



CHEMISTRY

BOOKS - KUMAR PRAKASHAN

CHEMICAL REACTIONS AND EQUATIONS

Questions And Answers

1. State the chemical reactions observed in daily life.

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2. What type of changes are observed during chemical reaction ?

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3. How is the chemical reaction written ? Explain it with suitable example.

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4. What is meant by unbalanced chemical equation ?

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5. State the importance of balanced chemical equation.

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6. Why is the balancing of chemical equation essential ?

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7. How do you balance the chemical equation ? Explain it stepwise with suitable illustration.

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Questions And Answers Intext Questions

1. Why should a magnesium ribbon be cleaned before burning in air ?

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2. Write the balanced equation for the chemical reactions :

Hydrogen + Chlorine Hydrogen chloride

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3. Write the balanced equation for the chemical reactions :

Barium chloride + Aluminium sulphate \rightarrow Barium sulphate +
Aluminium chloride

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4. Write the balanced equation for the chemical reactions :

Sodium + Water \rightarrow Sodium hydroxide + Hydrogen

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5. Write a balanced chemical equation with physical state symbols
for the reactions :

Solutions of barium chloride and sodium sulphate in water react

to give insoluble barium sulphate and the solution of sodium chloride.



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6. Write a balanced chemical equation with physical state symbols for the reactions :

Sodium hydroxide solution (in water) reacts with hydrochloric acid solution (in water) to produce sodium chloride solution and water.



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Questions And Answers Intext Questions Types Of Chemical Reactions

1. How is the product formed during the chemical reaction ?



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Questions And Answers Intext Questions Types Of Chemical Reactions Combination Reaction

1. What is meant by combination reaction ? Explain it with a suitable example.

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2. Mention the type of reactions involving burning of coal and formation of water. Write equation for the reactions.

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3. What is used for whitewashing of the walls of the house carried out ? Explain the preparation by showing chemical equation.



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4. What is called an exothermic chemical reaction ? Give examples.



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5. Respiration is an exothermic reaction (process). Why ?



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Questions And Answers Intext Questions Types Of Chemical Reactions Decomposition Reaction

1. What is meant by decomposition reaction ? Give examples.



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2. Which substance is used in the manufacture of cement ? Write its equation.

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3. Whose presence causes the decomposition reaction ? Can it be considered as an endothermic reaction ?

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4. What happens, when ammonium chloride is added to barium hydroxide ? What changes in temperature take place during the reaction ?

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5. A solution of a substance X is used for whitewashing.

(1) Name the substance .X. and write its formula.

(2) Write the reaction of the substance X named in (1) above with water.



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6. Why is the amount of gas collected in one of the test tubes in activity double of the amount collected in the other ? Name these gases.



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Questions And Answers Intext Questions Displacement Reaction

1. What is meant by displacement reaction ? Give examples.



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Questions And Answers Intext Questions Double Displacement Reaction

1. What is called double displacement reaction ? Give examples.



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2. What is meant by precipitation reaction ? Explain it with examples.



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Questions And Answers Intext Questions Oxidation And Reduction

1. What is meant by oxidation and reduction reactions ? Explain it by giving examples

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2. What is called redox reaction (or oxidation - reduction reactions)? Explain it with suitable examples.

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Questions And Answers Intext Questions Corrosion

1. Write a note on corrosion.

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Questions And Answers Intext Questions Rancidity

1. Write a note on .Rancidity..

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Questions And Answers Intext Questions

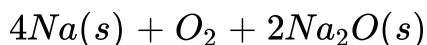
1. Why does the colour of copper sulphate solution change, when an iron nail is dipped in it ?

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2. Give an example of double displacement reaction other than the one given in activity 1.10.

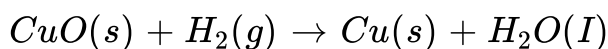
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3. Identify the substances that are oxidised and the substances that are reduced in the reaction :



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4. Identify the substances that are oxidised and the substances that are reduced in the reaction :



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1. Aim: To study the burning of magnesium ribbon in air.

Caution : It is necessary that this activity should be performed in the presence of a teacher. For safety purpose, teacher and student should wear goggles.

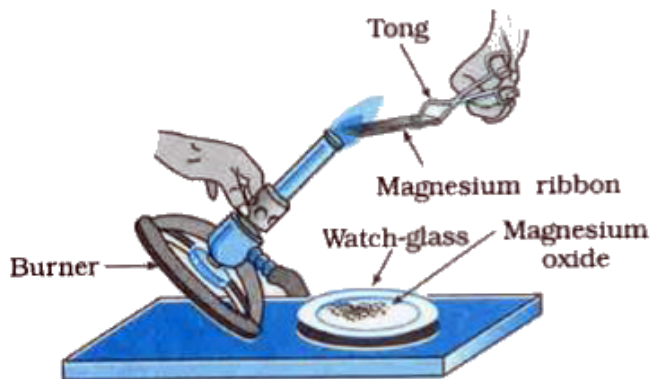
Activity:

Take approximately 3-4 cm long magnesium ribbon and make it clean by rubbing it with sand paper.

Hold it with a pair of tongs and heat on the flame of burner or spirit lamp and the ash being formed collects in the watch-glass as shown in the figure 1.1.

Collected ash in the watch-glass is of magnesium oxide.

Burn the magnesium ribbon. Keeping it away as far as possible from your eyes.



Why is magnesium ribbon selected ?

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2. Aim: To study the burning of magnesium ribbon in air.

Caution : It is necessary that this activity should be performed in the presence of a teacher. For safety purpose, teacher and student should wear goggles.

Activity:

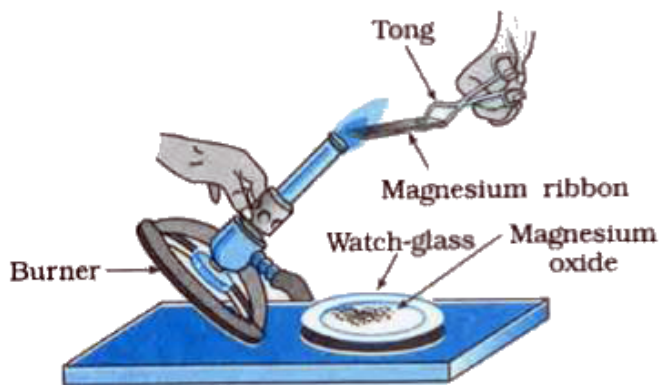
Take approximately 3-4 cm long magnesium ribbon and make it clean by rubbing it with sand paper.

Hold it with a pair of tongs and heat on the flame of burner or

spirit lamp and the ash being formed collects in the watch-glass as shown in the figure 1.1.

Collected ash in the watch-glass is of magnesium oxide.

Burn the magnesium ribbon. Keeping it away as far as possible from your eyes.



What is the colour of magnesium ribbon initially ?

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3. Aim: To study the burning of magnesium ribbon in air.

Caution : It is necessary that this activity should be performed in the presence of a teacher. For safety purpose, teacher and student

should wear goggles.

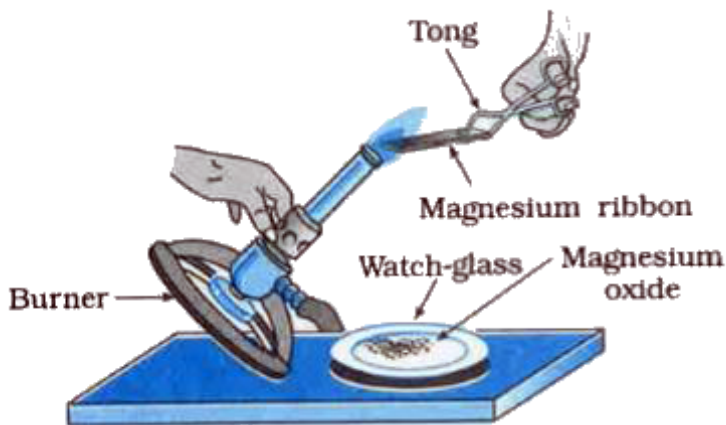
Activity:

Take approximately 3-4 cm long magnesium ribbon and make it clean by rubbing it with sand paper.

Hold it with a pair of tongs and heat on the flame of burner or spirit lamp and the ash being formed collects in the watch-glass as shown in the figure 1.1.

Collected ash in the watch-glass is of magnesium oxide.

Burn the magnesium ribbon. Keeping it away as far as possible from your eyes.



Which type of flame is formed during the burning of magnesium ribbon?



4. Aim: To study the burning of magnesium ribbon in air.

Caution : It is necessary that this activity should be performed in the presence of a teacher. For safety purpose, teacher and student should wear goggles.

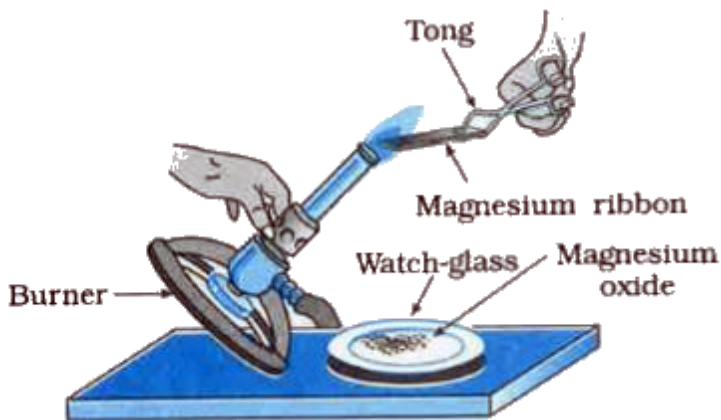
Activity:

Take approximately 3-4 cm long magnesium ribbon and make it clean by rubbing it with sand paper.

Hold it with a pair of tongs and heat on the flame of burner or spirit lamp and the ash being formed collects in the watch-glass as shown in the figure 1.1.

Collected ash in the watch-glass is of magnesium oxide.

Burn the magnesium ribbon. Keeping it away as far as possible from your eyes.



What is the composition of ash collected in watch-glass ?

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5. Aim: To study the burning of magnesium ribbon in air.

Caution : It is necessary that this activity should be performed in the presence of a teacher. For safety purpose, teacher and student should wear goggles.

