



CHEMISTRY

BOOKS - VGS BRILLIANT CHEMISTRY (TELUGU ENGLISH)

CHEMICAL REACTIONS AND EQUATIONS

Improve Your Learning Conceptual Understanding

1. a) What is corrosion?
- b) Write the effects of corrosion.
- c) How can prevent it ?



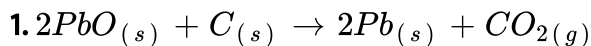
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2. Corrosion is enemy to metals. Can you support this statement ?



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Multiple Choice Questions



Which of the following statements are correct for the above chemical reaction ?

- i) Lead is reduced
- ii) Carbon dioxide is oxidized
- iii) Carbon is oxidized
- iv) Lead oxide is reduced

A. (i) and (ii)

B. (i) and (iii)

C. (i), (ii) and (iii)

D. all

Answer: B



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Essential Material For Examination Purpose

1. Which indicates the arrow mark in a chemical reaction?



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2. Label the parts in the given figure, for the given reaction,
acid + metal.



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3. Write some characteristics of the reactants and products to be expressed in the writing of chemical reaction to made more effective. Give one model of chemical equation.



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4. How do arrow marks simplify the chemical equation? Explain with examples.



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5. Differentiate exothermic and endothermic reactions.



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6. Which reaction does take place with oxygen ?



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7. Is the combustion reaction oxidation reaction ?



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8. c) Give one example of Combustion reaction.



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9. Ozone uses all oxygen atoms in the oxidation reaction with



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10. Dheeraj observed that the colour of walls of his house changes gradually into shiny white after white wash. He asked you, why did the colour changed into bright white.

Predict the reason.



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11. Collect the information and fill the table.





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12. Collect the information and note down in the given table.



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13. Draw a diagram of preparation and test of hydrogen gas in a conical flask.



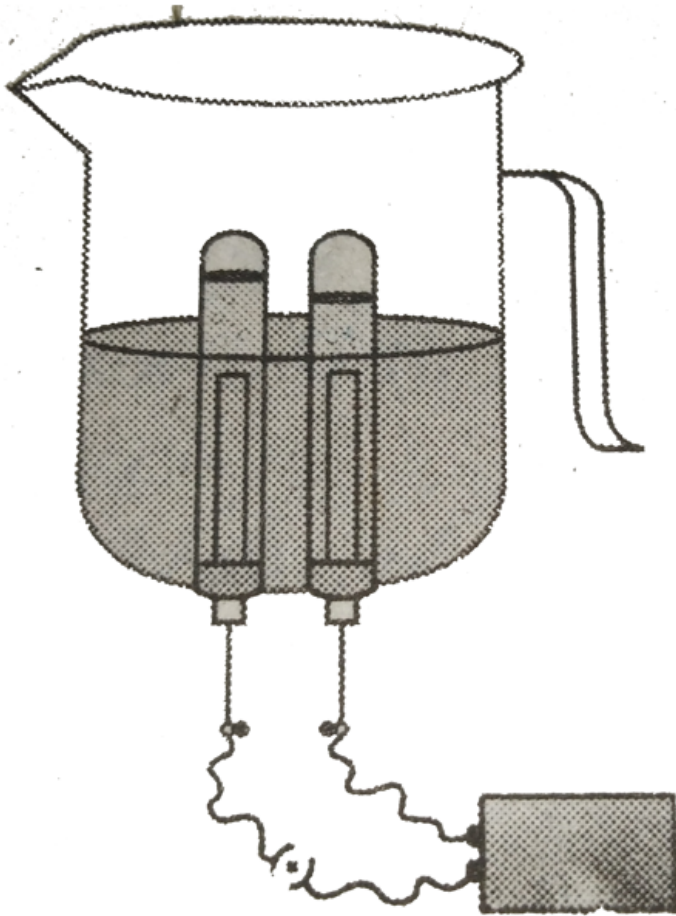
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14. Label the parts in the given diagram.



