

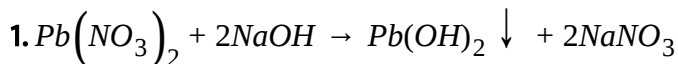


CHEMISTRY

BOOKS - VK JAISWAL CHEMISTRY (HINGLISH)

TYPES OF REACTIONS

LEVEL 1



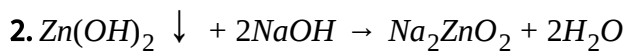
- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: A





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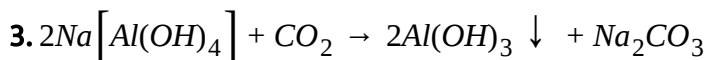


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: B



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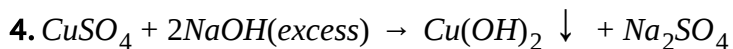


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: a

 [Watch Video Solution](#)



A. For precipitate formation reaction

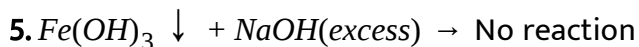
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: A

 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

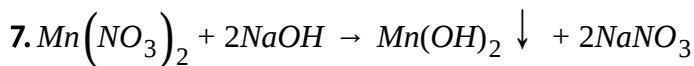
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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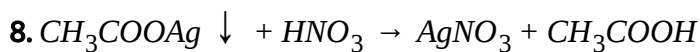


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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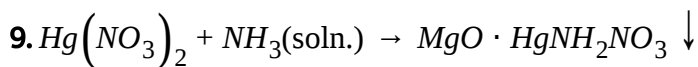


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

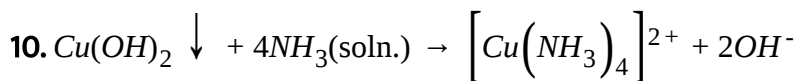
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

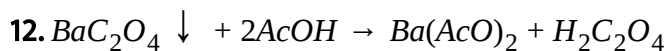
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11. $\text{CaC}_2\text{O}_4 \downarrow + \text{CH}_3\text{COOH} \rightarrow$ No reaction

- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b



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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

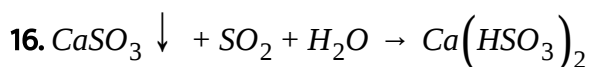
 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

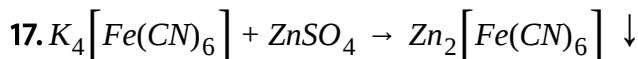
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)

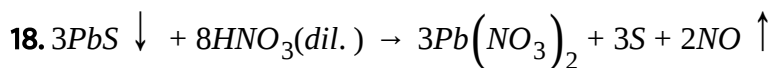


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



Watch Video Solution



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

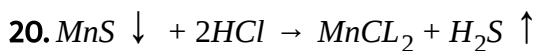
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

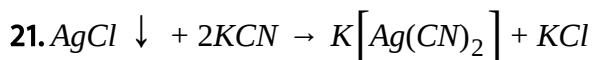
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

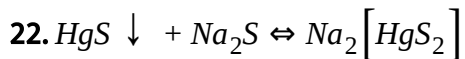
 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)

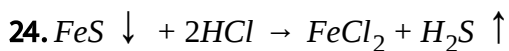


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: a

 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

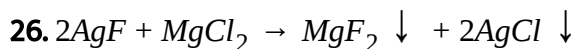
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

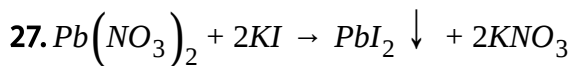
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

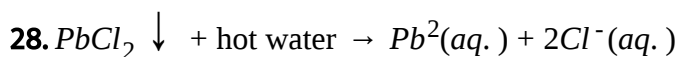


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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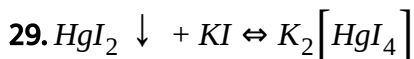


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

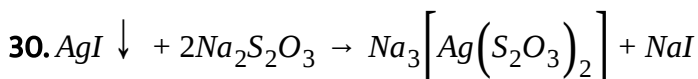
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

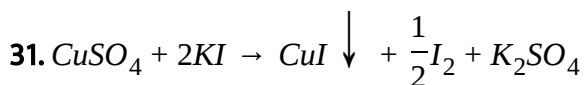
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)



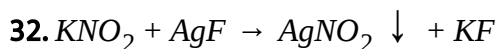
- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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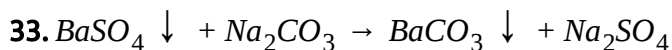


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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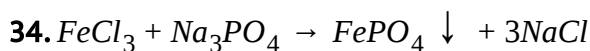


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: c

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A. For precipitate formation reaction

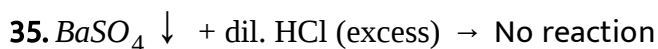
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: a

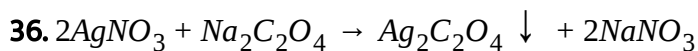
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

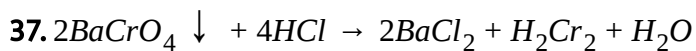
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

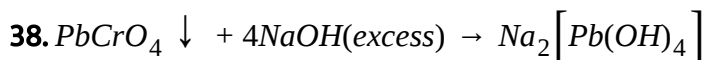
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

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39. $BaCrO_4 \downarrow + CH_3COOH(excess) \rightarrow$ No reaction

- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

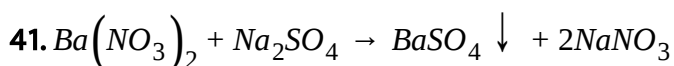
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40. $PbCl_2 \downarrow + H_2SO_4 \rightleftharpoons PbSO_4 \downarrow + 2HCl$

- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: c

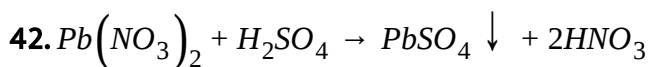
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

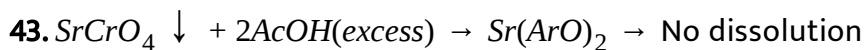
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

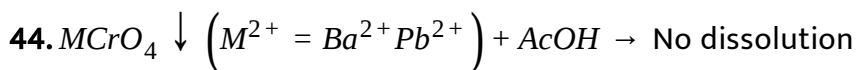


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

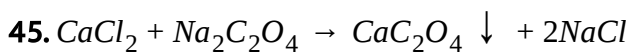
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

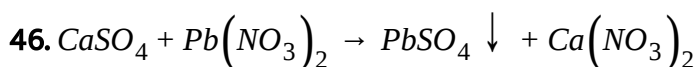
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

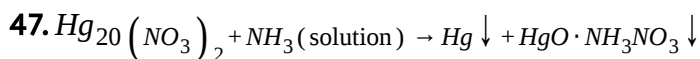
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

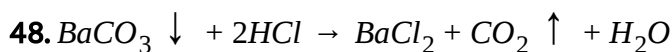
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

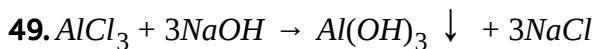


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: b

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A. For precipitate formation formation reaction

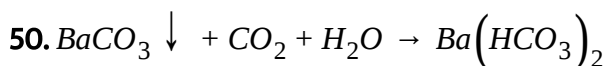
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: a

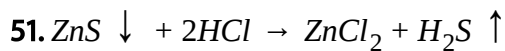
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

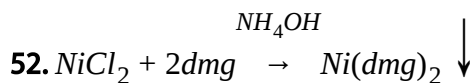
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

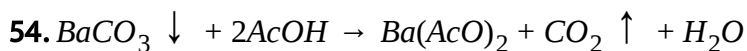


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: d

 [Watch Video Solution](#)



A. For precipitate formation reaction

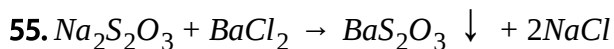
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: b

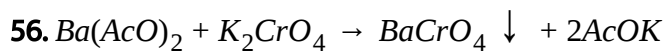
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

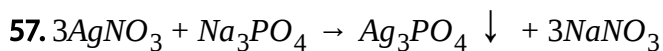
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

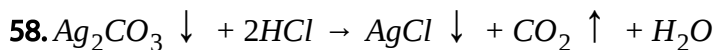


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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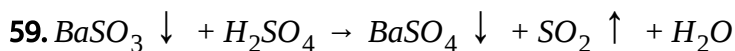


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: c

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A. For precipitate formation formation reaction

B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: c

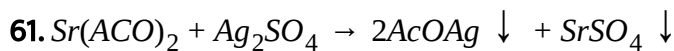
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

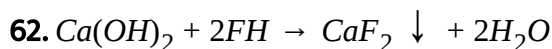
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

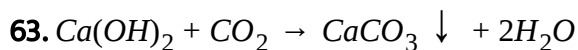
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

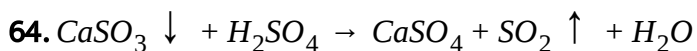


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: a

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A. For precipitate formation reaction

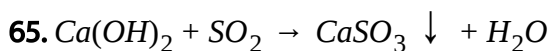
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: b

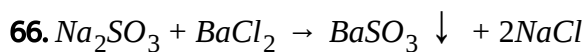
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

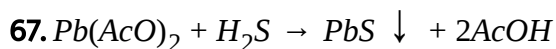
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

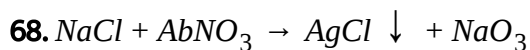
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)



- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: a

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A. For precipitate formation reaction

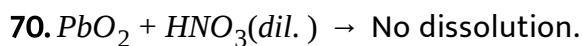
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: b

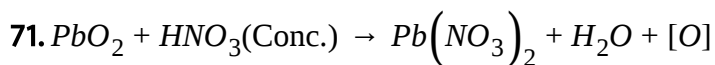
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

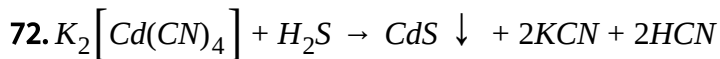
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

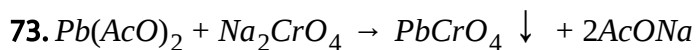
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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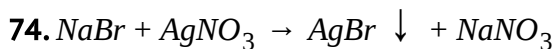


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction

D. For no reaction

Answer: a

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A. For precipitate formation reaction

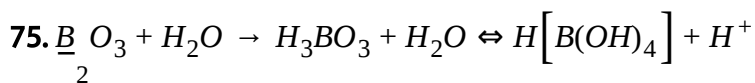
B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: a

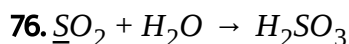
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

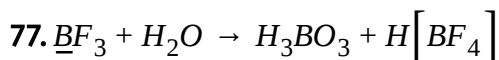
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

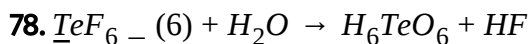
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

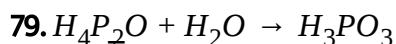
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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80. $\underline{C}O + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



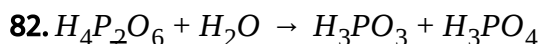
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81. $\underline{S}O_3 + H_2O \rightarrow H_2SO_4$

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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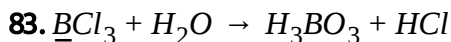


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c



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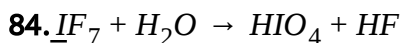


- A. If product is oxy acid with -ic suffix.
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- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



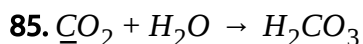
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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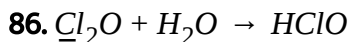


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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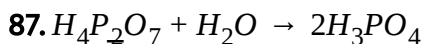


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

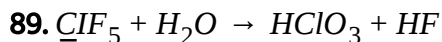
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88. $\underline{\text{C}}\text{Cl}_4 + \text{H}_2\text{O} \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

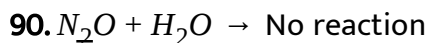
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

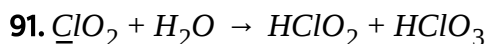
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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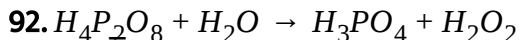


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c



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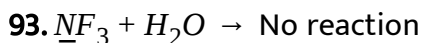


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



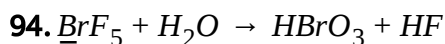
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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95. $\underline{N}O + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



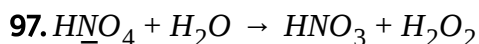
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96. $\underline{C}lO_3 + H_2O \rightarrow HClO_3 + HClO_4$

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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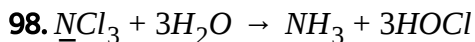


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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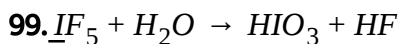


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



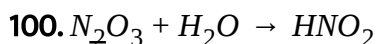
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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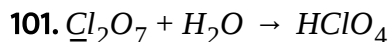


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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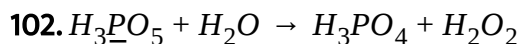


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



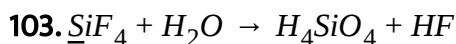
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- A. If product is oxy acid with -ic suffix.
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- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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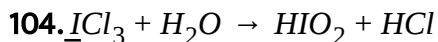


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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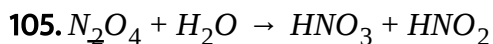


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- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



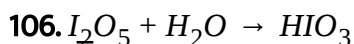
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c

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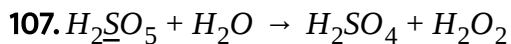


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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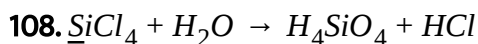


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



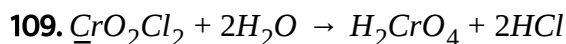
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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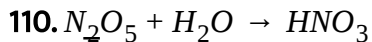


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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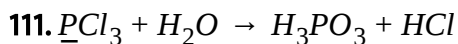


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



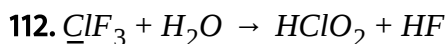
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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113. $\underline{\text{SiO}}_2 + \text{H}_2\text{O} \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



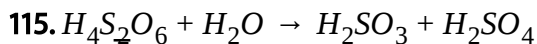
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114. $\text{H}_4\text{B}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow 2\text{H}_3\text{BO}_3$

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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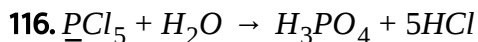


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c



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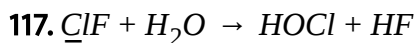


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



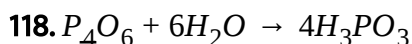
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

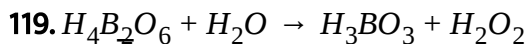
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

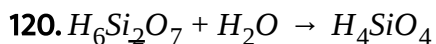
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

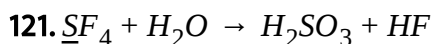
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

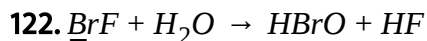
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

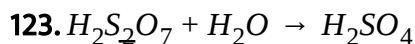
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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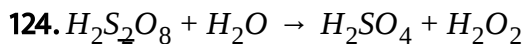


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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125. $\underline{S}F_6 + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

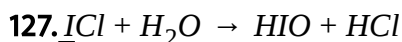
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126. $Icl + H_2O \rightarrow HIO + HCl$

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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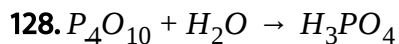


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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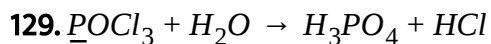


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



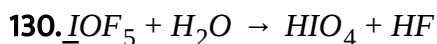
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

 [Watch Video Solution](#)

131. $P_4 + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

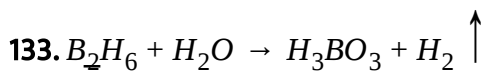
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132. $Na\underline{H} + H_2O \rightarrow NaOH + H_2 \uparrow$

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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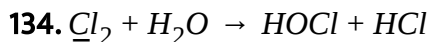


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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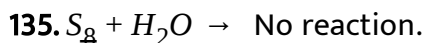


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



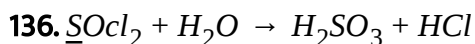
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- A. If product is oxy acid with -ic suffix.
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- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

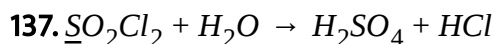
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

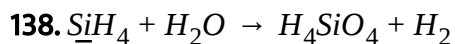
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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139. $I_2 + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

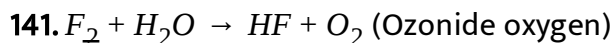
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

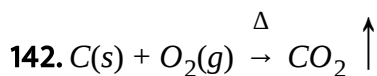
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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A. For disproportionation reaction.

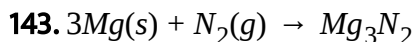
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

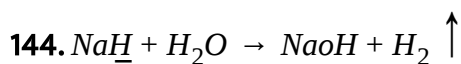
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

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A. For disproportionation reaction.

B. For comproportionation reaction.

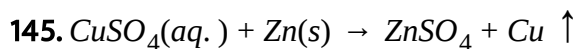
C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: bc



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A. For disproportionation reaction.

B. For comproportionation reaction.

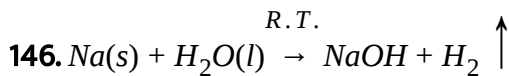
C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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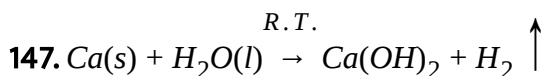


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



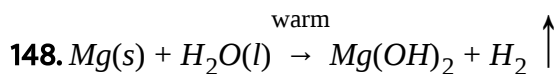
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

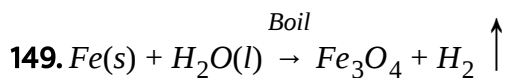
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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- A. For disproportionation reaction.
- B. For comproportionation reaction.
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- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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- A. For disproportionation reaction.
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- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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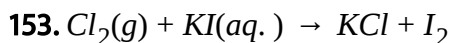


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



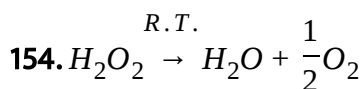
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
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- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

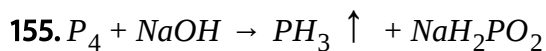
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

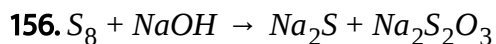
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

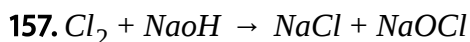
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

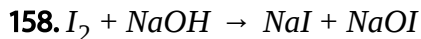
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

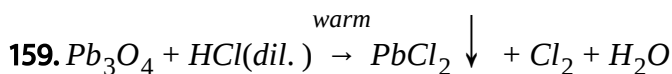
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

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A. For disproportionation reaction.

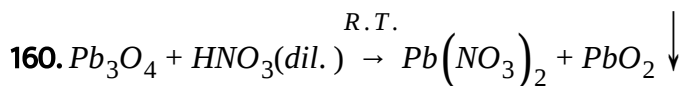
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

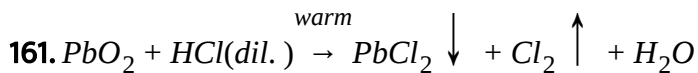
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

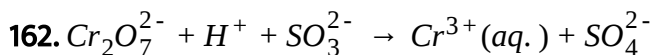
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

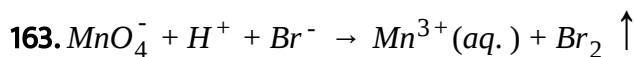
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

B. For comproportionation reaction.

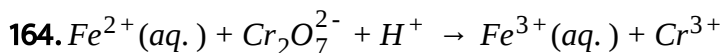
C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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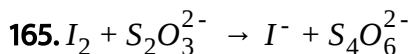


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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A. For disproportionation reaction.

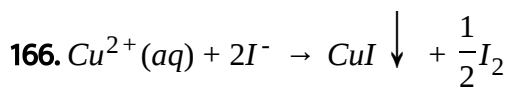
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

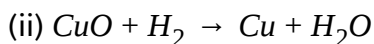
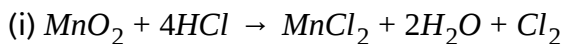
D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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167. What is redox reaction ? Identify the substance oxidised and the substance reduced in the following reactions :



A. For disproportionation reaction.

B. For comproportionation reaction.

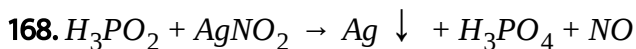
C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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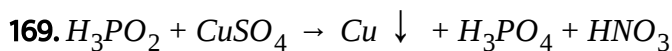


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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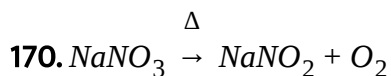


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

B. For comproportionation reaction.

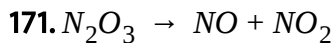
C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

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R. T.

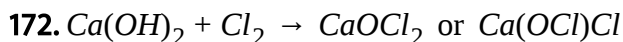


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad



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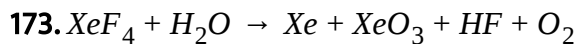


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

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A. For disproportionation reaction.

