



CHEMISTRY

BOOKS - SURA CHEMISTRY (TAMIL ENGLISH)

SOLUTIONS

Textbook Evaluation Choose The Correct Answer

1. A solution is a _____ mixture .

A. Homogeneous

B. heterogeneous

C. homogeneous and heterogeneous

D. non homogeneous

Answer: A



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2. The number of components in a binary solution is _____ .

A. 2

B. 3

C. 4

D. 5

Answer: A



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3. which of the following is the universal solvent ?

A. Acetone

B. Benzene

C. Water

D. Alcohol

Answer: C



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4. A solution in which no more solute can be dissolved in a definite amount of solvent at a given temperature is called _____

- A. Saturated solution
- B. Unsaturated solution
- C. super saturated solution
- D. Dilute solution

Answer: A



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5. Identify the non - aqueous solution.

- A. sodium chloride in water

B. glucose in water

C. Copper sulphate in water

D. sulphur in carbon - di - sulphide

Answer: D



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6. when pressure is increased at constant temperature the solubility of gases in liquid

-----.

A. No change

B. increases

C. decrease

D. no reaction

Answer: B



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7. Solubility of NaCl in 100 ml water is 36 g. If 25 g salt is dissolved in 100 ml of water how

much more salt is required for saturation ?

_____.

A. 12 g

B. 11 g

C. 16 g

D. 20 g

Answer: B



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8. A 25 % alcohol solution means

A. 25 ml alcohol in 100 ml of water

B. 25 ml alcohol in 25 of water

C. 25 ml alcohol in 75 ml of water

D. 75 ml alcohol in 25 ml of water

Answer: C



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9. Deliquescence is due to _____

A. strong affinity to water

B. less affinity to water

C. strong hatred to water

D. inerness to water

Answer: A



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10. Which of the following is hygroscopic in nature ?

A. Ferric chloride

B. copper sulphate pentahydrate

C. silica gel

D. none of the above

Answer: C



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Textbook Evaluation Fill In The Blanks

1. the component present in lesser amount , in a solution is called ____.



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2. Example for liquid in solid type solution is _____.



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3. solubility is the amount of solute dissolved in _____ g of solvent .



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4. Polar compounds are soluble in ____ solvents



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5. Volume percentage decreases with increases in temperature because _____.



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Textbook Evaluation Match The Following

1. 



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Textbook Evaluation True Or False If False Give The Correct Statement

1. Solutions which contain three components are called binary solution .



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2. In a solution the component which is present in lesser amount is called solvent .



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3. Sodium chloride dissolved in water forms a non - aqueous solution .



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4. The molecular formula of green vitriol is $MgSO_4 \cdot 7H_2O$.



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5. When silica gel is kept open , it absorbs moisture from the air , because it is hygroscopic in nature



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Textbook Evaluation Short Answer Questions

1. Define the term : solution .



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2. What is meant by binary solution ?



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3. Give an example each

gas in liquid



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4. Give an example each

solid in liquid





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5. Give an example each

solid in solid



[Watch Video Solution](#)

6. Give an example each

gas in gas



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7. What is aqueous and non - aqueous solution ?



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8. Define Volume percentage .



[Watch Video Solution](#)

9. The aquatic animals live more in cold region
Why ?



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10. Define Hydrated salt .



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11. A hot saturated solution of copper sulphate forms crystals as it cools . Why ?



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12. Classify the following substances into deliquescent , hygroscopic .

Conc . Sulphuric acid , copper sulphate penta hydrate , silica gel , Calcium chloride and Gypsum salt .



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Textbook Evaluation Long Answer Questions

1. Write notes no

Saturated solution



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2. Write notes no

unsaturated solution



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3. Write notes on various factors affecting solubility .



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4. What happens when $MgSO_4 \cdot 7H_2O$ is heated ? Write the appropriate equation



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5. Define solubility



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6. In what way hygroscopic substances differ from deliquescent substances .



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7. A solution is prepared by dissolving 45g of sugar in 180g of water. Calculate the mass percentage of solute.



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8. 3.5 litres of ethanol is present in 15 litres of aqueous solution of ethanol .Calculate volume percent of ethanol solution .



