



CHEMISTRY

BOOKS - TARGET CHEMISTRY (HINGLISH)

S-BLOCK ELEMENTS

Question

1. The long form of the periodle table is divided into _____ blocks

A. two

B. three

C. four

D. five

Answer: C



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2. Group 1 and group 2 elements are placed at the _____ of the long form of the modern periodic table .

A. right

B. left

C. middle

D. bottom

Answer: B



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3. The elements of group 1 are called alkali metals because

- A. They have one electron in their outermost orbital
- B. they have low melting point
- C. the solutions of their salts basic
- D. their oxides and hydroxides combine with water to form alkali

Answer: D



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4. Which of the following is NOT an alkali metal ?

A. Caesium

B. Strontium

C. Rubidium

D. Potassium

Answer: B



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5. Which of the following is the most alkaline earth metal ?

A. Beryllium

B. Magnesium

C. Calcium

D. Radium

Answer: D



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6. Amongst alkali metal cations ,_____ has the smallest size .

A. *Li*

B. *Ba*

C. *Cs*

D. *Ra*

Answer: A



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7. Which of the following is responsible for the solubility of lithium compounds in organic solvents ?

A. Covalent character of lithium compounds.

B. Ionic character of lithium compounds.

C. Basic character of lithium compounds.

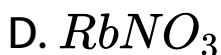
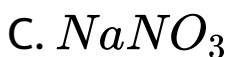
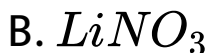
D. All of these

Answer: A



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8. Which of the following alkali metal nitrate on decomposition gives nitrogen dioxide as one of the product ?



Answer: B



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9. Which of the following compounds is NOT obtained in the solid form ?

- A. Calcium hydrogen carbonate
- B. potassium hydrogen carbonate
- C. Sodium hydrogen carbonate
- D. Lithium hydrogen carbonate

Answer: D



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10. The similarities between an element of a group and the element placed in the next period of the succeeding group is called ___.

A. Vertical relationship

B. horizontal relationship

C. diagonal relationship

D. periodic trend

Answer: C



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11. The metal showing resemblance with lithium is _____

A. Be

B. Mg

C. Al

D. Na

Answer: B



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12. Atomic radius of Li is _____

A. 152 pm

B. 265 pm

C. 376 pm

D. 520 pm

Answer: A



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13. Identify the CORRECT order of decreasing ionization enthalpies of alkali metals



D.

Answer: A



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14. The first ionization enthalpies of alkali metals decreases down the group because of _____.

- A. increase in atomic number
- B. increase in atomic size
- C. increase in screening effect
- D. all of these

Answer: D



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15. Sodium reacts with oxygen to form _____

A. Oxide

B. Peroxide

C. Superoxide

D. epoxide

Answer: B



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16. Alkali metals react oxygen to form _____

A. nitrogen gas

B. oxygen gas

C. chlorine gas

D. hydrogen gas

Answer: D



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17. Solvay process is also known as _____

A. soda process

B. caustic soda process

C. Brine process

D. ammonia soda process

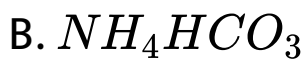
Answer: D



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18. In the solvay process, NH_4Cl is treated with milk of lime to recover _____.

A. $NaCl$



Answer: C



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19. Which of the following is used in qualitative as well as quantitative analysis ?



B. Na_2CO_3

C. $NaHCO_3$

D. CO_2

Answer: B



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20. The gases liberated at anode in the electrolysis of sodium acetate are

A. hydrogen

B. chlorine

C. oxygen

D. nitrogen

Answer: B



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21. The compound which is NOT present as impurity in crude NaCl is ____

A. $CaSO_4$

B. $CaCl_2$

C. $MgCl_2$

D. HCl

Answer: D



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22. Saturated solution of crude $NaCl \xrightarrow{x}$

Pure $NaCl$. In the reaction, X is ___

A. HCl gas

B. NHO_3 gas

C. NH_3 gas

D. H_2SO_4 gas

Answer: A



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23. NaCl when heated with conc :

H_2SO_4 and MnO_2 liberates _____

A. hydrogen gas

B. chlorine gas

C. Carbon dioxide gas

D. hydrogen chloride gas

Answer: B



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24. The aqueous solution of $NaHCO_3$ is _____

A. alkaline

B. neutral

C. acidic

D. amphoteric

Answer: A



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25. _____ is used as a reducing agent in the extraction of B and Si.