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15. Label all the parts in the given diagram.



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16. Draw the diagram of electrolysis of water in the lab and table it.



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17. What is the use of galvanising and alloying processes?



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18. Write an activity to each of the following chemical reaction.

A) Photo chemical reaction



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19. Write an activity about how you conduct an experiment to show that more reactive metals replace less reactive metals from their compounds.



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20. What is balanced chemical equation? Why should chemical equations be balanced?



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21. Write the steps involved in 'writing a word equation' of a reaction?



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22. A student was given the following substances and was asked to show types of chemical reactions through experiment. Write how he would have done that.

Copper sulphate solution, barium chloride solution, ferrous sulphate crystals, Iron nails, calcium oxide, water.



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23. a) Write a chemical equation for formation of water.

b) Balance the chemical equation by a systematic method.



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24. Write the steps involved in the balancing a chemical reaction. Give an examples balancing the chemical equation.



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25. a) What is a balancing equation?

b) Balance the equation $H_2 + O_2 \rightarrow H_2O$ step by step.



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26. a) What is a balancing equation?

b) Write the steps involved in the balancing a chemical reaction.

c) Balance the equation $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$ step by step.



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27. Chemical equation for combustion of propane



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28. Iron Oxide reacts with aluminium to form iron and aluminium trioxide.

Write the chemical equation to show the reaction and balance it.



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29. Balance the chemical equation $Fe_2O_3 + Al \rightarrow Fe + Al_2O_3$. Write the steps of balancing.



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30. What are the types of chemical reactions? Give examples to each.



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31. a) What is corrosion? Give examples.



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32. a) What is corrosion?

b) Write the effects of corrosion.

c) How can prevent it ?



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33. c) How can corrosion be prevented ?

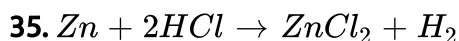


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34. Write an activity about how you conduct an experiment to show that more reactive metals replace less reactive metals from their compounds.



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- a) What type of reaction is it ?
- b) Which element was displaced?
- c) By which element?



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- 36.** a) If we dip the iron nail in the solution of copper sulphate, what changes can you observe?
- b) Write the chemical equation.
- c) What type of reaction is it?

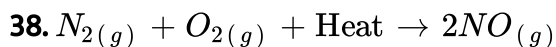


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37. Give one example to chemical displacement.



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What information do you get from the above equation ? Comment.



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39. What can be interpreting from a chemical equation ?



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40. What information does provide a chemical equation.



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41. Here some substances are given. Write them in a chemical equation form.

$C_6H_{12}O_6$, Fe , NH_3 , Na , H_2O , Cl_2 , O_2 , CO_2 , C_2H_5OH , N_2 , NH_4Cl , $NaOH$

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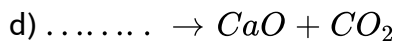
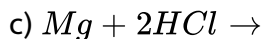
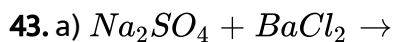


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42. Fill the table with suitable examples.



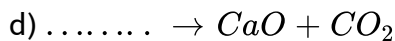
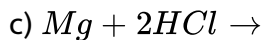
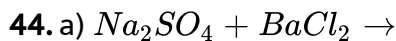
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1) Complete the above reactions.



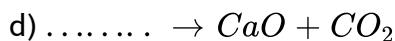
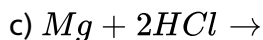
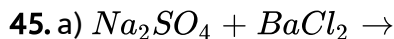
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2) Which of the above is a displacement reaction?



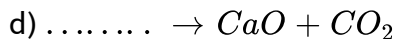
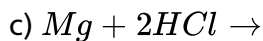
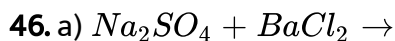
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3) What are the products obtained in the above thermal decomposition reaction ?



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4) How can you say $Na_2SO_4 + BaCl_2$ is a double displacement reaction?



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47. Draw the diagram of apparatus arrangement to decompose calcium carbonate in the lab.



