



CHEMISTRY

BOOKS - MCGROW HILL EDUCATION

CHEMISTRY (HINGLISH)

CHEMICAL REACTIONS

Elementary Questions

1. The valency of a phosphate ion is

A. -1

B. -2

C. 3

D. -3

Answer: D



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2. The reaction $2Na + Cl_2 \rightarrow 2NaCl$ is an example of

A. combination reaction

B. decomposition reaction

C. displacement reaction

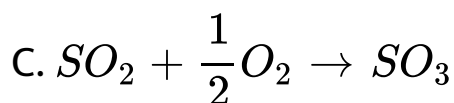
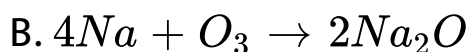
D. double decomposition reaction

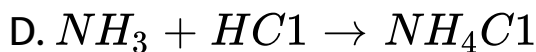
Answer: A



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3. Which of the following reaction involves the combination of two elements?



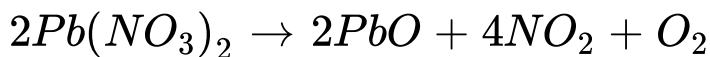


Answer: B



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4. When lead nitrate is heated it breaks down into lead monoxide, nitrogen dioxide and oxygen



The reaction is an example of

A. combination reaction

B. decomposition reaction

C. double decomposition reaction

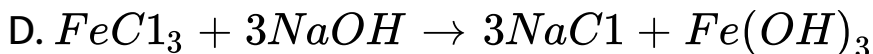
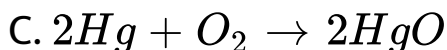
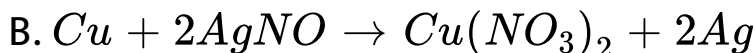
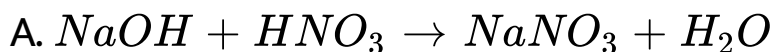
D. displacement reaction

Answer: B



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5. Which of the following is a displacement reaction?

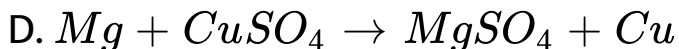
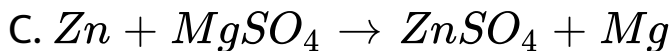
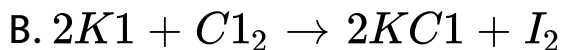
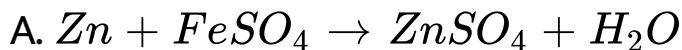


Answer: B



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6. Which of the following reaction will not take place?



Answer: C



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7. The reaction in which two compounds exchange their ions to form two compounds is called

- A. a displacement reaction
- B. a decomposition reaction
- C. an isomerization reaction
- D. a metathesis reaction

Answer: D



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8. When hydrogen sulphide gas is passed through a blue solution of copper sulphate, a black precipitate of copper sulphide is obtained and the sulphuric acid so formed remains in the solution. The reaction is an example of

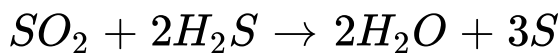
- A. a combination reaction
- B. a displacement reaction
- C. a decomposition reaction
- D. a double decomposition reaction

Answer: D



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9. When the gases sulphur dioxide and hydrogen sulphide mix in the presence of water, the reaction



Here hydrogen sulphide is acting as

- A. an oxidising agent
- B. a reducing agent
- C. a dehydrating agent
- D. a catalyst

Answer: B



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10. In the reaction: $2FeCl_2 + Cl \rightarrow 2FeCl_3$

Chlorine may be regarded as

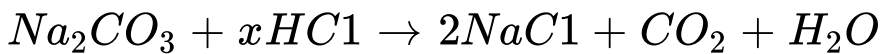
- A. an oxidising agent
- B. a reducing agent
- C. a catalyst
- D. providing an inert medium

Answer: A



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11. In the following equations



the value of x is

A. 1

B. 2

C. 3

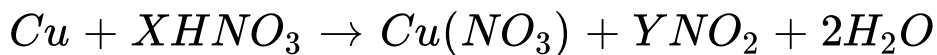
D. 4

Answer: B



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12. The equation



the values of X and Y are

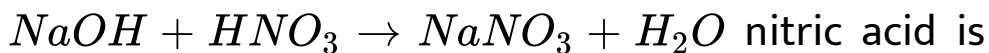
- A. 3 and 1
- B. 8 and 6
- C. 4 and 2
- D. 7 and 1 respectively

