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CHEMISTRY

BOOKS - VGS BRILLIANT CHEMISTRY (TELUGU ENGLISH)

ATOMS AND MOLECULES

Improve Your Learning

1. Explain the process and precautions in verifying law of conservation of mass

(OR)

Explain the procedure to prove in a chemical reaction the mass neither created nor destroyed.

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2. 0.24 g sample of compound of oxygen and boron was found by analysis to contain 0.144 g of oxygen and 0.096 g of boron. Calculate the percentage composition of the compound by weight.

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3. Lakshmi gives statement"CO and Co both represent element".Is it correct? State reason.

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4. What is the chemical name and formula of table salt?

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5. Find out the chemical names and formulae for the following common household substances.

a) common salt b) baking soda c) washing soda d) vinegar

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6. Which has more number of atoms 100 g of sodium or 100 g of iron? Justify your answer. Atomic mass of sodium =23 u, atomic mass of iron= 56 u)

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7. Complete the following table.

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8. Fill the following table.

Sl.No.	Name	Symbol/Formula	Molar mass	Number of particles present in molar mass
1	Atomic oxygen		16 g	6.022×10^{23} atoms of oxygen
2	Molecular oxygen			
3	Sodium			
4	Sodium ion		23 g	
5	Sodium chloride			6.022×10^{23} units of sodium chloride
6	Water			

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9. Carbon dioxide is added to 112 g of calcium oxide. The product formed is 200 g of calcium carbonate. Calculate the mass of carbon dioxide used. Which law of chemical combination will govern your answer?

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10. Imagine what would happens if we do not standard symbols for elements.

(OR)

is it necessary to use symbols for elements? Write your opinion.

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11. Draw a diagram to show the experimental setup for the law of conservation of mass.

(OR)

Draw the experimental arrangement used in verifying law of conservation of mass. Write the law of conservation of mass.

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Question Given In The Lesson

1. Did you observe any precipitate in the reaction ?

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2. Fill in the missing words in the blanck spaces in the following statements Cyclone may be caused due to differences in .....

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3. What do you observe from the table? .

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4. What difference do you observe in percentage of copper, carbon and oxygen in two samples? Are elements also made of atoms?

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5. Are elements also made of atoms ?

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6. Fill the blank spaces in the following statements. In india cyclones occur in the month of ..... and.....

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7. Observe the atomicity and fill the following table.

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8. What is valency? Write valency of hydrogen.

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9. Can you write the formula of carbon dioxide and carbon monoxide? Try to write formula for them as we have done in case of water molecule.

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Previous Summative Assassments Questions 2 Mark

1. Fill in the following table .

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Previous Summative Assassments Questions 4 Mark

1. Molecular mass of a molecule is the sum of the atomic masses of individual atoms in it.

(i) Find the molecular mass of Na_2CO_3.

(ii) IF the molecular weight of a compound is 44.

Which is made with carbon and oxygen.What its molecular formula?

(iii) What is the unit of atomic mass?

(iv) On the basis of molecular weights of NaOH and H_2O, which is heavier?

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2. Complete the table

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Essential Material For Examination Purpose 2 Mark

1. What are the Dalton's proposals about the nature of matter ?

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2. What is the proposal of Indian sage kanada, about atom?

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3. What is the use of symbols for elements?

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4. What are the characteristics of a symbol?

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5. In a reaction 5.3 g of sodium carbonate reached with 6 g of ethanoic acid. The products were 2.2 g of carbon dioxide , 0.9 g if water and 8.2 g of sodium ethanoate. Show that these observations

are in agreement with law of conservation of mass.

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6. Hydrogen and oxygen combine in the ratio of 1:8 by mass of form water.

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7. IF one moles of carbon atom weight 12 g. What is the mass of 1 atom of carbon?

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8. When 3 g of carbon is burnt in 8g of oxygen 11 g of carbob dioxide is produced. What mass of carbon dioxide is formed when 3 g of carbon is burnt in 50 g of oxygen.