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48. Draw the diagram of reduction of copper oxide to copper and label it.

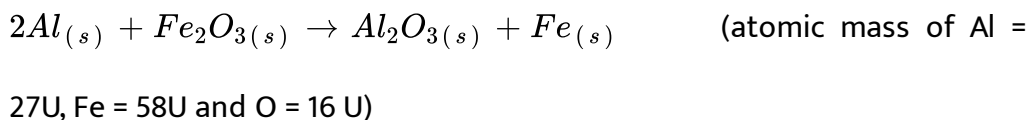


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49. How do you appreciate the chemical equations?

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50. Calculate the amount of aluminium, required to get 1120 kg of iron from the given chemical equation.

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51. Calculate the volume, mass and number of molecules of hydrogen liberated when 230 g of sodium reacts with excess of water at STP (atomic masses of Na = 23U, O = 16 U and H = 1U).

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52. Write some effects of oxidation reactions in daily life.

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53. Is oxidation good or evil or both ? Justify your answer with examples.



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Objective Type Questions

1. Why moist Cl_2 is act on bleaching agent ?

A. In presence of water, Cl_2 gives nascent hydrogen

B. In presence of water, Cl_2 gives nascent oxygen

C. Moist Cl_2 imparts colour to the substance

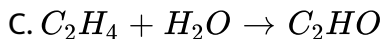
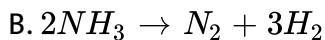
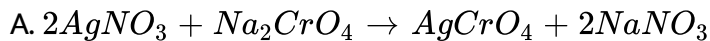
D. Moist chlorine liberates nascent chlorine.

Answer: B



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2. Choose correct matching.



Answer: D



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3. Which of the following statements is/are true?

A. Molecular mass is expressed in unified mass (u)

B. 22.4 l of any gas at STP contains 6.023×10^{23} molecules

C. 28g of N at STP occupies 22.4 litres of volume

D. All are correct

Answer: D



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4. The permanent change among the following is:

- A. Physical change
- B. Chemical change
- C. Both A and B
- D. None

Answer: B



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5. Not a chemical change in the following

- A. Coal is burnt

- B. Food get digested in our body
- C. Formation of the ice from water
- D. Milk converted into curd

Answer: C



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6. Crackers burnt is a

- A. Physical change
- B. Chemical change
- C. Both A and B
- D. None

Answer: B



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7. If new substances are formed with properties completely unlike those of the original substances are occurs in

- A. Physical change
- B. Chemical change
- C. Both A and B
- D. None

Answer: B



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8. Chemical change may be

- A. Exothermic
- B. Endothermic
- C. A and B
- D. A or B

Answer: D



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9. In a chemical formula, subscript indicates

- A. number of atoms
- B. number of compounds
- C. number of moles
- D. number of reactants

Answer: A



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10. which one is correct statement?

- A. The total mass of the products $>$ total mass of the reactants

- B. The total mass of the products $<$ total mass of the reactants
- C. the total mass of the products $=$ total mass of the reactants
- D. The total mass of the products \geq total mass of the reactants

Answer: C



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11. Which one accounts for the mass of any substance?

- A. Compound
- B. Atom
- C. Heat
- D. Electron

Answer: B



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12. The formula unit of H_2O is

A. 0

B. 1

C. 2

D. 3

Answer: B



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13. The formula unit of $MgBr_2$ is

A. 1

B. 2

C. 3

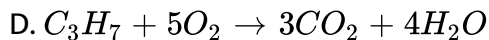
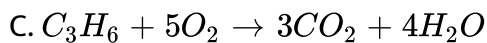
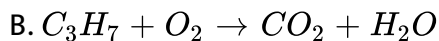
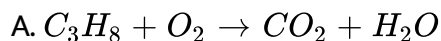
D. 4

Answer: A



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14. Chemical equation for combustion of propane



Answer: C



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15. 1 Gram molar volume of gas occupies

A. 2.24 litres

B. 22.4 litres

C. 4.22 litres

D. 42.2 litres

Answer: B



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16. Avagadaro's number

A. 3.02×10^{32}

B. 3.02×10^{22}

C. 6.02×10^{32}

D. 6.02×10^{23}

Answer: D



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17. 2g of hydrogen gas contains.....molecules.