A. Potassium

B. Calcium

C. Sodium

D. Lithium

Answer: C



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26. ____ plays an important role in Wurtz synthesis .

A. Li

B. Na

C. K

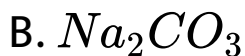
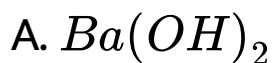
D. Cs

Answer: B



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27. _____ is used in the manufacture of soft soap .



C. $LiOH$

D. KOH

Answer: D



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28. _____ is used as a source of oxygen in submarine , space shuttles and in emergency breathing apparatus .

A. Sodium oxide

B. Potassium

C. Sodium sulphate

D. Potassium superoxide

Answer: D



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29. A typical 70 kg man contains about _____ of Na .

A. 70 g

B. 80 g

C. 90 g

D. 170 g

Answer: C



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30. Which of the following is CORRECT regarding human blood plasma ?

A. High concentration of K^+

B. Low concentration of K^+

C. Low concentration of K^+

D. None of these

Answer: B



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31. How does tubular secretion help in maintaining ionic and acid-base balance in body-fluids?

A. Carbonic acid

B. Carbonate

C. $Na^+ - K^+$ pump

D. Biscarbonate

Answer: D



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32. Alkaline earth metals belong to the

A. S- block in periodic table

B. p - block in periodic table

C. d- block in periodic table

D.

Answer: A



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33. The outer electronic configuration of alkaline earth metals is _____.

A. ns^2

B. ns^1

C. np^6

D. nd^{10}

Answer: A



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34. which of the following element forms compounds that are partly ionic and partly covalent ?

A. Ca

B. Sr

C. Ba

D. Mg

Answer: D



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35. The reaction of powdered Be with N_2 produces _____.

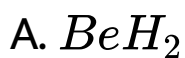


Answer: A



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36. Hydrolith is _____ .



B. CaH_2

C. BaH_2

D. $BeCl_2$

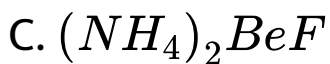
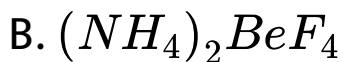
Answer: B



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37. BeF_2 is prepared in the presence of CO_2 by thermal decomposition of ___

A. $(NH_4)BeF_2$

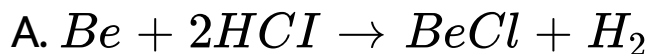


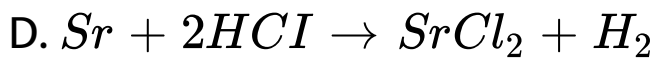
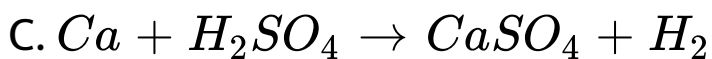
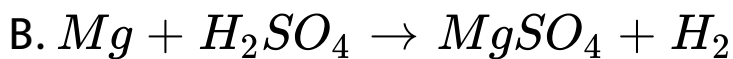
Answer: B



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38. Which of the following reaction do NOT occur ?





Answer: A



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39. Metallic _____ is used for making windows of X ray tubes .

A. Magnesium

B. beryllium

C. sulphur

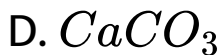
D. Magnalium

Answer: B



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40. Which of the following compound is used to scan the X - ray of the human digestive system ?