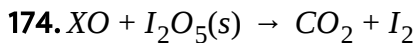
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

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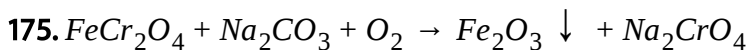


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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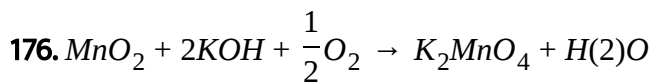


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)



A. For disproportionation reaction.

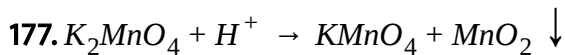
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

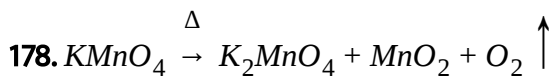
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

 [Watch Video Solution](#)



- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)



A. For disproportionation reaction.

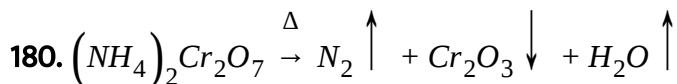
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

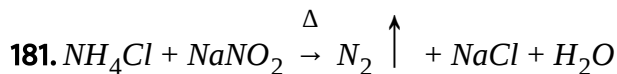
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

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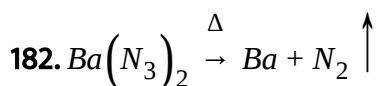


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: b

 [Watch Video Solution](#)



A. For disproportionation reaction.

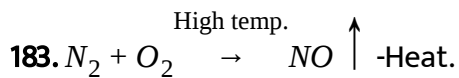
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: cd

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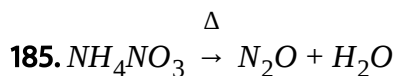


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)



A. For disproportionation reaction.

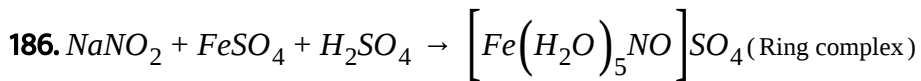
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: bd

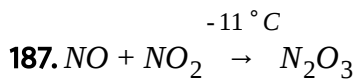
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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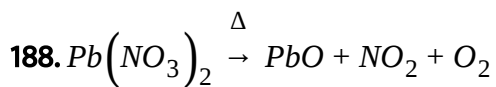


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: bd

 [Watch Video Solution](#)



A. For disproportionation reaction.

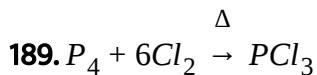
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)

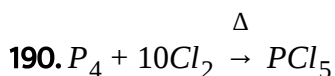


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: cd



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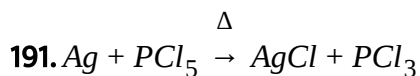


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: cd

 [Watch Video Solution](#)



A. For disproportionation reaction.

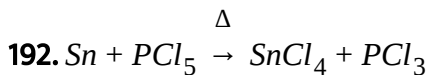
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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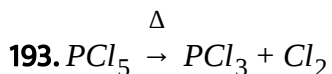


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)

194. Red P+Alkali $\rightarrow Na_4P_2O_6 + P_2H_4$

A. For disproportionation reaction.

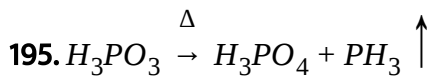
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

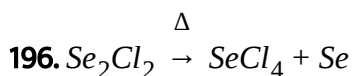
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

 [Watch Video Solution](#)

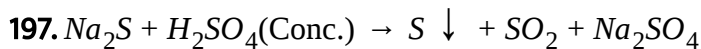


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

 [Watch Video Solution](#)



A. For disproportionation reaction.

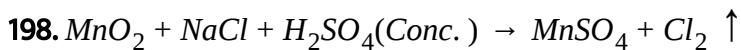
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

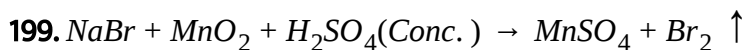
 [Watch Video Solution](#)



- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)

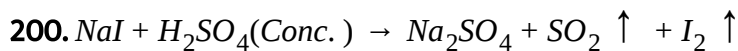


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)



A. For disproportionation reaction.

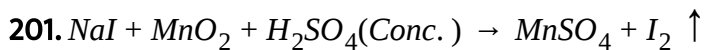
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

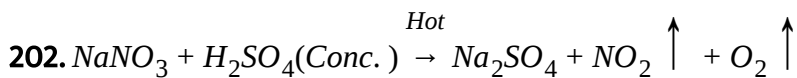
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)

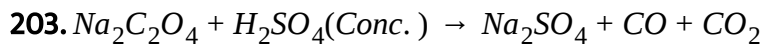


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)



A. For disproportionation reaction.

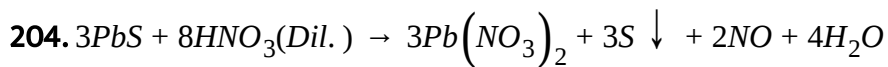
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

 [Watch Video Solution](#)

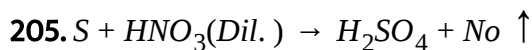


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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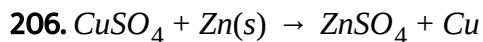


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)



A. For disproportionation reaction.

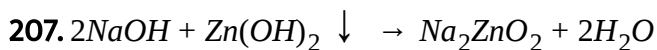
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

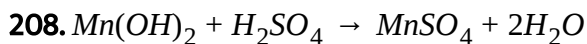
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c

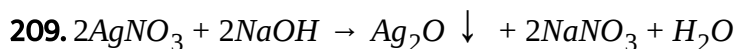
 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b

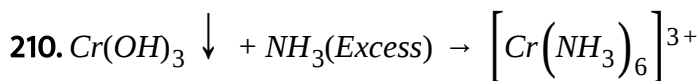
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt

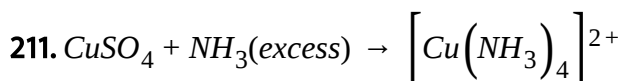
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

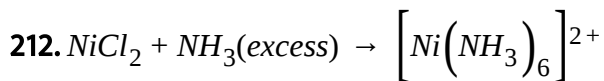
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b



Watch Video Solution



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a



Watch Video Solution



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c



Watch Video Solution



- A. For coloured ppt./Black ppt

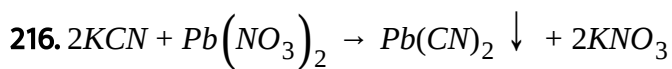
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

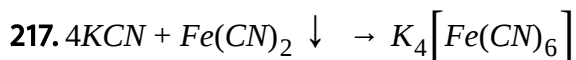
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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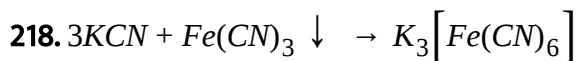


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b



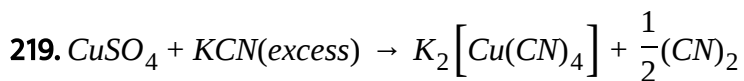
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b

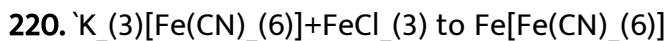
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt

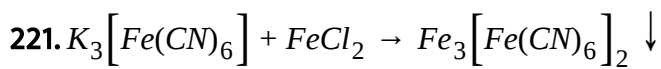
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

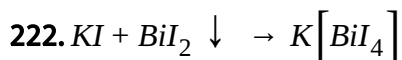
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: a

 [Watch Video Solution](#)

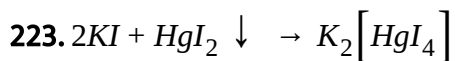


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b



Watch Video Solution



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c

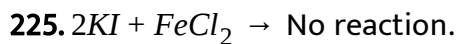
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt

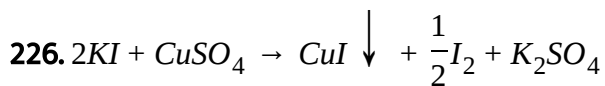
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

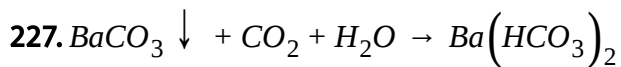
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

 [Watch Video Solution](#)

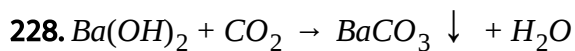


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c



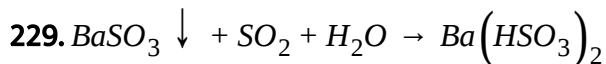
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: d

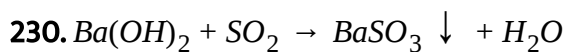
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt

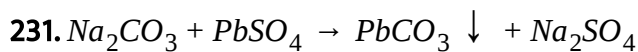
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

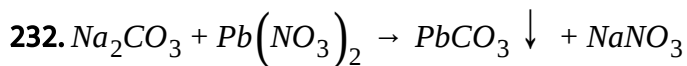
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: d



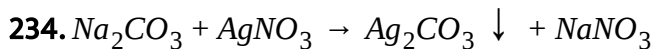
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c

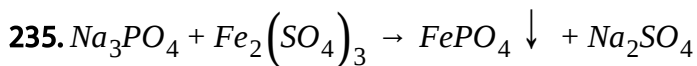
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: d

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- A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: a

 [Watch Video Solution](#)

236. $NiCl_2(\text{solution}) + NaNO_3(\text{solution}) \rightarrow$ No reaction

A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)

237. $\text{CuSO}_4(\text{solution}) + \text{ZnCl}_2(\text{solution}) \rightarrow$ No reaction

- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b



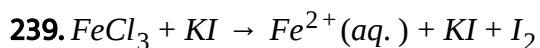
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238. $\text{FeSO}_4 + \text{Na}_2\text{S} \rightarrow \text{FeS} \downarrow$

- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a

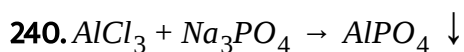
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b

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- A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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241. $CrCl_3(\text{solution}) + ZnSO_4(\text{Solution}) \rightarrow$ No reaction

A. For coloured ppt./Black ppt

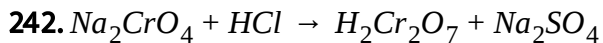
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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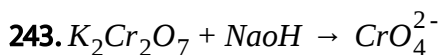


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b



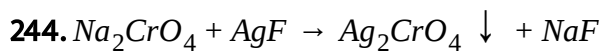
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b

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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a

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- A. For coloured ppt./Black ppt

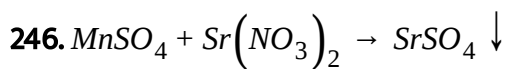
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

 [Watch Video Solution](#)

247. $ZnSO_4(\text{solution}) + MgCl_2(\text{solution}) \rightarrow$ No reaction

- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c



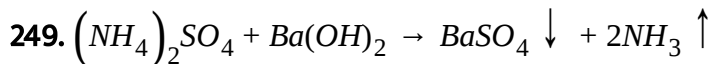
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248. $AgNO_3(\text{solution}) + NaF(\text{solution}) \rightarrow$ No reaction.

- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c

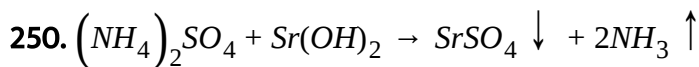
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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: d

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- A. For coloured ppt./Black ppt

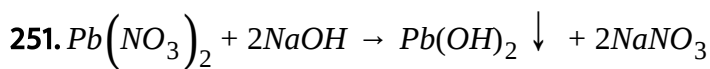
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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A. For precipitate formation reaction

B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: a

 [Watch Video Solution](#)

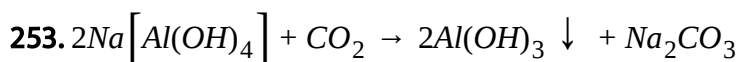


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b



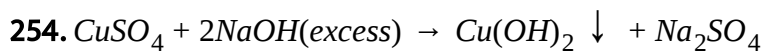
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

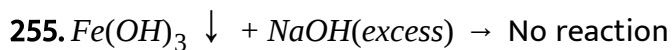
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)



- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

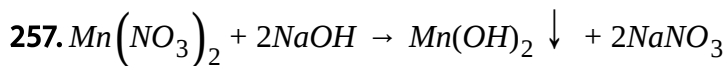
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)

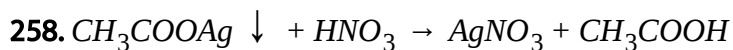


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



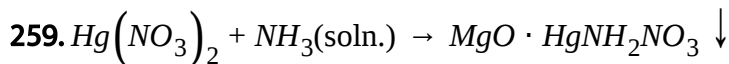
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

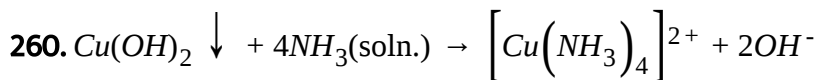
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)



- A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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261. $\text{CaC}_2\text{O}_4 \downarrow + \text{CH}_3\text{COOH} \rightarrow$ No reaction

- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

 [Watch Video Solution](#)



- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b



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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

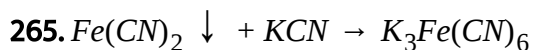
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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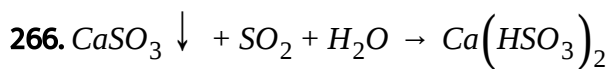


- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

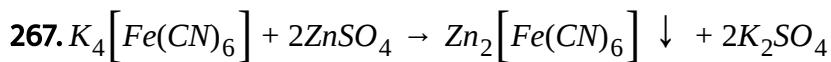
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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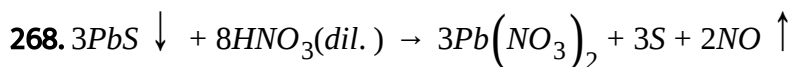


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



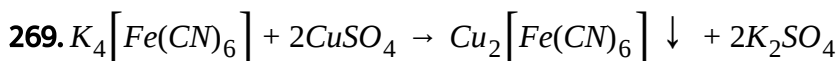
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)



- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

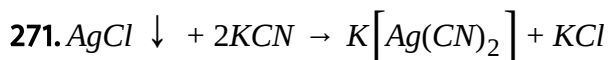


- A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

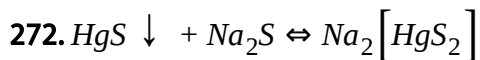
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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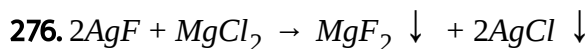


- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

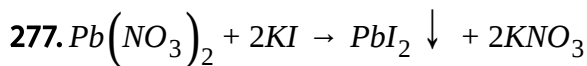
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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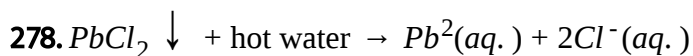


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



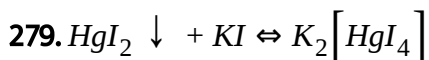
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

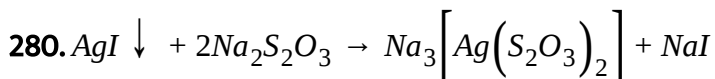
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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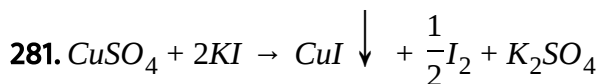


- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

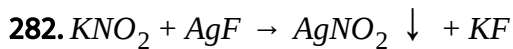
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

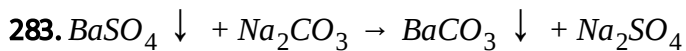
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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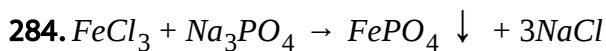


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: c



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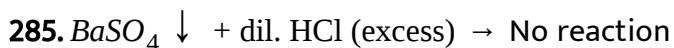


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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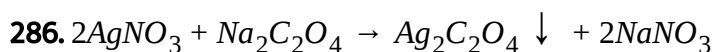


- A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

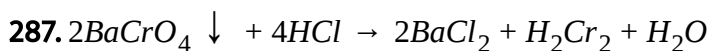
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

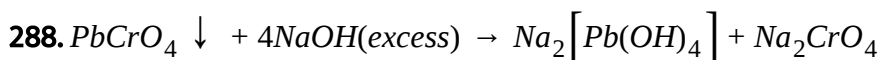
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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289. $BaCrO_4 \downarrow + CH_3COOH(excess) \rightarrow$ No reaction

A. For precipitate formation formation reaction

B. For precipitate dissolution reaction

C. For precipitate exchange reaction

D. For no reaction

Answer: d

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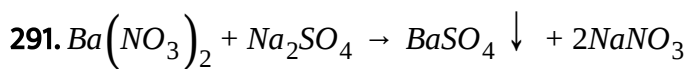
290. $PbCl_2 \downarrow + H_2SO_4 \Leftrightarrow PbSO_4 \downarrow + 2HCl$

A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: c

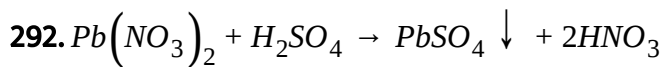
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

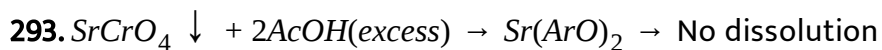
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

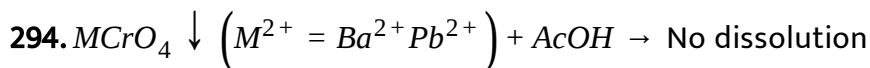
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

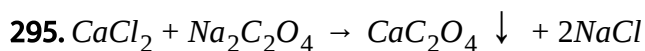
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

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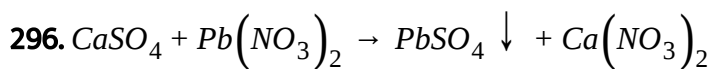


- A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

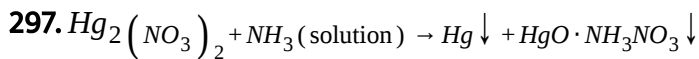
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

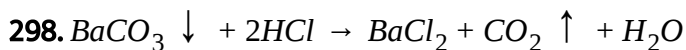


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



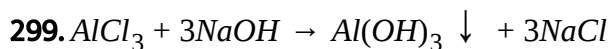
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

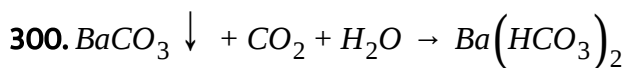
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)



- A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

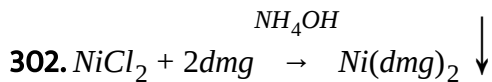
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

 [Watch Video Solution](#)



- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



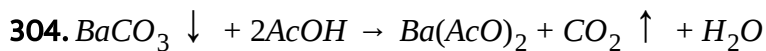
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

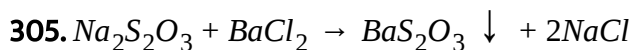
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

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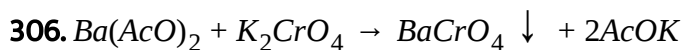


- A. For precipitate formation formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

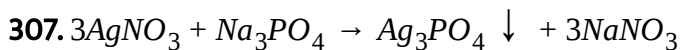
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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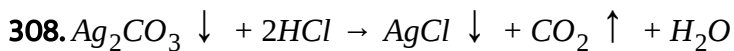


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



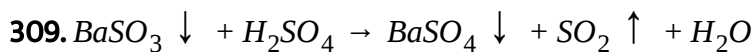
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: c

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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: c

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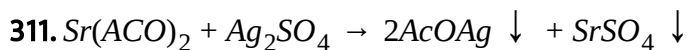


- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

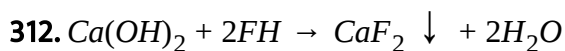
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

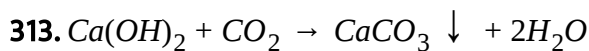
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

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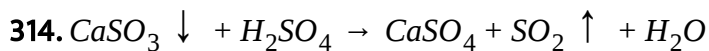


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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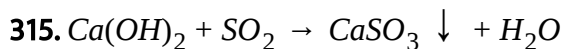


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b



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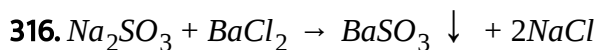


- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

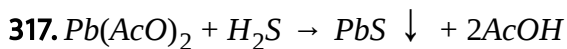
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

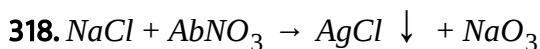


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



Watch Video Solution

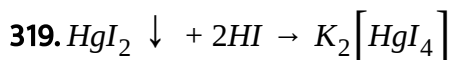


- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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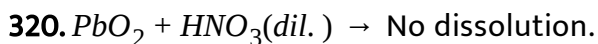


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b



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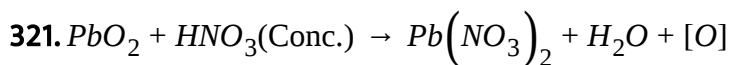


- A. For precipitate formation reaction

- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: d

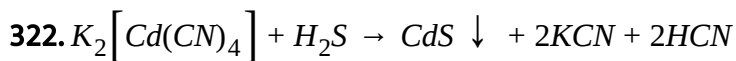
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- A. For precipitate formation formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: b

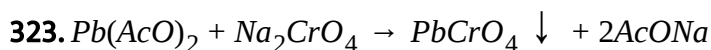
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- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a

 [Watch Video Solution](#)

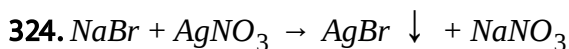


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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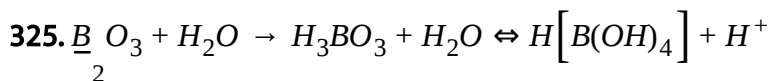


- A. For precipitate formation reaction
- B. For precipitate dissolution reaction
- C. For precipitate exchange reaction
- D. For no reaction

Answer: a



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- A. If product is oxy acid with -ic suffix.

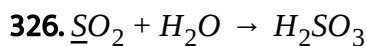
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

B. If product is oxy acid with -ous suffix

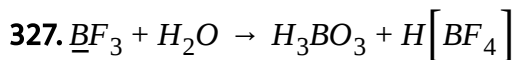
C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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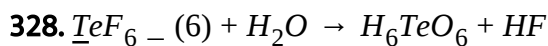


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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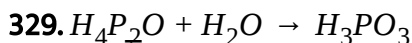
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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330. $\underline{C}O + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



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331. $\underline{S}O_3 + H_2O \rightarrow H_2SO_4$

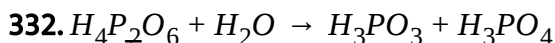
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

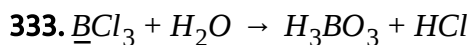
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c

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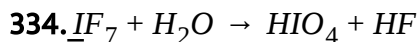


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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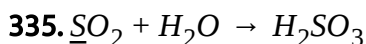
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

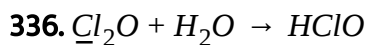
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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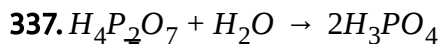


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b



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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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338. $\underline{C}Cl_4 + H_2O \rightarrow$ No reaction

A. If product is oxy acid with -ic suffix.

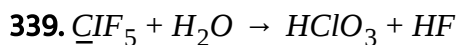
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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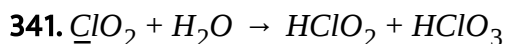
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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A. If product is oxy acid with -ic suffix.

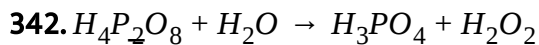
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c

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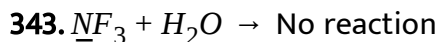


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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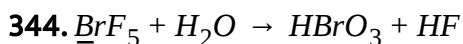
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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A. If product is oxy acid with -ic suffix.

B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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345. $\underline{N}O + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



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346. $\underline{C}lO_3 + H_2O \rightarrow HClO_3 + HClO_4$

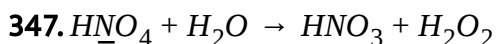
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

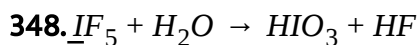
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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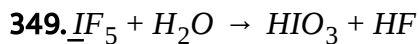


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



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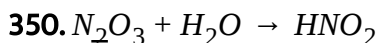
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

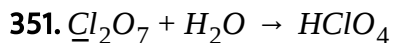
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

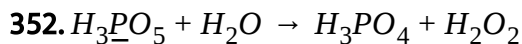
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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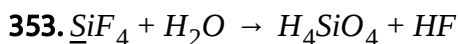
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

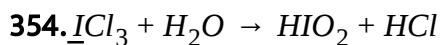
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

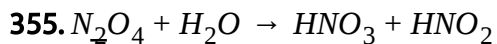
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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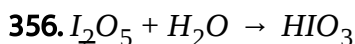
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c

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A. If product is oxy acid with -ic suffix.

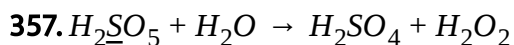
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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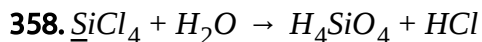


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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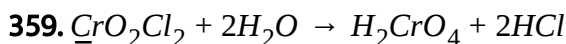
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

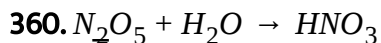
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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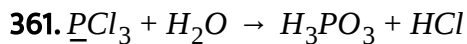


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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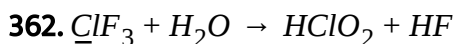
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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A. If product is oxy acid with -ic suffix.

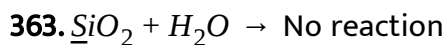
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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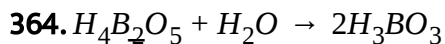


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



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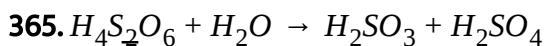
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

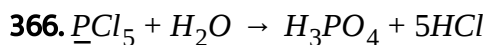
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c

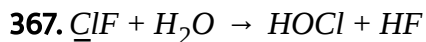
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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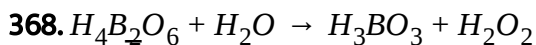
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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A. If product is oxy acid with -ic suffix.