Activity:

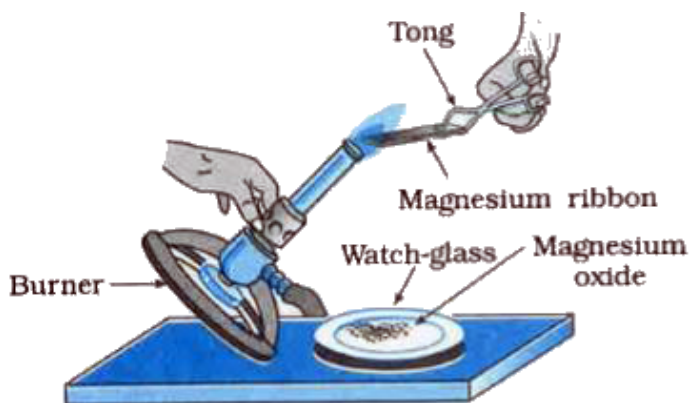
Take approximately 3-4 cm long magnesium ribbon and make it clean by rubbing it with sand paper.

Hold it with a pair of tongs and heat on the flame of burner or

spirit lamp and the ash being formed collects in the watch-glass as shown in the figure 1.1.

Collected ash in the watch-glass is of magnesium oxide.

Burn the magnesium ribbon. Keeping it away as far as possible from your eyes.



Write the chemical reaction of forming the magnesium oxide.

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Questions And Answers Activity 12

1. Aim : To study the reaction between lead nitrate and potassium iodide.

Activity:

Take lead nitrate solution in a test tube.

Add the solution of potassium iodide in it.

What is the colour of lead nitrate solution ?



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2. Aim : To study the reaction between lead nitrate and potassium iodide.

Activity:

Take lead nitrate solution in a test tube.

Add the solution of potassium iodide in it.

What is the colour of an aqueous solution of potassium iodide ?



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3. Aim : To study the reaction between lead nitrate and potassium iodide.

Activity:

Take lead nitrate solution in a test tube.

Add the solution of potassium iodide in it.

Write the balanced chemical equation for the reaction that takes place between lead nitrate and potassium iodide.

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4. Aim : To study the reaction between lead nitrate and potassium iodide.

Activity:

Take lead nitrate solution in a test tube.

Add the solution of potassium iodide in it.

What is the colour of PbI_2 ?



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5. Aim : To study the reaction between lead nitrate and potassium iodide.

Activity:

Take lead nitrate solution in a test tube.

Add the solution of potassium iodide in it.

Identify the type of the above reaction.



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Questions And Answers Activity 13

1. Aim: To study the reaction between zinc metal and dilute sulphuric acid.

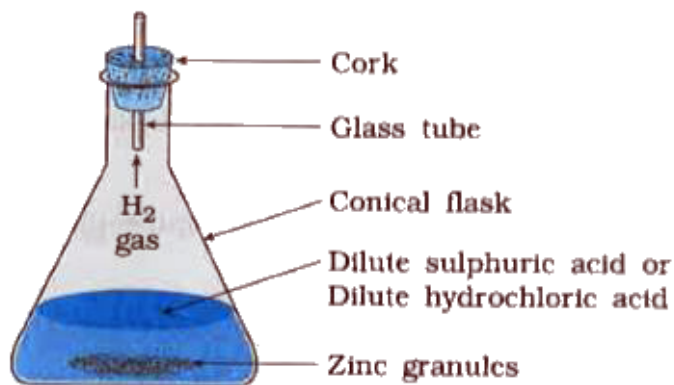
Caution: Use the acid with care.

Activity:

Take a conical flask.

Add a piece of zinc granules in it.

Then add dilute hydrochloric acid or dilute sulphuric acid.



What appears around the zinc granules ?

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2. Aim: To study the reaction between zinc metal and dilute sulphuric acid.

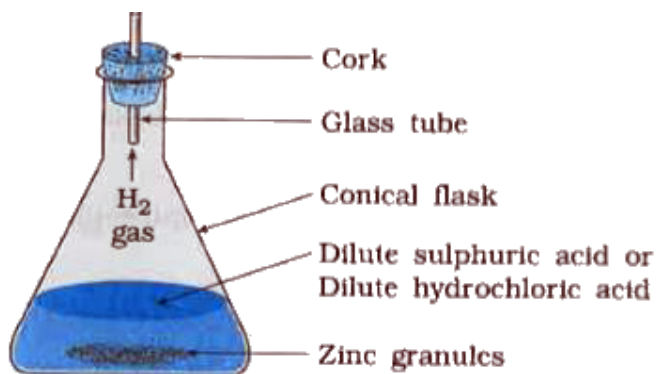
Caution: Use the acid with care.

Activity:

Take a conical flask.

Add a piece of zinc granules in it.

Then add dilute hydrochloric acid or dilute sulphuric acid.



What happens to the conical flask ?

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3. Aim: To study the reaction between zinc metal and dilute sulphuric acid.

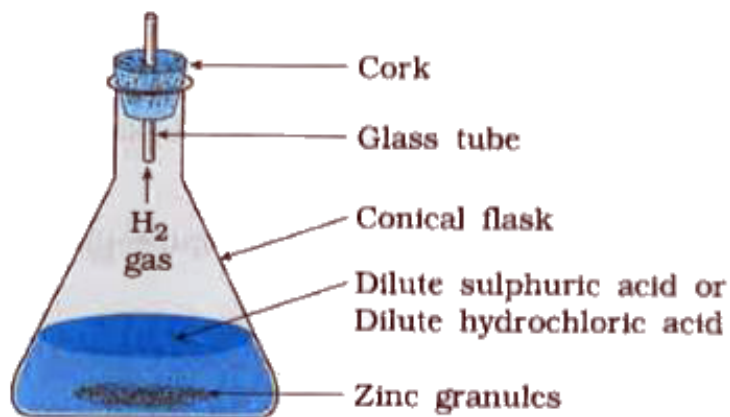
Caution: Use the acid with care.

Activity:

Take a conical flask.

Add a piece of zinc granules in it.

Then add dilute hydrochloric acid or dilute sulphuric acid.



State the reaction occurring between the pieces of zinc and dilute HCl.

[▶ Watch Video Solution](#)

4. Aim: To study the reaction between zinc metal and dilute sulphuric acid.

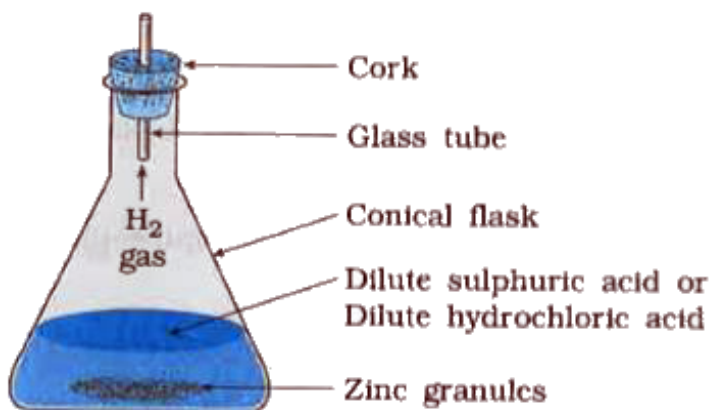
Caution: Use the acid with care.

Activity:

Take a conical flask.

Add a piece of zinc granules in it.

Then add dilute hydrochloric acid or dilute sulphuric acid.



Which type of reaction takes place between the pieces of zinc and dilute HCl ?

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Questions And Answers Activity 1 4

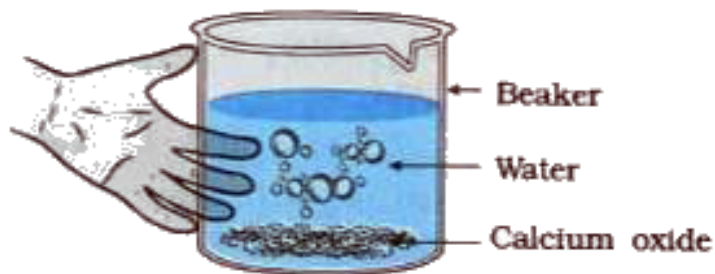
1. Aim: To study the reaction between calcium oxide and water.

Activity:

Take some quick lime (Calcium oxide - CaO) in a beaker. Add water

to it slowly.

Touch the beaker as shown in the figure 1.3.



What is formed by reaction of quick lime with water ? Write reaction.

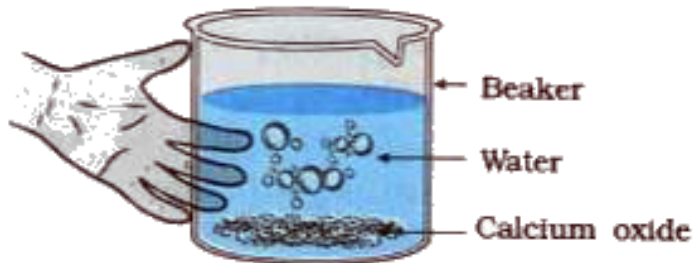
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2. Aim: To study the reaction between calcium oxide and water.

Activity:

Take some quick lime (Calcium oxide - CaO) in a beaker. Add water to it slowly.

Touch the beaker as shown in the figure 1.3.



What is called the reaction occurring between quick lime and water ?

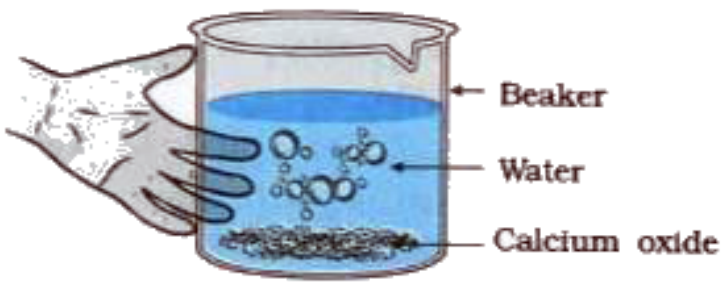
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3. Aim: To study the reaction between calcium oxide and water.

Activity:

Take some quick lime (Calcium oxide - CaO) in a beaker. Add water to it slowly.

Touch the beaker as shown in the figure 1.3.



What do you feel by touching the beaker outside ?

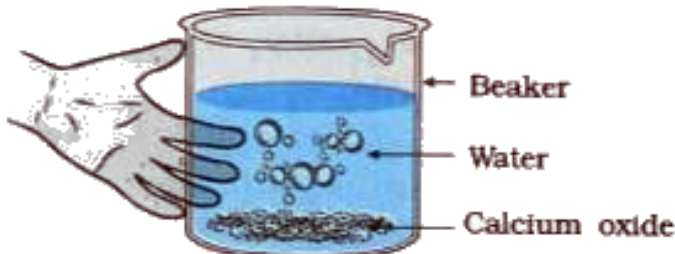
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4. Aim: To study the reaction between calcium oxide and water.