A. 6.04×10^{23}

B. 8.02×10^{23}

C. 6.02×10^{32}

D. 6.02×10^{23}

Answer: D



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18. 2g of hydrogen gas occupies.....litres.

A. 4.22 litres

B. 44.8 litres

C. 22.4 litres

D. 42.2 litres

Answer: C



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19. The chemical formula of marble is.....

A. CaO

B. Ca

C. CaCO_2

D. CaCO_3

Answer: D



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20. A reaction is carried out by..... Is called thermal decomposition reaction.

- A. heating
- B. lighting
- C. Catalyst
- D. Current passing

Answer: A



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21. All the decomposition reactions are....

- A. Exothermic
- B. Endothermic
- C. A or B
- D. None

Answer: B



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22. The number of molecules present in 1g of hydrogen

A. 6×10^{23}

B. 3×10^{23}

C. 1×10^{23}

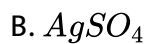
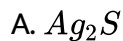
D. 2×10^{23}

Answer: B



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23. Corrosion on silver is



D. none

Answer: A



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24. Corrosion copper is...

A. FeO

B. AgO

C. CuO

D. Cu_2O

Answer: C



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25. Which of the following is an Anti oxidant

A. Vitamin C

B. Ferrous

C. Carbohydrates

D. Proteins

Answer: A



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26. Raising of dough with yeast depends onof sugars to CO_2 and water.

A. corrosion

B. alloying

C. bleaching

D. oxidation

Answer: D

27. Match the following .

Oxide Colour

1) MgO a) White

2) PbO b) Yellow

3) CuO c) Black

4) CaO d) Colourless

A. 1 - a, 2 - b, 3 - c, 4 - d

B. 1 - a, 2 - b, 3 - d, 4 - c

C. 1 - d, 2 - c, 3 - b, 4 - a

D. 1 - d, 2 - c, 3 - a, 4 - b

Answer: A

28. Assertion (A) : $CaO + H_2O \rightarrow Ca(OH)_2$ is a chemical change.

Reason (R) : The original substance does not loss their characteristic

properties in chemical change.

- A. Both 'A' and 'R' are correct and 'A' is supported by 'R'.
- B. Both 'A' and 'R' are correct, but 'A' is not supported by 'R'.
- C. 'A' is correct, but 'R' is incorrect.
- D. 'A' is incorrect, but 'R' is correct.

Answer: C



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29. Assertion (A) : The number of atoms of each element before and after reaction must be the same.

Reason (R) : Law of conservation of mass is followed by every chemical reaction.

- A. Both 'A' and 'R' are correct and 'A' is supported by 'R'.
- B. Both 'A' and 'R' are correct, but 'A' is not supported by 'R'.
- C. 'A' is correct, but 'R' is incorrect.

D. 'A' is incorrect, but 'R' is correct.

Answer: A



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30. Assertion (A) : All the chemical equations must be balanced.

Reason (R) : Atoms are either created or destroyed in chemical reaction.

- A. Both 'A' and 'R' are correct and 'A' is supported by 'R'.
- B. Both 'A' and 'R' are correct, but 'A' is not supported by 'R'.
- C. 'A' is correct, but 'R' is incorrect.
- D. 'A' is incorrect, but 'R' is correct.

Answer: C



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31. Assertion (A) : $2C_3H_8 + 10O_2 \rightarrow 6CO_2 + 8H_2O$ is not a balanced equation.

Reason (R) : The coefficients are not the smallest whole numbers.

- A. Both 'A' and 'R' are correct and 'A' is supported by 'R'.
- B. Both 'A' and 'R' are correct, but 'A' is not supported by 'R'.
- C. 'A' is correct, but 'R' is incorrect.
- D. 'A' is incorrect, but 'R' is correct.