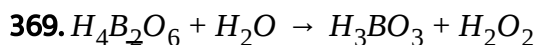
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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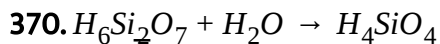


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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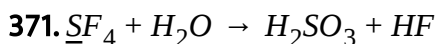
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

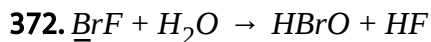
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

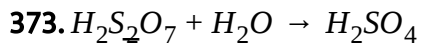
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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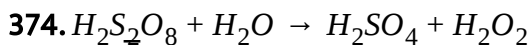
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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375. $\underline{S}F_6 + H_2O \rightarrow$ No reaction

- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



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376. $Icl + H_2O \rightarrow HIO + HCl$

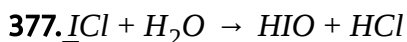
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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A. If product is oxy acid with -ic suffix.

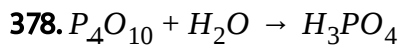
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: c

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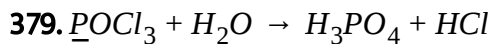


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a



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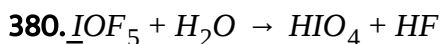
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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A. If product is oxy acid with -ic suffix.

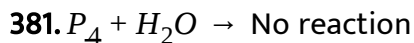
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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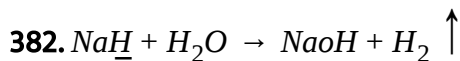


- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d



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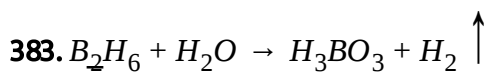
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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A. If product is oxy acid with -ic suffix.

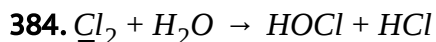
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

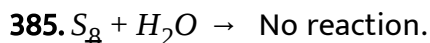
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

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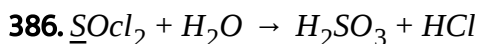
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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A. If product is oxy acid with -ic suffix.

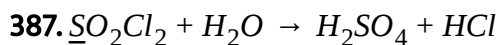
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: b

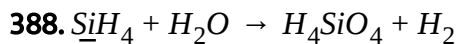
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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389. $I_2 + H_2O \rightarrow$ No reaction

A. If product is oxy acid with -ic suffix.

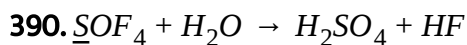
B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

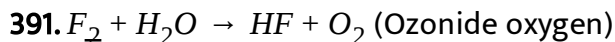
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- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix
- C. If product are two oxy acids one with -ic suffix and otherone with -ous suffix.
- D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: a

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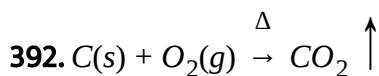
- A. If product is oxy acid with -ic suffix.
- B. If product is oxy acid with -ous suffix

C. If product are two oxy acids one with -ic suffix and other one with -ous suffix.

D. If product is not oxy acid, neither with -ic suffix nor with -ous suffix

Answer: d

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A. For disproportionation reaction.

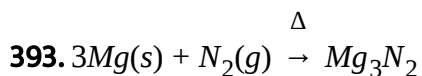
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

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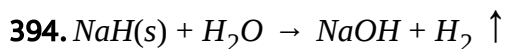


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d



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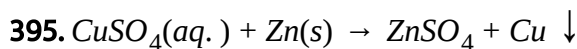
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: bc

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A. For disproportionation reaction.

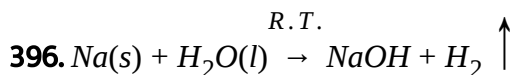
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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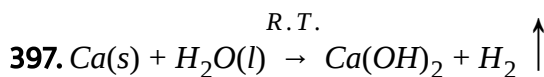


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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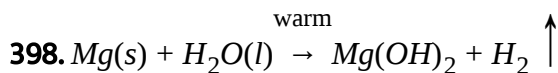
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

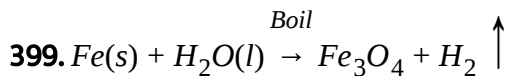
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

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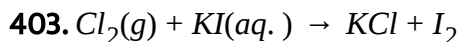


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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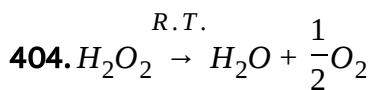
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

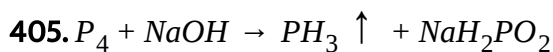
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

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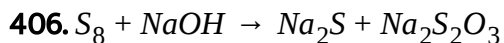


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a



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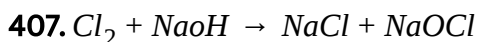
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

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A. For disproportionation reaction.

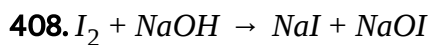
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

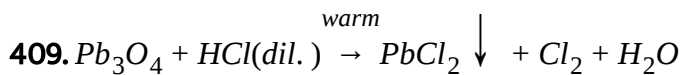
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

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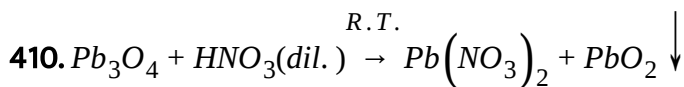
- A. For disproportionation reaction.
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C. For either intermolecular redox reaction or displacement reaction

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A. For disproportionation reaction.

B. For comproportionation reaction.

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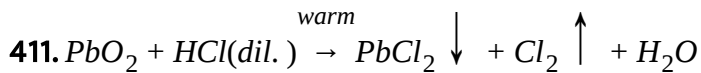
D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c





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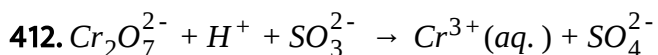


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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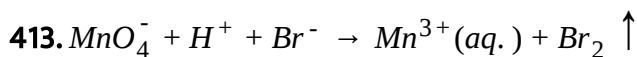
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

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A. For disproportionation reaction.

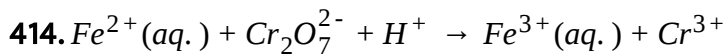
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

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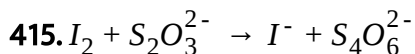


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
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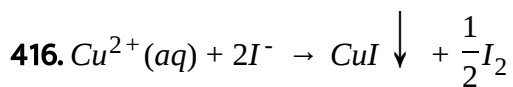
- A. For disproportionation reaction.
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Answer: c

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A. For disproportionation reaction.

B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

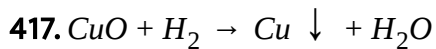
D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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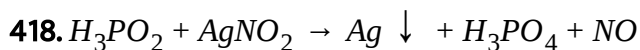


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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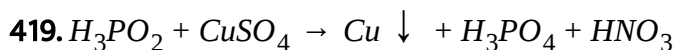
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)



A. For disproportionation reaction.

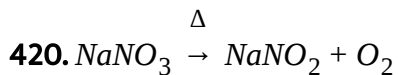
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

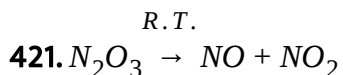
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)



- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

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422. $\text{Ca}(\text{OH})_2 + \text{Cl}_2 \rightarrow \text{CaOCl}_2$ or $\text{Ca}(\text{OCl})\text{Cl}$

A. For disproportionation reaction.

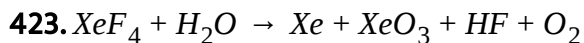
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

 [Watch Video Solution](#)

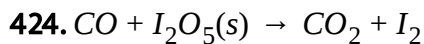


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a



Watch Video Solution



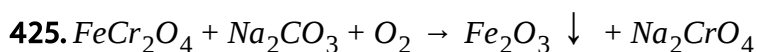
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)



A. For disproportionation reaction.

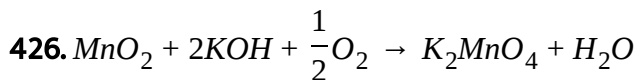
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)

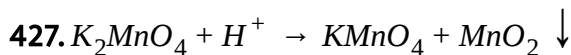


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c



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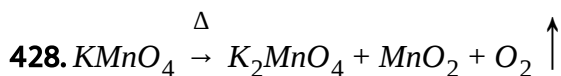
- A. For disproportionation reaction.
- B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

 [Watch Video Solution](#)



A. For disproportionation reaction.

B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

 [Watch Video Solution](#)

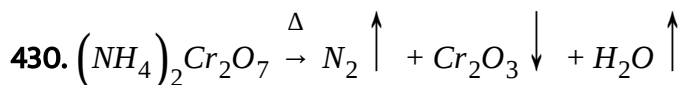


- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d



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- A. For disproportionation reaction.

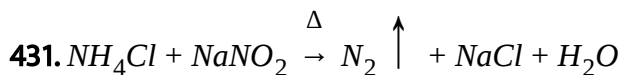
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

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A. For disproportionation reaction.

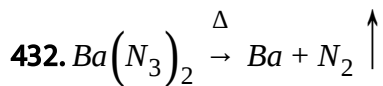
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: b

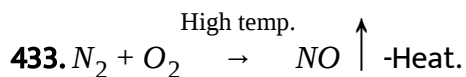
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

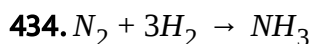
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: cd

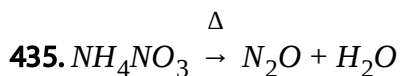
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

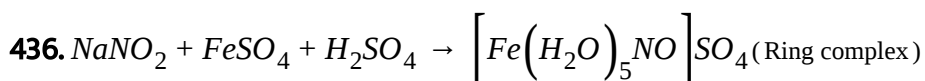
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: bd

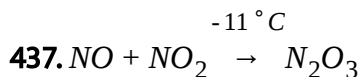
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

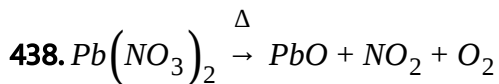
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: bd

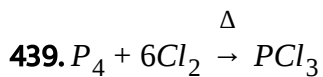
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

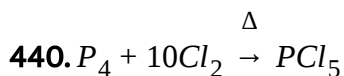
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: cd

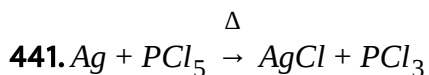
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: cd

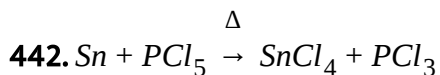
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

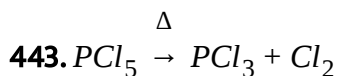
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

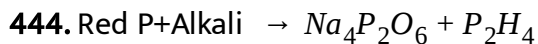
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

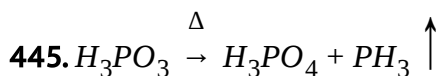
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

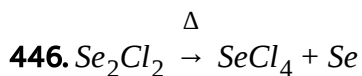
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

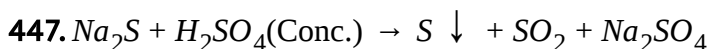
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: ad

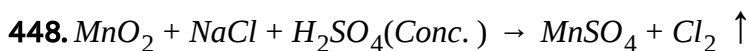
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

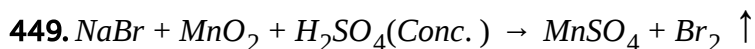
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

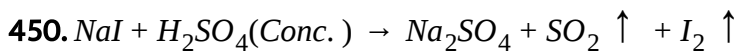
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

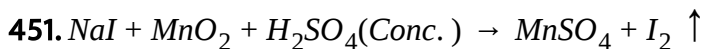
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For disproportionation reaction.

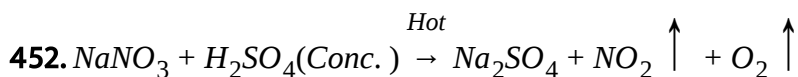
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

 [Watch Video Solution](#)



A. For disproportionation reaction.

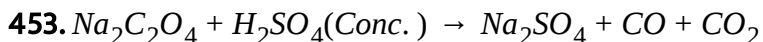
B. For comproportionation reaction.

C. For either intermolecular redox reaction or displacement reaction

D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: d

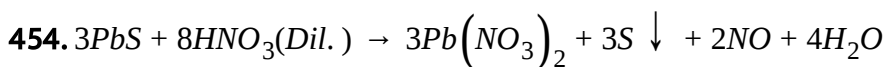
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: a

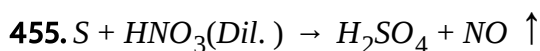
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

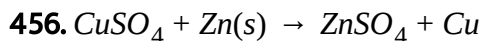
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

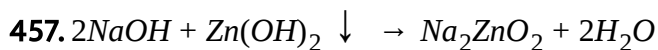
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- A. For disproportionation reaction.
- B. For comproportionation reaction.
- C. For either intermolecular redox reaction or displacement reaction
- D. For either thermal combination redox reaction or thermal decomposition redox reaction.

Answer: c

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A. For coloured ppt./Black ppt

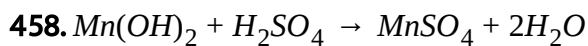
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: c

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

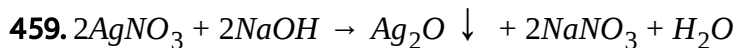
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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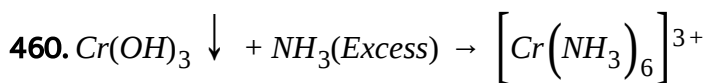


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a



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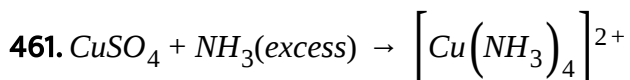


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

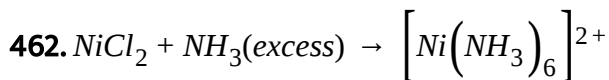
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

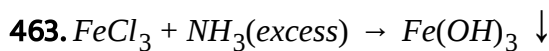
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: a

 [Watch Video Solution](#)



- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c



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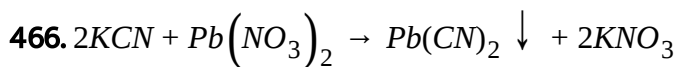


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

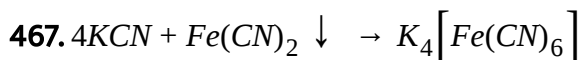
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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A. For coloured ppt./Black ppt

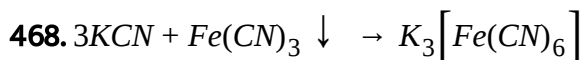
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

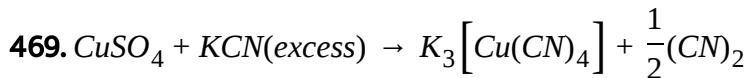
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c



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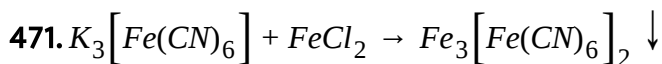


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

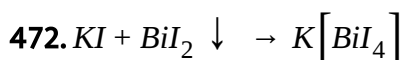
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: a

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A. For coloured ppt./Black ppt

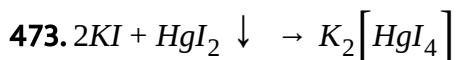
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: c

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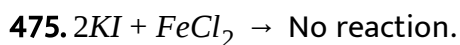


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a



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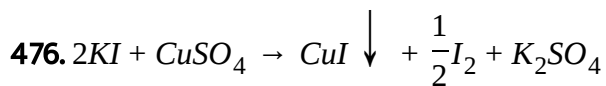


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

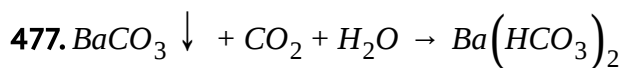
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

 Watch Video Solution



A. For coloured ppt./Black ppt

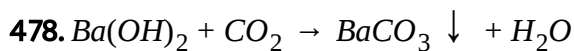
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: c

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

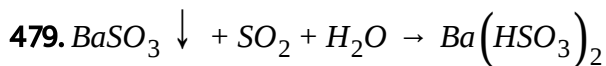
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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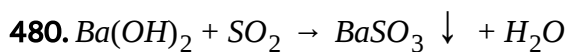


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: c



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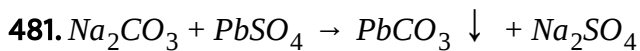


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: d

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

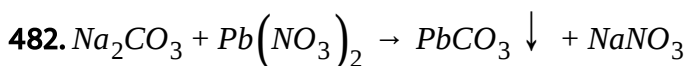
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

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483. $Na_2CO_3 + KNO_3 \rightarrow$ No reaction

A. For coloured ppt./Black ppt

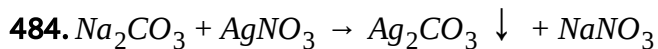
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: c

 [Watch Video Solution](#)



A. For coloured ppt./Black ppt

B. For coloured solution.

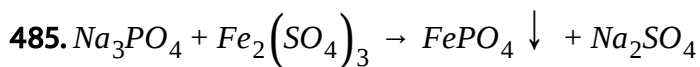
C. for clear/colourless solution

D. For white ppt.

Answer: d



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A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: a

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486. $NiCl_2(\text{solution}) + NaNO_3(\text{solution}) \rightarrow$ No reaction

A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

 [Watch Video Solution](#)

487. $CuSO_4(\text{solution}) + ZnCl_2(\text{solution}) \rightarrow$ No reaction

A. For coloured ppt./Black ppt

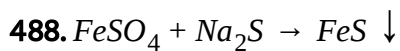
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

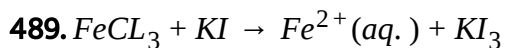
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: a

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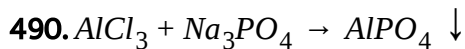


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: b



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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: d

 [Watch Video Solution](#)

491. $CrCl_3(\text{solution}) + ZnSO_4(\text{Solution}) \rightarrow$ No reaction

A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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492. $Na_2CrO_4 + HCl \rightarrow H_2Cr_2O_7 + Na_2SO_4$

A. For coloured ppt./Black ppt

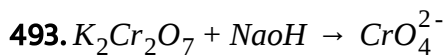
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

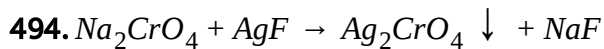
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: b

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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: a



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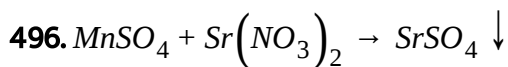


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

D. For white ppt.

Answer: b

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A. For coloured ppt./Black ppt

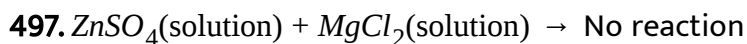
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: d

 [View Text Solution](#)



A. For coloured ppt./Black ppt

B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: c

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498. $AgNO_3(\text{solution}) + NaF(\text{solution}) \rightarrow$ No reaction.

A. For coloured ppt./Black ppt

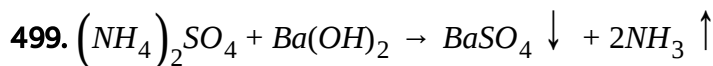
B. For coloured solution.

C. for clear/colourless solution

D. For white ppt.

Answer: c

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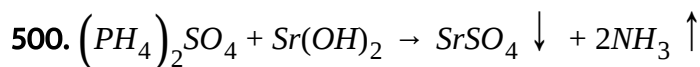


- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution
- D. For white ppt.

Answer: d



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- A. For coloured ppt./Black ppt
- B. For coloured solution.
- C. for clear/colourless solution

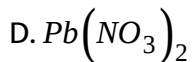
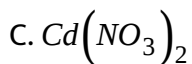
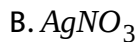
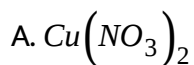
D. For white ppt.

Answer: d

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LEVEL 2

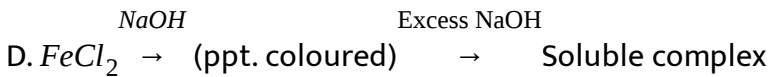
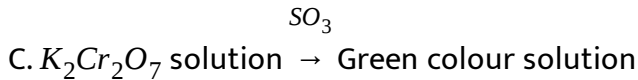
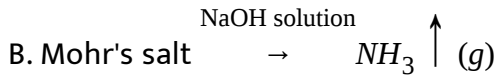
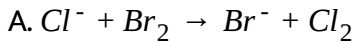
1. Which of the following metal nitrate produces gaseous product when reacts with KCN solution?



Answer: A

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2. Which of these reaction is correct?

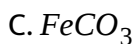
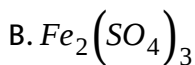
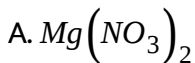


Answer: B



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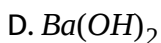
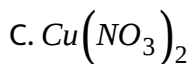
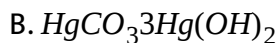
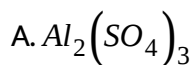
3. Compound which on heating produces paramagnetic acidic gas?



Answer: A

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4. Which compound on heating produces produces coloured metal oxide finally?

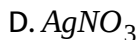
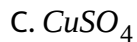
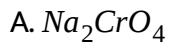


Answer: C

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5. P (Coloured solution) + $BaCl_2 \rightarrow Q \downarrow$ (White) + R (Coloured solution) ItBrgt

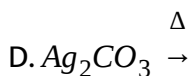
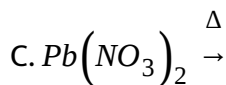
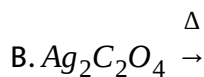
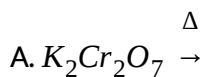
Then salt 'P' in above reaction is:



Answer: C

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6. Oxygen gas is not produced from the following decomposition reaction:

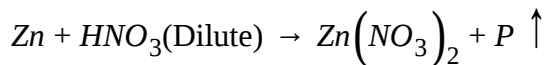


Answer: B



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7. Consider the following reaction and select incorrect statement about gas (P):



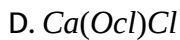
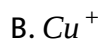
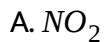
- A. Gives neutral solution in water
- B. Contains more O_2 than Air
- C. Forms brown ring with FeSO_4 solution
- D. None of these

Answer: C



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8. Which of the following ionic/molecular species does not disproportionate in water at room temperature?



Answer: D



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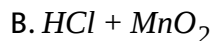
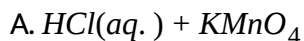
9. Which halogen oxidizes water at room temperature but does not undergo disproportionation into it?



Answer: A

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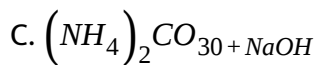
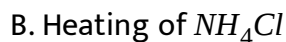
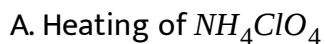
10. Which of the following combination doesn't evolve Cl_2 gas?

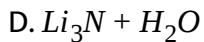


Answer: C

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11. Which of the following combination does not liberate NH_3 gas?



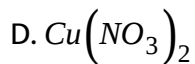
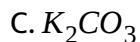
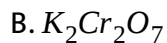
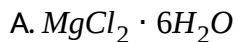


Answer: A



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12. Which of the following compound on heating does not produce metal oxide?

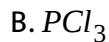


Answer: C



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13. Select the compound in which HCl is not the product of Hydrolysis:



Answer: A



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14. How many moles of H_2O are liberated when one mole hydrated MgCl_2 is heated?

A. 6

B. 5

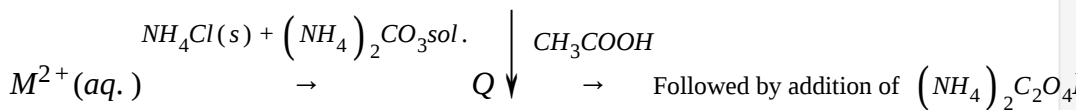
C. 4

D. 3

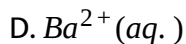
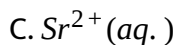
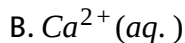
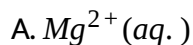
Answer: B

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15. Consider the following sequence of reaction:



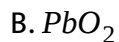
Which of the following cation can form ppt. Q but does not form ppt. 'R' ?



Answer: D

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16. Which of the following compound does not liberated oxygen gas on warming with conc. H_2SO_4 ?