Activity:

Take some quick lime (Calcium oxide - CaO) in a beaker. Add water to it slowly.

Touch the beaker as shown in the figure 1.3.



What is slaked lime ?



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Questions And Answers Activity 1 5

1. Aim : To study the decomposition of ferrous sulphate on heating

Activity:

Take approximately 2 g of ferrous sulphate crystals in a dry boiling tube.

Heat the boiling tube over the flame of a burner.

Observe the colour of the ferrous sulphate crystals carefully during the heating.

Observe the colour of the crystals after heating it.



What is the colour of crystals of ferrous sulphate ?

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2. Aim : To study the decomposition of ferrous sulphate on heating

Activity:

Take approximately 2 g of ferrous sulphate crystals in a dry boiling tube.

Heat the boiling tube over the flame of a burner.

Observe the colour of the ferrous sulphate crystals carefully during the heating.

Observe the colour of the crystals after heating it.



Which colour is observed on heating the crystals of ferrous sulphate ?

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3. Aim : To study the decomposition of ferrous sulphate on heating

Activity:

Take approximately 2 g of ferrous sulphate crystals in a dry boiling tube.

Heat the boiling tube over the flame of a burner.

Observe the colour of the ferrous sulphate crystals carefully during the heating.

Observe the colour of the crystals after heating it.



On heating ferrous sulphate in boiling tube, a gas evolved which has a characteristic smell. What is its reason ?



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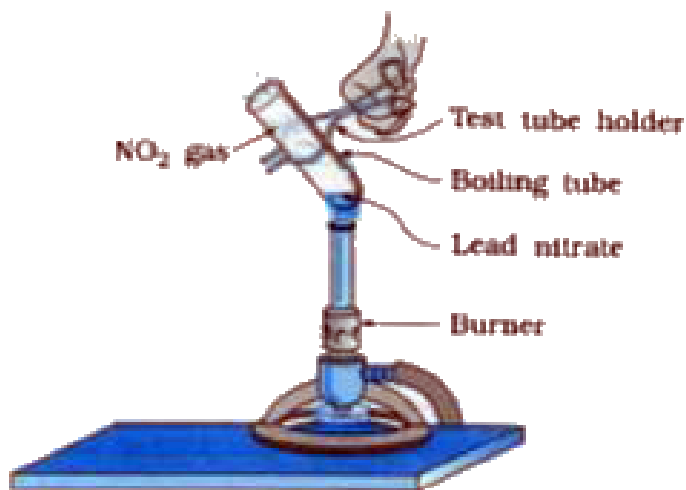
Questions And Answers Activity 16

1. Aim : To study the decomposition of lead nitrate.

Activity:

Take about 2g of lead nitrate powder in a boiling tube.

Hold the boiling tube with a pair of tongs and heat it over the flame of a burner.



Which gas is evolved from boiling tube on heating lead nitrate ?



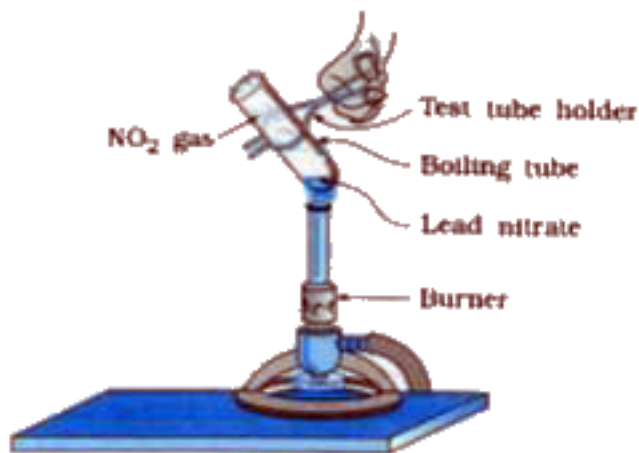
[View Text Solution](#)

2. Aim : To study the decomposition of lead nitrate.

Activity:

Take about 2g of lead nitrate powder in a boiling tube.

Hold the boiling tube with a pair of tongs and heat it over the flame of a burner.



Write the equation of reaction of heating lead nitrate.



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Questions And Answers Activity 17

1. Aim: To study electrolysis of water.

Activity:

Take a plastic mug. Drill two holes at the base of the mug. Fix two rubber corks in it, as shown in the figure 1.6.

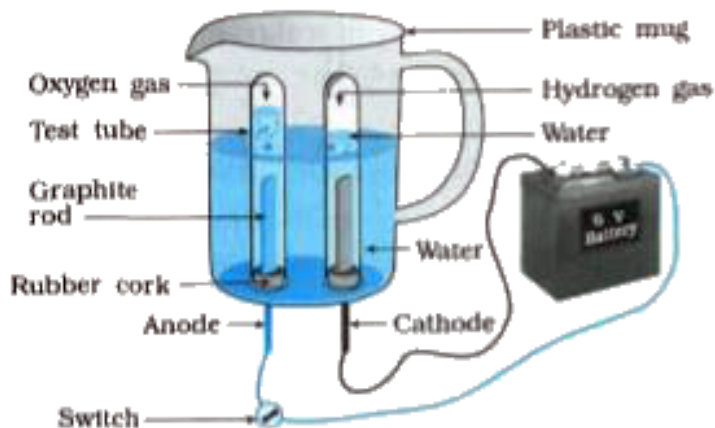
Arrange the test tubes in inverted position such that carbon electrodes remains in it as shown in the figure.

Add water in a mug such that electrodes are immersed.

Add a few drops of dilute sulphuric acid to water.

Connect electrodes to a 6 volt battery.

Now, switch on the current and leave the apparatus undisturbed for some time.



What is an electrode ?

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2. Aim: To study electrolysis of water.

Activity:

Take a plastic mug. Drill two holes at the base of the mug. Fix two rubber corks in it, as shown in the figure 1.6.

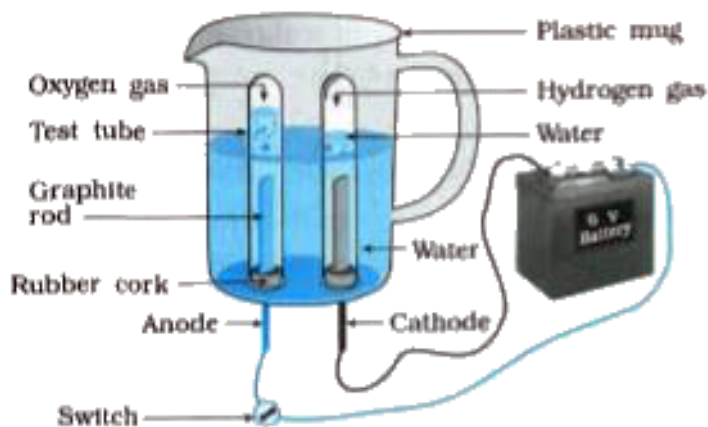
Arrange the test tubes in inverted position such that carbon electrodes remains in it as shown in the figure.

Add water in a mug such that electrodes are immersed.

Add a few drops of dilute sulphuric acid to water.

Connect electrodes to a 6 volt battery.

Now, switch on the current and leave the apparatus undisturbed for some time.



Which gas is evolved at anode during electrolysis of water ?

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3. Aim: To study electrolysis of water.

Activity:

Take a plastic mug. Drill two holes at the base of the mug. Fix two

rubber corks in it, as shown in the figure 1.6.

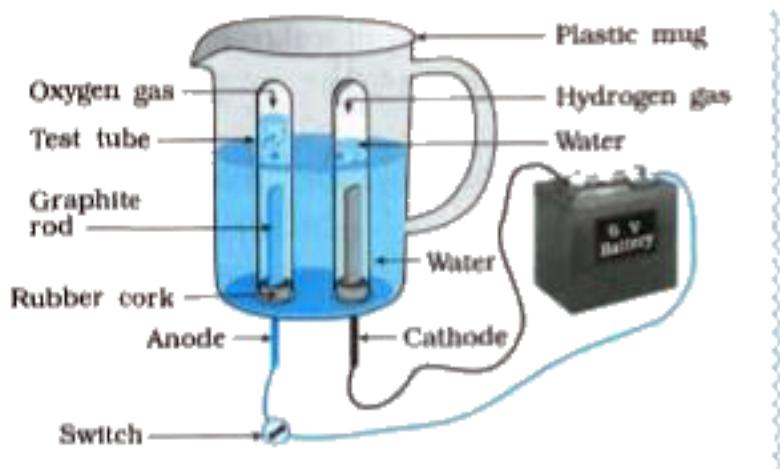
Arrange the test tubes in inverted position such that carbon electrodes remains in it as shown in the figure.

Add water in a mug such that electrodes are immersed.

Add a few drops of dilute sulphuric acid to water.

Connect electrodes to a 6 volt battery.

Now, switch on the current and leave the apparatus undisturbed for some time.



Which gas is evolved at cathode during electrolysis of water ?



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4. Aim: To study electrolysis of water.

Activity:

Take a plastic mug. Drill two holes at the base of the mug. Fix two rubber corks in it, as shown in the figure 1.6.

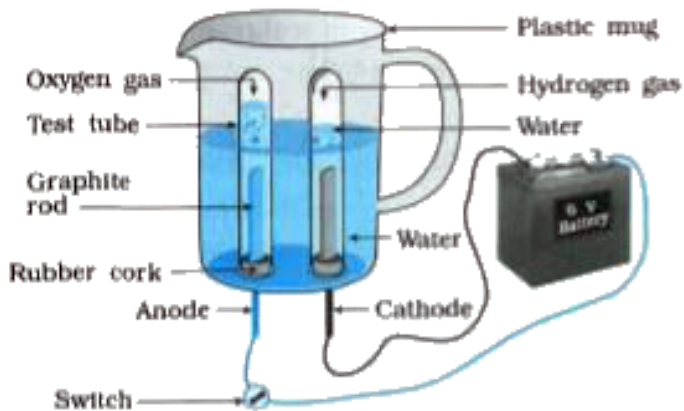
Arrange the test tubes in inverted position such that carbon electrodes remains in it as shown in the figure.

Add water in a mug such that electrodes are immersed.

Add a few drops of dilute sulphuric acid to water.

Connect electrodes to a 6 volt battery.

Now, switch on the current and leave the apparatus undisturbed for some time.



State the chemical equation of reaction of electrolysis of water.

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5. Aim: To study electrolysis of water.

Activity:

Take a plastic mug. Drill two holes at the base of the mug. Fix two rubber corks in it, as shown in the figure 1.6.