Answer: C



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13. In the equation



acting as

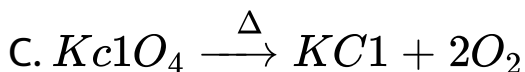
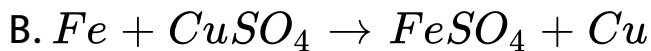
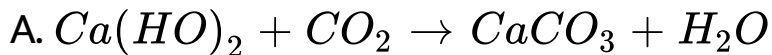
- A. an oxidising agent
- B. an acid
- C. a nitrating agent
- D. a dehydrating agent

Answer: B

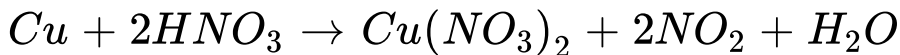


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14. Which of the following is not balanced equation?



D.



Answer: D



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15. Combustion of fuel is

- A. displacement reaction
- B. double displacement reaction
- C. oxidation Reaction
- D. isomerisation reaction

Answer: C



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16. $BaCl_2 + H_2SO_4 \rightarrow BaSO_4 + 2HCl$ is

- A. combination reaction
- B. decomposition reaction
- C. displacement reaction
- D. double decomposition reaction

Answer: D



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17. Zinc can displace copper from copper sulphate solution because zinc is

- A. more reactive than copper

B. less reactive than copper

C. more stable than copper

D. less stable than copper

Answer: A



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18. When Magnesium is burnt in air, a white ash remains as left over. What is this?

A. MgO_2

B. MgO

C. Mg

D. Mg_3O

Answer: B



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19. Carbon dioxide reacts with carbon to form Carbon monoxide. This is an example of a

A. displacement reaction

B. double displacement reaction

C. combination reaction

D. decomposition reaction

Answer: C



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20. $CaCO_3 \xrightarrow{Heat} CaO + CO_2$ is known as

A. Electrolysis

B. Thermal decomposition

C. Displacement

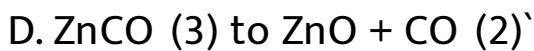
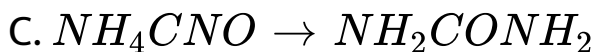
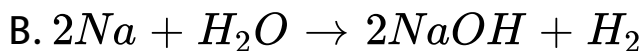
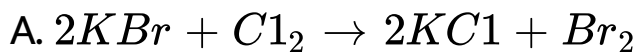
D. Combination

Answer: B



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21. Which of the following is a decomposition reaction?

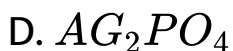
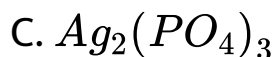
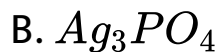
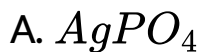


Answer: D



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22. the formula of Silver Phosphate is

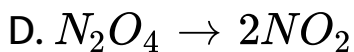
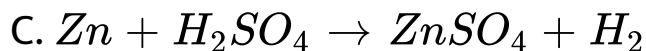
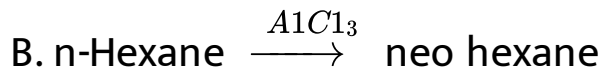
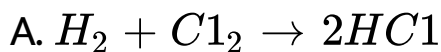


Answer: B



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23. Which of the following is example of combination reaction?



Answer: A



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24. Which of the following reactions is not a decomposition reaction?

A. starch is converted to glucose in our body

B. proteins are digested to amino acids

C. magnesium ribbon is burnt in air

D. electricity is passed through water

Answer: C



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25. the reaction between acid and base to form salt and water is an example of

A. decomposition reaction

B. combination reaction

C. displacement reaction

D. double displacement

Answer: D



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26. Which of the following statement is incorrect?