Answer: C



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41. Beryllium carbide on hydrolysis gives _____

A. Methane

B. ethane

C. beryllium

D. carbon dioxide

Answer: A



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42. Quick lime is _____

A. $Ca(OH)_2$

B. $CaCO_3$

C. CaO

D. $CaSO_4$

Answer: C



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43. which of the following is TRUE about commercial production of calcium oxide ?

A. Staked lime is used as reactant

B. the temperature is maintained at
around 773 K .

C. As soon as the byproduct is formed , is
removed to enable the reaction to proceed
to completion

D. All of these

Answer: C



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44. Slaking of lime is a / an _____

- A. exothermic reaction
- B. endothermic reaction
- C. substitution reaction
- D. elimination reaction

Answer: A



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45. $CaCO_3$ reacts with dilute acid to liberate

_____.

A. CO

B. CO_2

C. CO_3^{2-}

D. Cl_2

Answer: B



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46. Which of the following is INCORRECT about calcium carbonate ?

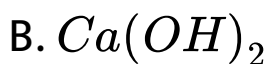
- A. It occurs in several forms like chalk ,
limestone , calcites , etc .
- B. It is a white fluffy power ,
- C. It is readily soluble in water
- D. It is used in the manufacture of quick
lime.

Answer: C



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47. _____ is used as mild abrasive in tooth paste .



Answer: C



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48. Deficiency of calcium causes _____.

A. Rickets

B. convulsion

C. hypertension

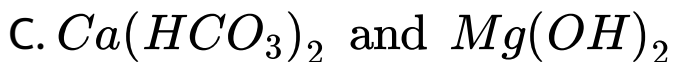
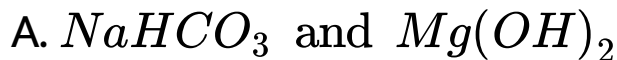
D. circulatory failure

Answer: A



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49. The pair whose both species are used in acid medicinal preparation is:



Answer: A



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50. Which of the following concerning is FALSE ?

A. Lithium and lithium nitrate decomposes to give the oxide on heating .

B. Lithium is the most electronegative element of group 1.

C. Lithium hydrogen carbonate cannot be isolated as a stable solid .

D. Lithium forms a peroxide but not a superoxide .

Answer: D



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51. which of the following CANNOT form peroxide ?

A. Cs

B. Mg

C. Be

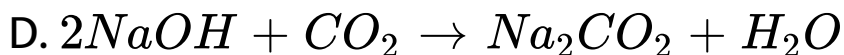
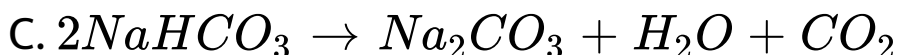
D. All of these

Answer: D



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52. Which one of the following reactions is not associated with the Solvay process of manufacture of sodium carbonate ?



Answer: D



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53. IN alkaline earth metals ,_____

- A. the cation is less stable than the atom
- B. the cation is smaller than the atom
- C. the cation and the atom have about the same size
- D. the cation is larger than the atom

Answer: B



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54. Which of the following does NOT belong to alkaline earth metals family ?

A. Mg

B. Ba

C. Rb

D. Ra

Answer: C



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55. _____ is useful in purification of bauxite .

A. Sodium carbonate

B. Sodium hydroxide

C. Sodium chloride

D. Sodium hydrogen carbonate

Answer: B



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56. Which of the following element belongs to alkaline earth metals ?

A. Boron

B. Carbon

C. Aluminium

D. Barium

Answer: D



57. Which of the following is the most abundant salt of sodium ?

A. Sodium phosphate

B. Sodium sulphate

C. sodium chloride

D. Sodium nitride

Answer: C



58. $AlCl_3$ is used as a catalyst in _____.

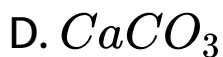
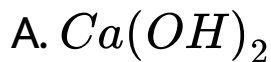
- A. Wurtz synthesis
- B. Saponification of esters
- C. Williamson synthesis
- D. Friedel-Crafts reaction

Answer: D



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59. _____ has wide application in purification of sugar .