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Textbook Evaluation Hot

1. Vinu dissolves 50 g of sugar in 250 ml of hot water, sarath dissolves 50g of same sugar in

250 ml of cold water. Who will get faster dissolution of sugar? And why?



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2. 'A' is a blue coloured crystalline salt . On heating it loses blue colour and to give 'B' when water is added ,'B' gives back to 'A' Identify A and B, write the equation .



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3. Will the cool drinks give more fizz at top of the hills or at the foot ? Explain .



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Government Exam Questions Answers

1. Common name of copper (II) sulphate pentahydrate is _____.

A. Green Vitriol

B. blue Vitriol

C. Gypsum

D. Epsom salt

Answer:



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2. 1.5 g of solute is dissolved in 15 g of water to form a saturated solution at 298K .Find out the solubility of the solute at the temperature .



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Additional Questions Answers Choose The Correct Answer

1. 40 G of sodium chloride in 100 g of water at 25° C forms _____ solution.

A. super saturated

B. Unsaturated

C. Saturated

D. Both (a) and (b)

Answer: B



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2. 8% of NaCl solution is _____.

A. 8g of NaCl in 100 g of water

B. 8g of NaCl in 92 g of water

C. 92 g NaCl in 8 g of water

D. 92 g of NaCl in 100 g of water

Answer: A



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3. The molecular formula of white vitriol is _____.



Answer: B



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4. Anhydrous copper sulphate is _____.

A. blue

B. bluish green

C. colourless

D. black

Answer: D



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5. Hygroscopic substances are used as _____.

A. oxidizing

B. reducing

C. decarbocyleting

D. drying

Answer: D



6. Solubility of a solute is governed by ____.

A. nature of solute and solvent

B. temperature

C. pressure

D. all the above

Answer: D



7. salt solution containing common salt in water is an example for _____.

- A. binary solution
- B. ternary solution
- C. suspension
- D. colloidal solution

Answer: A



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8. Under which of the following cases dissolution of sugar will be rapid ?

A. sugar crystal in hot water

B. sugar crystal in cold water

C. powdered sugar in hot water

D. powdered sugar in cold water

Answer: C



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9. A beaker contains a solutions of copper sulphate . Recipitaion of copper sulphate takes place when small amount of it added to _____ solution .

- A. saturated
- B. super saturated
- C. Unsaturated
- D. concentreated

Answer: A



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10. Quick lime is dissolved in water is a ____ process .

- A. exothermic
- B. endothermic
- C. reversible
- D. Both (a) and (b)

Answer: A



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11. Example for solid in solid _____

A. soda water

B. camphor in air

C. charcoal

D. alloy

Answer: D



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12. In exothermic process, as the temperature increases, solubility of the salt is _____

A. decreases

B. increases

C. no change

D. increase and then remains constant

Answer: A



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13. The solubility of gases in liquid increases with _____.

- A. increased volume
- B. increased pressure
- C. decreased pressure
- D. none of these

Answer: B



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14. The number of components in a binary solution is _____.

A. one

B. two

C. three

D. four

Answer: B



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15. Which is a non-aqueous solution ?

A. sugar in water

B. common salt in water

C. sulphur in carbon disulphide

D. none

Answer: C



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16. Non -aqueous solvent is _____.

A. benzene

B. ether

C. CS_2

D. All the above

Answer: D



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17. Which of the following is a saturated solution?

- A. 5g NaCl in 100 g water
- B. 10 g NaCl in 100 g water
- C. 20 g NaCl in 100 g water
- D. 36 g NaCl in 100 g water

Answer: D



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18. In which of the following solutions , both solute and solvent are solvent ?

A. cork

B. cheese

C. alloys

D. smoke

Answer: C



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19. An Example for a solution containing liquid solute in gas solvent is _____.

A. soda water

B. cloud

C. cork

D. smoke

Answer: B



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20. Which of the following factors affect solubility?

A. temperature

B. pressure

C. nature of solvent and solvent

D. all the above

Answer: D



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21. Solubility of KNO_3 _____ with the
increases in temperature .

A. increases

B. decreases

C. remains constant

D. None of these

Answer: A



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22. Solubility of CaO _____ with the increases in temperature .

