



CHEMISTRY

BOOKS - BHARATI BHAWAN

CHEMISTRY (HINGLISH)

ACIDS, BASES AND SALTS

Objectives Questions

1. Which of the following is a strong acid?

- A. Lactic acid
- B. Ascorbic acid
- C. Sulphuric acid
- D. Formic acid

Answer: C



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2. Which of the following is a strong base?

- A. Ammonium hydroxide

B. Sodium hydroxide

C. Magnesium hydroxide

D. Copper hydroxide

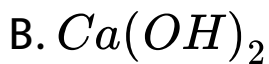
Answer: B



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3. Which of the following compounds is an acids?

A. Na_2O



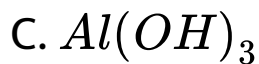
Answer: D



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4. Which of the following is a not a base?



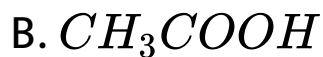
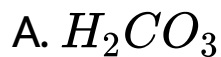


Answer: D



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5. Which of the following is a strong acid?



D. $HCOOH$

Answer: C



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6. Which of the following is a dibasic acid?

A. HCl

B. H_3PO_4

C. HNO_3

D. H_2SO_4

Answer: D



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7. Potash alum is a

A. simple salt

B. complex salt

C. acid salt

D. double salt

Answer: D



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8. Acetic acid is a weak because

A. its aqueous solution is acidic

B. it is highly ionized

C. it is weakly ionized

D. it contains $COOH$ group

Answer: C



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9. Dolomite is

A. an acid salt

B. a mixed salt

C. a normal salt

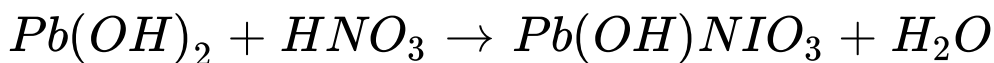
D. a double salt

Answer: D



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10. The reaction



shows that $Pb(OH)NO_3$ is

A. an acid salt

B. a basic salt

C. a base

D. an acid

Answer: B



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11. Partial neutralization of a polybasic acid gives

A. acid salt

B. basic salt

C. normal salt

D. double salt

Answer: A



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12. Strongest salt amongst the following is

A. $NaCl$

B. $CaCl_2$

C. $BaSO_4$

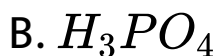
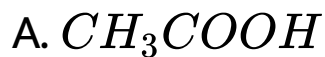
D. $LiCl$

Answer: A



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13. Which of the following can form more than one acid salt?

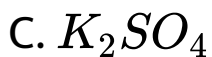
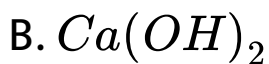


Answer: B



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14. Which of the following is not a base?



Answer: C



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15. A solution turns blue litmus red. The pH of the solution is probably.

A. 8

B. 10

C. 12

D. 6

Answer: D



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16. The type of medicine used to treat indigestion is

A. antihistamic

B. sulpha druig

C. antacid

D. antibiotic

Answer: C



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Fill In The Blanks

1. An acid salt can be formed by a.....acid.



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2.bases are called alkalis.



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3. An alkali reacts with an.....to give a.....





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4. The gas released when a metal reacts with an acid is.....



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5. When a carbonate reacts with an acid, the gas released is



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6. A base which is not a metallic oxide or hydroxide is.....



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7. When ammonium chloride is heated with caustic soda, the gas evolved is.....



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8. The acid used in lead storage battery is`



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9. A solution of $NaCl$ in water is....



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True False

1. Vinegar contains citric acid.



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2. Sulphuric acid is an hydracid.



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3. Carbonic acid is a weak acid.



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4. Sodium hydrogencarbonate is used in fire extinguishers.



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5. Potash alum is used to purify water.



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6. Calcium hydroxide is used to make soaps.



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7. Gastric juice contains hydrochloric acid.



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8. $Ph(OH)Cl$ is an acid salt.



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9. Dibasic acids are stronger than monobasic acids.



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10. The solution of a weak acid and a strong base is alkaline.



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Very Short Answer Question

1. Name a dibasic mineral acid.



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2. What is the process called in which an acid reacts with a base to produce salt and water ?



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3. A substance turns red litmus blue. What is the nature of the substance?



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4. Name an organic acid that is present in tomato.



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5. Which gas is evolved when sodium hydrogencarbonate is treated with hydrochloric acid?



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6. Name the acid which is used as a bathroom cleaner.



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7. What happens when calcium hydroxide solution is exposed to air?



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8. Name the gas whose aqueous solution is basic.



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9. Give the chemical name the of the salt which is an essential ingredient of our dialy food.



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10. What do you mean by replaceable hydrogen?



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11. Why is sodium bisulphate called an acid salt?



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12. What is the action of litmus on an aqueous solution of ammonium chloride?



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13. Write the balanced chemical equation for the reaction between aluminium powder and dilute hydrochloric acid.



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14. Why dry HCl gas does not show acidic behavior?



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15. Why is a basic substance used to treat a honey bee sting?



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Short Answer Question

1. Define an acid. Give some example of organic and inorganic acids.



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2. What do you mean by basicity of an acid?



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3. Define a base. What do you understand by the acidity of a base?





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4. What are strong acids and weak acids? Give two examples in each case.



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5. Name three acids which occur in nature. Also, state their sources.



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6. What is meant by alkali? Give two examples.



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7. What do you understand by a strong acid?

Give two examples of strong acids.



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8. What is the difference between a concentrated solution and a dilute solution of an acid?



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9. How would you define a salt? Give the names of any two normal salts.



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10. What happens when calcium hydroxide solution exposed to air?

A. hydrochloric acid reacts with caustic sodas?

B. NH_3 is dissolved in water?

C. a few drops of methyl orange are added to a solution of sodium carbonate?

D. CO_2 is passed into limewater?

Answer:



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11. Identify the acid salts and the normal salts among the following.

A. Sodium carbonate (Na_2CO_3)

B. Calcium sulphate ($CaSO_4$)

C. Disodium hydrogenphosphate

(Na_2HPO_4)

D. Sodium hydrogensulphate ($NaHSO_4$)

Answer:



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12. How would you show that magnesium metal reacts with dilute sulphuric acid to produce hydrogen gas?



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13. What happens when calcium carbonate is made to react with hydrochloric acid? Give the equation of the reaction also.



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Long Answer Questions

1. What are acid salts? Can you convert an acid salt into a normal salt?



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2. Point out three properties common to all acids.



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3. What is a basic salt? How does it differ from a normal salt?



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4. Explain acid, base and salt. Give two examples in each case.



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5. Give an example of a salt derived from a strong acid and a weak base. What would be

the behaviour of the salt towards water?



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6. State any four properties of a base.



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7. How are acids classified on the basis of elements present in them?



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8. Describe any three methods for preparing acids.



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9. What is the importance of pH in our daily life?



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