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9. Give the names of the elements present in the following compounds.

(i) quick lime (ii) Hydrogen bromide (iii) Baking soda (iv) Potassium compounds

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10. Calculate number of sulphur (S_6) present in 16 g of solid sulphur.

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11. Anitha wrote the formula for oxygen molecule as 2O. Is it correct or not? Why?

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Essential Material For Examination Purpose 4 Mark

1. How elements got their names? Explain all examples.

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2. Define the following terms.

Atomicity

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3. Define the following terms.

Ions

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4. Explain the method of writing a formula to a compound using criss-cross method, with the help of an example.

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5. Define the terms : Molecular mass.

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6. Define the terms : Formula unit mass

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7. Define the terms : Mole.

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8. What are the main postulates of Dalton's atomic theory? (OR)

Write the postulates of Dalton's Atomic theory.

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9. what is the defination of incident ray and reflected ray?.

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10. Illustrate the concept of mole through a diagram.

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11. How can you appreciate John Dalton for proposing his atomic theory?

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12. How do you feel after studying the symbols and formulae of different elements and compounds?

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13. Describe the experiment conducted by Joseph L., Proust, which lead him to purpose law of constant proportions.

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14. Write down the formulae of these compounds, using criss-cross methods?

sodium oxide

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15. Write down the formulae of these compounds, using criss-cross methods?

Aluminimum chloride

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16. Write down the formulae of these compounds, using criss-cross methods?

Sodium Sulphide

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17. Write down the formulae of these compounds, using criss-cross methods?

Magnesium hydroxide

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18. Write down the names of compounds represented by the following formulae and also write the anion cation.

(i) $Al_2(SO_4)^3$ (ii) $CaCl_2$ (iii) K_2SO_4 (iv) KNO_3 (v) $CaCO_3$

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19. What are the rules you have to follow while writing a chemical formula?

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20. Calculate the molecular weight $H_2, O_2, Cl_2, CO_2CH_4, C_2H_6, C_2H_4, NH_3, CH_3OH$.

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21. What is atomicity?

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22. Which mirrors are called spherical mirrors? how many types are they?

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23. Why do elements have different atomicities?

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Previous Summative Assessments

1. The correct order in the activity to observe law of conservation of mass is

P. Take 100 ml of silver nitrate solution in a chemical flask and 5 ml of sodium chloride in a test tube.

Q. Due to reaction between these two while precipitate will form.

R. Add sodium chloride solution to silver nitrate solutions

S. Weight of two and record the sum of their weights

T. After reaction weigh it again and it will be equal to the previous reading.

A. P,R,S,Q,T

B. P,S,R,Q,T

C. P,S,Q,R,T

D. P,S,R,T,Q

Answer: A

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2. Uma: mass is neither created nor destroyed during a chemical reaction

Rama: matter is conserved in a chemical reaction.

A. Uma is correct,Rama is incorrect

B. Both Uma and Rama are correct.

C. Both uma and rama is incorrect

D. Uma is incorrect and Rama is correct

Answer: A

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3. $\square$: Sodium chloride :: $Al_2(SO_4)_3$: Aluminium sulphate

A. $NaCO_3$

B. $Na_2(CO_3)_2$

C. $Na(CO_3)_2$

D. Na_2CO_3

Answer: D

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4. David: By transfer of electrons atoms- can attain octet

Dravid: By sharing of electrons atoms- can attain oclet

A. David incorrect,Dravid correct

B. Both david and dravid are incorrect

C. Both david and dravid are correct

D. David correct,dravid incorrect

Answer: C

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5. A given chemical substance always contains the same elements combined in a fixed proportion by mass. Guess the compounds formed in the ratio of $1:1$ and $1:2$ by carbon and oxygen.