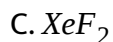


Answer: A



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17. One of the hydrolysed product of the following compound does not react with silica of glass vessel:



D. SF_4

Answer: A

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gas 'N' changes colour of $FeSO_4$ solution into yellow solution then salt M in above reaction is

A. BaS_2O_3

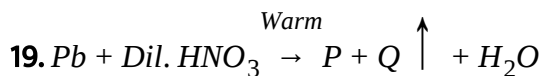
B. Ag_2SO_3

C. $AgNO_2$

D. $Pb(NO_3)_2$

Answer: C

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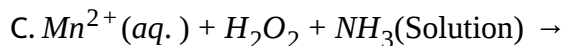
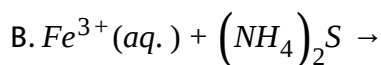
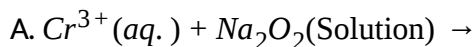
Incorrect statement for Q is:

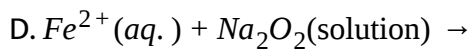
- A. Paramagnetic colourless gas
- B. It is oxidized to paramagnetic coloured gas by air
- C. It combines with $Fe_2(SO_4)_3$
- D. It can be also obtained by disproportionation of HNO_2

Answer: C

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20. In which of the following redox reaction precipitate is not formed?





Answer: A



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21. Which metal sulphide is soluble in excess NH_3 solution?

A. ZnS

B. MnS

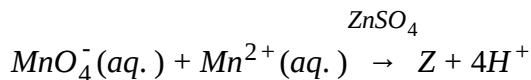
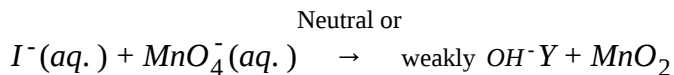
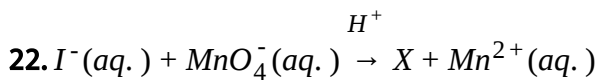
C. FeS

D. Cr_2S_3

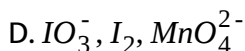
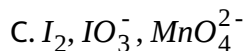
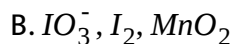
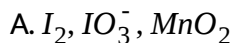
Answer: D



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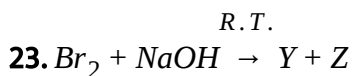


Product X, Y and Z are respectively.

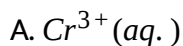


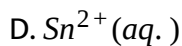
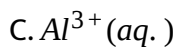
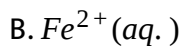
Answer: A

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If Y gives precipitate with $AgNO_3$, then Z does not undergo reaction with:



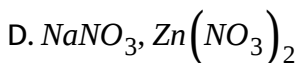
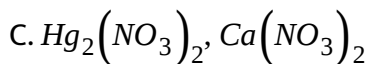
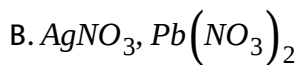
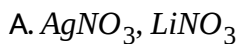


Answer: C

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P & X are respectively:



Answer: B



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25. Iodine is not oxidized to iodic acid/iodic anhydride by:

A. conc. HNO_3

B. conc. H_2SO_4

C. Excess Cl_2 water

D. conc. H_3PO_4

Answer: D



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26. Colourless gas that has oxidising as well as reducing properties:

A. CO_2

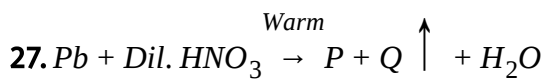
B. SO_2

C. NO_2

D. SO_3

Answer: B

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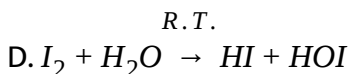
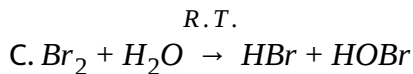
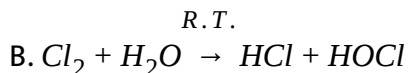
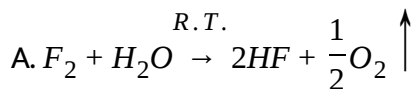
Incorrect statement for Q is:

- A. Paramagnetic colourless gas
- B. It is oxidized to paramagnetic coloured gas by air
- C. It combines with $Fe_2(SO_4)_3$
- D. It is also obtained by disproportionation of HNO_2

Answer: C

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28. Which reaction has positive value of ΔG° ?

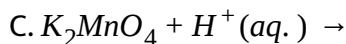
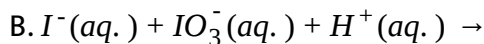
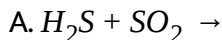


Answer: D



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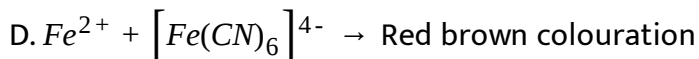
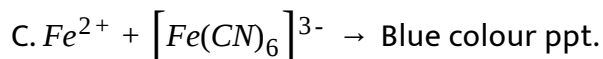
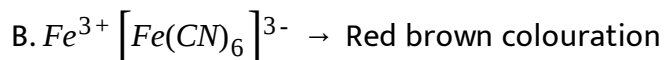
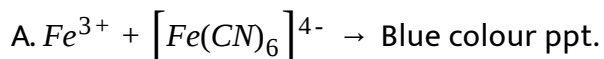
29. Which does not undergo comproportionation reaction?



Answer: C

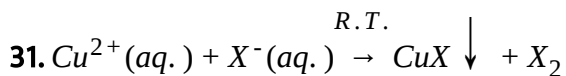
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30. Select the incorrect match:



Answer: D

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'X' cannot be:

A. $Cl^-(aq.)$

B. $I^-(aq.)$

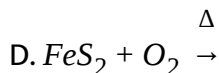
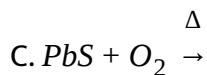
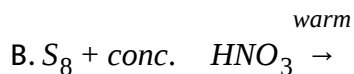
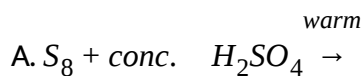
C. $CN^-(aq.)$

D. $SCN^-(aq.)$

Answer: A

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32. In which of the following reaction SO_2 gas is not produced?



Answer: B

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33. Which metal gives NH_4NO_3 , when react with dilute HNO_3 acid?

A. *Zn*

B. *Pb*

C. *Cu*

D. *Au*

Answer: A



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34. Select the salt whose aqueous solution is not green:

A. $FeSO_4$

B. $CrCl_3$

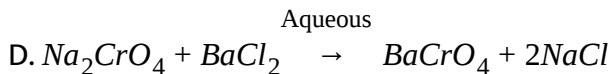
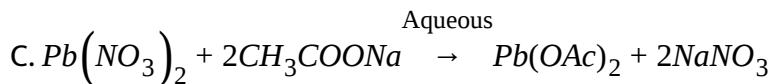
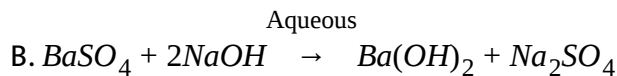
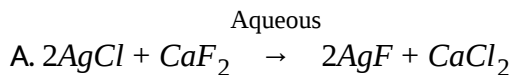
C. $NiCl_2$



Answer: D

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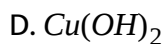
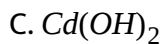
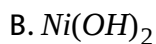
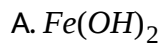
35. Select the ion exchange reaction, which proceeds to forward direction in aqueous medium:



Answer: D

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36. Which of the following metal hydroxide is not soluble in excess of NH_3 solution?

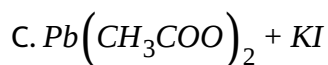
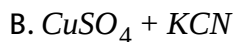
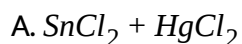


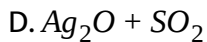
Answer: A



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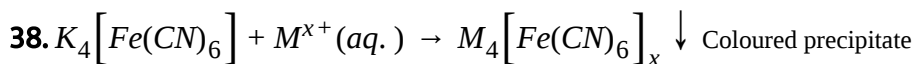
37. Which of the following combination of reagents does not undergo redox reaction in aqueous medium?



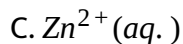
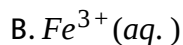
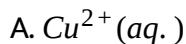


Answer: C

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Which of the following cation does not respond to the above reaction?



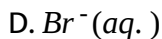
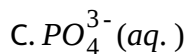
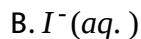
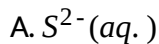
D. None of these

Answer: C

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39. Sodium salt solution + $AgNO_3$ soln. \rightarrow Coloured precipitate.

If coloured precipitate is soluble in both dil. HNO_3 and excess conc. NH_3 solution then which of the following anion is present in the salt solution?

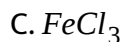
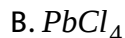


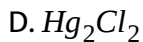
Answer: C



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40. Chlorine gas is not produced by heating:

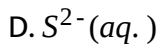
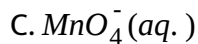
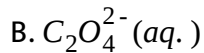
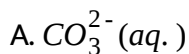




Answer: D

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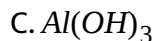
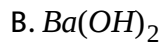
41. Which of the following anion does not produce precipitate with $BaCl_2$ solution however gives precipitate with $AgNO_3$?



Answer: D

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42. Which of the following compound is completely water soluble?

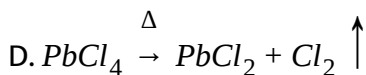
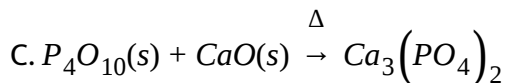
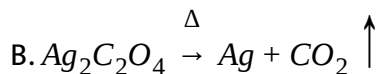
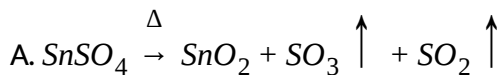


Answer: B



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43. Which chemical reaction contains incorrect products?

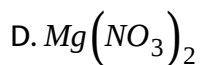
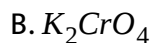
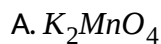


Answer: A



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44. Which of the following compound undergoes disproportionation in presence of SO_3 gas?

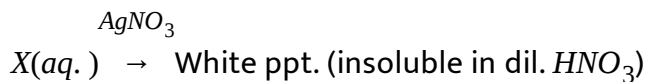
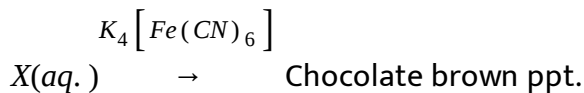


Answer: A

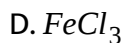
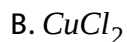


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45. Consider the following reaction:



Then 'X' will be:

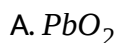


Answer: B



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46. Which of the following reagent does not oxidize HCl ?



B. conc. H_2SO_4

C. MnO_2

D. $K_2Cr_2O_7/H^+$

Answer: B

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47. Select correct match:

Anions

(a) CO_3^{2-} , SO_3^{2-}

(b) CO_3^{2-} , HCO_3^-

(c) SO_3^{2-} , SO_4^{2-}

(d) Cl^- , Br^-

Separated by reagent

$BaCl_2$

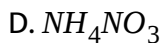
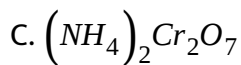
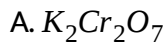
$CaCl_2$

$(CH_3COO)_2Pb$

$AgNO_3$

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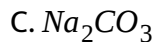
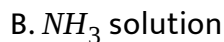
48. Which of the following compound does not produce green coloured product on thermal decomposition?



Answer: D

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49. Aqueous solution of $FeSO_4$ does not produce precipitate with:

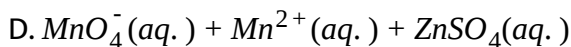
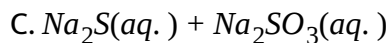
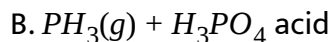
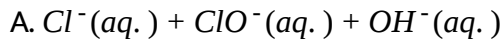


D. None of these

Answer: D

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50. Comproportionation occurs between:



Answer: D



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51. Colour of $CrO_4^{2-}(aq.)$ is not changed by

A. dil. HCl

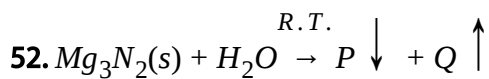
B. NH_3 solution

C. CH_3COOH

D. NO_2 gas

Answer: B

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Excess 'Q' gas does not form coloured complex with:

A. $\text{Ni}^{2+}(\text{aq.})$

B. $\text{Zn}^{2+}(\text{aq.})$

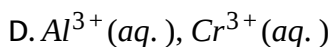
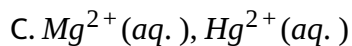
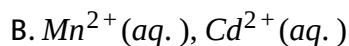
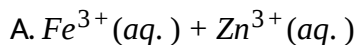
C. $\text{Cr}^{3+}(\text{aq.})$

D. $\text{Cu}^{2+}(\text{aq.})$

Answer: B

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53. Which of the following pair of cations can be separated by excess NaOH solution?

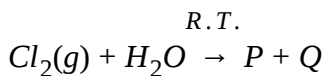


Answer: A



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54. Consider following reaction:



If molecular weight of P is less than Q then incorrect statement is:

A. On warming 'P' can form deep red coloured vapours with CrO_3

B. Q' exhibits bleaching property

C. MnO_2 can change 'P' into Cl_2 gas on warming

D. P reacts with H_2S gas while 'Q' does not

Answer: D

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55. Which of the following reagent can dissolves precipitate of $HgS \downarrow$

A. NH_3 solution

B. conc. HCl

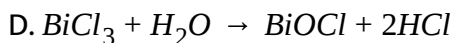
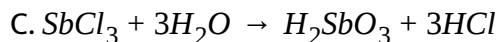
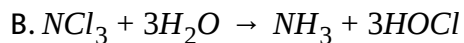
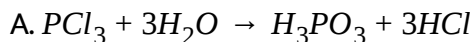
C. conc. HNO_3

D. Na_2S solution

Answer: D

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56. Which of the following reaction is incorrect?

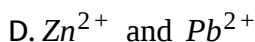
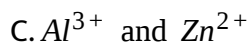
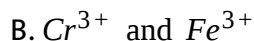
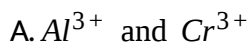


Answer: C



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57. Concentrated sodium hydroxide can separate a mixture of:



Answer: B



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58. Select correct set of species which can't react with water but react with NaOH,

(i) NO_2

(ii) P_4

(iii) Al

(iv) I_2

A. Only (iv)

B. (iii) and (iv)

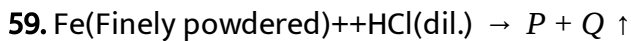
C. (ii), (iii) and (iv)

D. all (i), (ii), (iii) and (iv)

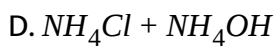
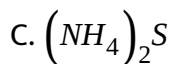
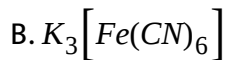
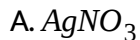
Answer: C



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Compound 'P' does not precipitate with:

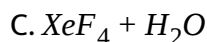
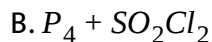
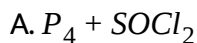


Answer: D



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60. Which combination gives maximum number of products?

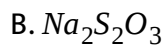
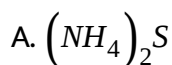


D. $NH_4NO_3 + Zn + \text{Excess Excess NaOH}$

Answer: C

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61. $Cu^{2+}(aq.)$ does not undergo redox reaction with solution:



Answer: A

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62. Hydrolysis of which of the following compound liberates acidic gas?

A. Li_2NH

B. Al_2S_3

C. CaC_2

D. $CaNCN$

Answer: B



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63. The non-metal which does not react with water but reacts with alkali?

A. Boron

B. Bromine

C. P_4

D. Fluorine

Answer: C



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64. A very dilute acidic solution of Cd^{2+} & Ni^{2+} gives only yellow ppt. of CdS on passing H_2S , this is due to:

- A. solubility product (K_{sp}) of CdS is more than that of NiS.
- B. Solubility product (K_{sp}) of CdS is less than that of NiS.
- C. Cd^{2+} belong to II B group while Ni^{2+} belongs to IVth group
- D. CdS is insoluble in yellow ammonium sulphide (YAS).

Answer: B



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65. Thermal decomposition of which of the salt listed below yield a basic and acidic oxides simultaneously?

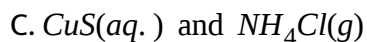
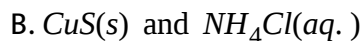
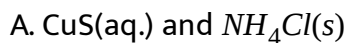




Answer: B

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66. What are formed products, when aqueous solution of CuCl_2 and $(\text{NH}_4)_2\text{S}$ are mixed?



Answer: B

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67. Which of the following compound does not react with cold and dil.

HNO_3 ?

A. PbO

B. PbO_2

C. $FeSO_4$

D. $PbCl_2$

Answer: B



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68. The incorrect order of solubility in water is:

A. $Ca(OH)_2 < Sr(OH)_2 < Ba(OH)_2$

B. $Li_2CO_3 < Na_2CO_3 < K_2CO_3$

C. $CsNO_3 < RbNO_3 < KNO_3$

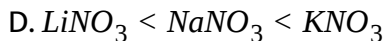
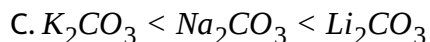
D. $BeS_2O_3 < MgS_2O_3 < CaS_2O_3$

Answer: D



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69. The correct order of increasing solubility in water is:

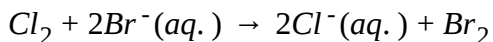


Answer: B



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70. Bromine is commercially prepared from sea water by displacement reaction



Br_2 gas thus formed is dissolved into solution of Na_2CO_3 and then pure Br_2 is obtained by treatment of the solution with :

A. $Ca(OH)_2$

B. $NaOH$

C. H_2SO_4

D. HI

Answer: C



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71. Which of the following metal on burning in moist air does not give smell of ammonia?

A. Mg

B. Ca

C. K

D. Li

Answer: C



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72. Gas that can not be collected over water is:

A. N_2

B. O_2

C. SO_2

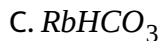
D. PH_3

Answer: C



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73. Compound having lowest thermal stability is:



Answer: A

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74. Which of the following statement is incorrect regarding Fe^{2+} and Fe^{3+} cations?

A. Fe^{3+} gives brown colour solution with potassium ferricyanide

B. Fe^{2+} gives blue precipitate with potassium ferricyanide

C. Fe^{3+} gives red colour solution with potassium thiocyanate

D. Fe^{2+} gives brown colour with ammonium thiocyanate

Answer: D

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75. $(NH_4)_2Cr_2O_7$ on heating liberates a gas. The same gas will be obtained by

- A. Heating NH_4NO_2
- B. Heating NH_4NO_3
- C. Heating $(NH_4)_2SO_4$
- D. Treatment Mg_3N_2 with H_2O

Answer: A

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76. Which of the following compound liberates acidic gas during its hydrolysis?

- A. Ca_3P_2

B. AlN

C. Al_2S_3

D. CaH_2

Answer: C

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77. Which of the following combination does not evolve Cl_2 gas?

A. $HCl(aq.) + KMnO_4$

B. $HCl + MnO_2$

C. $HCl + Br_2$

D. $HCl + F_2$

Answer: C

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78. NH_3 gas does not liberate by which of the following combination?

A. Heating of NH_4ClO_4

B. Heating of NH_4Cl

C. $(NH_4)_2CO_3 + NaOH$

D. $Li_3N + H_2O$

Answer: A



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79. If salt Q undergoes redox reaction with H_2S in acidic medium then which of the following species can not be possible product?

A. MnO_4^{2-} (aq.)

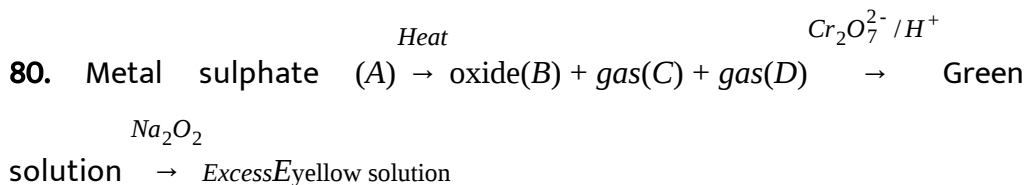
B. S

C. MnO_2

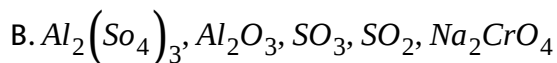
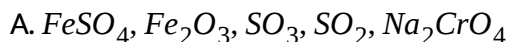
D. both (a) and (c)

Answer: D

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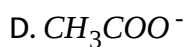
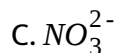
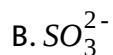
Compound A, B, C, D are E are respectively:



Answer: A

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81. Which of the following radical does not liberate gas with (Zn+dil. HCl) on warming?

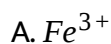


Answer: C



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82. Which of the following cation does not give precipitate with H_2S in neutral medium?



D. Ag^+

Answer: A

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83. $NaCl(\text{solid}) + K_2Cr_2O_7(\text{solid}) + \text{conc. } H_2SO_4 \xrightarrow{\text{warm}}$ Reddish brown fumes of 'X'.

The oxidation state of central atom in compound 'X' is:

A. +6

B. +3

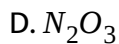
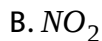
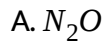
C. +2

D. zero

Answer: A

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84. Diamagnetic gas neutral towards water is:

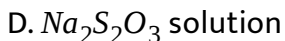
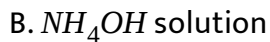
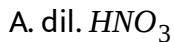


Answer: A



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85. Which of the following reagent can be used to separate $AgCl$ and AgI ?



Answer: B

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86. When PbO_2 reacts with conc. HNO_3 then evolved gas is:

A. NO_2

B. O_2

C. N_2

D. N_2O

Answer: B

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87. In a closed container there is a mixture of SO_2 , CO_2 and O_2 gas, which sequence of reagent can be helpful to separate them ?

(I) Limewater

(II) Acidified potassium dichromate

(III) Alkaline pyragallo.

A. (I),(II) and (III)

B. (II), (I), (III)

C. (III),(II), (I)

D. (III), (I), (II)

Answer: B



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88. Which salt is colourless?

A. $KMnO_4$

B. $BaSO_4$

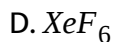
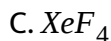
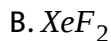
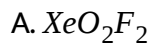
C. Na_2CrO_4

D. $CoCl_2$

Answer: B

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89. Which of the following Xenon compound does not produce explosive XeO_3 on its complete hydrolysis?



Answer: B

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90. $FeSO_4 \cdot 7H_2O$ (Green vitriol) salt on thermal decomposition does not produce:

A. SO_2

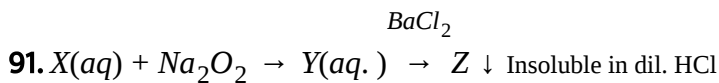
B. O_2

C. SO_3

D. H_2O vapour

Answer: B

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X and Y are different sodium salts, then anion present in the salt (X) is:

A. $Cr_2O_7^{2-}$

B. $C_2O_4^{2-}$

C. SO_3^{2-}

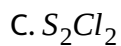
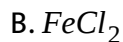
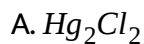
D. SO_4^{2-}

Answer: C



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92. Which of the following chloride does not react with PCl_5 on heating?

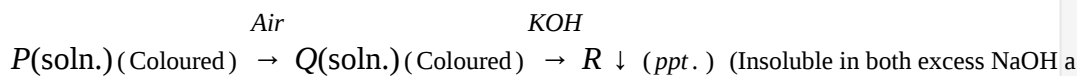


Answer: D

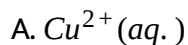


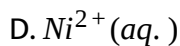
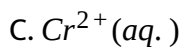
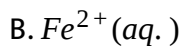
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93.