Answer: C



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60. In the periodic table , the element with atomic number 38 belongs to _____

A. Fourth period and group 2

B. Fourth period and group 4

C. Third period and group 4

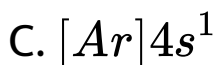
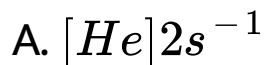
D. fifth period and group 2

Answer: D



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61. Electronic configuration of potassium with respect to nearest noble gas is



Answer: C



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62. The common oxidation state for alkaline earth metals is

A. +4

B. +1

C. +2

D. +3

Answer: C



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63. Fr has half life period of

A. 21 secs

B. 21 minutes

C. 21 hours

D. 21 days

Answer: B



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64. Which of the following statements is INCORRECT ?

A. Alkali metals form monovalent M^+ ion

B. General electrons condifuration of outermost shell of alkaline earth metals is ns^2

C. In Alkali metals , the differentiating electron enters ns orbital of the atom .

D. Alkaline earth metals occur in nature in free state .

Answer: D



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65. Chloride of metal Y was found to be deliquescent and on crystallisation formed a deliquate .The metal Y is _____.

A. Na

B. Li

C. K

D. either (B) or (C)

Answer: B



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66. Which of the following hydroxide is a weak base ?

A. LiOH

B. KOH

C. NaOH

D. CsOH

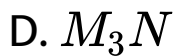
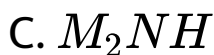
Answer: A



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67. Alkali metals (other than lithium) form amides of the general formula _____.

A. MNH



Answer: B



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68. Which of the following is FALSE ?

A. LiCl is soluble in ethanol

B. MgO combines with excess of oxygen .

C. Carbonates of lithium and magnesium decompose on heating to form their corresponding oxides and CO_2

D. Magnesium does not form solid hydrogen carbonate

Answer: B



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69. Alkali metal do NOT form dipositive ions because they have high _.

A. Nuclear change

B. electron affinity

C. Second ionization ebnthalpy

D. Third ioziation enthalpy

Answer: C



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70. Alkali metals tarnish in dry air because _____.

A. they react with oxygen to form corresponding oxides

B. they react with nitrogen to form corresponding nitrides

C. they hydrogen to form corresponding hydrides

D. they react with halogens to form corresponding halides

Answer: A



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71. Above 373 K , washing soda changes to a white power called _____.

- A. Baking soda
- B. hydrous washing soda
- C. charcoal ash
- D. soda ash

Answer: D



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72. In Castner -Kellner used for preparation of sodium carbonate ,_____ acts as an anode .

- A. Nickel rod
- B. graphite rod
- C. copper rod
- D. zinc rod

Answer: B



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73. The substance used pure fats and oils is

_____.

- A. Sodium carbonate
- B. sodium chloride
- C. sodium hydroxide
- D. sodium sulphate

Answer: C



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74. Which gas is evolved when $NaHCO_3$ is heated ?

A. CO

B. CO_2

C. NO_2

D. O_2

Answer: B



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75. Toughened bearings are made using ____ alloy

A. *Na – Pb*

B. *Li – pb*

C. *Li – Al*

D. *Li – Mg*

Answer: B



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76. _____ is an excellent absorbent of carbon dioxide .

A. $NaOH$

B. KCl

C. KOH

D. $KHCO_3$

Answer: C



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77. the material used in photoelectric cells contains ,

A. Cs

B. Li

C. Be

D. Mg

Answer: A



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78. Sodium and potassium ions maintain ____ .

A. constant vapour pressure on either side
of cell wall

B. constant osmotic pressure on either
side of cell wall

C. the sensitivity of nerves

D. both (B) and (C)

Answer: D



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79. Which of the following is FALSE ?

A. sodium ion is the major cation in the blood plasma of vertebrates .

B. Na^+ ions depress the activity of muscle enzyme .

C. K^+ ions permit the heart muscles to relax between the beats.

D. Na^+ is especially an enzyme activator .

Answer: D

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80. In humans ,NaCl is a source of _____ for gastric juices .

