)_2$ by boiling ?



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19. H_2O_2 acts as an oxidizing agent as well as a reducing agent. Why?



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20. Why do lakes freeze from top towards bottom?



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21. A mixture of H_2O_2 and hydrazine with copper (II) is used as a rocket propellant. Why ?



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22. Water cannot be used to extinguish petrol fires. Why ?



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23. What is the volume strength of H_2O_2 ?



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24. Hard water is softened before using in boilers. Explain.



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25. What are the uses of Heavy Water ?



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26. Show the oxidising property of H_2O_2 with PbS.



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27. Discuss the position of hydrogen in the periodic table is not justified.



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28. Explain property of H_2O_2 with MnO_4^- in acidic medium.



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29. How does H_2O_2 reduces iodine in reducing property ?



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30. Describe the structure of common form of ice.



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31. Distinguish clearly between (a) hard and soft water, (b) temporary hardness and permanent hardness.



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32. What is the structure of H_2O_2 ? Draw a schematic diagram indicating the shape of the molecule clearly?



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33. Sea water can't be used in boiler. Explain given chemical equations.



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1. Explain occurrence of hydrogen.



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2. Explain the structure of water molecule ?



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3. In what respects does hydrogen resemble alkali metals ? How does it resemble halogen ?





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4. In what respects H_2 differs from alkali metals and halogens ?



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5. Dihydrogen is considered as fuel Explain.



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6. Explain hydrogen economy.



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7. What different methods are used for softening the hard water ? Explain the principle used in each method ? (Write any two)



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8. Give four uses of hydrogen peroxide.



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9. How is hydrogen peroxide prepared industrially ? Explain why it is stored in coloured wax-lined glass or plastic bottles ?



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10. Distinguish between temporary hardness permanent hardness.



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11. What happens when lead sulphide is reacted with hydrogen peroxide solution ?



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12. Why hard water does not form lather with soap ? What advantage has soap over detergents ?



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13. Advantage of soap over detergent : Soap are biodegradable whereas detergents are non-biodegradable. Difference in chemical behaviour of compound of hydrogen with elements of atomic number 17 and 20.





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14. The mixture of hydrazine and hydrogen peroxide with copper (II) catalyst is used as a rocket propellant. Why ? Write the reactions involved.



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15. Hydrogen peroxide acts both as an oxidizing agent as a reducing agent in alkaline solution towards certain first row transition

metal ions. Illustrate both these properties of H_2O_2 using chemical equations.



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