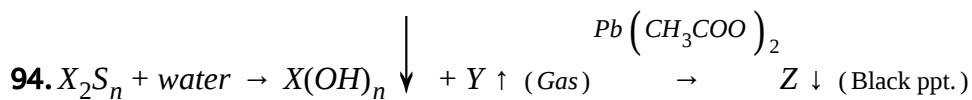
then P contains:



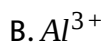
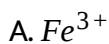


Answer: B

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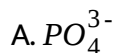
Then (X) cation can not be:



Answer: A

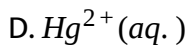
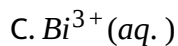
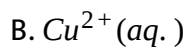
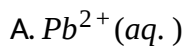
95. $X(\text{satl}) + \text{AgNO}_3(\text{aq.}) \rightarrow Y \downarrow$ (yellow ppt.) (soluble in excess of NH_3 solution)

Salt X, does not contain:



Answer: C

96. $M^{n+}(\text{aq.}) + \text{KI} \rightarrow X \downarrow$ ppt. $\xrightarrow{\text{Excess KI}}$ KI ppt. remains insoluble in excess KI solution. Then cation $M^{n+}(\text{aq.})$ can be:

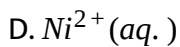
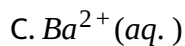
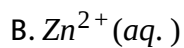
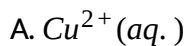


Answer: B



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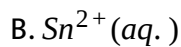
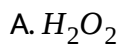
97. Aqueous solution of which of the following cation gives precipitate with potash alum?



Answer: C

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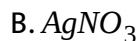
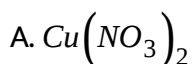
98. Colour of acidified $K_2Cr_2O_7$ is not changed by:

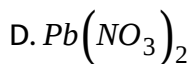
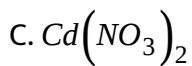


Answer: C

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99. Which of the following metal nitrate produces gaseous product when reacts with KCN solution?

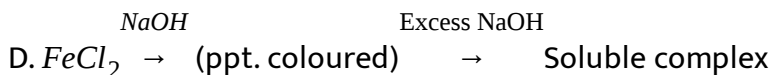
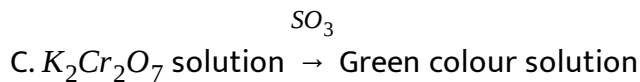
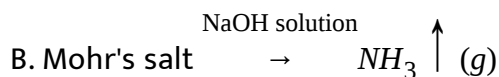
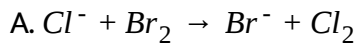




Answer: A

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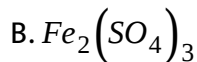
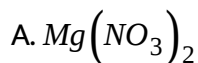
100. Which of these reaction is correct?



Answer: B

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101. Compound which on heating produces paramagnetic acidic gas?

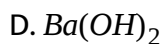
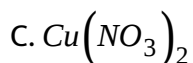
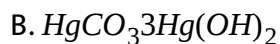
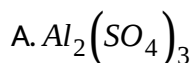


Answer: A



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102. Which compound on heating produces produces coloured metal oxide finally?

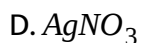
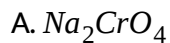


Answer: C

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103. P (Coloured solution) + $BaCl_2 \rightarrow Q \downarrow$ (White) + R (Coloured solution) ItBrgt

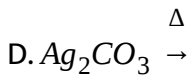
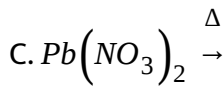
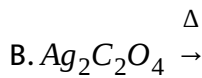
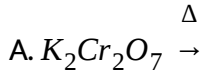
Then salt 'P' in above reaction is:



Answer: C

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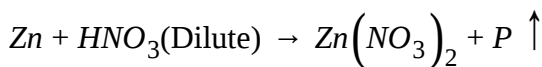
104. Oxygen gas is not produced from the following decomposition reaction:



Answer: B

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105. Consider the following reaction and select incorrect statement about gas (P):

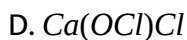
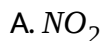


- A. Gives neutral solution in water
- B. Contains more O_2 than Air
- C. Forms brown ring with $FeSO_4$ solution
- D. None of these

Answer: C

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106. Which of the following ionic/molecular species does not disproportionate in water at room temperature?



Answer: D

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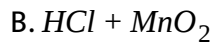
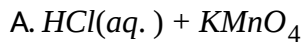
107. Which halogen oxidizes water at room temperature but does not undergo disproportionation into it?



Answer: A

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108. Which of the following combination doesn't evolve Cl_2 gas?



Answer: C

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109. Which of the following combination does not liberated NH_3 gas?

A. Heating of NH_4ClO_4

B. Heating of NH_4Cl

C. $(NH_4)_2CO_3 + NaOH$

D. $Li_3N + H_2O$

Answer: A



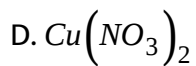
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110. Which of the following compound on heating does not produce metal oxide?

A. $MgCl_2 \cdot 6H_2O$

B. $K_2Cr_2O_7$

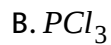
C. K_2CO_3



Answer: C

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111. Select the compound in which HCl is not the product of Hydrolysis:



Answer: A

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112. How many moles of H_2O are liberated when one mole hydrated $MgCl_2$ is heated?

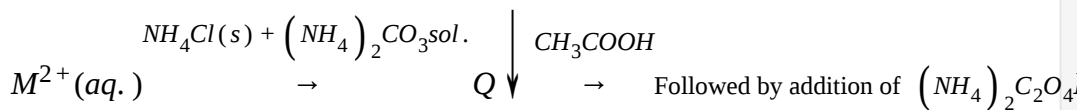
- A. 6
- B. 5
- C. 4
- D. 3

Answer: B



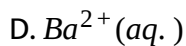
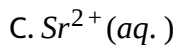
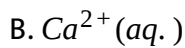
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113. Consider the following sequence of reaction:



Which of the following cation can form ppt. Q but does not form ppt. 'R' ?

- A. $Mg^{2+}(aq.)$



Answer: D

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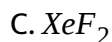
114. Which of the following compound does not liberated oxygen gas on warming with conc. H_2SO_4 ?



Answer: A

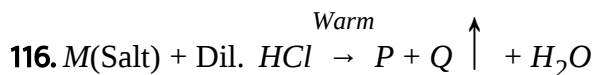
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115. One of the hydrolysed product of the following compound does not react with silica of glass vessel:



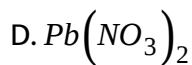
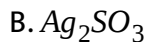
Answer: A

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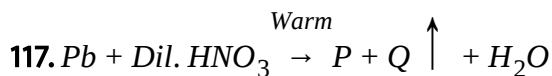
gas 'Q' changes colour of $FeSO_4$ solution into yellow solution then salt M in above reaction is





Answer: C

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Incorrect statement for Q is:

A. Paramagnetic colourless gas

B. It is oxidized to paramagnetic coloured gas by air

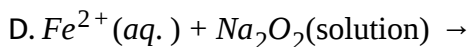
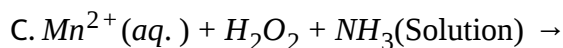
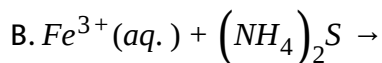
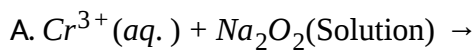
C. It combines with $Fe_2(SO_4)_3$

D. It can be also obtained by disproportionation of HNO_2

Answer: C

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118. In which of the following redox reaction precipitate is not formed?



Answer: A



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119. Which metal sulphide is soluble in excess NH_3 solution?

A. ZnS

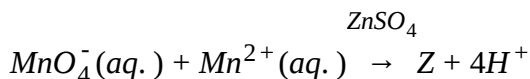
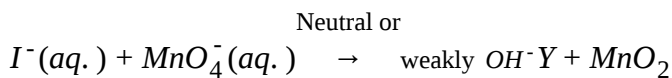
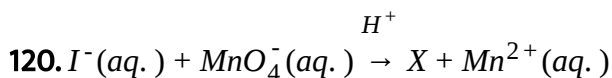
B. MnS

C. FeS

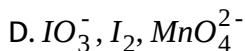
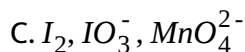
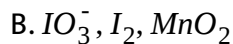
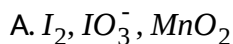
D. Cr_2S_3

Answer: D

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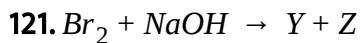
Product X, Y and Z are respectively.



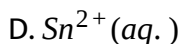
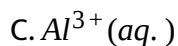
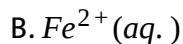
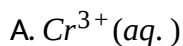
Answer: A

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R.T.

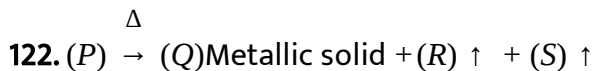


If Y gives precipitate with $AgNO_3$, then Z does not undergo reaction with:

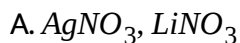


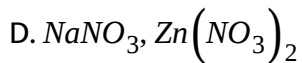
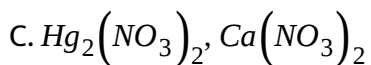
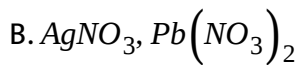
Answer: C

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P & X are respectively:



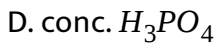
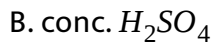
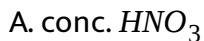


Answer: B



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123. Iodine is not oxidized to iodic acid/iodic anhydride by:



Answer: D



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124. Colourless gas that has oxidising as well as reducing properties:

A. CO_2

B. SO_2

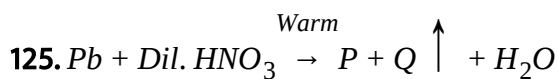
C. NO_2

D. SO_3

Answer: B



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Incorrect statement for Q is:

A. Paramagnetic colourless gas

B. It is oxidized to paramagnetic coloured gas by air

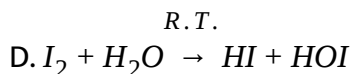
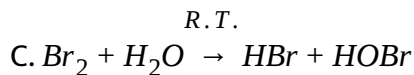
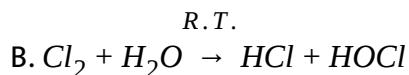
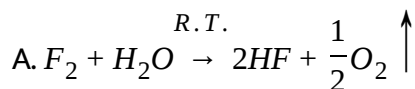
C. It combines with $Fe_2(SO_4)_3$

D. It is also obtained by disproportionation of HNO_2

Answer: C

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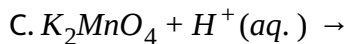
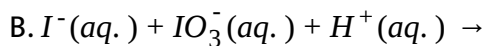
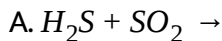
126. Which reaction has positive value of ΔG° ?



Answer: D

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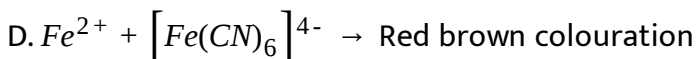
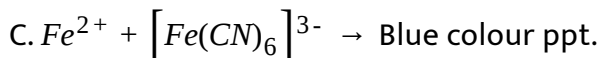
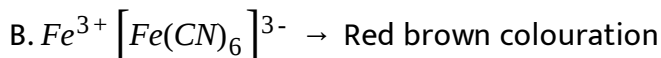
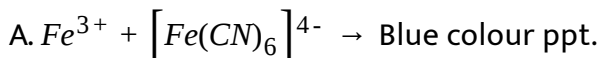
127. Which does not undergo comproportionation reaction?



Answer: C

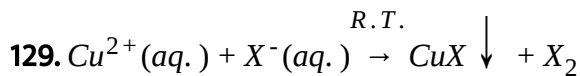
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128. Select the incorrect match:

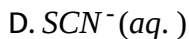
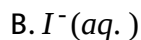
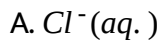


Answer: D

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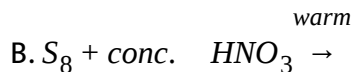
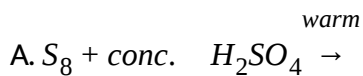
'X' cannot be:

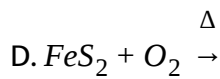
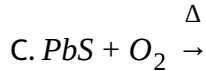


Answer: A

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130. In which of the following reaction SO_2 gas is not produced?





Answer: B

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131. Which metal gives NH_4NO_3 , when react with dilute HNO_3 acid?

A. *Zn*

B. *Pb*

C. *Cu*

D. *Au*

Answer: A

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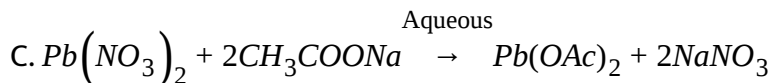
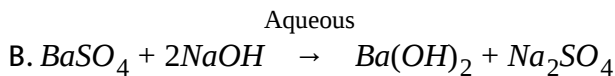
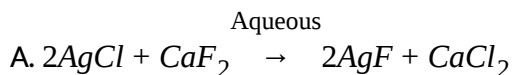
132. Select the salt whose aqueous solution is not green:



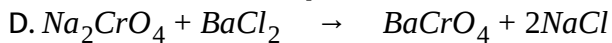
Answer: D

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133. Select the ion exchange reaction, which proceeds to forward direction in aqueous medium:



Aqueous

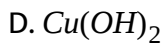
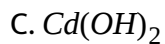
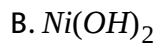
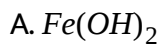


Answer: D



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134. Which of the following metal hydroxide is not soluble in excess of NH_3 solution?

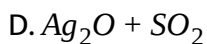
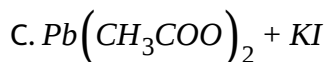
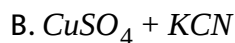
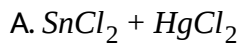


Answer: A



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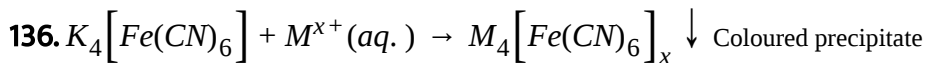
135. Which of the following combination of reagents does not undergo redox reaction in aqueous medium?



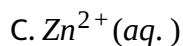
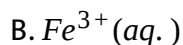
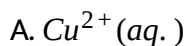
Answer: C



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Which of the following cation does not respond to the above reaction?



D. None of these

Answer: C

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137. Sodium salt solution + $AgNO_3$ soln. \rightarrow Coloured precipitate.

If coloured precipitate is soluble in both dil. HNO_3 and excess conc. NH_3 solution then which of the following anion is present in the salt solution?

A. $S^{2-}(aq.)$

B. $I^-(aq.)$

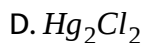
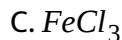
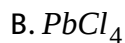
C. $PO_4^{3-}(aq.)$

D. $Br^-(aq.)$

Answer: C

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138. Chlorine gas is not produced by heating:

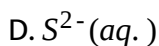
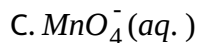
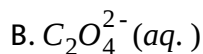
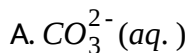


Answer: D



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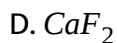
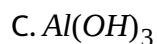
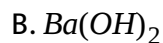
139. Which of the following anion does not produce precipitate with BaCl_2 solution however gives precipitate with AgNO_3 ?



Answer: D

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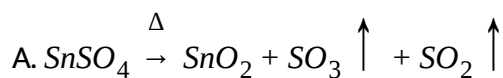
140. Which of the following compound is completely water soluble?

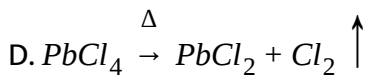
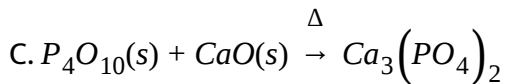
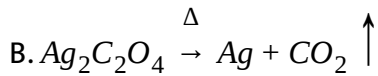


Answer: B

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141. Which chemical reaction contains incorrect products?

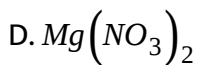
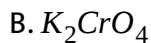
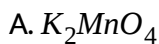




Answer: A

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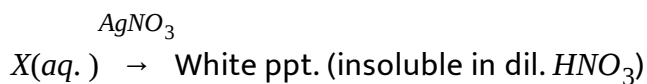
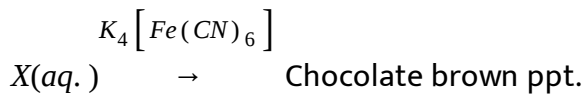
142. Which of the following compound undergoes disproportionation in presence of SO_3 gas?



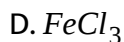
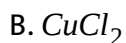
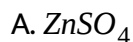
Answer: A

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143. Consider the following reaction:



Then 'X' will be:



Answer: B



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144. Which of the following reagent does not oxidize HCl ?

A. PbO_2

B. conc. H_2SO_4

C. MnO_2

D. $K_2Cr_2O_7/H^+$

Answer: B

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145. Select correct match:

Anions

(a) CO_3^{2-} , SO_3^{2-}

(b) CO_3^{2-} , HCO_3^-

(c) SO_3^{2-} , SO_4^{2-}

(d) Cl^- , Br^-

Separated by reagent

$BaCl_2$

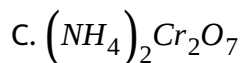
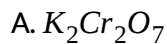
$CaCl_2$

$(CH_3COO)_2Pb$

$AgNO_3$

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146. Which of the following compound does not produce green coloured product on thermal decomposition?

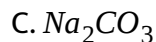
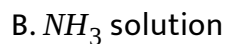


Answer: D



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147. Aqueous solution of $FeSO_4$ does not produce precipitate with:

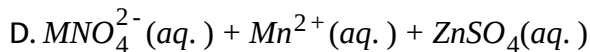
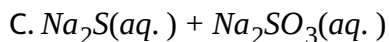
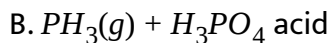
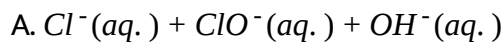


D. None of these

Answer: D

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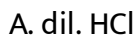
148. Comproportionation occurs between:



Answer: D

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149. Colour of $CrO_4^{2-}(aq.)$ is not changed by



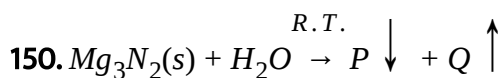
B. NH_3 solution

C. CH_3COOH

D. NO_2 gas

Answer: B

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Excess 'Q' gas does not form coloured complex with:

A. $Ni^{2+}(aq.)$

B. $Zn^{2+}(aq.)$

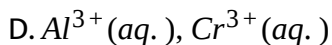
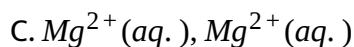
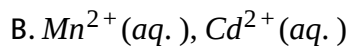
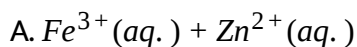
C. $Cr^{3+}(aq.)$

D. $Cu^{2+}(aq.)$

Answer: B

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151. Which of the following pair of cations can be separated by excess NaOH solution?

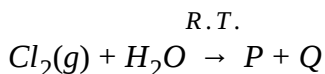


Answer: A



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152. Consider following reaction:



If molecular weight of P is less than Q then incorrect statement is:

A. On warming 'P' can form deep red coloured vapours with CrO_3

B. Q' exhibits bleaching property

C. MnO_2 can change P into Cl_2 gas on warming

D. P' reacts with H_2S gas while 'Q' does not

Answer: D

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153. Which of the following reagent can dissolves precipitate of $HgS \downarrow$

A. NH_3 solution

B. conc. HCl

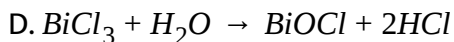
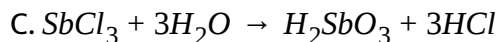
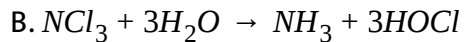
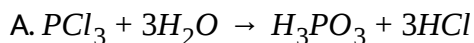
C. conc. HNO_3

D. Na_2S solution

Answer: D

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154. Which of the following reaction is incorrect?

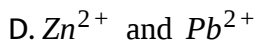
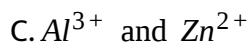
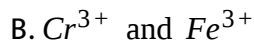
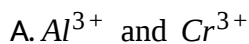


Answer: C



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155. Concentrated sodium hydroxide can separate a mixture of:



Answer: B



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156. Select correct set of species which can't react with water but react with NaOH,

(i) NO_2

(ii) P_4

(iii) Al

(iv) I_2

A. Only (iv)

B. (iii) and (iv)

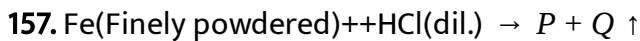
C. (ii), (iii) and (iv)

D. all (i), (ii), (iii) and (iv)

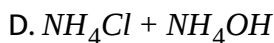
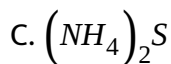
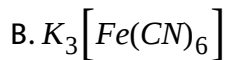
Answer: C



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Compound 'P' does not precipitate with:

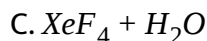
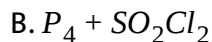
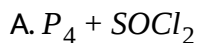


Answer: D



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158. Which combination gives maximum number of products?

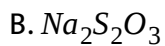
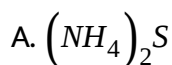


D. $NH_4NO_3 + Zn + \text{Excess Excess NaOH}$

Answer: C

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159. $Cu^{2+}(aq.)$ does not undergo redox reaction with solution:



Answer: A

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160. Hydrolysis of which of the following compound liberates acidic gas?

A. Li_2NH

B. Al_2S_3

C. CaC_2

D. $CaNCN$

Answer: B

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161. The non-metal which does not react with water but reacts with alkali?

A. Boron

B. Bromine

C. P_4

D. Fluorine

Answer: C

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162. A very dilute acidic solution of Cd^{2+} & Ni^{2+} gives only yellow ppt. of CdS on passing H_2S , this is due to:

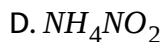
- A. solubility product (K_{sp}) of CdS is more than that of NiS.
- B. Solubility product (K_{sp}) of CdS is less than that of NiS.
- C. Cd^{2+} belong to II B group while Ni^{2+} belongs to IVth group
- D. CdS is insoluble in yellow ammonium sulphide (YAS).

Answer: B

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163. Thermal decomposition of which of the salt listed below yield a basic and acidic oxides simultaneously?

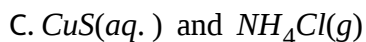
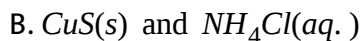
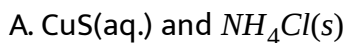
- A. NH_4ClO_4
- B. $CaCO_3$



Answer: B

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164. What are formed products, when aqueous solution of CuCl_2 and $(\text{NH}_4)_2\text{S}$ are mixed?



Answer: B

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165. Which of the following compound does not react with cold and dil.

HNO_3 ?

A. PbO

B. PbO_2

C. $FeSO_4$

D. $PbCl_2$

Answer: B



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166. The incorrect order of solubility in water is:

A. $Ca(OH)_2 < Sr(OH)_2 < Ba(OH)_2$

B. $Li_2CO_3 < Na_2CO_3 < K_2CO_3$

C. $CsNO_3 < RbNO_3 < KNO_3$

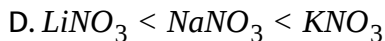
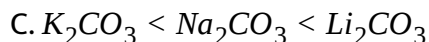
D. $BeS_2O_3 < MgS_2O_3 < CaS_2O_3$

Answer: D



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167. The correct order of increasing solubility in water is:

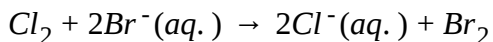


Answer: B



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168. Bromine is commercially prepared from sea water by displacement reaction



Br_2 gas thus formed is dissolved into solution of Na_2CO_3 and then pure Br_2 is obtained by treatment of the solution with :

A. $Ca(OH)_2$

B. $NaOH$

C. H_2SO_4

D. HI

Answer: C



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169. Which of the following metal on burning in moist air does not give smell of ammonia?

A. Mg

B. Ca

C. K

D. Li

Answer: C



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170. Gas that can not be collected over water is:

A. N_2

B. O_2

C. SO_2

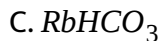
D. PH_3

Answer: C



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171. Compound having lowest thermal stability is:



Answer: A

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172. Which of the following statement is incorrect regarding Fe^{2+} and Fe^{3+} cations?