A. Polynucleic acid

B. chlorine gas

C. Hydrochloric acid

D. protein

Answer: C



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81. Which of the follows indicates deficiency of potassium in human body ?

A. Paralysis of muscles

B. Retarded bone growth

C. Reduction in fat growth .

D. Low blood pressure

Answer: A



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82. The alkaline earth metals *Ba*, *Sr*, *Ca* and *Mg* may be arranged in the order of their decreasing first ionization enthalpy as ____.

A. *Mg, Ca, Sr, Ba*

B. *Ca, Sr, Ba, Mg*

C. *Sr, Ba, Mg, Ca*

D. *Ba, Mg, Ca, Sr*

Answer: A



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83. The first ionization enthalpies of alkaline earth metals are higher than corresponding alkali metals due to their _____.

A. large size

B. Density

C. Small size

D. Atomic mass

Answer: C



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84. Which of the following alkaline earth metal does NOT react with water ?

A. Be

B. Mg

C. Ca

D. Sr

Answer: A



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85. BeH_2 is prepared reducing $BeCl_2$ with
_____.

A. HCl

B. $LiAlH_4$

C. H_2

D. $NaBH_4$

Answer: B



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86. Which of the following statements is INCORRECT ?

A. The increasing order of the affinity of alkaline earth metals towards oxygen is



B. The increasing order of reactivity of alkaline earth metals with water is



C. All the hydrides of alkaline earth metals are ionic solids .

D. All the alkaline earth metals combine with halogens at high temperature to

form corresponding halides .

Answer: C



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87. An oxide of alkaline earth metal on heating with carbon in the presence of chlorine gives $BeCl_2$. The oxide is _____.

A. BeO

B. BeO_2

C. Be_2O

D. Be_2O_2

Answer: A



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88. which of the following has highest melting and boiling point ?

A. Ba

B. Sr

C. Mg

D. Be

Answer: D



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89. Oxides and hydroxides of alkaline earth metals EXCEPT beryllium are _____ in nature .

A. Acidic

B. basic

C. amphoteric

D. neutral

Answer: B



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90. An alloy containing 1- 15 % Mg and 85-99%

Al is known as _____ .

A. elektron

B. Magnalium

C. alnico

D. Duralumin

Answer: B



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91. Which of the following metal is used to remove sulphur from petroleum ?

A. Beryllium

B. Calcium

C. Barium

D. Magnesium

Answer: B



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92. An element X is used in the extraction of metal from their oxides where carbon is ineffective. The element X is _____.

A. Magnesium

B. Beryllium

C. Barium

D. Calcium

Answer: D



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93. The first member of the group -2- elements exhibits diagonal relationship with ___ of the modern periodic table .

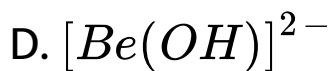
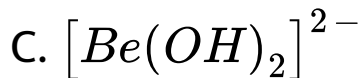
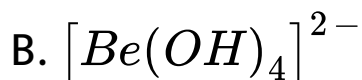
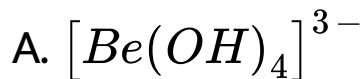
- A. the first member of group -3
- B. the first member of group -13
- C. the first member of group - 13
- D. the second member of group -14

Answer: C



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94. which of the following represents beryllate ion ?



Answer: B



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95. Quick lime is a/ an _____.

A. Basic oxide

B. acidic oxide

C. amphoteric oxide

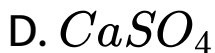
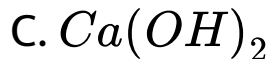
D. neutral oxide

Answer: A



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96. Compound of calcium (X) is the cheapest form of alkali, which can be decomposed to give back X at very high temperature. X is _____.



Answer: B



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97. Which compound has low lattice energy ?



B. KH

C. NaH

D. CsH

Answer: D



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98. Photoesctric effct is maximum in _____.