A. increases

B. decreases

C. remains constant

D. none of these

Answer: B



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23. Solubility of CO_2 gas in water _____ with
the increase in pressure

A. increaes

B. decreases

C. remains constant

D. none of these

Answer: A



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24. Which of the following is a dehydrating agent (absorbs moisture) ?

A. sodium hydroxide

B. anhydrous calcium chloride

C. sugar

D. none of these

Answer: B



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Additional Questions Answers Fill In The Blanks

1. Iodine dissolved in carbon tetrachloride is an example of _____.



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2. The effect of pressure on the solubility of a gas in liquid is given by _____.



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3. Volume pressure on the solubility of a gas in liquid is given by _____.



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4. Blue vitriol contains _____ molecules of water of crystallisation.



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5. Blue vitriol is _____.



[Watch Video Solution](#)

6. Magnesium sulphate heptahydrate is called

_____.



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7. Deliquescent substances are _____.



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8. Conc H_2SO_4 is _____ in nature .



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9. Caustic potash is an example of _____
substance .



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10. The number of water molecules found in
the crystalline substance is called _____.



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11. The substance present in lesser amount , in a solution is called _____



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12. A solution that contain more solute than the saturated solution at a given temperature _____ solution .



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13. Benzene is an example of _____.



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14. Molecular formula of white vitriol is _____.



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15. substances that absorb moisture from atmosphere , they dissolve in the absorbed

water and form solution is called _____.



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16. The colour of Iron (II) sulphate is _____.



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17. Example of liquid in gas ____.



Watch Video Solution

18. In endothermic process , solubility _____ with increase in temperature .



Watch Video Solution

19. IN exothermic process , solubility ____ with increase in temperature .



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20. A solution containing less amount of solute is known in temperature .



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21. Concentration of a solution is amount of solute dissolved in _____.



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22. Solution in a _____ mixture of two or more substances .



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23. The solution containing two components is called _____.



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24. True solution containing two components is called ____.



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25. Polar compound dissolves in _____.





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26. Polar compound dissolves in _____.



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27. _____ is an example for solid in gas .



Watch Video Solution

28. _____ is an example for liquid in liquid .





Watch Video Solution

29. _____ is an example of liquid in gas .



Watch Video Solution

30. The mixture of gases used by deep - sea diver is _____.



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31. When we burn wood , the smoke released is a mixture of solid carbon and gases like _____ and _____.



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32. Air is a mixture of gases like _____, _____, _____ and other gases .



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33. Sulphur dissolves in _____.



Watch Video Solution

34. The IUPAC name of Epsom salt is _____.



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35. 36 g of sodium chloride in 100 g of water at $25^{\circ}C$ forms _____ solution .



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36. The IUPAC name of blue Vitriol is ____.



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37. _____ is a solution of liquid in gas .



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38. _____ are example for solubility of gas in liquid .



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39. Deliquescence is maximum when the atmosphere is ____.



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40. 10 g or 20 g or 30 g of sodium chloride in 100 g of water at 25°C forms an ____ solution .



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41. salt in water forms _____.



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42. Solubility of solid solute in a liquid solvent increases with increase in _____.



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43. ____ substances absorb moisture without changing their physical state .





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44. Mass percentage is independent of _____.



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45. The solubility of sodium chloride in water is _____ g/100 g water .



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46. Naturally existing solutions are ____.



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47. Solubility of gases in liquid ____ with
increases in temperature .



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48. Phosphorous pentoxide chloride is an
($CaCl_2$) example of a ____ substance .



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49. Common salt in water is an example of an _____ solution .



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50. Solubility is mathematically expressed as _____.



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51. The percentage by mass of the solute present in the solution is called _____.



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52. _____ substance absorbs moisture without changing their physical state.



[Watch Video Solution](#)

53. When a solute gets distributed uniformly through the solvent it forms a _____ mixture .



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54. Hygroscopic substances may be _____.



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55. Greater amount of sugar will dissolve in _____ water than in _____ water .



Watch Video Solution

56. Non-polar compounds do not dissolved in _____ solvents.



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57. When blue coloured copper sulphate crystals are heated , it becomes _____ copper sulphate .



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58. _____ acts as a dissolving medium in a solution .



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59. Solubility of _____ in water is more at low temperatures .



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60. soda water is a solution of _____.