A. Fe^{3+} gives brown colour solution with potassium ferricyanide

B. Fe^{2+} gives blue precipitate with potassium ferricyanide

C. Fe^{3+} gives red colour solution with potassium thiocyanate

D. Fe^{2+} gives brown colour with ammonium thiocyanate

Answer: D

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173. $(NH_4)_2Cr_2O_7$ on heating liberates a gas. The same gas will be obtained by

- A. Heating NH_4NO_2
- B. Heating NH_4NO_3
- C. Heating $(NH_4)_2SO_4$
- D. Treatment Mg_3N_2 with H_2O

Answer: A

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174. Which of the following compound liberates acidic gas during its hydrolysis?

- A. Ca_3P_2

B. AlN

C. Al_2S_3

D. CaH_2

Answer: C

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175. Which of the following combination does not evolve Cl_2 gas?

A. $HCl(aq.) + KMnO_4$

B. $HCl + MnO_2$

C. $HCl + Br_2$

D. $HCl + F_2$

Answer: C

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176. NH_3 gas does not liberate by which of the following combination?

A. Heating of NH_4ClO_4

B. Heating of NH_4Cl

C. $(NH_4)_2CO_3 + NaOH$

D. $Li_3N + H_2O$

Answer: A



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177. If salt Q undergoes redox reaction with H_2S in acidic medium then which of the following species can not be possible product?

A. MnO_4^{2-} (aq.)

B. S

C. MnO_2

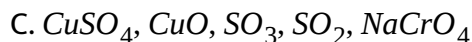
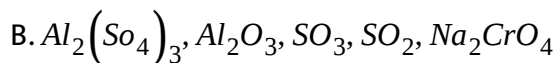
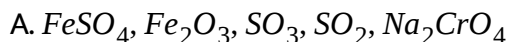
D. both (a) and (c)

Answer: D

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178. Metal sulphate $(A) \xrightarrow{\text{Heat}}$ oxide $(B) + \text{gas}(C) + \text{gas}(D) \xrightarrow{\text{Cr}_2\text{O}_7^{2-} / \text{H}^+}$ Green
solution $\xrightarrow{\text{Na}_2\text{O}_2}$ Excess E yellow solution

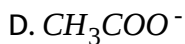
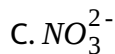
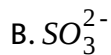
Compound A, B, C, D are E are respectively:



Answer: A

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179. Which of the following radical does not liberate gas with (Zn+dil. HCl) on warming?

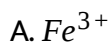


Answer: C



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180. Which of the following cation does not give precipitate with H_2S in neutral medium?



D. Ag^+

Answer: A

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181. $NaCl(solid) + K_2Cr_2O_7(solid) + conc. H_2SO_4 \xrightarrow{warm}$ Reddish brown fumes of 'X'.

The oxidation state of central atom in compound 'X' is:

A. +6

B. +3

C. +2

D. zero

Answer: A

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182. Diamagnetic gas neutral towards water is:

A. N_2O

B. NO_2

C. NO

D. N_2O_3

Answer: A

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183. Which of the following reagent can be used to separate $AgCl$ and AgI ?

A. dil. HNO_3

B. NH_4OH solution

C. KCN solution

D. $Na_2S_2O_3$ solution

Answer: B

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184. When PbO_2 reacts with conc. HNO_3 then evolved gas is:

A. NO_2

B. O_2

C. N_2

D. N_2O

Answer: B

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185. In a closed container there is a mixture of SO_2 , CO_2 and O_2 gas, which sequence of reagent can be helpful to separate them ?

(I) Limewater

(II) Acidified potassium dichromate

(III) Alkaline pyragallo.

A. (I),(II) and (III)

B. (II), (I), (III)

C. (III),(II), (I)

D. (III), (I), (II)

Answer: B



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186. Which salt is colourless?

A. $KMnO_4$

B. $BaSO_4$

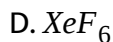
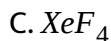
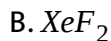
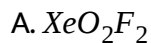
C. $NaCrO_4$

D. $CoCl_2$

Answer: B

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187. Which of the following Xenon compound does not produce explosive XeO_3 on its complete hydrolysis?



Answer: B

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188. $FeSO_4 \cdot 7H_2O$ (Green vitriol) salt on thermal decomposition does not produce:

A. SO_2

B. O_2

C. SO_3

D. H_2O vapour

Answer: B



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189. $X(aq) + Na_2O_2 \rightarrow Y(aq.) \rightarrow Z \downarrow$ Insoluble in dil. HCl

X and Y are different sodium salts, then anion present in the salt (X) is:

A. $Cr_2O_7^{2-}$

B. $C_2O_4^{2-}$

C. SO_3^{2-}

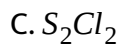
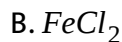
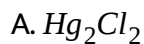
D. SO_4^{2-}

Answer: C



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190. Which of the following chloride does not react with PCl_5 on heating?

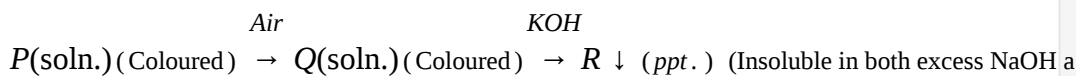


Answer: D

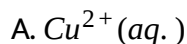


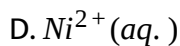
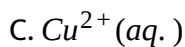
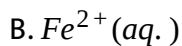
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191.



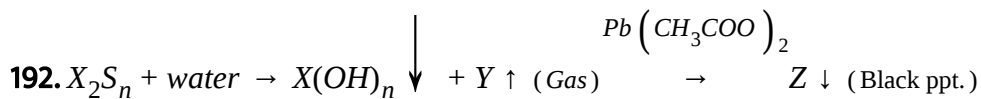
then P contains:





Answer: B

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Then (X) cation can not be:

A. Fe^{3+} gives brown colour solution with potassium ferricyanide

B. Al^{3+}

C. Cr^{3+}

D. Mg^{2+}

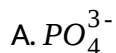
Answer: A



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193. $X(\text{satl}) + \text{AgNO}_3(\text{aq.}) \rightarrow Y \downarrow$ (yellow ppt.) (soluble in excess of NH_3 solution)

Salt X, does not contain:

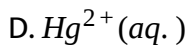
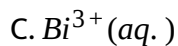
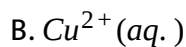
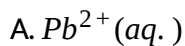


Answer: C



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194. $M^{n+}(\text{aq.}) + \text{KI} \rightarrow X \downarrow \text{ppt.} \xrightarrow{\text{Excess KI}} \text{KI ppt.}$ remains insoluble in excess KI solution. Then cation $M^{n+}(\text{aq.})$ can be:

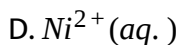
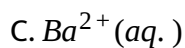
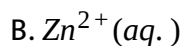
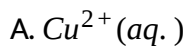


Answer: B



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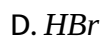
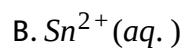
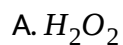
195. Aqueous solution of which of the following cation gives precipitate with potash alum?



Answer: C

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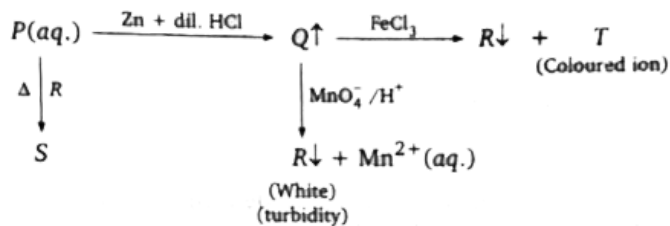
196. Colour of acidified $K_2Cr_2O_7$ is not changed by:



Answer: C

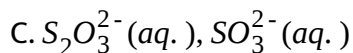
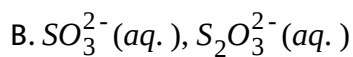
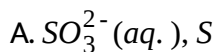
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LEVEL 3



1.

Q. Species P and S are respectively:

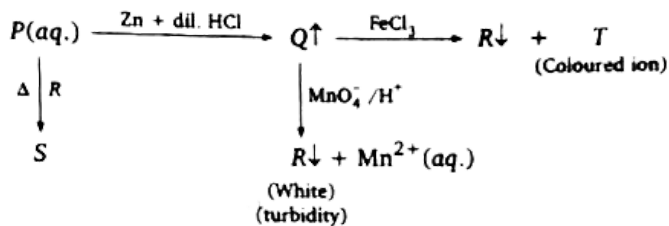


D. None of these

Answer: B



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2.

Q. 'T' cannot be identify by:

A. NH_3 solution

B. NH_4SCN

C. $(\text{NH}_4)_2\text{S}$

D. Excess KCN

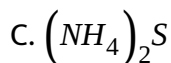
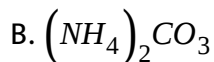
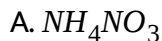
Answer: B

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3. Consider three P, Q, R, salts among them P and Q salts have different cations and also have different coloured polyatomic anion due to charge transfer phenomenon while P and R salts have same cation but have different anions. Salts R decomposes into an acidic gas and a basic

gas.

Q. Salt R can not be:



Answer: A

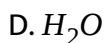
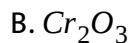


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4. Consider three P, Q, R, salts among them P and Q salts have different cations and also have different coloured polyatomic anion due to charge transfer phenomenon while P and R salts have same cation but have different anions. Salts R decomposes into an acidic gas and a basic gas.

Q. Salt P decomposes on heating into a coloured solid, neutral gas and a

neutral vapour, then which of the following can not be the product of salt P after decomposition?



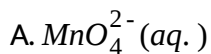
Answer: C



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5. Consider three P, Q, R, salts among them P and Q salts have different cations and also have different coloured polyatomic anion due to charge transfer phenomenon while P and R salts have same cation but have different anions. Salt R decomposes into an acidic gas and a basic gas.

Q. If salt Q undergoes redox reaction with H_2S in acidic medium then which of the following species can not be possible product?



B. S



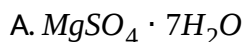
D. Both (a) and (c)

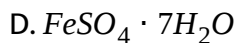
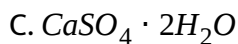
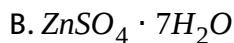
Answer: D

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6. Three compound X, Y and Z were taken into three different laboratory vessels and they are carried out by a chemist in his car. The car caught fire due to short circuit and the chemist came out of the car and noticed following observations:

Q. Compound X changes into white substance along with liberation of neutral oxide and then white substance decomposed into three products among which two are acidic oxides. among these oxides non-polar oxide can undergo polar cyclic polymer on cooling. the compound X will be:





Answer: D

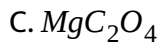


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7. Three compound X, Y and Z were taken into three different laboratory vessels and they are carried out by a chemist in his car. The car caught fire due to short circuit and the chemist came out of the car and noticed following observations:

Q. Compound Y produced two oxides, among these one oxide turns anhydrous $CuSO_4$ into blue and other gas slows down fire in the car, then Y is





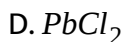
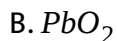
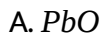
Answer: B



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8. Three compound X, Y and Z were taken into three different laboratory vessels and they are carried out by a chemist in his car. The car caught fire due to short circuit and the chemist came out of the car and noticed following observations:

Q. Which of the following compound does not react with cold and dil. HNO_3 ?

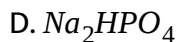
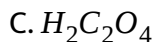
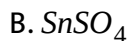


Answer: B

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9. In salts of polyatomic anion, as polarising power of cation increases, thermal stability of the salt decreases, and decomposed species may further undergo redox reaction.

Q. Which of the following species undergoes non-redox thermal decomposition reaction on heating?

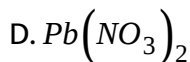
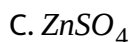
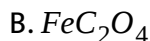
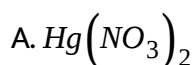


Answer: D

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10. In salts of polyatomic anion, as polarising power of cation increases, thermal stability of the salt decreases, and decomposed species may further undergo redox reaction.

Q. Water soluble salt (x) was heated into three products A, B and C and B and C are two different paramagnetic gases. A is red in hot condition, then salt (x) is:



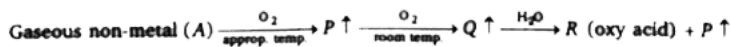
Answer: D



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11. Dioxygen directly reacts with nearly all metals and non-metals except some metals (e.g., Au, Pt) and some noble gases and form oxide(s). Oxides

can be simple (e.g., MgO , Al_2O_3) or mixed (Pb_3O_4 , Fe_3O_4). Simple oxides can be classified on the basis of their acidic, basic or amphoteric character. An oxide that combines with water to give an acid is termed an acidic oxide (i.e., SO_2 , Cl_2O_7 , CO_2 , N_2O_5). For example, SO_2 combines with water to give H_2SO_3 , an acid.



Then select correct statement with respect to gas 'Q'?

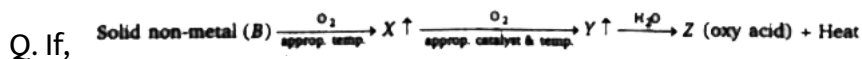
- A. Paramagnetic gas
- B. Neutral oxide
- C. Colourless gas
- D. Diatomic gas

Answer: A



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12. Dioxygen directly reacts with nearly all metals and non-metals except some metals (e.g., Au, Pt) and some noble gases and form oxide(s). Oxides can be simple (e.g., MgO, Al_2O_3) or mixed (Pb_3O_4 , Fe_3O_4). Simple oxides can be classified on the basis of their acidic, basic or amphoteric character. An oxide that combines with water to give an acid is termed acidic oxide (i.e., SO_2 , Cl_2O_7 , CO_2 , N_2O_5). For example, SO_2 combines with water to give H_2SO_3 , an acid.



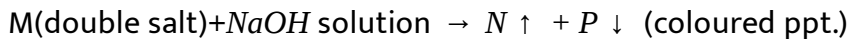
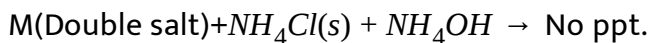
Then select incorrect statement with respect to gas 'X'

- A. burning sulphur smell
- B. Reacts with Cl_2
- C. Residue of sulphur with H_2S
- D. Does not react with $Ca(OCl)Cl$

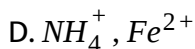
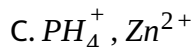
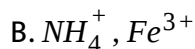
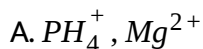
Answer: D

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13. Consider the following reactions and answer the following questions.



Q. Which of the following pair of cations are present in salt M?

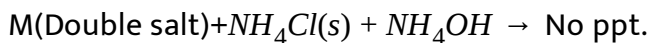


Answer: D



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14. Consider the following reactions and answer the following questions.



Q. $P \downarrow + \text{conc. } HCl \rightarrow Q(\text{coloured solution})$

Incorrect statement about Q is:

- A. It can exist in dimeric form
- B. Its aqueous solution is acidic
- C. It is used in methylene blue test for H_2S
- D. On passing Cl_2 gas colour of aqueous solution of Q changes

Answer: C



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15. Consider the following reactions and answer the following questions.

$M(\text{Double salt}) + NH_4Cl(s) + NH_4OH \rightarrow \text{No ppt.}$

$M(\text{double salt}) + NaOH \text{ solution} \rightarrow N \uparrow + P \downarrow (\text{coloured ppt.})$

Q. Reaction does not occur with salt M and gas N:

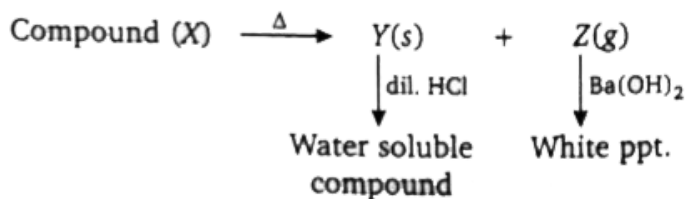
A. $NaNO_2 + \text{dil. } H_2SO_4 + M(\text{salt solution}) \rightarrow$

B. $HgI_2 + N \uparrow \rightarrow$



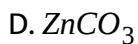
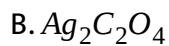
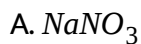
Answer: C

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16.

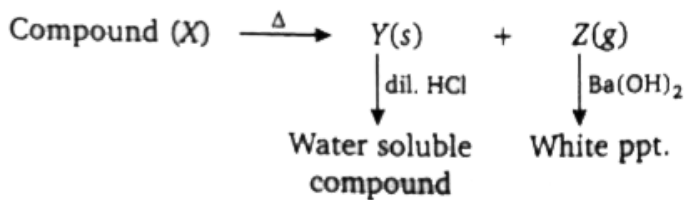
Q. Compound 'X' is:



Answer: D



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17.

Q. Incorrect statement 'Y' changes on heating:

- A. Colour of 'Y' changes on heating
- B. Z' is anhydride of H_2CO_3
- C. Y' can react with NaOH
- D. Z' does not act as Lewis acid

Answer: D

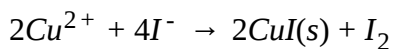


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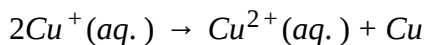
18. The unique behaviour of Cu, having a positive E° (reduction potential) accounts for its inability to liberate H_2 from acids. Only oxidising acids (nitric acid and hot concentrated sulphuric acid) react with Cu. The high energy of transform Cu(s) to $Cu^{2+}(aq.)$ is not balanced by its hydration enthalpy.

On the other hand, All Cu(II) halides are known except iodide. in this case,

Cu^{2+} oxidises I^- to I_2 :

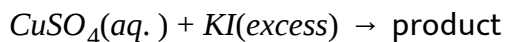


However, copper (I) compounds are unstable in aqueous solution and undergo disproportionation.



The stability of $Cu^{2+}(aq.)$ rather than $Cu^+(aq.)$ is due to the much more negative Δ_{Hyd} of $Cu^{2+}(aq.)$ than $Cu^+(aq.)$

Q. Consider the following transformation:



Select the correct statement:

A. Product contains $[Cu(H_2O_4)]^{2+}$ ion.

B. Presence of brown colouration in product is due to I_3^- ion

C. Oxidation state of sulphur in reactant and product is different

D. white ppt. of CuI_2 is observed in product

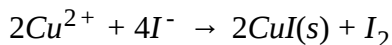
Answer: B

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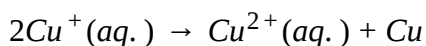
19. The unique behaviour of Cu, having a positive E° (reduction potential) accounts for its inability to liberate H_2 from acids. Only oxidising acids (nitric acid and hot concentrated sulphuric acid) react with Cu. The high energy of transform Cu(s) to $Cu^{2+}(aq.)$ is not balanced by its hydration enthalpy.

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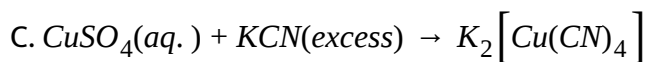
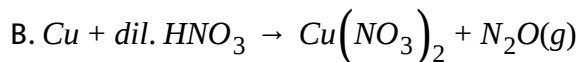
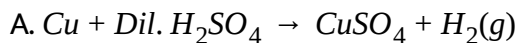
However, copper (I) compounds are unstable in aqueous solution and undergo disproportionation.



The stability of $Cu^{2+}(aq.)$ rather than $Cu^+(aq.)$ is due to the much more

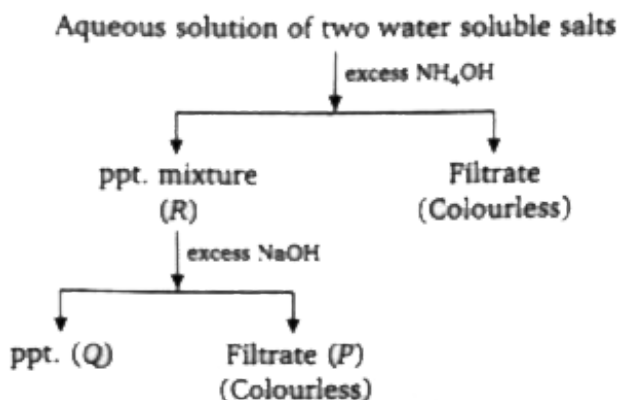
negative Δ_{Hyd} of $Cu^{2+}(aq.)$ than $Cu^+(aq.)$

Q. Select the correct chemical change:



Answer: D

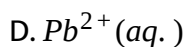
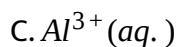
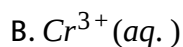
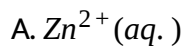
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20.

Q. When H_2S gas was passed into filtrate (P), a coloured precipitate was

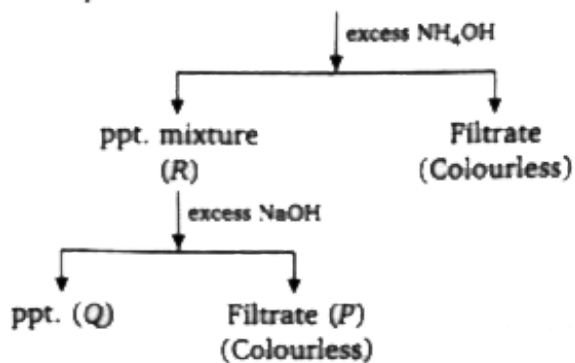
obtained, then cation present in the filtrate is:



Answer: D

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Aqueous solution of two water soluble salts



21.

Q. Precipitate (Q) was treated with dil. HCl and a coloured solution was obtained. On passing H_2S gas into this solution no precipitate was

obtained but colour of the solution changes, then cation present in the precipitate (Q) can be identified by:

A. $Na_2S_2O_3$ solution

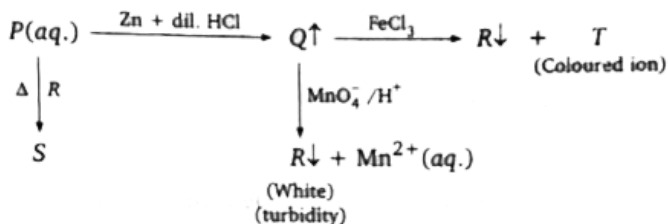
B. $KI + Starch$

C. $K_4[Fe(CN)_6]$

D. All

Answer: D

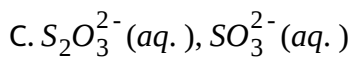
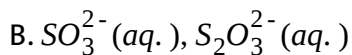
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22.

Q. Species P and S are respectively:

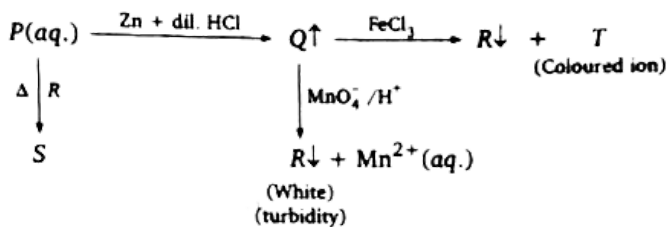
A. $SO_3^{2-}(aq.)$, S



D. None of these

Answer: B

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23.

Q. 'T' cannot be identify by:

A. NH_3 solution

B. NH_4SCN

C. $(\text{NH}_4)_2\text{S}$

D. Excess KCN

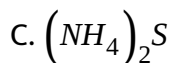
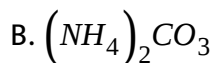
Answer: B



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24. Consider three P, Q, R, salts among them P and Q salts have different cations and also have different coloured polyatomic anion due to charge transfer phenomenon while P and R salts have same cation but have different anions. Salt R decomposes into an acidic gas and a basic gas.

Q. Salt R can not be:



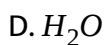
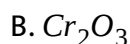
Answer: A



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25. Consider three P, Q, R, salts among them P and Q salts have different cations and also have different coloured polyatomic anion due to charge transfer phenomenon while P and R salts have same cation but have different anions. Salts R decomposes into an acidic gas and a basic gas.