Answer: A



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32. Assertion (A) : All the decomposition reactions are endothermic.

Reason (R) : Decomposition reactions require energy in the form of heat, light or electricity.

- A. Both 'A' and 'R' are correct and 'A' is supported by 'R'.

B. Both 'A' and 'R' are correct, but 'A' is not supported by 'R'.

C. 'A' is correct, but 'R' is incorrect.

D. 'A' is incorrect, but 'R' is correct.

Answer: A



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33. Assertion (A) : The element zinc has displaced hydrogen from hydrochloric acid.

Reason (R) : Generally metals which are less active than hydrogen displaces it from an acid.



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34. Assertion (A) : $Fe + CuSO_4 \rightarrow FeSO_4 + Cu$.

Reason : (R) Copper is more reactive than iron.



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35. Assertion (A) : Oxidation takes place in metals only.

Assertion (B) : Corrosion is an oxidation process.

- A. A and B are correct.
- B. A is correct, B is incorrect.
- C. A is incorrect, B is correct
- D. Both A and B are incorrect.

Answer: C



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36. Which one is not an oxidation process ?

- A. Rancidity
- B. alloying
- C. Corrosion

D. Combustion

Answer: B



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37. Assertion (A) : Chips packets are filled with Vitamin 'C' and Vitamin 'E'.

Reason (R) : Vitamin 'C' and Vitamin 'E' prevents oxidation.

- A. 'A', 'R' are correct, A is supported by 'R'.
- B. 'A', 'R' are correct, A is not supported by 'R'.
- C. 'A' is correct, 'R' is incorrect.
- D. 'A' is incorrect, 'R' is correct.

Answer: D



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38. Match it.

- | | |
|-------------------|----------|
| a) Corrosion | 1) Feal |
| b) Combustion | 2) Food |
| c) Rancidity | 3) Plant |
| d) Photosynthesis | 4) Iron |

A. $a - 2, b - 1, c - 3, d - 4$

B. $a - 3, b1, c - 2, d - 4$

C. $a - 4, b - 1, c - 2, d - 3$

D. $a - 4, b - 2, c - 3, d - 1$

Answer: C



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39. Match it and select the correct option.

- | | |
|------------|--------------|
| i) Copper | A) Fe_2O_3 |
| ii) Silver | B) Ag_2S |
| iii) Iron | C) CuO |

A. $i - A, ii - B, iii - C$

B. $i - B, ii - C, iii - A$

C. $i - C, ii - A, iii - B$

D. $i - C, ii - B, iii - A$

Answer: D



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40. Select the correct option given below.

a) Chemical combination is always endothermic.

b) Chemical decomposition is always exothermic.

A. 'a' and 'b' are correct.

B. 'a' is correct and 'b' is incorrect.

C. 'a' is incorrect and 'b' is correct.

D. both 'a' and 'b' are incorrect.

Answer: D

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41. The layer which gives shiny finish to the wall after whitewash

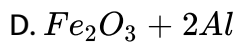
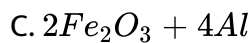
- A. calcium layer
- B. milk layer
- C. marble layer
- D. glass layer

Answer: C

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42. Product of a reaction is $\rightarrow 2Fe + Al_2O_3$, Choose the reactants

- A. $Fe_2O_3 + Al$
- B. $2Fe_2O_3 + 2Al$



Answer: D



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43. On heating lead nitrate brown fumes of Are liberated.