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61. _____ solution is one that contains less solute than that of the saturated solution at a

given temperature .



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62. Non- polar compounds are solute in _____ 'Solvents .



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63. Solutions which contain 3 components are called __.



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64. The solution in which any liquid , other than water acts as a solvent is called _____ solution .



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65. Solution contains lesser amount of solute per given amount of solvent is called a _____ solution .



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66. 10 % by volume of the solution of ethanol in water , means _____ of ethanol in 100 ml of solution .



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67. when two solutes are dissolved in one solvent a _____ solution is formed .



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68. the solubility of sodium iodide in water is _____ g/ 100 g water .



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69. The common name of calcium sulphate dihydrate is _____.



[Watch Video Solution](#)

70. Super saturated solutions are _____.



[Watch Video Solution](#)

71. Iodine dissolved in carbon tetrachloride forms a _____ solution.



[Watch Video Solution](#)

72. _____ is a solutionn of liquid in solid .



[Watch Video Solution](#)

73. Iron (II) sulphate heptahydrate is called as _____.



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74. As temperature increases , solubility of gases in liquid _____.



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75. The process of uniform distribution of solute into solvent is called _____.



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76. The percentage by volume of solute present in the given volume of the solution _____.



[Watch Video Solution](#)

77. Ionic compounds are soluble in ___ and forms aqueous solution .



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78. Hygroscopic substances are used as _____.



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79. Solutions which are made of one solute and one solvent (two components) Are

called _____.



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80. The number of molecules in blue vitriol is 5, so the water of crystallization is _____.



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81. Polar compounds do not dissolve in _____.



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82. Solubility of a solid solute in a liquid solvent increases with increase in _____.



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83. Sulphur dissolved in carbon disulphide is an example of _____ solution .



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84. When the pressure is increased , solubility of gas in liquid _____.



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85. the nature of the solute and solvent plays an important role in _____.



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86. _____ is the primary factor which determine the characteristics of the solution .



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87. The water of crystallization of Epsom salt is

_____.



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88. The solubility of ammonia in water is ___ g/100 g water .



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89. ___ absorb enough water and get completely dissolved .



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90. The solubility of glucose in water is _____ g

/ 100 g water



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91. On heating, blue vitriol changes its colour

from _____.



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92. Deliquescence is maximum when when the temperature is _____.



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93. In a solution , the component which is present in a larger amount , is called _____.



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94. On heating these hydrated crystalline salts, lose their water of crystallization and become ___.



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95. Alcohol, benzene, ether, acetone are used as ___ solvents.



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96. In solution , the component present in larger amount is called as _____.



Watch Video Solution

97. Calculate the percentage of salt for 20 g of salt in 80 g of water .



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98. The solubility of sodium bromide in water is _____ g / 100 g water .



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99. sand in water forms _____.



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100. The solution in which water acts a solvent in called _____ solution.



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101. The solubility of calcium carbonate in water is g/ 100 water .



[Watch Video Solution](#)

102. The molecules formula of white vitriol is _____.



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Additional Questions Answers State Whether The Following Statements Are True Or False Correct The False Statement

1. Sea water is an example of solution .



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2. Supar saturated solutions are highly stable .



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3. The solution with higher amount of solute is called a dilute solution .



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4. In exothermic process , solubility decreases with increase in temperature .



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5. Mass percentage is independent of temperature .



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6. Deliquescent substances do not changes its state on exposure to air



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7. The common rule for solubility is "Like dissolves Like".



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8. Volume percentage is expressed when solute is a solid and solvent is a liquid.



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9. In ointments , the concentration of solutions are expressed as W/W.



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10. If few drops of water is added to anhydrous $CuSO_4$, it turns blue in colour.



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11. when we burn wood , the smoke released is a mixture of liquid .



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12. Air is a mixture of gases like nitrogen oxygen , carbon dioxide and other gases .



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13. Non-polar compounds are soluble in non - polar solvents



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14. Mass percentage is expressed as W/W.



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Additional Questions Answers Assertion And Reason

1. Assertion : Air is a solution .

Reason : It is a homogenous mixture of nitrogen , oxygen , carbon di oxide and other gases .