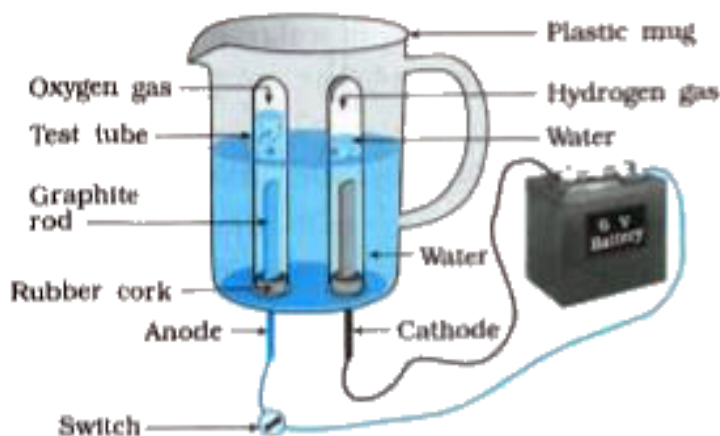
Arrange the test tubes in inverted position such that carbon electrodes remains in it as shown in the figure.

Add water in a mug such that electrodes are immersed.

Add a few drops of dilute sulphuric acid to water.

Connect electrodes to a 6 volt battery.

Now, switch on the current and leave the apparatus undisturbed for some time.



What happens when $O_2(g)$ and $H_2(g)$ evolved are brought close to the burning candle ?

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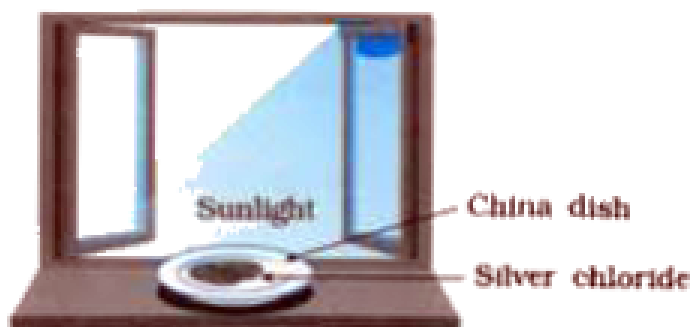
Questions And Answers Activity 18

1. Aim : To study photochemical decomposition of silver chloride.

Activity:

Take about 2 g of silver chloride in a china dish.

Put this china dish in sunlight for some time.



What was the colour of silver chloride before exposure to sunlight?



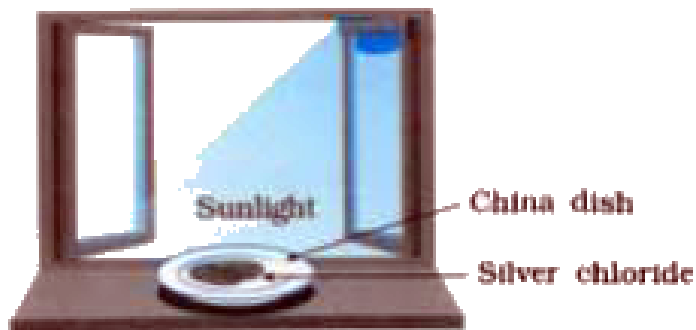
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2. Aim : To study photochemical decomposition of silver chloride.

Activity:

Take about 2 g of silver chloride in a china dish.

Put this china dish in sunlight for some time.



What is the change in colour of silver chloride during its exposure to sunlight for sometime ?

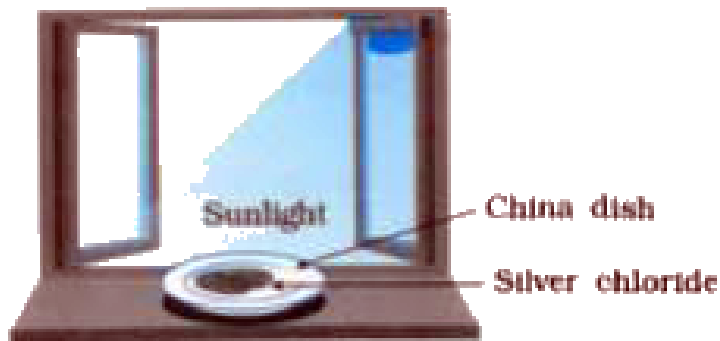
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3. Aim : To study photochemical decomposition of silver chloride.

Activity:

Take about 2 g of silver chloride in a china dish.

Put this china dish in sunlight for some time.



State the chemical equation of decomposition reaction of silver chloride and silver bromide in presence of sunlight.

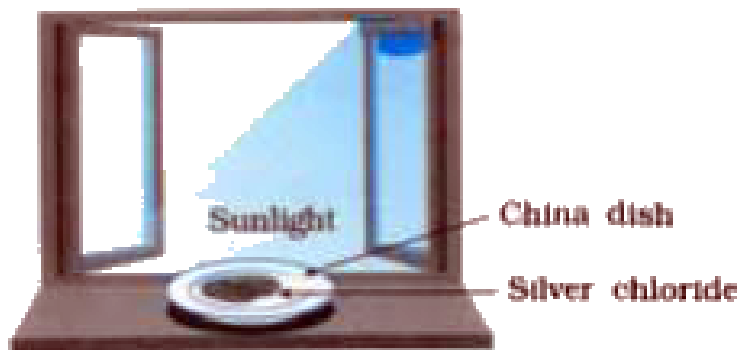
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4. Aim : To study photochemical decomposition of silver chloride.

Activity:

Take about 2 g of silver chloride in a china dish.

Put this china dish in sunlight for some time.



State the uses of AgCl and AgBr .

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Questions And Answers Activity 19

1. Aim : To study the displacement reaction taking place between iron nail and solution of copper sulphate.

Activity:

Take three iron nails and clean their surface by rubbing them with a sand paper.

Take two test tubes labelled as (A) and (B). Take about 10 mL

solution of copper sulphate in each test tube.

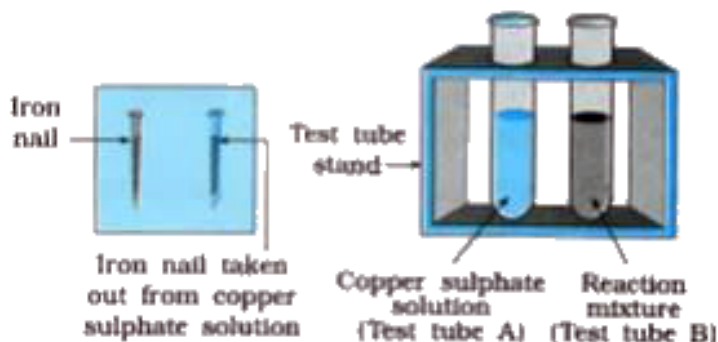
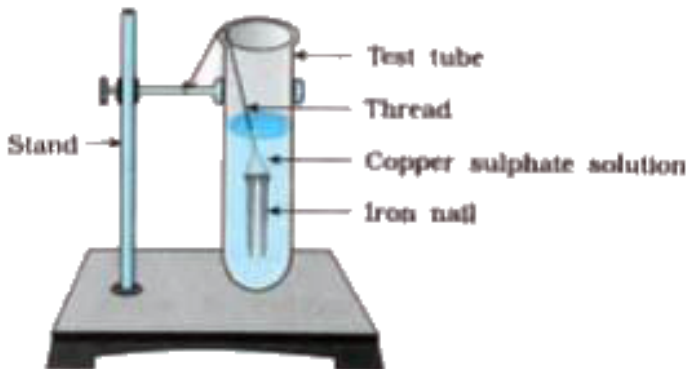
Tie two iron nails with a thread and immerse them in copper sulphate solution for about 20 minutes.

Keep one iron nail aside for comparison.

Take out the iron nails from the copper sulphate solution after 20 minutes.

Compare the colour of both iron nails with the nail kept aside.

Compare the intensity of the colour of copper sulphate solutions of both the test tubes, (A) and (B).



What would be the colour of iron nail placed in the solution of $CuSO_4$?

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2. Aim : To study the displacement reaction taking place between iron nail and solution of copper sulphate.

Activity:

Take three iron nails and clean their surface by rubbing them with a sand paper.

Take two test tubes labelled as (A) and (B). Take about 10 mL solution of copper sulphate in each test tube.

Tie two iron nails with a thread and immerse them in copper sulphate solution for about 20 minutes.

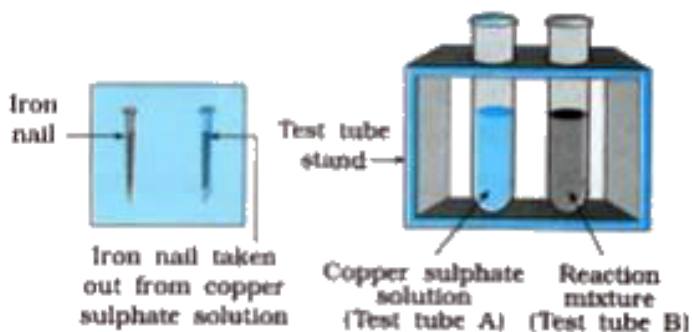
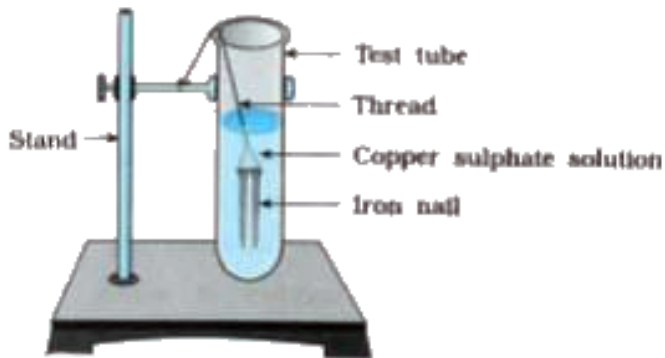
Keep one iron nail aside for comparison.

Take out the iron nails from the copper sulphate solution after 20 minutes.

Compare the colour of both iron nails with the nail kept aside.

Compare the intensity of the colour of copper sulphate solutions of both the test tubes, (A) and (B).

and (B).



What would be the change in colour of solution of copper sulphate ?

[View Text Solution](#)

3. Aim : To study the displacement reaction taking place between iron nail and solution of copper sulphate.

Activity:

Take three iron nails and clean their surface by rubbing them with a sand paper.

Take two test tubes labelled as (A) and (B). Take about 10 mL solution of copper sulphate in each test tube.