A. Cs

B. Na

C. k

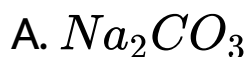
D. Li

Answer: A



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99. On prolonged exposure to air, sodium finally change to :



C. $NaOH$

D. $NaHCO_3$

Answer: A



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100. $LiAlH_4$ is used as _____.

A. an oxidising agent

B. a reducing agent

C. a moderator

D. a water softener

Answer: B



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101. The valence shell electronic configuration of *Rb* is _____.

A. $3s^1$

B. $4s^1$

C. $5s^1$

D. $7s^1$

Answer: C



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102. The valence shell electronic configuration of alkali metals is

A. ns^2np^1

B. ns^1

C. $(n - 1)p^6ns^2$

D. $(n - 1)d^2ns^2$

Answer: B



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103. Which of the following does not illustrate the anomalous properties of lithium?

A. The melting point and boiling point of Li are comparatively high .

B. Li is much softer than the other group 1 metals .

C. Li forms a nitride Li_3N unlike other group 1 metals .

D. the ion of Li and its compounds are more heavily hydrated than those of the rest of the group .

Answer: B



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104. Lithium shows similarities to magnesium in its chemical behaviour because.

- A. Similar size , greater electronegativity and similar polarizing power
- B. Similar size , same electronegativity and lower polarizing power
- C. similar size , same electronegativity and similar high polarizing power
- D. None of these

Answer: C



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105. Alkali metals lose electrons in _____.

- A. s-orbitals
- B. p- orbitals
- C. d-orbitals
- D. f- orbitals

Answer: A



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106. Which one of the following properties of alkali metals increases in magnitude as the atomic number rises ?

- A. Electronegativity
- B. Ionic radius
- C. First ionization energy
- D. Melting point

Answer: B



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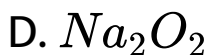
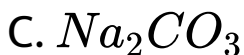
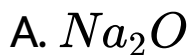
107. Which of the following properties is not true for an alkali metal ?

- A. Low atomic volume
- B. Low ionization energy
- C. Low density
- D. Low electronegativity

Answer: A



108. When sodium is heated with moist air, then the product obtained is



Answer: B



109. Which one of the alkali metals forms only the normal oxide, M_2O , on heating in air ?

A. Rb

B. K

C. Li

D. N

Answer: C



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110. Potassium is kept in

A. alcohol

B. water

C. kerosene

D. liquid ammonia

Answer: C



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111. The alkali metal that reacts with nitrogen directly to form nitride is

A. Li

B. Na

C. K

D. Rb

Answer: A



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112. Which of the following reacts with water with high rate ?

A. Li

B. K

C. Na

D. Rb

Answer: D



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113. The alkali metals form salt like hydrides by the direct synthesis at elevated temperature. The thermal stability of these hydrides decreases in which of the following orders ?



Answer: B



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114. The commercial production of sodium carbonate is done by _____.

- A. Lead - chamber process
- B. Haber's process
- C. Solvay process
- D. Castner's process

Answer: C



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115. Sodium carbonate is manufactured by Solvay process, the products that are recycled are

A. CO_2 and NH_3

B. CO_2 and NH_4Cl

C. $NaCl$ and CaO

D. $CaCl_2$ and CaO

Answer: A



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116. Na_2CO_3 can be manufactured by Solva's process but K_2CO_3 cannot be prepared because

A. K_2CO_3 is more soluble

B. K_2CO_3 is less soluble

C. $KHCO_3$ is more soluble than

$NaHCO_3$

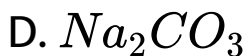
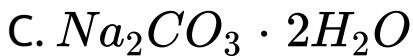
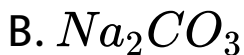
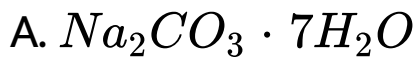
D. $KHCO_3$ is less soluble than $NaHCO_3$

Answer: C



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117. Washing soda is _____



Answer: B



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118. When sodium chloride solution is electrolysed, the gas that is liberated at the cathode is _____.