A. Metals like Cu, Ag, Au cannot displace H from acids

B. In reactivity series metals are arranged in order of increasing reactivity

C. Silver cannot displace Cu from $Cu(NO_3)_2$

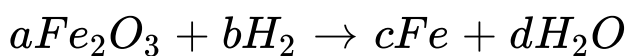
D. Zince displaces Cu from $CuSO_4$

Answer: B



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27. In the balanced reaction



a, b, c and d, respectivel, are

A. 1123

B. 1111

C. 1323

D. 1223

Answer: C



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28. Which of the following reactions are exothermic in nature?

A. Combustion of Carbon

B. Bond breaking

C. Bond formation

D. Both (a) and (c)

Answer: D



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29. Which of the following changes is not a physical change?

A. glowing of filament in bulb

B. combustion

C. boiling of water

D. sublimation

Answer: B



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30. The substance that loses electrons is called as

A. oxidizing agent

B. reducing agent

C. catalyst

D. none of above

Answer: B



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31. In the reaction $Mg + Cl_2 \rightarrow MgCl_2$

- A. magnesium is oxidized and Cl_2 is reduced
- B. magnesium is reduced and Cl_2 is oxidized
- C. magnesium gains 2 electrons
- D. Chlorine loses 1 electron

Answer: A



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32. When Zn changes to Zn^{+2} it

A. loses 2 electrons

B. loses 1 electron

C. gains 1 electron

D. gains 2 electrons

Answer: A



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33. The process of reduction involves

A. removed of hydrogen

B. gain electrons

C. addition of oxygen

D. loss of electrons

Answer: B



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34. In the reaction
 $3MnO_2 + 4Al \rightarrow 3Mn + 2Al_2O_3$ the oxidizing agent is

A. MnO_2

B. Al

C. $Al_2 - (3)$

D. Mn

Answer: A



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35. In the reaction $H_2S + Cl_2 \rightarrow 2HCl + S$, the oxidizing agent is

A. H_2S

B. Cl_2

C. HCl

D. S

Answer: B



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36. Which of the following statement is correct?

A. oxidation involves gain of electron

B. substance which is reduced is reducing agent

C. exothermic process involves absorption of
heat

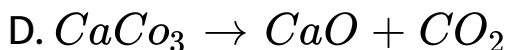
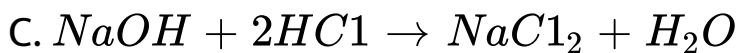
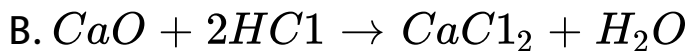
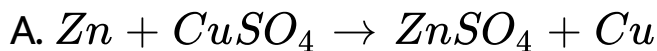
D. oxidation involves loss of electrons

Answer: D



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37. Select Redox reaction from the following

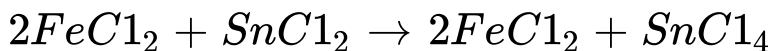


Answer: A



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38. In the reaction



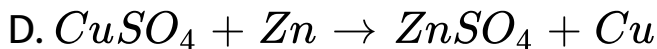
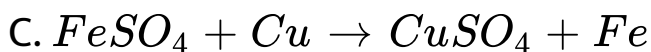
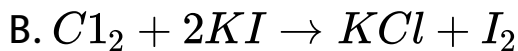
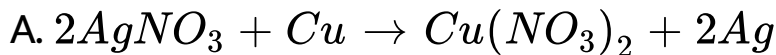
- A. Fe^{+2} reduced to Fe^{+3}
- B. Sn^{+2} is reduced to Sn^{+4}
- C. Sn^{+2} oxidized to Sn
- D. Fe^{+2} gains two electrons

Answer: A



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39. Which of the following reactions is not correct?

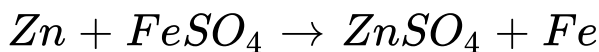


Answer: C



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40. In the reaction of



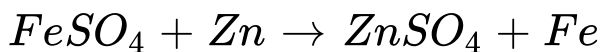
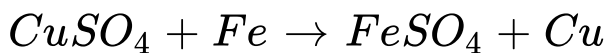
- A. Zn gets oxidized
- B. Fe gets oxidized
- C. Zn is oxidising agent
- D. Zn and Fe both get oxidized

Answer: A



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41. Consider the reaction



- A. Zn is most reaction, Fe is least reactive
- B. Fe is most reaction and Cu is least reactive
- C. Zn is most reaction and Cu is least reactive
- D. Cu is most reaction, Fe is least reactive

Answer: C



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42. Which of the following reaction is not endothermic in nature?