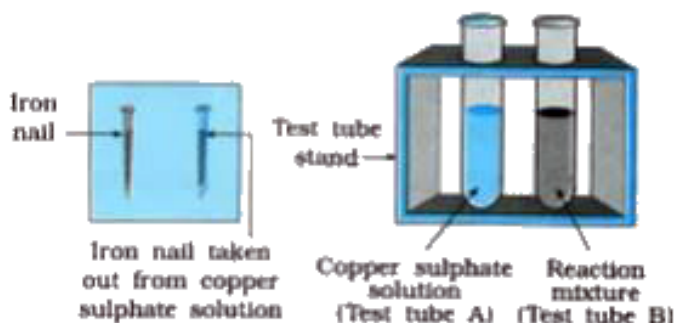
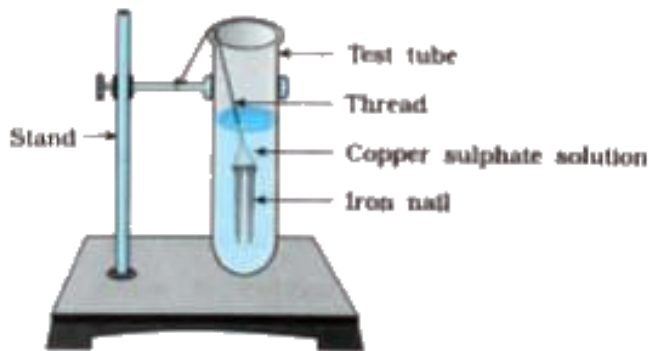
Tie two iron nails with a thread and immerse them in copper sulphate solution for about 20 minutes.

Keep one iron nail aside for comparison.

Take out the iron nails from the copper sulphate solution after 20 minutes.

Compare the colour of both iron nails with the nail kept aside.

Compare the intensity of the colour of copper sulphate solutions of both the test tubes, (A) and (B).



Which reaction takes place when iron nail is dipped in the solution of copper sulphate ?

[View Text Solution](#)

4. Aim : To study the displacement reaction taking place between iron nail and solution of copper sulphate.

Activity:

Take three iron nails and clean their surface by rubbing them with a sand paper.

Take two test tubes labelled as (A) and (B). Take about 10 mL solution of copper sulphate in each test tube.

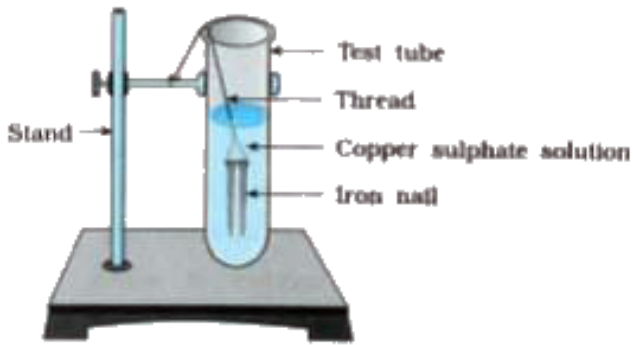
Tie two iron nails with a thread and immerse them in copper sulphate solution for about 20 minutes.

Keep one iron nail aside for comparison.

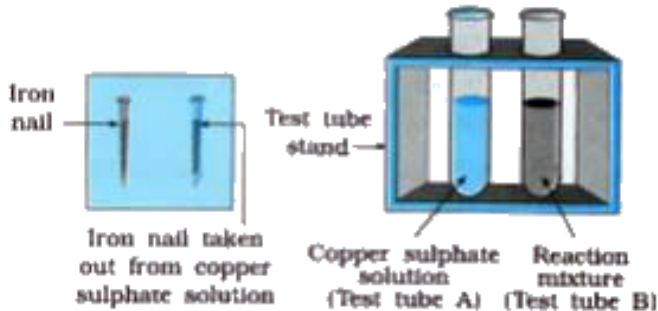
Take out the iron nails from the copper sulphate solution after 20 minutes.

Compare the colour of both iron nails with the nail kept aside.

Compare the intensity of the colour of copper sulphate solutions of both the test tubes, (A) and (B).



[Fig. 1.8 (a) : Iron nails dipped in copper sulphate solution]



What type of chemical reaction occur, when iron nail is dipped in copper sulphate solution ?

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Questions And Answers Activity 1 10

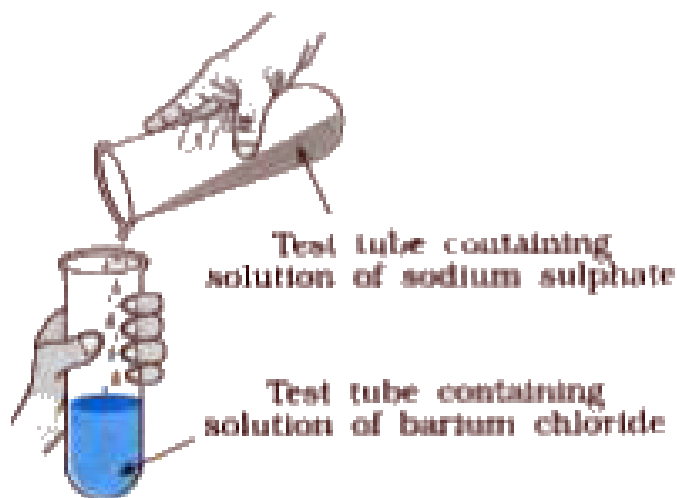
1. Aim: To study double displacement reaction between barium chloride and sodium sulphate.

Activity:

Take about 3 mL of sodium sulphate solution in a test tube.

In another test tube, take about 3 mL of barium chloride solution.

Mix the two solutions as shown in the figure.



What is the colour of sodium sulphate and barium chloride solutions ?

 [View Text Solution](#)

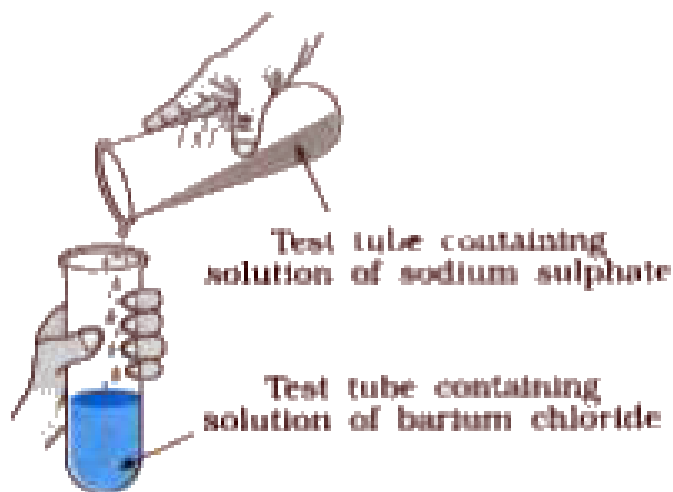
2. Aim: To study double displacement reaction between barium chloride and sodium sulphate.

Activity:

Take about 3 mL of sodium sulphate solution in a test tube.

In another test tube, take about 3 mL of barium chloride solution.

Mix the two solutions as shown in the figure.



Which precipitate is obtained by mixing the solution of sodium sulphate and barium chloride ? Mention its colour.

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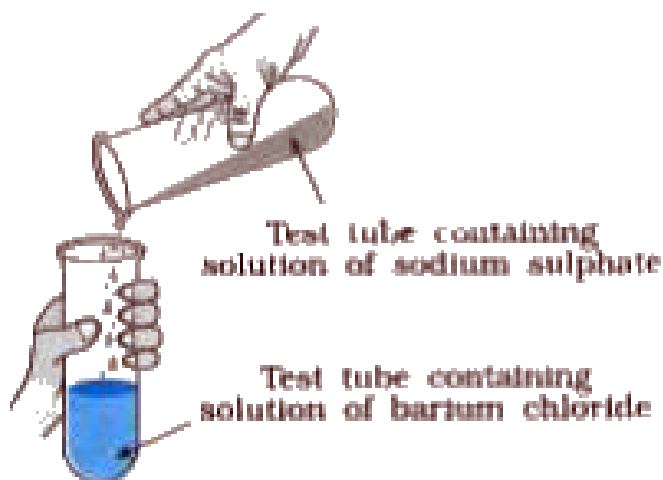
3. Aim: To study double displacement reaction between barium chloride and sodium sulphate.

Activity:

Take about 3 mL of sodium sulphate solution in a test tube.

In another test tube, take about 3 mL of barium chloride solution.

Mix the two solutions as shown in the figure.



Write a balanced chemical equation for the reaction between barium chloride and sodium sulphate.

 [View Text Solution](#)

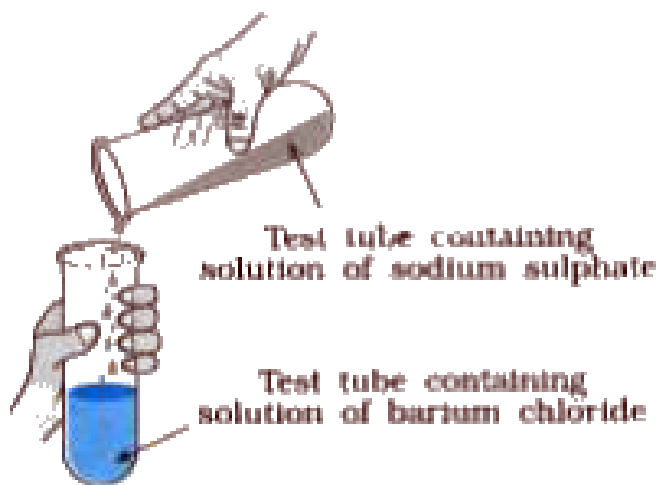
4. Aim: To study double displacement reaction between barium chloride and sodium sulphate.

Activity:

Take about 3 mL of sodium sulphate solution in a test tube.

In another test tube, take about 3 mL of barium chloride solution.

Mix the two solutions as shown in the figure.



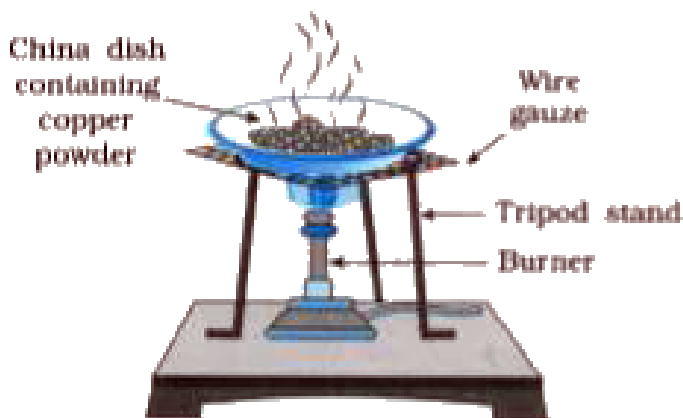
What type of chemical reaction takes place between barium chloride and sodium sulphate ?