Answer: D



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44. Heating of $Pb(NO_3)_2$ gives two gases 'X' and 'Y'. The colour of X and Y is

A. X = Colourless, Y = Colourless

B. X = Red, Y = Brown

C. X = Brown, Y = Colourless

D. X = Green, Y = Yellow

Answer: C



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45. $xH_2 + yO_2 \rightarrow zH_2O$. The values of x, y, z are

A. x = 1, y = 1, z = 1

B. x = 2, y = 1, z = 2

C. x = 2, y = 2, z = 2

D. x = 2, y = 1, z = 1

Answer: B



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46. Potato chips are spoiled in a chips packet, predict the reason

- A. O_2 not filled
- B. N_2 not filled
- C. CO_2 not filled
- D. Cl_2 not filled

Answer: B



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47. Iron grill was not rusted, predict the proper reason

- A. It may be painted

- B. It may be washed
- C. It may be put in the closed room
- D. It may be covered with cloth

Answer: A



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48. A green coating is observed on a metal pipe. The pipe may be made with

- A. Iron
- B. Aluminium
- C. Copper
- D. Silver

Answer: C



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49. Brown coloured metal powder is changed into black colour on heating. That metal powder may be powder.

- A. Iron
- B. Aluminium
- C. Copper
- D. Silver

Answer: C



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50. An yellow coloured powder changed into gray colour on exposing in the sunlight that powder may be

- A. Aluminium bromide
- B. Silver bromide

C. Copper bromide

D. Magnesium bromide

Answer: B



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51. Gas in a cylinder was put off the burning stick with a pop sound. The gas may be

A. O_2

B. N_2

C. H_2

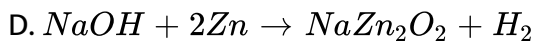
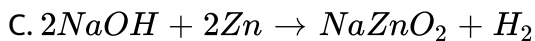
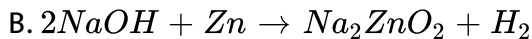
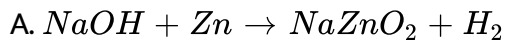
D. Cl_2

Answer: C



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52. The balanced chemical equation, among the following, is

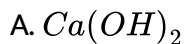


Answer: B



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53. Which one is necessary to test for a hydrogen gas ?



B. Battery

C. Burning stick

D. Litmus paper

Answer: C



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54. Barium sulphate is formed when these are added in beaker

- A. Sodium sulphate, barium sulphate
- B. Sodium chloride, barium sulphate
- C. Sodium sulphate, barium chloride
- D. Barium, sulphur

Answer: C



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55. Which gas is produced when zinc granules react with Hydrochloric acid ?

A. CO_2

B. O_2

C. Cl_2

D. H_2

Answer: D



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56. Heating of which of the following compound gives acidic and basic oxide?

A. $Pb(NO_3)_2$

B. $AgCl$

C. $CaCO_3$

D. All of the above

Answer: C

57. What do you notice when water added to quick lime?

- A. releases heat
- B. absorbs heat
- C. releases H_2
- D. absorbs O_2

Answer: A

58. What substance will be formed when Na_2SO_4 and $BaCl_2$ are added in a beaker?

- A. $BaSO_4$
- B. SO_2

C. NaCl

D. Both A and C

Answer: D



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59. What do you notice when a burning matchstick put in a hydrogen ?

A. put off matchstick

B. releases enormous amount of heat

C. produces dazzling white flame

D. leaves white powder.

Answer: A



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60. What do you notice when $AgCl$ is exposed to sunlight ?

- A. It releases chlorine
- B. Oxidation takes place
- C. Photosynthesis takes place
- D. None of the above

Answer: A



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61. Which one is observed while a magnesium ribbon burning ?

- A. White powder
- B. Dazzling light
- C. Black dust
- D. Both A and B

Answer: D



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62. What do you observe when lead nitrate powder is heated in a boiling tube?

- A. Yellow precipitate
- B. White fumes
- C. Brown fumes
- D. Colour fade

Answer: C



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63. What will you observe when $Pb(NO_3)_2$ is added to KI in a test tube?

- A. A white precipitate
- B. A yellow precipitate
- C. A brick red precipitate
- D. A brown precipitate

Answer: B



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64. What will you notice when Na_2SO_4 is added to $BaCl_2$ in a test tube?

- A. A white precipitate
- B. A yellow precipitate
- C. A brick red precipitate
- D. A brown precipitate

Answer: A



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65. When copper powder is heated ' CuO ' would be formed. What type of colour changes do you notice?