A. CO_2, CO

B. CO, CO_3

C. CO_2, CO_3

D. CO, CO_2

Answer: D

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6. A neutral metal atom becomes dipositive ion by loosing two electrons. In the same way a neutral non-metal atom becomes uni-negative ion by gaining one electron. The formula of a compound which is formed by these two ions is

A. m_2X

B. m_2X_1

C. m_2X_3

D. m_2x_2

Answer: D

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7. We can test the law of conservation of mass by the reaction between barium chloride and sodium sulphate.While doing the experiment, one of the precautions you have taken is

A. Reactants should be taken in solid state

B. Cork was not fitted while the reaction is taking place

C. common balance is used to weigh the substances

D. Flask was tilted and swirled for the mixing of two solutions.

Answer: D

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8. The apparatus which is not required in the experiment mentioned in question number (8) is

A. Test tube

B. Beaker

C. Stand

D. Conical flask

Answer: B

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9. The molecular mass of a substances is the sum of the atomic masses of all the atoms in the molecule of the substance . If the atomic masses of hydrogen, sulphur and oxygen are 1,32 ,16

successively then the molecular mass of sulphuric acid is

A. 98 U

B. 49U

C. 97 U

D. 89 U

Answer: A

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10. Some symbols of elements are based on their English names, Some other are based on their Latin names.Which of the following elements have symbols based on Latin names?

(i) Sodium (ii) Neon (iii) Gold (iv) Argon

A. (i) and (iii)

B. (i) and (iv)

C. (ii) and (iv)

D. (i) and (ii)

Answer: A

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11. IF valency of MG is '+2' and valency of Sulphate (SO_4) is '-2', then the formula of compound formed by these

A. Mg_2SO_4

B. $Mg(SO_4)_2$

C. $MgSO_4$

D. $Mg_3(SO_4)_2$

Answer: A

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12. Homogeneous molecules from the following

A. H_2O

B. N_2

C. N_2O_3

D. $FeSO_4$

Answer: B

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13. The common name of hydrogen monoxide is

A. Water

B. Salt

C. Washing soda

D. Baking soda

Answer: A

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14. Valencies of nitrogen and hydrogen and 3 and 1 respectively. The formula of ammonia molecules formed by these two atoms

A. NH_3

B. NH_4

C. N_3H

D. N_4H

Answer: A

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15. P: the atmoicity of Oxygen is 3

Q: The formula of ozone is O_2

A. P-correct,Q-incorrect

B. P-incorrect,Q-correct

C. Both P & Q incorrect

D. Both P & Q correct

Answer: D

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16. One of the precautions to be taken in the conduction of experiment on law of conservation of mass is

A. test tube should not be tilted

B. test tube should be tilted

C. Test tube should be immersed in the conical flask

D. Test tube should be kept outside the conical flask

Answer: B

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17. Ornaments that we wear contains.....

A. Mercury

B. Sodium

C. Calcium

D. Gold

Answer: D

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18. The latin name of the element Tungsten is

A. Aurum

B. Plumbum

C. kalium

D. Wolfram

Answer: D

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19. Molecular formula of Ozone is.....

A. O_3

B. O_2

C. O

D. O_8

Answer: A

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Objective Type Question

1. When the particles of a substance contain only one type of atoms, that substance is called a/an

A. molecule

B. element

C. compound

D. atom

Answer:

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2. Combination of atoms of same elements is known as

A. molecule

B. element

C. compound

D. atom

Answer: A

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3. IF atoms of different elements combine then..... is formed.

A. molecule

B. element

C. compound

D. atom

Answer: C

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4. Symbol of potassium

A. Pb

B. Na

C. Fe

D. K

Answer: D

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5. The latin name of the element Tungsten is

A. Natrium

B. Kalium

C. Wolfram

D. Cuprum

Answer:

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6. Correct method of representing a symbol

A. BE

B. he

C. al

D. Cr

Answer: D

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7. Which of the following is octatomic?