A. oxygen

B. hydrogen

C. chlorine

D. Air

Answer: B



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119. Molten sodium is used in nuclear reactors to _____.

A. Absorb neutrons in order to control the chain reaction

B. Slow down the fast neutrons

C. absorb the heat generated by nuclear fission

D. extract radio -isotopes produced in the reactor

Answer: C



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120. Excess of Na^+ ions in our system causes

A. High B.P

B. Low B.P

C. diabetes

D. anaemia

Answer: A



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121. Which of the following can be represented by the configuration $[Kr]5s^2$?

A. Ca

B. Sr

C. Ba

D. Ra

Answer: B



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122. The most electropositive amongst the alkaline earth metals is:

A. Beryllium

B. Magnesium

C. calcium

D. Barium

Answer: D



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123. Among the alkaline earth metals, the element forming predominantly covalent compound is

A. Be

B. Mg

C. Sr

D. Ca

Answer: A



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124. Which of the following is formed when calcium combines with oxygen ?



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125. Which among the following elements of groups-2 exhibits anomalous properties ?

A. Be

B. Mg

C. Ca

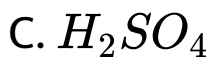
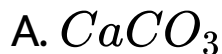
D. Ba

Answer: A



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126. A substance absorbs CO_2 and violently reacts with water. The substance is

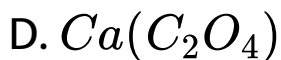
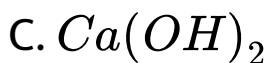
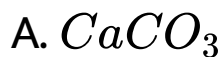


Answer: B



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127. Slaked lime is obtained when water is added to



Answer: C



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128. Bleaching powder is obtained by the action of chlorine gas and

A. conc.solution of $Ca(OH)_2$

B. dilute solution of $Ca(OH)_2$

C. Dry calcium oxide

D. dry slaked lime

Answer: D



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129. Which of the following metal is present in green colouring pigment chlorophyll of plants ?

A. *Fe*

B. *Mg*

C. *Na*

D. *Al*

Answer: B



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130. Which metal plays an important role in neuromuscular functions , internuronal transimission cell membrane integrity and blood coagulation ?

A. *Be*

B. *Mg*

C. *Ca*

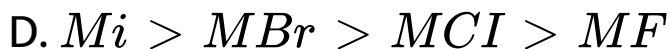
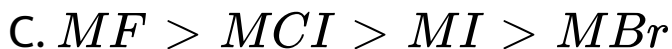
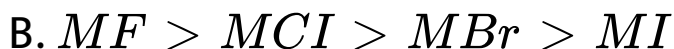
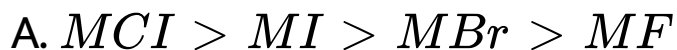
D. *Ba*

Answer: C



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131. In the case of alkali metals, the covalent character decreases in the order.



Answer: D



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132. Which of the following alkali metals has the biggest tendency of the half reaction



A. Lithium

B. Sodium

C. Cesium

D. potassium

Answer: A



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133. The metal which reacts with water at room temperature is _____.

A. copper

B. iron

C. magnesium

D. sodium

Answer: D



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134. Among the alkali metals , cesium is the most reactive because _____.

A. Its incomplete shell is nearest to the nucleus

B. it has a single electron metal

C. It is the heaviest alkali metal

D. the outermost electron is more loosely bound than the outermost electron of the other alkali metals

Answer: D



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135. Peroxide band is present in _____.

A. MgO

B. CaO

C. Li_2O

D. BaO_2

Answer: D



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136. Amongst the following select the element having highest ionization enthalpy.

A. Sodium

B. potassium

C. beryllium

D. magnesium

Answer: C



137. The most electropositive element in alkali metals is _____.

A. *Na*

B. *K*

C. *Rb*

D. *Cs*

Answer: D



138. Which alkali metal is most metallic in character?

A. K

B. Cs

C. Na

D. Li

Answer: B



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139. In the preparation of sodium carbonate ,
which of the following is used ?