Q. Salt P decomposes on heating into a coloured solid, neutral gas and a neutral vapour, then which of the following can not be the product of salt P after decomposition?



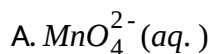
Answer: C



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26. Consider three P, Q, R, salts among them P and Q salts have different cations and also have different coloured polyatomic anion due to charge transfer phenomenon while P and R salts have same cation but have different anions. Salts R decomposes into an acidic gas and a basic gas.

Q. If salt Q undergoes redox reaction with H_2S in acidic medium then which of the following species can not be possible product?



B. S



D. Both (a) and (c)

Answer: D

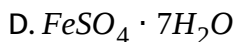
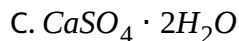
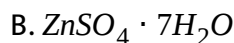
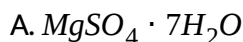


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27. Three compound X, Y and Z were taken into three different laboratory vessels and they are carried out by a chemist in his car. The car caught

fire due to short circuit and the chemist came out of the car and noticed following observations:

Q. Compound X changes into white substance along with liberation of neutral oxide and then white substance decomposed into three products among which two are acidic oxides. among these oxides non-polar oxide can undergo polar cyclic polymer on cooling. the compound X will be:



Answer: D



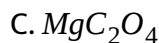
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28. Three compound X, Y and Z were taken into three different laboratory vessels and they are carried out by a chemist in his car. The car caught fire due to short circuit and the chemist came out of the car and noticed

following observations:

Q. Compound Y produced two oxides, among these one oxide turns anhydrous $CuSO_4$ into blue and other gas slows down fire in the car, then

Y is



Answer: B



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29. Three compound X, Y and Z were taken into three different laboratory vessels and they are carried out by a chemist in his car. The car caught fire due to short circuit and the chemist came out of the car and noticed following observations:

Q. Which of the following compound does not react with cold and dil.

HNO_3 ?

A. PbO

B. PbO_2

C. $FeSO_4$

D. $PbCl_2$

Answer: B

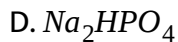
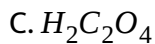
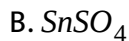


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30. In salts of polyatomic anion, as polarising power of cation increases, thermal stability of the salt decreases, and decomposed species may further undergo redox reaction.

Q. Which of the following species undergoes non-redox thermal decomposition reaction on heating?

A. $FeSO_4$



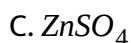
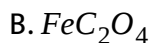
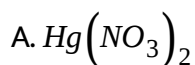
Answer: D

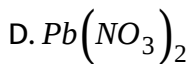


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31. In salts of polyatomic anion, as polarising power of cation increases, thermal stability of the salt decreases, and decomposed species may further undergo redox reaction.

Q. Water soluble salt (x) was heated into three products A, B and C and B and C are two different paramagnetic gases. A is red in hot condition, then salt (x) is:

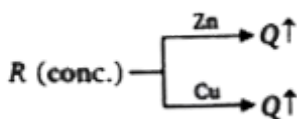
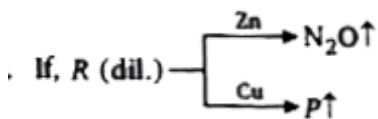
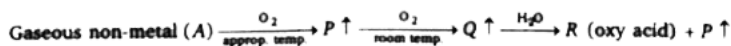




Answer: D

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32. Dioxygen directly reacts with nearly all metals and non-metals except some metals (e.g., Au, Pt) and some noble gases and form oxide(s). Oxides can be simple (e.g., MgO, Al_2O_3) or mixed (Pb_3O_4 , Fe_3O_4). Simple oxides can be classified on the basis of their acidic, basic or amphoteric character. An oxide that combines with water to give an acid is termed acidic oxide (i.e., SO_2 , Cl_2O_7 , CO_2 , N_2O_5). For example, SO_2 combines with water to give H_2SO_3 , an acid.



Then select correct statement with respect to gas 'Q'?

A. Paramagnetic gas

B. Neutral oxide

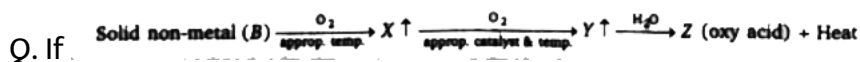
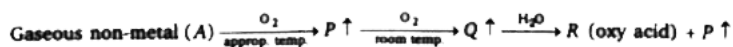
C. Colourless gas

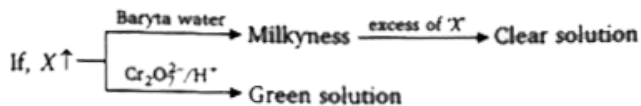
D. Diatomic gas

Answer: A

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33. Dioxygen directly reacts with nearly all metals and non-metals except some metals (e.g., Au, Pt) and some noble gases and form oxide(s). Oxides can be simple (e.g., MgO , Al_2O_3) or mixed (Pb_3O_4 , Fe_3O_4). Simple oxides can be classified on the basis of their acidic, basic or amphoteric character. An oxide that combines with water to give an acid is termed acidic oxide (i.e., SO_2 , Cl_2O_7 , CO_2 , N_2O_5). For example, SO_2 combines with water to give H_2SO_3 , an acid.





Then 'X' is

- A. NO
- B. CO_2
- C. SO_2
- D. SO_3

Answer: C



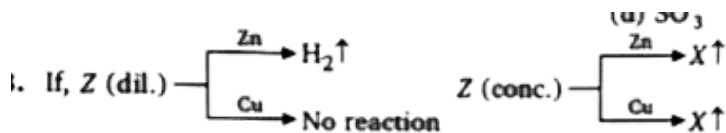
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34. Dioxygen directly reacts with nearly all metals and non-metals except some metals (e.g., Au, Pt) and some noble gases and form oxide(s). Oxides can be simple (e.g., MgO , Al_2O_3) or mixed (Pb_3O_4 , Fe_3O_4). Simple oxides can be classified on the basis of their acidic, basic or amphoteric character. An oxide that combines with water to give an acid is termed acidic oxide (i.e., SO_2 , Cl_2O_7 , CO_2 , N_2O_5). For example, SO_2 combines with

water to give H_2SO_3 , an acid.



Q. If, $\text{Solid non-metal (B)} \xrightarrow[\text{approp. temp.}]{O_2} X \uparrow \xrightarrow[\text{approp. catalyst \& temp.}]{O_2} Y \uparrow \xrightarrow{H_2O} Z \text{ (oxy acid)} + \text{Heat}$



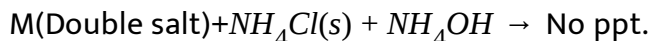
Then select incorrect statement with respect to gas 'X'

- A. burning sulphur smell
- B. Reacts with Cl_2
- C. Residue of sulphur with H_2S
- D. Does not react with $Ca(OCl)Cl$

Answer: D

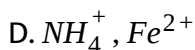
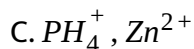
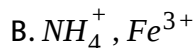
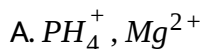
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35. Consider the following reactions and answer the following questions.



M(double salt)+NaOH solution $\rightarrow N \uparrow + P \downarrow$ (coloured ppt.)

Q. Which of the following pair of cations are present in salt M?

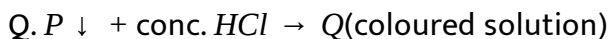
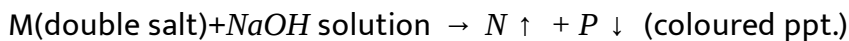
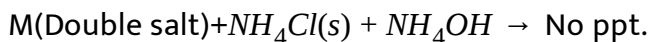


Answer: D



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36. Consider the following reactions and answer the following questions.



Incorrect statement about Q is:

A. It can exist in dimeric form

B. Its aqueous solution is acidic

C. It is used in methylene blue test for H_2S

D. On passing Cl_2 gas colour of aqueous solution of Q changes

Answer: C

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37. Consider the following reactions and answer the following questions.

$M(\text{Double salt}) + NH_4Cl(s) + NH_4OH \rightarrow \text{No ppt.}$

$M(\text{double salt}) + NaOH \text{ solution} \rightarrow N \uparrow + P \downarrow \text{ (coloured ppt.)}$

Q. Reaction does not occur with salt M and gas N:

A. $NaNO_2 + \text{dil. } H_2SO_4 + M(\text{salt solution}) \rightarrow$

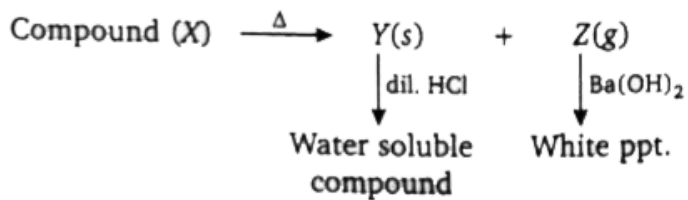
B. $HgI_2 + N \uparrow \rightarrow$

C. $M(\text{salt solution}) + H_2S \rightarrow$

D. $M(\text{salt solution}) + Br_2 \rightarrow$

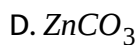
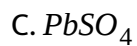
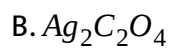
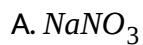
Answer: C

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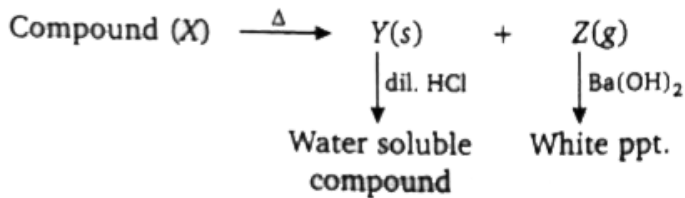
38.

Q. Compound 'X' is:



Answer: D

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39.

Q. Incorrect statement 'Y' changes on heating:

- A. Colour of 'Y' changes on heating
- B. Z' is anhydride of H_2CO_3
- C. Y' can react with NaOH
- D. Z' does not act as Lewis acid

Answer: D

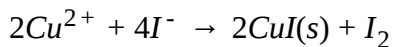
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40. The unique behaviour of Cu, having a positive E° (reduction potential) accounts for its inability to liberate H_2 from acids. Only oxidising acids (nitric acid and hot concentrated sulphuric acid) react with Cu. The high energy of transform Cu(s) to $Cu^{2+}(aq.)$ is not balanced

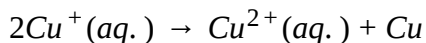
by its hydration enthalpy.

On the other hand, All Cu(II) halides are known except iodide. in this case,

Cu^{2+} oxidises I^- to I_2 :

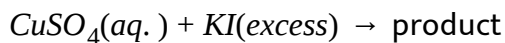


However, copper (I) compounds are unstable in aqueous solution and undergo disproportionation.



The stability of $Cu^{2+}(aq.)$ rather than $Cu^+(aq.)$ is due to the much more negative Δ_{Hyd} of $Cu^{2+}(aq.)$ than $Cu^+(aq.)$

Q. Consider the following transformation:



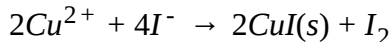
Select the correct statement:

- A. Product contains $[Cu(H_2O_4)]^{2+}$ ion.
- B. Presence of brown colouration in product is due to I_3^- ion
- C. Oxidation state of sulphur in reactant and product is different
- D. white ppt. of CuI_2 is observed in product

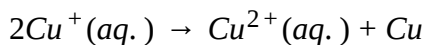
Answer: B

41. The unique behaviour of Cu, having a positive E° (reduction potential) accounts for its inability to liberate H_2 from acids. Only oxidising acids (nitric acid and hot concentrated sulphuric acid) react with Cu. The high energy of transform Cu(s) to $Cu^{2+}(aq.)$ is not balanced by its hydration enthalpy.

On the other hand, All Cu(II) halides are known except iodide. In this case, Cu^{2+} oxidises I^- to I_2 :

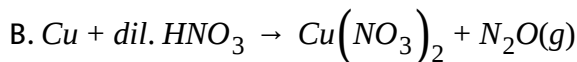
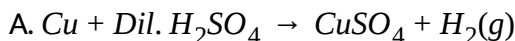


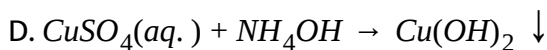
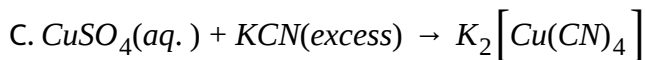
However, copper (I) compounds are unstable in aqueous solution and undergo disproportionation.



The stability of $Cu^{2+}(aq.)$ rather than $Cu^+(aq.)$ is due to the much more negative Δ_{Hyd} of $Cu^{2+}(aq.)$ than $Cu^+(aq.)$

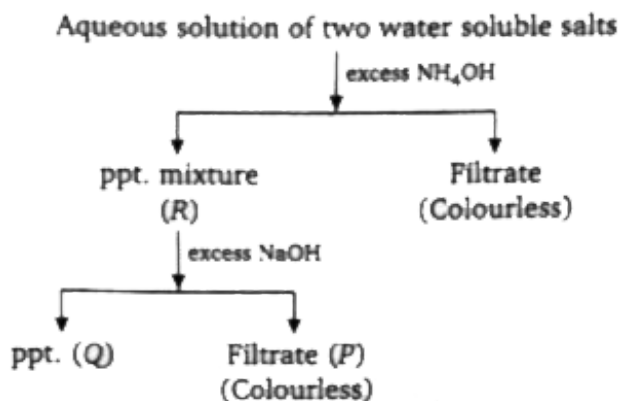
Q. Select the correct chemical change:





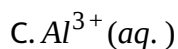
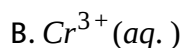
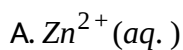
Answer: D

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42.

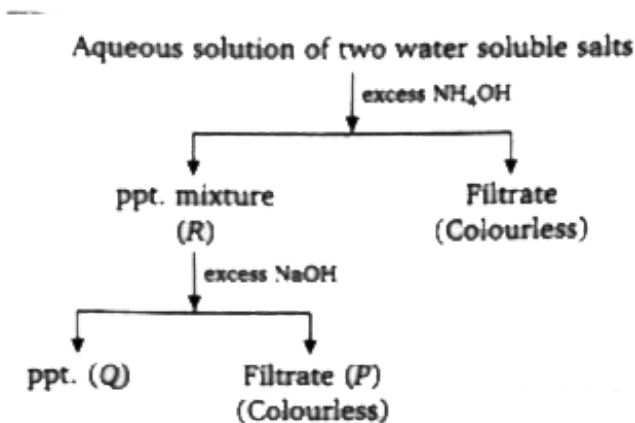
Q. When H_2S gas was passed into filtrate (P), a coloured precipitate was obtained, then cation present in the filtrate is:



D. $Pb^{2+}(aq.)$

Answer: D

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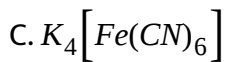


43.

Q. Precipitate (Q) was treated with dil. HCl and a coloured solution was obtained. On passing H_2S gas into this solution no precipitate was obtained but the colour of the solution changes, then the cation present in the precipitate (Q) can be identified by:

A. $Na_2S_2O_3$ solution

B. $KI + Starch$



D. All

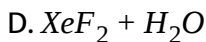
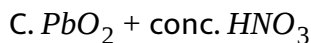
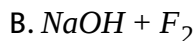
Answer: D



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ONE OR MORE ANSWERS IS/ARE CORRECT

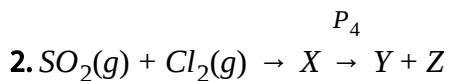
1. Which of the following combination of species can evolve O_2 ?



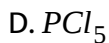
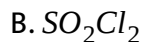
Answer: A::B::C::D



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then X, Y and Z can be :

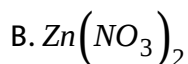
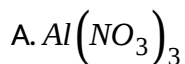


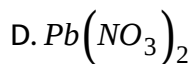
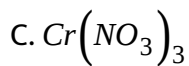
Answer: B::C::D



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3. Which of the following nitrate salt solution neither produce ppt. with excess NaOH nor with excess NH_4OH solution?

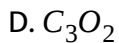
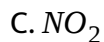
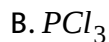
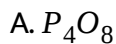




Answer: B::C

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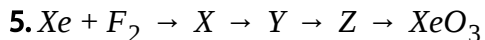
4. Which of the following compound(s) give two acids on dissolution in H_2O ?



Answer: A::B::C

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1:20 H_2O H_2O H_2O



Select correct option(s) for X, Y, Z and given chemical change:

- A. X, Y and Z are in same oxidation state
- B. All have equal number of lone pair on central atom
- C. All are non-planar
- D. All have equal number of covalent bonds

Answer: A::B::C::D



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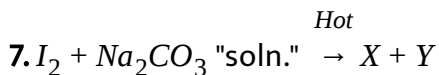
6. Which of the following sulphide(s) does/do not liberate H_2S on warming with dil. HCl?

- A. HgS
- B. ZnS
- C. FeS

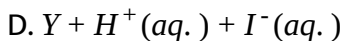
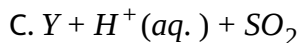
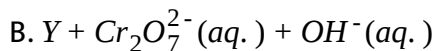
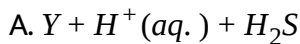
D. CuS

Answer: A::D

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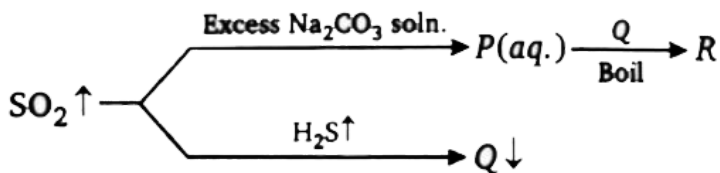
If 'X' gives coloured ppt. with $Pb(CH_3COO)_2$ solution, then 'Y' will respond to which of the following ?



Answer: A::C::D

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8.



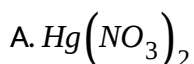
Incorrect statement about 'R' is

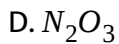
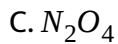
- A. Antichlor agent
- B. Fixing agent in photography
- C. Forms ppt. with CaCl_2 solution
- D. Reduces $\text{Cu}^{2+}(\text{aq})$ cation

Answer: C

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9. NO_2 gas evolves on thermal decomposition of which of the following compound(s)?

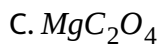
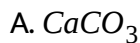




Answer: A::C::D

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10. Which of the following precipitate(s) is/are dissolved to colourless solution on adding sufficient amount of dilute HCl?



Answer: A::C

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11. Which of the following combination of reagent(s) produce observable change in aqueous medium?

- A. $Ba(OH)_2$ solution + $SO_2(g)$
- B. AgF solution + $NaNO_3$ solution
- C. $Pb(OAc)_2$ solution + Na_2CO_3 solution
- D. $CuCl_2$ solution + NH_3 (excess)

Answer: A::C::D



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12. Which of the following species is/are not liberating oxygen gas on reaction with water at $25^\circ C$?

- A. Na_2O_2
- B. Cl_2
- C. P_4

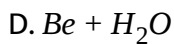
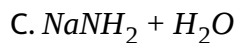
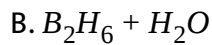
D. KO_2

Answer: B::C



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13. Hydrogen gas is not evolved by:

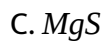
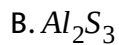
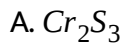


Answer: A::C::D



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14. Which of the following metal sulphide does not undergo hydrolysis?



Answer: D

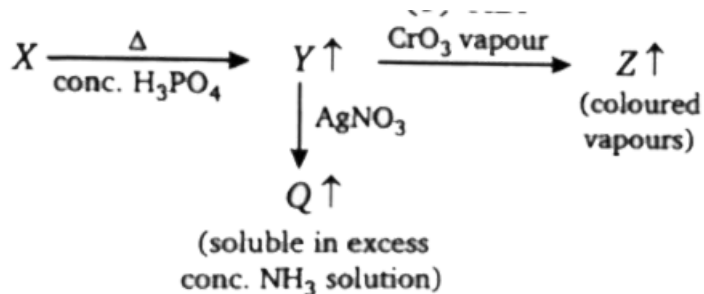
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15. Which of the following gas is not dried by conc. H_2SO_4 ?



Answer: B::C

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16.

Which of the following anion cannot be in X?

A. F^-

B. Cl^-

C. Br^-

D. I^-

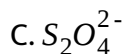
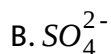
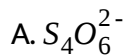
Answer: A::D



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17. When ozone reacts with an excess of potassium iodide solution buffered with a borate buffer (pH 9.2) iodine is liberated which can be

titrated against a standard solution of sodium thiosulphate, this is a quantitative method for estimating O_3 gas. when liberated I_2 and sodium thiosulphate will react, then product is/are:

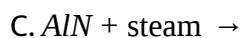
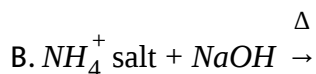
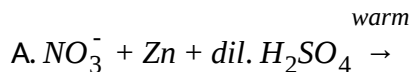


Answer: A



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18. In which of the following reactions NH_3 gas evolution occurs?

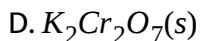
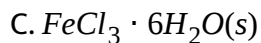
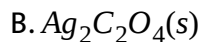




Answer: B::C::D

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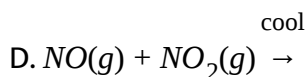
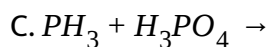
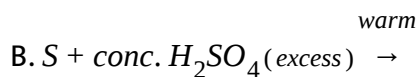
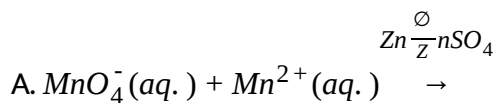
19. Which of the following compound() during heating undergo redox decomposition reaction?



Answer: A::B::D

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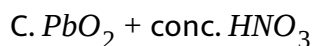
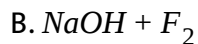
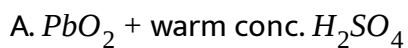
20. Which of the following combination of species undergo(es) comproportionation?

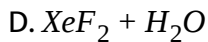


Answer: A::D

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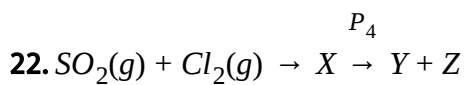
21. Which of the following combination of species can evolve O_2 ?



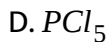
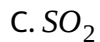
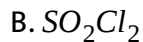


Answer: A::B::C::D

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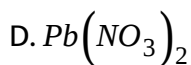
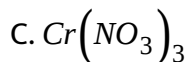
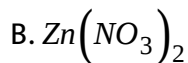
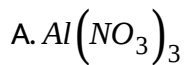
then X, Y and Z can be :



Answer: B::C::D

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23. Which of the following nitrate salt solution neither produce ppt. with excess NaOH nor with excess NH_4OH solution?

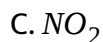
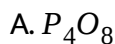


Answer: B::C



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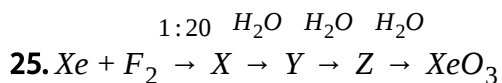
24. Which of the following compound(s) give two acids on dissolution in H_2O ?



D. C_3O_2

Answer: A::B::C

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Select correct option(s) for X, Y, Z and given chemical change:

- A. X, Y and Z are in same oxidation state
- B. All have equal number of lone pair on central atom
- C. All are non-planar
- D. All have equal number of covalent bonds

Answer: A::B::C::D

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26. Which of the following sulphide(s) does/do not liberate H_2S on warming with dil. HCl?

A. HgS

B. ZnS

C. FeS

D. CuS

Answer: A::D



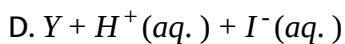
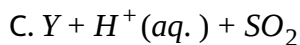
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27. $I_2 + Na_2CO_3$ "soln." \xrightarrow{Hot} $X + Y$

If 'X' gives coloured ppt. with $Pb(CH_3COO)_2$ solution, then 'Y' will respond to which of the following ?