A. nitrogen

B. oxygen

C. carbon

D. sulphur

Answer: D

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8. The valency of sulphate is

A. 2-

B. 2+

C. 3-

D. 3+

Answer: A

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9. In NH_4Cl the cation is

A. Cl-

B. NH_4+

C. NH_4Cl

D. None

Answer:

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10. Chemical formula of aluminium sulphate is

A. Al_2SO_4

B. $(Al_2)_2(SO_4)_3$

C. $Al_2(SO_4)_3$

D. $AlSO_4$

Answer: C

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11. Molecular mass of H_2SO_4 is

A. 98 u

B. 89 u

C. 49 u

D. 106 u

Answer: A

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12. Molar mass of 1.5055×10^{23} number of calcium atoms

A. 20 gm

B. 40 gm

C. 10 gm

D. 30 gm

Answer: C

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13. Convert into moles : 8 gm of Mg.

A. 0.3

B. 3

C. 2

D. 0.2

Answer:

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14. Among the following which element has more number of atoms?

A. Sulphur

B. Calcium

C. Nitrogen

D. Carbon

Answer: D

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15. Mass is neither created nor destroyed during a chemical reaction " This is known as .....

A. Law of constant proportions

B. Law of multiple proportions

C. Law of conservation of mass

D. Law of conservation of energy

Answer: C

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16. Dalton proposed atomic theory based on.....

A. Law of conservation of mass

B. Law of constant proportions

C. Both A & B

D. Neither A nor B

Answer: C

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17. According to Dalton atom is .....particle.

A. indivisible

B. tiny

C. both A & B

D. none of these

Answer: A

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18. The word atom is derived from a Greek word " atmos" means.....

A. indivisible

B. divisible

C. Both A & B

D. none of these

Answer: A

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19. ..... are the building blocks of all matter.

A. Atoms

B. Molecules

C. Elements

D. Compounds

Answer:

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20. Latin word for water is.....

A. hydro

B. oxy

C. helios

D. none of these

Answer:

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21. latin word for atom is.....

A. hydro

B. oxy

C. helios

D. atmos

Answer: D

22. Symbol of Beryllium is.....

A. Ba

B. Be

C. Br

D. B

Answer: B

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23. Symbol of nitrogen is.....

A. Nl

B. Na

C. N

D. No

Answer: C

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24. Symbol for gold.....

A. G

B. Ga

C. Ge

D. Au

Answer:

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25. Cl_2 is the formula for.....

A. Chlorine

B. Cadmium

C. Chromium

D. Calcium

Answer:

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26. We are protected by shielding the earth from some harmful rays of the sun by.....

A. atmosphere

B. ozone layer

C. clouds

D. sky

Answer: B

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27. The number of atoms constituting a molecule is known as its.....

A. valency

B. atomicity

C. atomic number

D. mass number

Answer: B

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28. The atomicity of sodium is.....

A. 1

B. 2

C. 3

D. 4

Answer: B

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29. The combining capacity of one atoms with another atom is its.....

A. valency

B. atomicity

C. atomic number

D. mass number

Answer:

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30. Valency of Argon is.....

A. 0

B. 1

C. 2

D. 3

Answer:

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31. Valency of carbon is .....

A. 1

B. 2

C. 3

D. 4

Answer:

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32. A positively charged ion is called.....

A. radical

B. anion

C. cation

D. multiple ion

Answer:

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33. A negatively charged ion is called.....

A. radical

B. anion

C. cation

D. multiple ion

Answer:

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34. In NH_4OH the anion is.....

A. OH^-

B. NH_4^+

C. NH_3^+

D. NH^-

Answer:

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35. Standard reference for measuring atomic masses of other elements is .....

A. atomic mass of carbon -12

B. atomic mass of carbon-14

C. atomic mass of oxygen-16

D. atomic mass of oxygen -18

Answer: A

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36. The number of times one atom of given element is heavier than $1/12$ th part of atomic mass of carbon-12 is called as.....

A. valency

B. atomicity

C. atomic mass

D. atomic number

Answer:

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37. Atomic mass of Magnesium is.....

A. 8

B. 10

C. 12

D. 24

Answer:

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38. Formula of silver nitrate.....