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1. Aim: To study oxidation of copper to copper oxide.

Activity:

Take 1 g of copper powder in a china dish and heat it as shown in the figure.



What happens on heating the copper powder ?

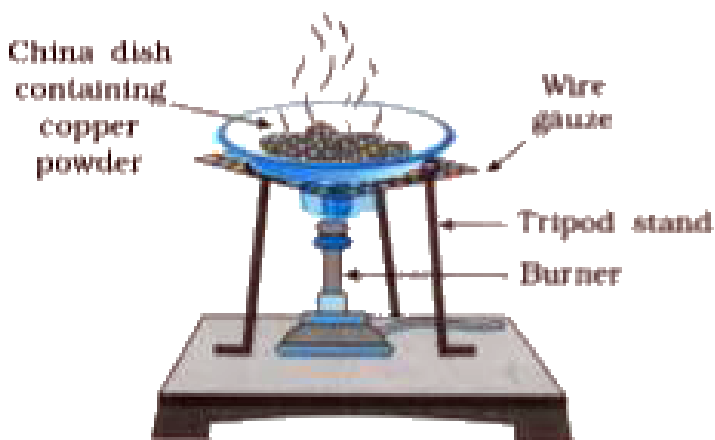


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2. Aim: To study oxidation of copper to copper oxide.

Activity:

Take 1 g of copper powder in a china dish and heat it as shown in the figure.



State the colour of copper oxide.

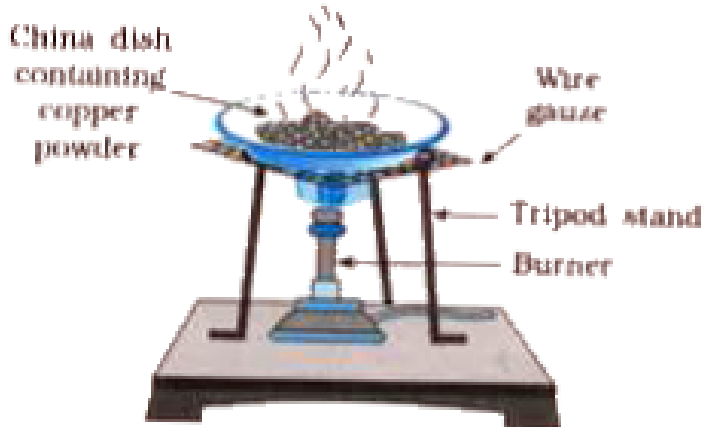


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3. Aim: To study oxidation of copper to copper oxide.

Activity:

Take 1 g of copper powder in a china dish and heat it as shown in the figure.



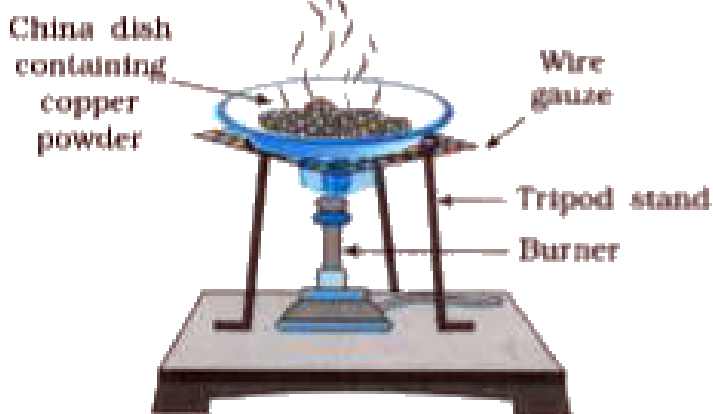
What type of reaction is represented by the reaction of formation of copper oxide from copper ?

[View Text Solution](#)

4. Aim: To study oxidation of copper to copper oxide.

Activity:

Take 1 g of copper powder in a china dish and heat it as shown in the figure.



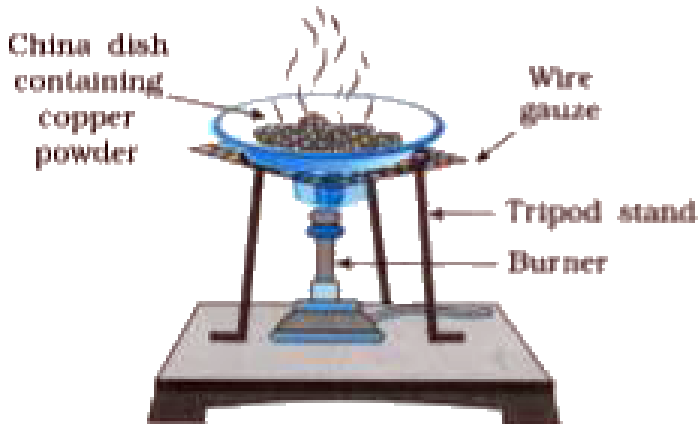
Which product is obtained by passing hydrogen gas over hot copper oxide ?

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5. Aim: To study oxidation of copper to copper oxide.

Activity:

Take 1 g of copper powder in a china dish and heat it as shown in the figure.



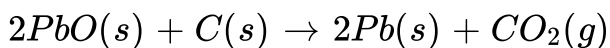
Name the reaction of forming copper from copper oxide.

[View Text Solution](#)

Textual Exercise

1. Which of the statements about the reaction below are incorrect

?



(a) Lead is getting reduced.

(b) Carbon dioxide is getting oxidised.

(c) Carbon is getting oxidised.

(d) Lead oxide is getting reduced.

A. (a) and (b)

B. (a) and (c)

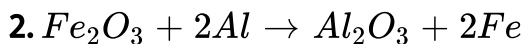
C. (a), (b) and (c)

D. all

Answer: A



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The above reaction is an example of a

A. combination reaction.

B. double displacement reaction.

C. decomposition reaction.

D. displacement reaction,

Answer: D



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3. What happens when dilute hydrochloric acid is added to iron fillings?

A. Hydrogen gas and iron chloride are produced

B. Chlorine gas and iron hydroxide are produced

C. No reaction takes place.

D. Iron salt and water are produced.

Answer:



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4. What is a balanced chemical equation ? Why should chemical equations be balanced ?

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5. Translate the statement into chemical equations and then balance them:

Hydrogen gas combine with nitrogen to form ammonia.

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6. Translate the statements into chemical equations and then balance them:

Hydrogen sulphide gas burns in air to give water and sulphur dioxide.



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7. Translate the statements into chemical equations and then balance them:

Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.



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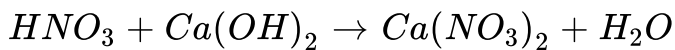
8. Translate the statements into chemical equations and then balance them:

Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.



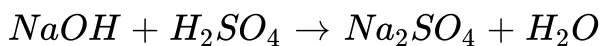
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9. Balance the chemical equations :



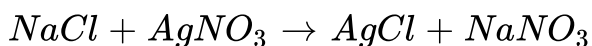
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10. Balance the chemical equations :



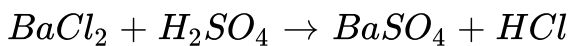
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11. Balance the chemical equations :



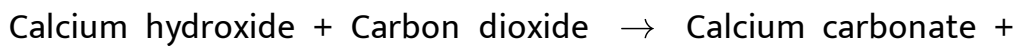
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12. Balance the chemical equations :



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13. Write the balanced chemical equation for the reactions :

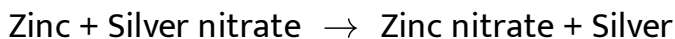


Water



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14. Write the balanced chemical equation for the reactions :



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15. Write the balanced chemical equation for the reactions :

Aluminium + Copper chloride \rightarrow Aluminium chloride + Copper

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16. Write the balanced chemical equation for the reactions :

Barium chloride + Potassium sulphate \rightarrow Barium sulphate +
Potassium chloride

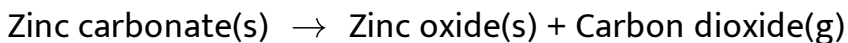
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17. Write the balanced chemical equation and identify the type of reaction :

Potassium bromide(aq) + Barium iodide(aq) \rightarrow Potassium
iodide(aq) + Barium bromide(s)

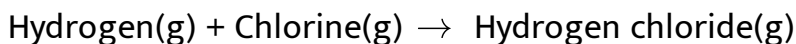
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18. Write the balanced chemical equation and identify the type of reaction :



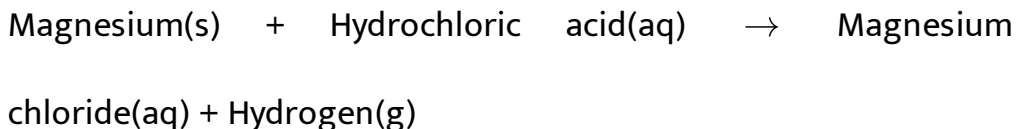
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19. Write the balanced chemical equation and identify the type of reaction :



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20. Write the balanced chemical equation and identify the type of reaction :





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21. What does one mean by exothermic and endothermic reactions ? Give examples.



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22. Why are decomposition reactions called the opposite of combination reactions ? Write equations for these reactions.



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23. Write one equation each for decomposition reactions where energy is supplied in the form of heat, light or electricity.



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24. What is the difference between displacement and double displacement reactions ? Write equations for these reactions.

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25. In the refining of silver, the recovery of silver from silver nitrate solution involved displacement by copper metal. Write down the reaction involved.

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26. What do you mean by a precipitation reaction ? Explain by giving examples.

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27. Explain the following in terms of gain or loss of oxygen with two examples each:

(a) Oxidation (b) Reduction



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28. A shiny brown coloured element .X. on heating in air becomes black in colour. Name the element .X. and the black coloured compound formed.



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29. Why do we apply paint on iron articles ?



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30. Oil and fat containing food items are flushed with nitrogen.

Why ?

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Additional Questions And Answers

1. State the type of reaction in

Vegetable matter changing into compost.