A. Both A and R is true and R is the correct explanation of A

B. Both A and R is is true but R is not the correct explanation of A .

C. A is false but R is false

D. A is false but R is true .

Answer:



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2. Assertion : sulphur dissolves in water .

Reason : Non polar substances are soluble in non polar solvents

A. Both A and R is true and R is the correct explanation of A

B. Both A and R is is true but R is not the correct explanation of A .

C. A is false but R is false

D. A is false but R is true .

Answer:



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3. Assertion : Aquatic animals live more in cold regions .

Reason .

Solubility of gases in liquids decrease with increase in temperature .

- A. Both A and R is true and R is the correct explanation of A
- B. Both A and R is is true but R is not the correct explanation of A .
- C. A is false but R is false
- D. A is false but R is true .

Answer:



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4. Assertion : Sodium chloride (table salt) is dissolved in water

Reason : Polar substances are soluble in polar solvents.

A. Both A and R is true and R is the correct explanation of A

B. Both A and R is true but R is not the correct explanation of A .

C. A is false but R is false

D. A is false but R is true .

Answer:



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5. Assertion : when silica gel is kept open , it absorbs moisture from the air .

Reason : It is deliquescent nature .

A. Both A and R is true and R is the correct explanation of A

B. Both A and R is is true but R is not the correct explanation of A .

C. A is false but R is false

D. A is false but R is true .

Answer:



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**Additional Questions Answers Analogy Type
Questions Identify The First Words And Their
Relationship And Suggest A Suitable Word For
The Fourth Blank**

1. Homogeneous solution : salt + water ::

Heterogeneous solution : _____.



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2. Higher amount of solute : concentrated

solution :: Lesser amount of solute : _____.



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3. Copper sulphate is heated : anhydrous (colour less) :: copper sulphate is cooled : _____.



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4. aqueous solution : Water :: Non - aqueous solution : _____.



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5. Green vitriol : Iron (II) sulphate heptahydrate :: white vitriol : _____.



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6. Endothermic process : solubility increases :: Exothermic process : _____.



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7. Binary solutions : Two components :: Ternary solutions : _____.



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8. Hygroscopic : silica gel :: deliquescence : _____.



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9. Epsom salt : $MgSO_4 \cdot 7H_2O$:: Gypsum : _____.



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Additional Questions Answers Very Short Answers

1. Define solute and solvent .



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2. Define dissolution .



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3. What is super saturated solution ? Give an example .



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4. Define mass percentage .



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5. What is blue vitriol



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6. what is Epsom ?



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7. Find the concentration of solution in terms of weight percent if 20 g of sugar is dissolved in 40 g of water .



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8. What are the important conditions for maximum deliquescence to occur ?



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9. Give an example of (i) gas in liquid (ii) liquid in liquid .



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10. 5g of copper sulphate are dissolved in 100 g of water to form a saturated solution at 298 K .Find out the solubility of the solute at the temperature .



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11. What are the liquids present in human body ?



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12. Define ternary solution .



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13. 7.5 litre of ethanol is present in 15 litre of aqueous solution of ethanol . Calculate Volume percent of ethanol solution .



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14. Why is it budding when water is boiled ?





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15. Define ionic substances are dissolved in water to make their saturated aqueous solution , their ions attract water molecules which then attached chemically in certain ratio .This process is called hydration .



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16. What is water of crystallization ?



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Additional Questions Answers Short Answers

1. Calculate the percent by mass of glucose in a solution made by dissolving 50 g of glucose in 500 g of water .



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2. We often find small packets of silica gel in food packs leather and electronic goods .

What is silica gel ?



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3. We often find small packets of silica gel in food packs leather and electronic goods .

Why are these little packets of silica gel kept in leather products ?



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4. We often find small packets of silica gel in food packs leather and electronic goods .

What property of silica gel is the reason for its extensive use ?



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5. We often find small packets of silica gel in food packs leather and electronic goods .

What are hygroscopic substances ?



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6. 60 grams of NaOH is dissolved in 120 grams of water at $25^{\circ}C$ to form a saturated solution. Find mass percentage of solute and solvent .



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Additional Questions Answers Long Answers

1. explain solid solution ,Liquid solution and Gaseous solution .



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Additional Questions Answers Higher Order Thinking Skills Hots

1. Butter is an example of one type of colloidal solution .Name it .



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