A. $AgNO_2$

B. $AgSO_4$

C. $AgNO_3$

D. $Ag(NO_3)_2$

Answer:

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39. Molecular mass of sodium carbonate is.....

U.

A. 108

B. 104

C. 110

D. 106

Answer:

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40. Avogadro constant N_A=.....

A. 6.022×10^{20}

B. 6.022×10^{21}

C. 6.022×10^{22}

D. 6.022×10^{23}

Answer:

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41. Molar mass of water=.....U.

A. 16

B. 18

C. 20

D. 22

Answer:

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42. Number of particles present in 32 g of oxygen molecule are.....

A. 6.022×10^{20}

B. 3.011×10^{23}

C. 6.022×10^{22}

D. 6.022×10^{23}

Answer:

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43. Convert in to moles:

22 g of carbon dioxide

A. 1

B. 0.25

C. 0.75

D. 0.5

Answer:

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44. The valency of copper in Cu_2O is.....

A. $+1$

B. $+2$

C. $+3$

D. -1

Answer: A

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45. Number of particles in 7.75 g of phosphorous is.....

A. 6.022×10^{23}

B. 3.011×10^{23}

C. 1.5055×10^{23}

D. 6.022×10^{22}

Answer:

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46. Symbol of nitrate is.....

A. NO_3^-

B. NO_2^-

C. N^{3-}

D. N^-

Answer:

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47. Statement -I The symbol of chloride ion is Cl^-.

Statement -II: The symbol of ammonium ion is NH_4^+.

A. Both the statements are true.

B. Statement-I is true and statement-II is false.

C. Statement-I is false and stament-II is true

D. Both the statements are false

Answer:

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48. Symbol of sodium ion is.....

A. Na

B. Na^{2+}

C. Na^{3+}

D. Na^{+}

Answer: D

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49. The formula of Aluminium Carbonate is.....

A. $AlCO_3$

B. Al_2CO_3

C. $Al_2(CO_3)_3$

D. $Al(CO_3)_2$

Answer:

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50. The symbol of Zinc ion is.....

A. Zn

B. Zn^{+}

C. Zn^{2+}

D. Zn^{3+}

Answer:

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51. Which of the following statement is not given by Dalton? Statement (A) : If mass was to be conserved, then all elements must be made up of

extremely small particles. Statement (B) : If law of constant proportion is to be followed, the particles of same substance be either same or different.

A. A only

B. B only

C. Both A & B

D. none

Answer:

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52. Who is correct?

Srinu: Elements are made up of atoms.

Mohan: Atoms are made of elements.

A. Srinu

B. Mohan

C. Both

D. Neither srinu and mohan

Answer:

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53. Assertion (A): The atomic mass of magnesium is 24.

Reason (R): Magnesium atom is 24 times heavier than that of $1/12$ of mass of carbon.

A. A and R are true.A is supported by R

B. A and R are true , A is not supported by R

C. A and R are false

D. A is true but R is false

Answer:

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54. Assertion (A): Atomic mass has no units.

Reason (R): Atomic mass is an expression of ratio.

A. A and R are true.A is supported by R

B. A and R are true , A is not supported by R

C. A and R are false

D. A is true but R is false

Answer:

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55. An iron rod changed into iron oxide by rusting. Guess how the weights before and after rusting of the iron rod are changed.

If a= weight of the iron rod

b= weight of the iron oxide

A. $a > b$

B. $b < a$

C. a=b

D. we cannot say

Answer:

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56. In the experiment of "To prove law of conservation of mass", Manoj found that weight of 'Mg' in a closed test tube is not equal to the weight of 'MgO' after burning. The reason may be

A. Law of conservation of mass is not applicable to some chemical changes

B. Some gas was escaped in this experiment

C. He was used simple balance in weight

D. Above all

Answer:

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57. $CO : 1 : 1 : : CO_2 :$ .....?