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2. State the type of reaction in

Burning of natural gas

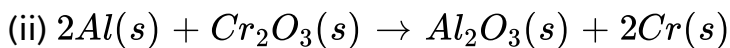
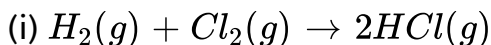
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3. State the type of reaction in

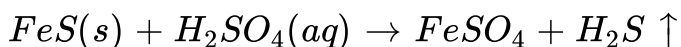
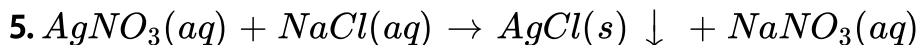
Adding water to quick lime to form slaked lime.

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4. Identify the substance which acts as an oxidising agent and reducing agent in the following reaction :



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6. Distinguish between :

Endothermic reaction and Exothermic reaction

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7. Distinguish between :

Oxidation reaction and Reduction reaction

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8. What happens when carbon dioxide and water react in the same ratio ?

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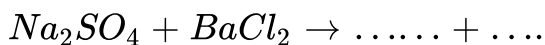
9. How can the black coating of copper oxide be removed chemically?

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10. When quick lime is added to water, a hissing sound is produced. Write the chemical reaction and state the type of reaction that takes place.

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11. Write the name of the products obtained and type of reaction given below:

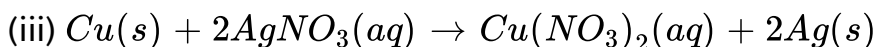
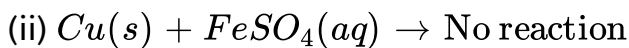
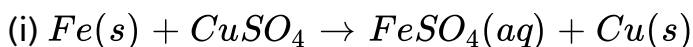


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12. Reactants A and B react together and forms zinc chloride and hydrogen gas. Identify A and B. Write the chemical equation.

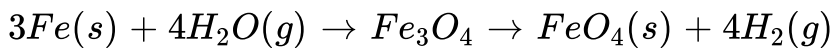
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13. State the ascending order of reactivity for Cu, Ag and Fe metals based on the reaction given below:



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14. State the true and false option for the chemical reaction given below:

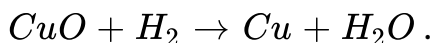


- (i) Fe is being oxidised.
- (ii) Water is being reduced.
- (iii) Water acts as reducing agent.
- (iv) Water acts as oxidising agent.



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15. Choose the true and false statements for the reaction



- (i) CuO is an oxidising agent.
- (ii) H_2 is being oxidised.
- (iii) The reaction is a displacement reaction.
- (iv) The valency of Cu is not changing.



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16. Three test tubes are taken and marked as .X., .Y. and .Z.. In test tube X, iron nail is dipped in water. In test tube Y, iron nail is dipped in mixture of water and oil. In test tube Z, iron nail is added with dry $CaCl_2$. In which test tube, the iron nail will rust ?

Why ?



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17. A metal .M. kept in air turns green and when it is heated, it turns black. Name the metal and the compound formed in both cases.



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18. State the equations of zinc and lead, where it displaces copper from its compounds.



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19. Write balanced equations for thermal decomposition and photochemical decomposition reactions.



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20. Name the methods used to prevent rusting.



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21. Why does most of the metal articles become dull, when left in an open air ?



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22. Give two examples of a reaction which is both endothermic and decomposition in nature.



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23. What is called rancidity ? What is the general name of chemicals which are added to fat and oil containing food to prevent the rancidity ?



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24. Why is photosynthesis considered as an endothermic reaction ?



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25. What is meant by electrolysis ? Mention its two uses with examples.

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26. Write the definition of displacement reaction. Give one example of it and explain, how is it different from double displacement reaction.

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27. How does a chemical equation makes a reaction more informative ?

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28. Give the balanced chemical equations for the

Reaction used in black and white photography



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29. Give the balanced chemical equations for the

Reaction of oxidation of glucose



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30. Give the balanced chemical equations for the

Formation of water from H_2 and O_2



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31. What is a redox reaction ? When a magnesium ribbon is burnt in air, it burns with a dazzling white flame and white powder is formed. Is magnesium oxidised or reduced ? Why ?



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32. Name different types of chemical reactions, and explain it with suitable example.



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33. What is called rancidity ?



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34. Suggest two methods to reduce the effect of rancidity.



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35. How is corrosion different from rusting ?



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36. Water contains hydrogen and oxygen in the ratio of 2:1. Prove it by an activity.



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37. What is meant by chemical equation ? Explain its characteristics with example.



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Additional Questions And Answers Give Scientific Reasons For The Following Statements

1. Food items should be stored in air tight closed containers.

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2. A thin layer of zinc is applied on the plates of steamer to prevent its rusting.

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Objective Questions And Answers Answer The Questions In Short

1. Write the chemical reaction taking place between lead nitrate and potassium iodide. Identify the type of reaction.





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2. What is meant by skeletal equation ? Give an example.



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3. To which sides of a chemical equation the reactants and products are represented ?



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4. Write the name and molecular formula of the substance used in whitewashing the walls of the house.



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5. Blue crystals of copper sulphate on heating in a dry test tube become colourless. Why ?

 [View Text Solution](#)

6. Why does the solution of $AgNO_3$ stored in dark brown coloured bottle ?

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7. What is meant by reversible reaction ?

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8. State the law of conservation of energy.

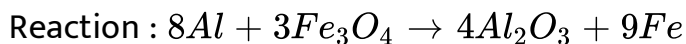
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9. Which carbohydrate substances, on decomposition form glucose ?

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10. Which substance is reduced in the following reaction ?



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11. Which compound is used in black and white photography ?

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12. What characteristics of food containing oil and fat change when exposed to air for a long time ?

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Objective Questions And Answers Fill In The Blanks

1. Respiration is an type of reaction.

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2. Matter can neither be created, nor be destroyed. It is called the law of

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3. Quick lime is also known as

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4. Burning of coal and formation of water is type of reaction.

 [View Text Solution](#)

5. When iron nail is dipped in a solution of copper sulphate, then, the colour of the solution becomes after sometime.

 [View Text Solution](#)

6. Reaction, $CaCO_3 \xrightarrow{\Delta} CaO + CO_2$ is a reaction.

 [View Text Solution](#)

7. is used in black and white photography.

 [View Text Solution](#)

8. Reducing agent undergoes

 [View Text Solution](#)

9. $Pb(s) + \dots\dots\dots ToPbCl_2(aq) + Cu(s)$

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Objective Questions And Answers Choose The Correct Option From Those Given Below Each Question

1. 6g of hydrogen is burnt in the presence of excess oxygen. The mass of water formed is:

- A. 54 g
- B. 108 g
- C. 36 g
- D. 18 g

Answer: A



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2. Which information not obtained by the balanced chemical equation ?

- A. Physical states of reactants and products.

B. Symbols and formulae of all the substances involved in a particular reaction.

C. Number of atoms / molecules of the reactants and products formed.

D. Whether a particular reaction is actually feasible or not.

Answer: D



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3. Which of the following are exothermic reactions ?

(i) Evaporation of water

(ii) Dilution of H_2SO_4

(iii) Reaction of water with quick lime

(iv) Sublimation of crystalline camphor

- A. (i) and (ii)
- B. (iii) and (iv)
- C. (i) and (iv)
- D. (ii) and (iii)

Answer:



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4. Reaction : $4NH_3(g) + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(g)$ is an example of:

- (i) displacement reaction
- (ii) combustion reaction
- (iii) redox reaction
- (iv) neutralisation reaction

- A. (i) and (iv)

B. (ii) and (iii)

C. (iii) and (iv)

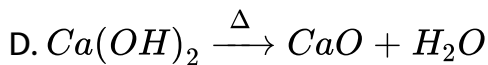
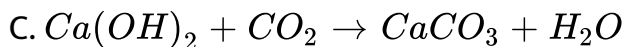
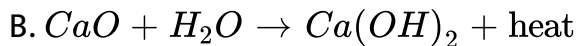
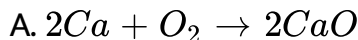
D. (i) and (ii)

Answer:



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5. Which of the following reaction occurs in whitewashing of walls ?



Answer:

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6. What happens when crystals of lead nitrate are heated strongly in a dry test tube ?

- A. Crystals melt immediately.
- B. Brown fumes are obtained.
- C. White fumes get formed in the test tube.
- D. Yellow precipitates are obtained.

Answer:

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7. Dilute hydrochloric acid is added to test tube containing pieces of zinc. The following observations are recorded. Identify the correct observation.

- A. The surface of metal becomes lustrous.
- B. The reaction mixture becomes milky.
- C. Odour of pungent smelling gas is experienced.
- D. A colourless and odourless gas is formed.

Answer:



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8. Rancidity can be prevented by .

- A. adding antioxidants.

- B. storing the food in freeze.
- C. keeping food away from the light.
- D. All of the given.

Answer:



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9. Which of the following is not a single displacement reaction ?

- A. $CuO + H_2 \rightarrow H_2O + Cu$
- B. $Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$
- C. $4NH_3 + 5O_2 \rightarrow 4NO + 6H_2O$
- D. $Zn + 2HCl \rightarrow ZnCl_2 + H_2$

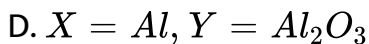
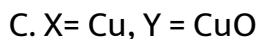
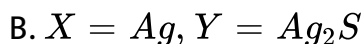
Answer:





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10. An element X on exposure to moist air turns reddish brown and new compound Y is formed. Identify X and Y.

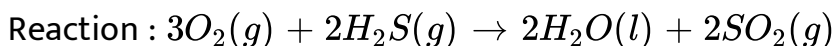


Answer:



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11. Identify the reducing agent in the following reaction :



A. O_2

B. H_2S

C. H_2O

D. SO_2

Answer:



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12. Both H_2 and CO_2 gases are...

A. heavier than air.

B. colourless.

C. acidic in nature.

D. soluble in water.

Answer:

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13. When reddish brown copper metal is heated it forms a black solid surface. Which of the following statement is incorrect ?

- A. Black solid substance is CuO .
- B. It is redox reaction,
- C. It is precipitation reaction.
- D. Copper undergo oxidation.

Answer:

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14. Silver chloride is stored in dark coloured bottle because...

- A. it is a white solid.
- B. it gives redox reaction.
- C. to avoid the effect of sunlight.
- D. None of these

Answer:



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15. On immersing the Zn rod in the solution of copper sulphate, you will observe ...

- A. deposition of Cu on Zn
- B. deposition of Zn on Cu.

C. Cu^{2+} oxidises.

D. blue coloured solution become more dark.

Answer:



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16. The reaction of H_2 gas with oxygen gas forms water. This reaction is an example of:

A. combination reaction

B. redox reaction

C. an exothermic reaction

D. all of these reactions

Answer:



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17. For the reaction, $CuO + H_2 \rightarrow Cu + H_2O$, choose the correct statement.