- A. breaking in bonds

B. digestion of food

C. combustion of Carbon

D. evaporation of water

Answer: C



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43. Choose the incorrect statement.

A. physical change in reversible

B. physical change results is formation of new
substance

C. chemical change is permanent

D. physical change is accompanied by energy change

Answer: B



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44. Which of following is fast reaction?

A. reaction between H_2 and O_2 from H_2O

B. reaction between acid and base to form salt and water

C. hydrolysis of ester

D. hydrolysis of sugar to glucose

Answer: B



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Higher Order Thinking Questions

1. The following reaction



is an example of

A. decomposition reaction

B. redox reaction

C. none of these

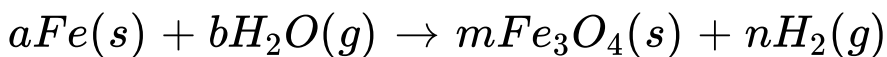
D. both decomposition and redox reaction

Answer: D



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2. In the reaction



a, b, m and n respectively are

A. 6, 8, 2, 8,

B. 3, 4, 1, 4

C. 3, 4, 4, 1

D. 2, 3, 4, 4

Answer: B



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3. Copper reacts with conc. HNO_3 to form copper nitrate, water and gives brown fumes of nitrogen oxide and the name of the reaction is

A. NO_2 , combination reaction

B. NO , redox reaction

C. NO_2 , redox reaction

D. N_2O combination reaction

Answer: C



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4. When an iron nail placed in copper sulphate solution, the reddish brown coating on the nail is

A. soft and dull

B. hard and flaky

C. smooth and shining

D. rough and granular

Answer: B



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5. The mixing of aqueous solution of Na_2SO_4 and $BaCl_2$ results in the formation of

A. white precipitate of sodium chloride and is a combination reaction

B. yellow precipitate of sodium chloride and is a double displacement reaction

C. yellow precipitate of barium sulphate and is a combination reaction

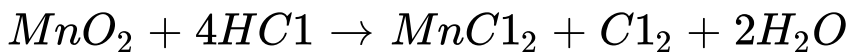
D. white precipitate of barium sulphate and is a double displacement reaction

Answer: D



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6. In the reaction



the substance getting reduced is

A. MnO_2

B. HCl

C. MnCl_2

D. Cl_2

Answer: A



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7. The given reaction

$NaOH + HCl \rightarrow NaCl + H_2O$ is an example of

- A. neutralisation reaction
- B. double displacement reaction
- C. exothermic reaction
- D. all of these

Answer: D



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8. The valency of silver in silver carbonate Ag_2CO_3 is

A. 1

B. 2

C. 3

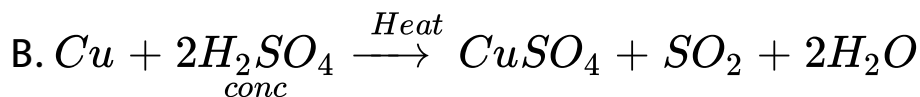
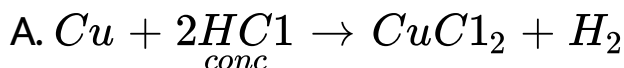
D. zero

Answer: A

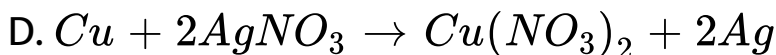
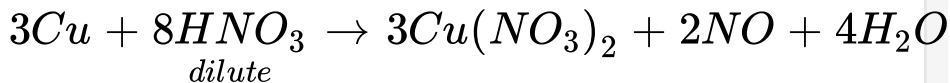


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9. Which of the following reaction does not take place?



C.



Answer: A



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10. When magnesium is burnt in air, a white ash remains as left over. Aqueous solution of this white ash turns red litmus blue. The formula of the white ash and type of reaction respectively are

- A. MgO , non-redox reaction
- B. MgO , redox reaction
- C. $Mg(OH)_2$, decomposition reaction
- D. MgO , displacement reaction

Answer: B



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