A. $1 : 1$

B. $2 : 1$

C. $1 : 2$

D. $2 : 3$

Answer:

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58. Molecules in water is H_2O then molecules in the Hydrogen is

A. H

B. H_2

C. A or B

D. No molecules

Answer:

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59. Why did symbol of copper taken as 'Cu'?

A. latin name of copper is cuprum

B. original spelling of copper is cupper

C. Symbols of all metals should be taken in two letters

D. Above all

Answer:

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60. Statement (A) : Symbol of carbon is 'C' whereas symbol of calcium is 'Ca'. Reason (R) : Eventhough carbon and calcium have 'C' as its initial, carbon occurs first in periodic table. So, its symbol was taken as 'C' and calcium symbol was taken as 'Ca'.

A. Both A and R are true

B. Both A and R are false

C. A is true and but R is false

D. A is false but R is true

Answer:

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61. Generally all inert gases like , He,Ne,Ar,Kr,Xe are available as monoatomic elements.

Guess the reason.

A. They are more reactive

B. Their valency is zero

C. Their valency is less than 8

D. They are unstable.

Answer:

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62. IF $'x^{-1}'$ and Na'x' then 'x' may be

A. carbon

B. cromium

C. chlorine

D. copper

Answer:

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63. The valencies of 'Mg' and 'O' are.....respectively in MgO.

A. 1,1

B. 2,2

C. 1,2

D. 2,1

Answer:

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64. ' X_2 Y','X' H 'Y', 'X' OH

Guess X and Y.

A. X=Na,Y=OH

B. X=Na,Y=CO_3

C. $X = CO_3$,Y=Zn

D. X=Zn, Y= CO_3

Answer:

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65. By using Ag^+, Cl^-, Na^+, OH^-

How many substance can be formed?

A. 2

B. 3

C. 4

D. 6

Answer:

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66. Molar mass of molecule x= 2 gr

Molar mass of molecule y= 32 gr

Molar mass of molecule x_2y=18 gr

Predict x and y.

A. $x = H_2, y = O_2$

B. $x = O_2, y = H_2$

C. $x = H_2, y = Cl_2$

D. $x = Cl_2, y = H_2$

Answer:

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67. No. of molecules in 44 grams of CO_2 is equal to

A. no. of molecules in 18 gr of H_2O

B. no. of molecules in 2 grams of H_2

C. no. of molecules in 32 grams of O_2

D. any one of the above

Answer:

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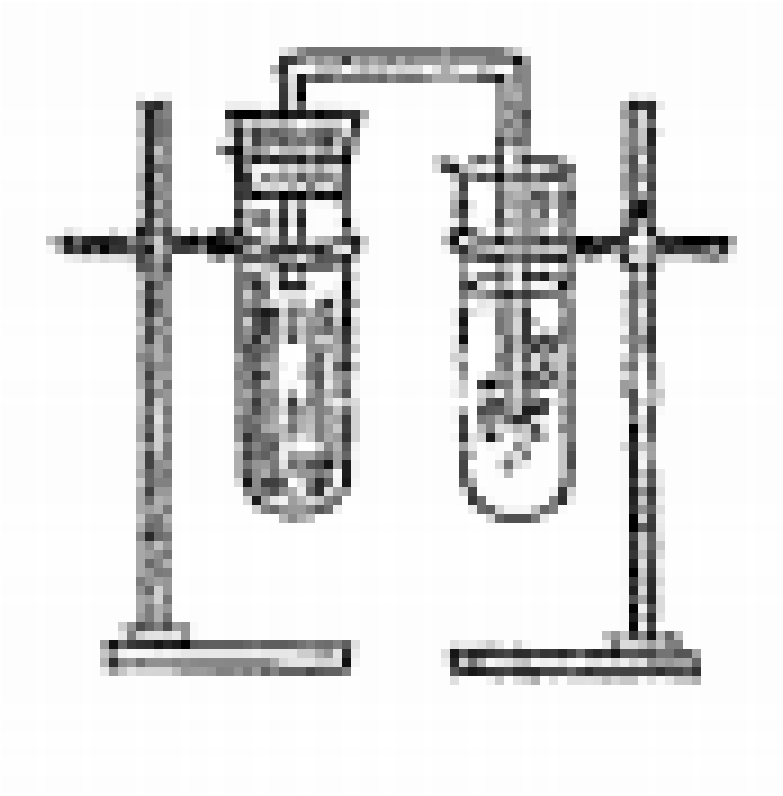
68. In this experiment which gas is produced?