- A. CuO is an oxidising agent.
- B. H_2 , undergo oxidation.
- C. It is a displacement reaction.
- D. All of the given

Answer:



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18. Select the proper option for the following statements :

Statement 1: Burning of magnesium ribbon in air is a redox

reaction.

Statement 2: Oxidation number of oxygen in its metal oxide is -1.

- A. Statement 1 is correct.
- B. Statement 2 is incorrect.
- C. Statement 1 is correct, but statement 2 is incorrect.
- D. Statement 1 and statement 2 both are incorrect.

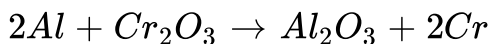
Answer:



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Objective Questions And Answers Choose More Than One Correct Options From Those Given Below Each Question

1. Identify the type of reaction for



(i) Oxidation (ii) Reduction (iii) Redox

A. Only (i)

B. Only (ii)

C. (i), (ii), (iii)

D. None of these

Answer:



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2. What is correct for redox reaction ?

(i) Reducing agent undergoes oxidation.

(ii) Reducing agent undergoes reduction.

(iii) Reduction reaction always occurs at cathode.

(iv) Oxidising agent undergo oxidation.

A. Only (i), (ii)

B. Only (i), (iii)

C. Only (iii)

D. Given all

Answer:



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3. Quick lime when added in water produces hissing sound. State the type of reaction.

(i) Combination (ii) Endothermic

(iii) Exothermic (iv) Redox

A. Only (i),

B. Only (iii)

C. Only (i), (ii), (iii)

D. Only (iv)

Answer:

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Objective Questions And Answers Answer The Following Questions In One Word

1. Write the formula of rust.

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2. Name the reaction observed during rancidity of food.

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3. Write the formula of lead sulphate.

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4. What is obtained by the combustion of methane?

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5. What is called the insoluble substance formed during the chemical reaction ?

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6. Name two metals which do not corrode.

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7. Which gas burns with popping sound ?

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8. Name the product formed, when silver bromide is exposed to sunlight.

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9. Which compound is used to detect the formation of carbon dioxide gas ?

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10. Name the gas evolved when lead nitrate is heated.



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11. Write the example of decomposition reaction which occurs in nature.



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12. Write the name of chemical used in black and white photography.



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13. Which ions are present in barium sulphate ?



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14. Mention one use of quick lime.



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15. Give one example of decomposition reaction in which solid and gas are formed as products.



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16. Give an example of antioxidant.



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17. Write the reaction in which hydrogen acts as a reducing agent.



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18. Name the reaction in which two compounds exchange their ions to form two new compounds.

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19. What would be the effect on lime water, when carbon dioxide gas is passed through it ?

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20. State the molecular formula and chemical name of iron (III) oxide.

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21. State different forms of energy required for breaking down the molecules of reactant in decomposition reaction.

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22. What does (II) indicate in iron (II) oxide ?

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23. Give two examples of noble metals.

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24. Mention two examples of exothermic reactions.

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25. When does the breaking and forming of bonds of chemical substances take place ?

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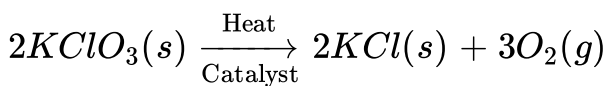
26. Name the reaction which forms insoluble salts.

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27. Write the formula for two oxides of sulphur.

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28. Identify the type of reaction for the following reaction :





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29. On immersing an iron nail in $CuSO_4$ solution for few minutes, what will you observe ?



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Objective Questions And Answers Match The Following

1. Match the following:

Column I (Type)	Column II (Reactions)
1. Neutralisation reaction	a. $Zn + Cu^{2+} \rightarrow Zn^{2+} + Cu$ b. $CaCO_3 \rightarrow CaO + CO_2$
2. Redox reaction	c. $HCl + NaOH \rightarrow NaCl + H_2O$



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2. Match the following:

Column I (Name)	Column II (Molecular Formula)
1. Quick lime	a. CaCO_3
2. Slaked lime	b. CaO
	c. Ca(OH)_2



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3. Match the following:

Column I [Colour]	Column II (Substance)
1. Green	a. Copper sulphate
2. Blue	b. Barium sulphate
	c. Ferrous sulphate



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4. Match the following:

Column I (Reaction)	Column II (Type)
1. $\text{CaO} + \text{H}_2\text{O} \rightarrow$ $\text{Ca(OH)}_2 + \text{heat}$	a. Displacement
2. $\text{Zn} + \text{CuSO}_4 \rightarrow$ $\text{ZnSO}_4 + \text{Cu}$	b. Double displacement
	c. Combination



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Objective Questions And Answers Complete The Following Chemical Reactions

1. Complete the chemical reaction



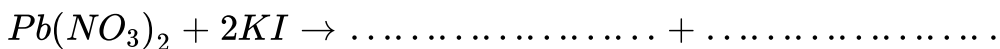
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2. Complete the chemical reaction



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3. Complete the chemical reaction



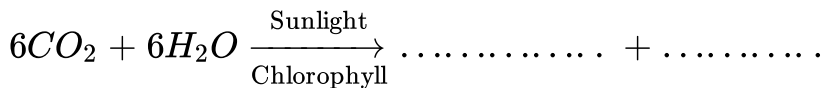
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4. Complete the chemical reaction



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5. Complete the chemical reaction



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6. Complete the chemical reaction



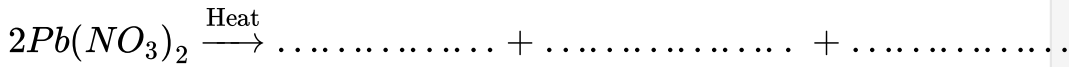
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7. Complete the chemical reaction



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8. Complete the chemical reaction



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9. Complete the chemical reaction



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10. Complete the chemical reaction



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11. Complete the chemical reaction



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12. Complete the chemical reaction



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13. Complete the chemical reaction



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14. Complete the chemical reaction



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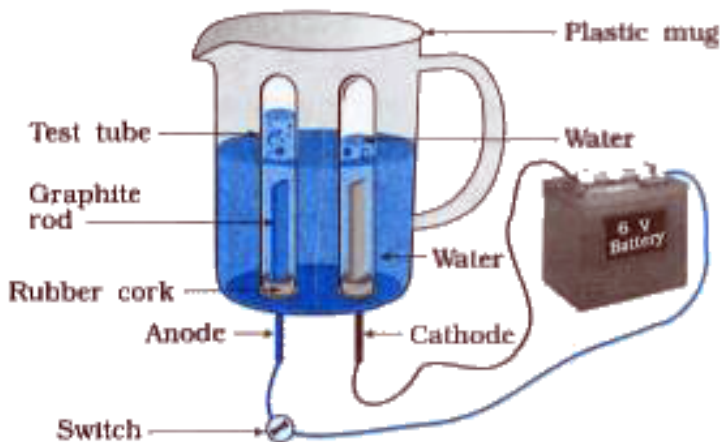
15. Complete the chemical reaction



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Objective Questions And Answers Carefully Observe The Given Diagram And Answer The Questions Related With It

1. Which gas is collected in a test tube at anode as shown in figure ?



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2. Does the reaction occur in the test tube as shown in the figure ?

State the change in colour.



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3. Take a small amount of ammonium chloride in a beaker as shown in figure. Add water to it. What change do you observe in

temperature ? Which type of reaction is this ?



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4. As shown in figure, the strong smell of which gas will the student experience ?



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Value Based Questions With Answers

1. Hitarthi saw her grandmother storing pickles in ceramic pots. She also knew that pickles should not be stored in metal containers. Her friend often brought pickles wrapped in

aluminium foil. Hitarthi advised her not to wrap pickles in foil.

Why pickles should not be stored in metal containers ?



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2. Hitarthi saw her grandmother storing pickles in ceramic pots.

She also knew that pickles should not be stored in metal containers. Her friend often brought pickles wrapped in aluminium foil. Hitarthi advised her not to wrap pickles in foil.

Which substances present in pickles reacts with the metal containers ?



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3. Hitarthi saw her grandmother storing pickles in ceramic pots.

She also knew that pickles should not be stored in metal containers. Her friend often brought pickles wrapped in

aluminium foil. Hitarthi advised her not to wrap pickles in foil.

What value of Hitarthi is seen in the above act ?



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4. Surekha was very upset as her silver jewellery had turned black, and lost its lustre. Her father is a science teacher and he washed and cleaned the jewellery using toothpaste and brought the shine back.

Why does the silver jewellery tarnishes when left open ?



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5. Surekha was very upset as her silver jewellery had turned black, and lost its lustre. Her father is a science teacher and he washed and cleaned the jewellery using toothpaste and brought the shine

back.

How had the toothpaste got the shine of silver back ?



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6. Surekha was very upset as her silver jewellery had turned black, and lost its lustre. Her father is a science teacher and he washed and cleaned the jewellery using toothpaste and brought the shine back.

What value of Surekha's father is seen in this act ?



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7. Rakesh visited the government hospital to meet his cousin. He observed the medicine in dark bottles were not stored properly. They were not kept away from light and heat. Rakesh immediately reported the same to the medical superintendent and made sure

that all medicines are stored properly.

Why are some medicines stored at cool places in dark bottles ?



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8. Rakesh visited the government hospital to meet his cousin. He observed the medicine in dark bottles were not stored properly. They were not kept away from light and heat. Rakesh immediately reported the same to the medical superintendent and made sure that all medicines are stored properly.

Why do some medicines are kept in refrigerator ?



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9. Rakesh visited the government hospital to meet his cousin. He observed the medicine in dark bottles were not stored properly. They were not kept away from light and heat. Rakesh immediately

reported the same to the medical superintendent and made sure that all medicines are stored properly.

What value of Rakesh is seen in this act ?

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Practical Skill Based Questions With Answers

1. What would happen if you add zinc coated iron nail into the solution of copper sulphate ? Give reason for your prediction.

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2. How will you differentiate sodium metal and zinc metal given in the test tubes in the laboratory ? You are advised not to touch any of the metals. Identify the type of reaction used.

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3. State the example of the combination reaction that is also exothermic by nature. How will you show its exothermic nature in the laboratory ?

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4. The teacher wishes to show in laboratory that all the crystals will decompose to give water on heating, which is collected on the upper surface inside the test tube. Name any four compounds the teacher should use in the laboratory to perform this activity.

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5. A student wants to study the decomposition reaction of iron sulphate in laboratory. What care the student should take and why?



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