- A. black to brown
- B. brown to yellow
- C. brown to black
- D. yellow to brown

Answer: C



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66. Match the colours of the given substances.

- 1) $2Fe_2O_3$ A) Black
- 2) CuO B) White
- 3) PbI_2 C) Yellow
- 4) $BaSO_4$ D) Brown

A. 1 – D, 2 – A, 3 – C, 4 – B

B. $1 - A, 2 - D, 3 - B, 4 - C$

C. $1 - B, 2 - D, 3 - A, 4 - C$

D. $1 - D, 2 - C, 3 - A, 4 - B$

Answer: A



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67. Which apparatus do we need to decomposes $AgCl_2$?

A. test tube

B. plate

C. conical flask

D. boiling tube

Answer: B



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68. $C_6H_{12}O_6 \rightarrow C_2H_5OH + CO_2$ ischemical reaction.

- A. combination
- B. decomposition
- C. displacement
- D. double displacement

Answer: B



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69. The iron nail dipped in Copper Sulphate solution becomes brown and the blue colour of the Copper Sulphate solution fades. Which type of reaction is this ?

- A. chemical combination
- B. chemical decomposition
- C. Double decomposition

D. Displacement

Answer: D



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70. The process of preparing slaked lime by adding water to quicklime is this type of chemical reaction.

A. Decomposition reaction

B. Exothermic reaction

C. Endothermic reaction

D. Displacement reaction

Answer: B



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71. $Zn + 2HCl \rightarrow ZnCl_2 + H_2$ is an example for

- A. chemical combination
- B. chemical decomposition
- C. Chemical displacement
- D. chemical double displacement

Answer: C



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72. If the gas liberated in an experiment allows the burning splinter to continue burning more brightly in its presence, the gas is

- A. oxygen
- B. nitrogen
- C. hydrogen
- D. carbon dioxide.

Answer: A



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73. The reaction that takes place when quicklime is added to water is

.....

- A. displacement reaction
- B. gas liberating reaction
- C. heat liberating reaction
- D. combustion reaction

Answer: C



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74. Which of the following combination is wrong ?

A. $2N_2O \rightarrow 2N_2 + O_2$ – decomposition

B. $Ca + 2H_2O \rightarrow Ca(OH)_2$ – combination

C. $Zn + 2HCl \rightarrow ZnCl_2 + H_2$ – double decomposition

D. $Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2$ – displacement

Answer: C



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75. $2PbO + C \rightarrow 2Pb + CO_2$

Which of the following statements are correct for the above chemical reaction?

a) Lead is reduced

b) CO_2 is oxidized

c) Carbon is oxidized

d) Lead oxide is oxidized

A. (a) and (b)

B. (a) and (c)

C. (a), (b) and (c)

D. All

Answer: B



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76. Why food materials are packed in air tight containers ?

A. To preserve the nutrients

B. To prevent the spoilage by oxidation

C. To prevent the spoilage by reduction

D. To improve the taste

Answer: B



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77. Flush bags contains potato chips filled with

A. O_2

B. N_2

C. SO_2

D. CO_2

Answer: B



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78. Which of the following is an oxidation process observed in daily life ?

A. Black coating on silver

B. Formation of green compound on copper

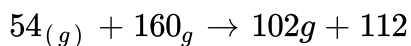
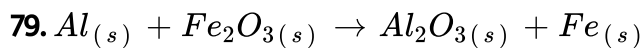
C. Fermentation

D. All the above

Answer: D



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was given, Then, the amount of aluminium required to get 1120 grams of iron is

A. 540 gr

B. 540 kg

C. 112 gr

D. 1120 gr

Answer: A



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80. 10 g of hydrogen contain molecules. If 2g of hydrogen contain 6.02×10^{23} molecules.

A. 3.01×10^{12}

B. 3.01×10^{24}

C. 6.02×10^{24}

D. 6.02×10^{23}

Answer: B



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