A. put the cork tightly

B. measure components accurately

C. while measuring leave the apparatus freely

D. Above all

Answer:

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69. Law of conservation of mass was proposed by

A. Antoine Lavoisier

B. Joseph L.Proust

C. John dalton

D. Landolt

Answer:

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70. Law of conservation of mass was experimentally verified by

A. Antoine Lavoisier

B. Joseph L.Proust

C. John dalton

D. Landolt

Answer:

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71. Law of constant proportions was proposed by

A. Antoine Lavoisier

B. Joseph L.Proust

C. John dalton

D. Landolt

Answer:

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72. What is the original name of Kanada, an Indian sage, who proposed anu and paramanu?

A. Vaishesika Sutra

B. Rishi

C. Kasyapa

D. Bhaskara

Answer:

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73. Initial letter of an element written in capitals should represent that a particular element. This was proposed by

A. John dalton

B. Landolt

C. John Berzelius

D. Wilhelm Ostwald

Answer:

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74. The word 'mole' was first introduced by

A. John Berzelius

B. Wilhelm Ostwaid

C. John dalton

D. Avogadro

Answer:

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75. Number of particles present in 9 grams of aluminium

A. 2.007×10^{23}

B. 3.011×10^{23}

C. 18.066×10^{23}

D. 6.022×10^{23}

Answer:

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76. Father of modern chemistry is.....

A. Levoisier

B. Proust

C. Dalton

D. Landolt

Answer:

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77. Greek name of sun is.....

A. hydro

B. oxy

C. helios

D. atmos

Answer:

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78. The latin name of Mercury is

A. Aurum

B. Cuprum

C. kalium

D. Hydragyrum

Answer:

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79. O_2 From the given substance select the wrong statement.

A. It is molecule of oxygen

B. It has two elements

C. IT contains two oxygen atoms

D. Is is not a compound

Answer:

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80. $2H_2O$ By using this, select the wrong statement.

A. The atomicity of water molecule is 6

B. Water molecule contains 3 atoms

C. The above indicates 2 molecules of water

D. It is not exists because of unstable.

Answer:

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81. Find the odd one according to their net charge of ion.

A. Hydrogen,sodium,potassium

B. Magnesium,calcium,zinc

C. Aluminium,iron,silver

D. Ammoniu,copper,silver

Answer:

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82. Antione Lavoisier is appreciable because

A. He proposed the law of conservation of mass

B. He is the father of modern chemistry

C. He proposed the law of constant proportions

D. A and B

Answer:

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83. Atomic mass can be measured accurately by using the device.

A. Mass spectrometer

B. Light spectrometer

C. Electronic balance

D. None

Answer:

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84. No. of atoms in 16 g of oxygen.

A. 6.022×10^{23}

B. 3.011×10^{23}

C. 12.044×10^{23}

D. None

Answer:

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85. In a soda water there is 44 grams of CO_2 is mixed with 18 grams of water. The no. of

molecules of H_2CO_3 (Soda water) is

A. 6.022×10^{23}

B. 3.011×10^{23}

C. 12.044×10^{23}

D. None

Answer:

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86. Avogadro is appreciable, because he found that one mole of any substance contains.....

molecules.

A. 6.2×10^{22}

B. 6.4×10^{19}

C. 6.02×10^{23}

D. Countless

Answer:

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87. The great mole concept was introduced by.....

And he is appreciable.

A. Avogadro

B. Ostwaid

C. Dalton

D. Lavoisier

Answer:

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