A. Slaked lime

B. Quick lime

C. Lime stone

D. $NaOH$

Answer: A



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140. Consider the following statements and choose the CORRECT statement .

(i) Cs^+ ion is more highly hydrated than other alkali metal ions.

(ii) Among the alkali metals , only lithium forms a stable nitride by direct combination with nitrogen .

(iii) Among alkali metal Li , Na , K , Rb , the metal , Rb has the highest melting point

(iv) Among alkali metals Li, Na , K , Rb , only Li forms peroxide when heated with oxygen ,

(iv)

A. I

B. II

C. III

D. IV

Answer: B



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141. Ionic mobility of which of the following alkali metal ions is lowest when aqueous

solution of their salts are put under an electric field ?

A. Na

B. K

C. Rb

D. Li

Answer: D



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142. When washing soda is heated.

A. Co is released

B. $\text{CO} + \text{CO}_2$ is released

C. CO_2 is released

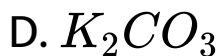
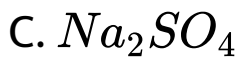
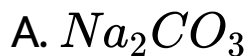
D. Water vapour is released

Answer: D



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143. Baking soda is



Answer: B



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144. As compared to potassium ,sodium has _____.

- A. Lower electronegativity
- B. higher ionization potential
- C. greater atomic radius
- D. Lower melting point

Answer: B



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145. Which of the following has density greater than water?

A. Li

B. Na

C. K

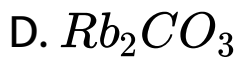
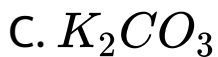
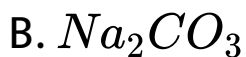
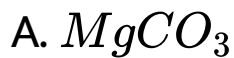
D. Rb

Answer: D



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146. Which one decomposes on heating



Answer: A



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147. Which one of the following samples of metals can produce the most hydrogen by reacting with water ?

A. 0.5 Mol Of Na

B. 0.5 Mol of Ba

C. 0.5 Mol of Li

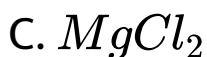
D. 0.5 Mol of K

Answer: B



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148. In which of the following the ratio of the sizes of cation to anion is smallest ?



Answer: C



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149. The main oxides formed on combustion of Li, Na and K in excess of air respectively are

- A. Li_2O , NaO and KO_2
- B. LiO_2 , Na_2O_2 and K_2O
- C. Ki_2O , NaO_2 and KO_2
- D. Li_2O , Na_2O_2 and KO_2

Answer: D



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150. Suspension of slaked lime in water is known as

A. aqueous solution of slaked lime

B. lime water

C. quick lime

D. milk of lime

Answer: D



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151. In context with beryllium, which one of the following statements is incorrect ?

A. Its hydride is electron -deficient and polymeric

B. It is rendered passive by nitric acid

C. It forms Be_2C

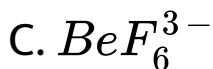
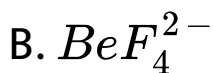
D. Its salt rarely hydrolyze

Answer: D



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152. Which of the following species do NOT exist ?



Answer: C



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153. Which of the following is most electropositive ?

A. Na

B. Na^+

C. K

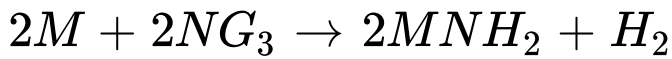
D. K^+

Answer: C



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154. An alkali metal (M) reacts with ammonia to form amide (MNH_2).



the reaction does NOT hold true when M is _____.

A. sodium

B. Potassium

C. Lithium

D. Caesium

Answer: C



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155. An element having electronic configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ will form:

- A. acidic oxide
- B. basic oxide
- C. amphoteric oxide
- D. neutral oxide

Answer: B



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156. Which of the following pair of elements do NOT show diagonal relationship ?

A. *Li, Mg*

B. *Be, Al*

C. *B, Si*

D. *Na, Mg*

Answer: D



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