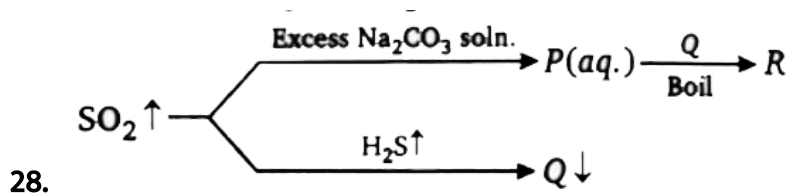
A. $Y + H^+(aq.) + H_2S$

B. $Y + Cr_2O_7^{2-}(aq.) + OH^-(aq.)$



Answer: A::B::D

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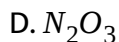
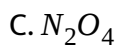
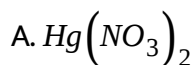
Incorrect statement about 'R' is

- A. Antichlor agent
- B. Fixing agent in photography
- C. Forms ppt. with $CaCl_2$ solution
- D. Reduces $Cu^{2+}(aq)$ cation

Answer: C

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29. NO_2 gas evolves on thermal decomposition of which of the following compound(s)?



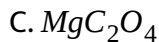
Answer: A::C::D



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30. Which of the following precipitate(s) is/are dissolved to colourless solution on adding sufficient amount of dilute HCl ?

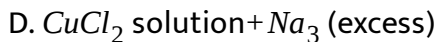
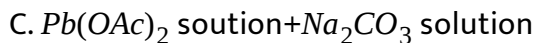
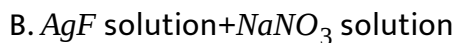
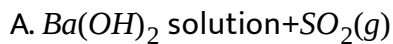




Answer: A::C

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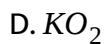
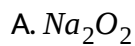
31. Which of the following combination of reagent(s) produce observable change in aqueous medium?



Answer: A::C::D

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32. Which of the following species is/are not liberating oxygen gas on reaction with water at 25°C ?

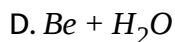
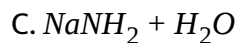
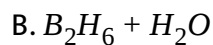


Answer: B::C



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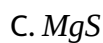
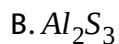
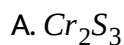
33. Hydrogen gas is not evolved by:



Answer: A::B::D

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34. Which of the following metal sulphide does not undergo hydrolysis?



Answer: D

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35. Which of the following gas is not dried by conc. H_2SO_4 ?



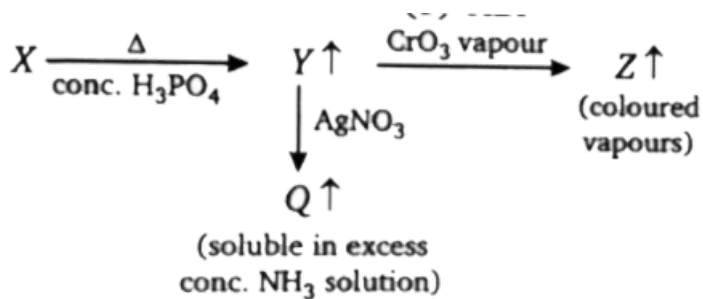
B. HBr

C. H_2S

D. SO_2

Answer: B::C

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36.

Which of the following anion cannot be in X?

A. F^-

B. Cl^-

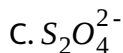
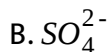
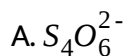
C. Br^-

D. I^-

Answer: A::D

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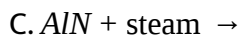
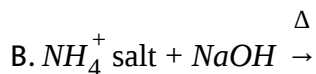
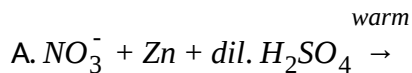
37. When ozone reacts with an excess of potassium iodide solution buffered with a borate buffer (pH 9.2) iodine is liberated which can be titrated against a standard solution of sodium thiosulphate, this is a quantitative method for estimating O_3 gas. when liberated I_2 and sodium thiosulphate will react, then product is/are:



Answer: A

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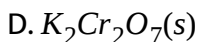
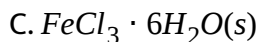
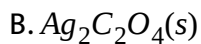
38. In which of the following reactions NH_3 gas evolution occurs?



Answer: B::C::D

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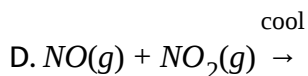
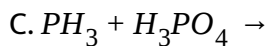
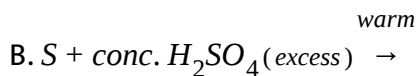
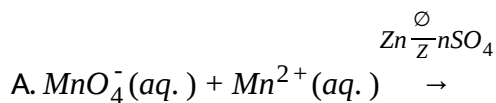
39. Which of the following compound() during heating undergo redox decomposition reaction?



Answer: A::B::D

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40. Which of the following combination of species undergo(es) comproportionation?



Answer: A::D

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MATCH THE COLUMN

Column-I
(Ionic Compounds)

- (A) HgCO_3
- (B) FeSO_4
- (C) $\text{BeCl}_2 \cdot \text{O}_2$
- (D) AgNO_3

Column-II

(Possible observations on thermal decomposition)

- (P) Acidic gas evolves
- (Q) Metallic residue is obtained as final product
- (R) Metal cation of salt undergoes redox reaction
- (S) Metallic oxide can be obtained
- (T) Neutral gas is evolved

1.

 [View Text Solution](#)

Column-I

- (A) $\text{Na}_2\text{S}_2\text{O}_3 + \text{dil. HCl}$
- (B) $\text{ICl}_3 + \text{H}_2\text{O}$
- (C) $\text{FeCl}_3 + \text{H}_2\text{S}/\text{H}^+$
- (D) $\text{H}_2\text{SO}_3 \xrightarrow{\Delta}$

Column-II

- (P) Disproportionation reaction
- (Q) Yellow ppt.
- (R) Redox reaction
- (S) One of the product gives white fumes with NH_3

2.

 [View Text Solution](#)

Column-I
(Halide compound)

- (A) PCl_3
- (B) NF_3
- (C) SbCl_3
- (D) BF_3

Column-II
(Characteristics)

- (P) Can act as π -acid ligand
- (Q) Final hydrolysed product is a proton donor oxyacid
- (R) Can act as classical/normal ligand
- (S) Undergoes partial hydrolysis
- (T) Final hydrolysed product has ($p\pi-p\pi$) bond

3.

 [View Text Solution](#)

**Column-I
(Anions)**

- (A) SO_3^{2-}
- (B) CO_3^{2-}
- (C) Cl^-
- (D) NO_2^-

**Column-II
[Reaction of anion(s) with dil.
 $\text{HCl}/\text{conc. H}_2\text{SO}_4$]**

- (P) Colourless volatile product is formed
- (Q) Coloured volatile product is formed
- (R) Volatile product forms precipitate with $\text{Ba}(\text{OH})_2$ solution
- (S) Volatile product forms precipitate with AgNO_3 solution
- (T) Formed volatile product decolourizes $\text{MnO}_4^-/\text{H}^+$ solution

4.

 [View Text Solution](#)

Column-I

(Reaction with Salt/Radical)

- (A) $\text{Zn} + \text{dil. H}_2\text{SO}_4$
- (B) dil. HCl
- (C) NaOH (excess)
- (D) KI

Column-II

(Salt/Radical)

- (P) $\text{Pb}(\text{NO}_2)_2$
- (Q) $(\text{NH}_4)_2\text{S}$
- (R) $\text{MnO}_4^-(\text{aq.})$
- (S) $\text{Hg}_2^{2+}(\text{aq.})$
- (T) $\text{Bi}^{3+}(\text{aq.})$

5.

 [View Text Solution](#)

Column-I
(Acidic Radicals)

- (A) S^{2-} (aq.)
- (B) SO_3^{2-} (aq.)
- (C) NO_2^- (aq.)
- (D) $S_2O_3^{2-}$ (aq.)

Column-II
(Observations)

- (P) Redox reaction with alkaline Br_2
- (Q) Evolution of diamagnetic gas with dil. HCl on warming
- (R) White ppt. with $Pb(CH_3COO)_2$ and ppt. remains white even after boiling
- (S) Evolution of gas with (Al + NaOH solution).
- (T) Evolution of same gas with dil. HCl as well as with conc. H_2SO_4 on warming

6.

 [View Text Solution](#)

Column-I

- (A) Undergoes hydrolysis via. S_N2 mechanism
- (B) Undergoes hydrolysis via. S_N1 mechanism
- (C) Hybridisation of central atom in transition state changes during hydrolysis
- (D) Proton donor oxy acid is formed as final hydrolysed product

Column-II

- (P) BCl_3
- (Q) NCl_3
- (R) SOF_2
- (S) $POCl_3$
- (T) ClF_3

7.

 [View Text Solution](#)

Column-I
(Ionic Compounds)

- (A) HgCO_3
- (B) FeSO_4
- (C) BeC_2O_4
- (D) AgNO_3

Column-II

(Possible observations on thermal decomposition)

- (P) Acidic gas evolves
- (Q) Metallic residue is obtained as final product
- (R) Metal cation of salt undergoes redox reaction
- (S) Metallic oxide can be obtained
- (T) Neutral gas is evolved

8.

 [View Text Solution](#)

Column-I

- (A) $\text{Na}_2\text{S}_2\text{O}_3 + \text{dil. HCl}$
- (B) $\text{ICl}_3 + \text{H}_2\text{O}$
- (C) $\text{FeCl}_3 + \text{H}_2\text{S}/\text{H}^+$
- (D) $\text{H}_2\text{SO}_3 \xrightarrow{\Delta}$

Column-II

- (P) Disproportionation reaction
- (Q) Yellow ppt.
- (R) Redox reaction
- (S) One of the product gives white fumes with NH_3

9.

 [View Text Solution](#)

Column-I
(Halide compound)

- (A) PCl_3
- (B) NF_3
- (C) SbCl_3
- (D) BF_3

Column-II
(Characteristics)

- (P) Can act as π -acid ligand
- (Q) Final hydrolysed product is a proton donor oxyacid
- (R) Can act as classical/normal ligand
- (S) Undergoes partial hydrolysis
- (T) Final hydrolysed product has ($p\pi-p\pi$) bond

10.

 [View Text Solution](#)

**Column-I
(Anions)**

- (A) SO_3^{2-}
- (B) CO_3^{2-}
- (C) Cl^-
- (D) NO_2^-

Column-II

[Reaction of anion(s) with dil. HCl/conc. H_2SO_4]

- (P) Colourless volatile product is formed
- (Q) Coloured volatile product is formed
- (R) Volatile product forms precipitate with $\text{Ba}(\text{OH})_2$ solution
- (S) Volatile product forms precipitate with AgNO_3 solution
- (T) Formed volatile product decolourizes $\text{MnO}_4^-/\text{H}^+$ solution

11.

 [View Text Solution](#)

**Column-I
(Reactions)**

- (A) $(\text{NH}_4)_2\text{Cr}_2\text{O}_7 \xrightarrow{\Delta}$
- (B) $\text{FeSO}_4 \xrightarrow{\Delta}$
- (C) $\text{Pb}(\text{NO}_3)_2 \xrightarrow{\Delta}$
- (D) $\text{P}_4 \xrightarrow[\Delta]{\text{NaOH}}$

Column-II

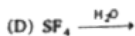
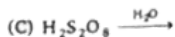
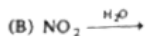
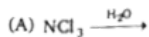
(Characteristics of any one product)

- (P) Amphoteric species
- (Q) Basic species
- (R) Non-polar gas
- (S) Polar acidic gas
- (T) Coloured residue

12.

 [View Text Solution](#)

Column-I
(Complete hydrolysis)



Column-II
(Characteristics of any hydrolysed product/hydrolysis)

(P) Dibasic acid

(Q) Can act as flexidentate ligand

(R) Can act as both oxidising and reducing agent

(S) Can act as monodentate ligand

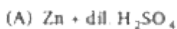
(T) Non-redox hydrolysis

13.

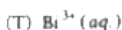
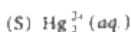


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Column-I
(Reaction with Salt/Radical)



Column-II
(Salt/Radical)



14.



[View Text Solution](#)

Column-I

(A) Disproportionation in alkaline medium

(B) Oxidizing agent

(C) Reacts with water

(D) Basic gas evolves on heating

Column-II



15.





Column-I

- (A) NO_2
- (B) SOF_2
- (C) XeF_4
- (D) ClF_5

Column-II

- (P) Hydrolysis occurs through redox reaction
- (Q) Hydrolysed product can undergo tautomeric change
- (R) All hydrolysed products are acids
- (S) Hybridization of central atom remains same in final hydrolysed product
- (T) One of the hydrolysed product reacts with glass

16.

Column-I
(Acidic Radicals)

- (A) $\text{S}^{2-}(\text{aq.})$
- (B) $\text{SO}_3^{2-}(\text{aq.})$
- (C) $\text{NO}_2^-(\text{aq.})$
- (D) $\text{S}_2\text{O}_3^{2-}(\text{aq.})$

Column-II
(Observations)

- (P) Redox reaction with alkaline Br_2
- (Q) Evolution of diamagnetic gas with dil. HCl on warming
- (R) White ppt. with $\text{Pb}(\text{CH}_3\text{COO})_2$ and ppt. remains white even after boiling
- (S) Evolution of gas with $(\text{Al} + \text{NaOH}$ solution).
- (T) Evolution of same gas with dil. HCl as well as with conc. H_2SO_4 on warming

17.

Column-I	Column-II
(A) Undergoes hydrolysis via S_N2 mechanism	(P) BCl_3
(B) Undergoes hydrolysis via S_{NA} mechanism	(Q) NCl_3
(C) Hybridisation of central atom in transition state changes during hydrolysis	(R) SOF_2
(D) Proton donor oxy acid is formed as final hydrolysed product	(S) $POCl_3$
	(T) ClF_3

18.

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SUBJECTIVE PROBLEMS

1. Find total number of reagents which can produce I_2 from KI solution.

Conc. H_2SO_4 , $Hg(NO_3)_2$ solution, $CuSO_4$ solution, Conc. H_3PO_4 ,

$K_2Cr_2O_7/H^+$, Cl_2 Water, $Pb(CH_3COO)_2$ solution, $Ca(OCl)Cl/H^+$, $NaNO$

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2. Find total number of metal cations which are ppted as metal sulphide on passing H_2S gas through metal salt solution.

$Pb^{2+}(aq.)$, $Mn^{2+}(aq.)$, $Sn^{2+}(aq.)$, $Cr^{3+}(aq.)$, $Mg^{2+}(aq.)$, $Hg^{2+}(aq.)$

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3. Consider the following reaction $P_4 + KOH \rightarrow PH_3 + X$

How many P-H bonds are present in species X?

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4. Which of the following species/reagent can reduce $Fe^{3+}(aq.)$ into $Fe^{2+}(aq.)$ at normal conditions?

$(NH_4)_2S$, HI , $Sn^{2+}(aq.)$, $CN^-(aq.)$, $NaNO_2$, SO_2 , $Na_2S_2O_3$, SCN^-

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5. Find out number ionic compound(s) which is/are water insoluble at room temperature

$BaSO_4$, $AgNO_3$, $PbCO_3$, $CaCl_2$, $Mg(OH)_2$, $KMnO_4$, CH_3COOAg , C

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6. Find the value of expression $|x-y|$ for following compounds.

where,

x = total number of water insoluble salts.

y = total number of salts, which can liberate non-olar acidic gas during their complete thermal decomposition.

$BaCO_3$, $PbSO_4$, $AgNO_3$, CaC_2O_4 , $CsHCO_3$, Na_3PO_4 , CH_3COOAg ,

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7. Find out total number of coloured compound(s) from following:

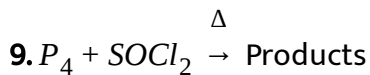
$BaCO_3$, HgO , $PbSO_4$, Ag_2S , HgI_2 , PbO , CdS , $AgNO_2$, $PbCrO_4$

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8. Find out total number of cation(s) that produce precipitate with aqueous solution of Na_2CO_3 .

$\text{Cu}^{2+}(\text{aq.}), \text{Mg}^{2+}(\text{aq.}), \text{Fe}^{3+}(\text{aq.}), \text{Pb}^{2+}(\text{aq.}), \text{Al}^{3+}(\text{aq.}), \text{Hg}^{2+}(\text{aq.}), \text{Zn}^{2+}(\text{aq.})$

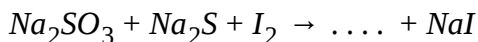
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Find out total number of non-planar and polar molecules of products in balanced equation for one mole of P_4 .

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10. What is average oxidation state state of sulphur in product formed in given reaction?



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11. find out total number of coloured/black water insoluble compound(s) from following substances:

$Ag_2O, HgI_2, FeS, Ag_3PO_4, Ba(MnO_4)_2, Na_2CrO_4, PbI_2, AgNO_2, Ag_2C_2O_4$

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12. Find out total number of compounds which on heating undergo redox reactions.

$PbCl_4, Mg(NO_3)_2, HgC_2O_4, Ag_2CO_3, Pb(CN)_4, Al(OH)_3, Cu(CN)_2$

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13. How many following Ammonium salts will evolve N_2 gas on heating?

$(NH_4)_2CO_3, (NH_4)_2Cr_2O_7, NH_4NO_2, NH_4ClO_4, NH_4Cl, (NH_4)_2S, (NH_4)_2C_2O_4$

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14. How many following metals evolve NO (Nitric oxide) gas with dil. HNO_3 (20%)?

$Hg, Cu, Pb, Zn, Fe, Al, Ag, Au, Mn$

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15. Find number of basic radicals among the following cations, which can form soluble complex on adding excess of NH_4 solution.

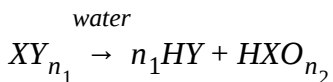
$Cd^{2+}(aq.)$, $Pb^{2+}(aq.)$, $Ni^{2+}(aq.)$, $Mn^{2+}(aq.)$, $Zn^{2+}(aq.)$, $Ag^+(aq.)$, $Hg^{2+}(aq.)$

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16. Calculate difference between oxidation state of chromium (Cr) in blue and green coloured chromium species formed during the following given transformation.

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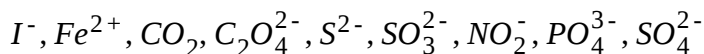
17. If hydrolysis of interhalogen compound can be represented by following general reaction:



If given interhalogen compound is polarr and non-planar, then calculate value of $n_1 + n_2$.

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18. Total number of species that can be oxidized by acidic permanganate ion (MnO_4^-/H^+).



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19. How many following metals evolve N_2O gas with dil. HNO_3 (20%)

Cr, Cu, Pb, Zn, Fe, Al, Ag, Au, Mn.

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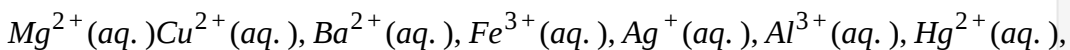
20. How many following ammonium salts will evolve NH_3 gas on heating?





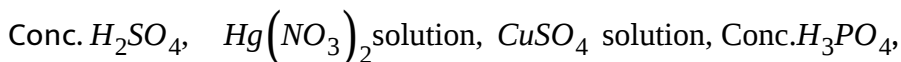
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21. Find out the number of cation (s) which form(s) black ppt. (soluble in hot and dilute HNO_3) on passing H_2S gas into their salt solution?



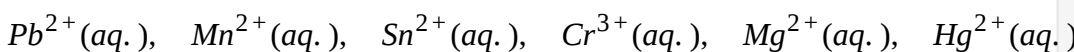
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22. Find total number of reagents which can produce I_2 from KI solution.



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23. Find total number of metal cations which are ppted as metal sulphide on passing H_2S gas through metal salt solution.



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24. Consider the following reaction $P_4 + KOH \rightarrow PH_3 + X$

How many P-H bonds are present in species X?

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25. Which of the following species/reagent can reduce $Fe^{3+}(aq.)$ into $Fe^{2+}(aq.)$ at normal conditions?

$(NH_4)_2S$, HI , $Sn^{2+}(aq.)$, $CN^-(aq.)$, $NaNO_2$, SO_2 , $Na_2S_2O_3$, SCN^-

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26. Find out number ionic compound(s) which is/are water insoluble at room temperature

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27. Find the value of expression $|x-y|$ for following compounds.

where,

x = total number of water insoluble salts.

y = total number of salts, which can liberate non-olar acidic gas during their complete thermal decomposition.

$BaCO_3$, $PbSO_4$, $AgNO_3$, CaC_2O_4 , $CsHCO_3$, Na_3PO_4 , CH_3COOAg ,

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28. Find out total number of coloured compound(s) from following:

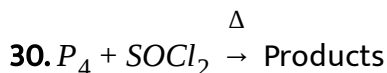
$BaCO_3$, HgO , $PbSO_4$, Ag_2S , HgI_2 , PbO , CdS , $AgNO_2$, $PbCrO_4$

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29. Find out total number of cation(s) that produce precipitate with aqueous solution of Na_2CO_3 .

$Cu^{2+}(aq.)$, $Mg^{2+}(aq.)$, $Fe^{3+}(aq.)$, $Pb^{2+}(aq.)$, $Al^{3+}(aq.)$, $Hg^{2+}(aq.)$, $Zn^{2+}(aq.)$

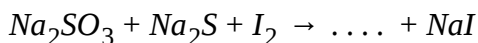
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Find out total number of non-planar and polar molecules of products in balanced equation for one mole of P_4 .

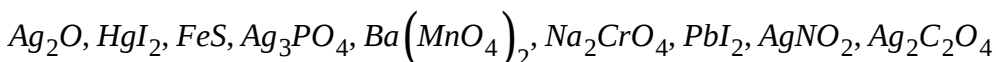
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31. What is average oxidation state state of sulphur in product formed in given reaction?



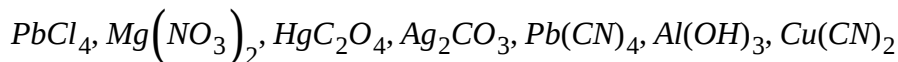
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32. find out total number of coloured/black water insoluble compound(s) from following substances:



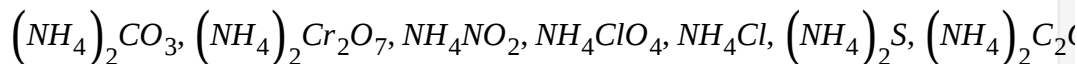
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33. Find out total number of compounds which on heating undergo redox reactions.



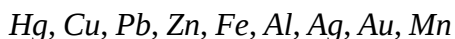
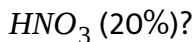
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34. How many following Ammonium salts will evolve N_2 gas on heating?



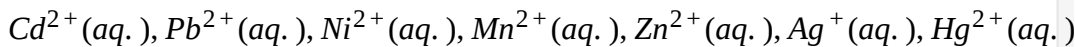
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35. How many following metals evolve NO (Nitric oxide) gas with dil.



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36. Find number of basic radicals among the following cations, which can form soluble complex on adding excess of NH_4 solution.

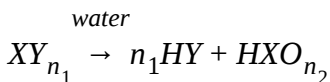


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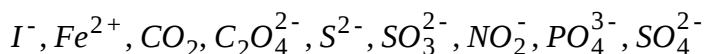
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40. How many following metals evolve N_2O gas with dil. HNO_3 (20%)

Cr, Cu, Pb, Zn, Fe, Al, Ag, Au, Mn.

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41. How many following ammonium salts will evolve NH_3 gas on heating?



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42. Find out the number of cation (s) which form(s) black ppt. (soluble in hot and dilute HNO_3) on passing H_2S gas into their salt solution?

$Mg^{2+}(aq.)$, $Cu^{2+}(aq.)$, $Ba^{2+}(aq.)$, $Fe^{3+}(aq.)$, $Ag^+(aq.)$, $Al^{3+}(aq.)$, $Hg^{2+}(aq.)$,



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