



CHEMISTRY

BOOKS - BHARATI BHAWAN

CHEMISTRY (HINGLISH)

COMPOUNDS OF COMMON USE

Fill In The Blanks

1. Fill in the following blanks:

(a) Common salt is obtained from sea-water by

the process of _____

(b) Rock salt is mined just like _____

(c) Chemical formula of washing soda is _____

(d) Sodium hydrogencarbonate is _____ soda
whereas sodium carbonate is _____ soda.

(e). The chemical formula of plaster of Paris
is _____



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2. Chemical formula of washing soda is _____



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3. sodium bicarbonate is _____ soda, whereas sodium carbonate is _____ soda.



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4. Bleaching powder reacts with dilute acids to produce _____



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5. Sodium hydrogencarbonate is used as an _____



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6. Baking powder is a mixture of _____ and sodium bicarbonate.



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7. Plaster of Paris is obtained by heating _____



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8. _____ is the common name of sodium hydroxide.



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Yes Or No

1. Is common salt a major chemical present in seawater ?



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2. Is the presence of oxygen necessary during the preparation of bleaching powder ?



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3. Does sodium carbonate dissolved in water give a basic solution ?



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4. Can bleaching powder be used for disinfecting water ?



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5. Can plaster of Paris be obtained by heating calcium sulphate at a temperature beyond control ?



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6. Can quicklime be used for making cement and glass ?



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True T Or False F

1. State whether the following statement is true or false:

Copper sulphate crystals are always wet due

to the presence of water of crystallisation in them.



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2. A crystal of barium chloride contains two molecules of water of crystallization.



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3. For making cake, baking powder is taken. If at home your mother uses baking soda

instead of baking powder in cake.

(a) How will it affect the taste of the cake and why?

(b) How can baking soda be converted in to baking powder ?

(c) What is the role of tartaric acid added to baking soda?



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4. Chlorine is used to prepare PVC and freon.



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5. Hydrogen is used to prepare vegetable ghee.



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6. Sodium chloride always contains water of crystallization .



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7. Bleaching powder is prepared by passing chlorine over quicklime.



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8. Lime is prepared by the thermal decomposition of calcium carbonate.



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9. Sodium hydroxide is prepared by chlor-alkali process. (True or False)



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10. State whether an aqueous solution of washing soda is acidic or alkaline.



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MCQ

1. Which of the following substances can be obtained by using common salt as the raw material ?

A. Caustic soda

B. Glauber's salt

C. Chlorine

D. Washing soda

Answer: A::C::D



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2. Which of the following substances is (are) obtained as by-product(s) during the preparation of sodium hydroxide by chlor-alkali process ?

A. Hydrogen

B. sodium bicarbonate

C. Common salt

D. Chlorine

Answer: A::D



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3. The property by which a crystal loses its water of crystallization is called

A. deliquescence

B. efflorescence

C. hygroscopy

D. diffusion

Answer: B



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4. The number of water molecules present in a molecule of copper sulphate crystal is

A. 2

B. 4

C. 7

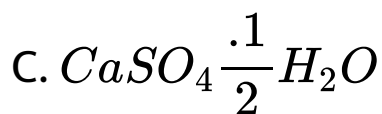
D. 5

Answer: D



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5. The chemical formula of plaster of Paris is

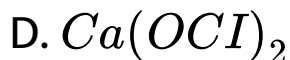
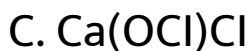
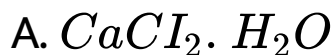


Answer: C



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6. The chemical formula of bleaching powder is



Answer: C



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1. Sodium chloride itself is not deliquescent, but it absorbs moisture when exposed to air.

Why ?



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2. Name the constituents of baking powder.



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3. What is the chemical formula of caustic soda ?



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4. (a) Explain the process of preparation of soap in laboratory.

(b) why is common salt (sodium chloride) added during the preparation of soap?

(c) Why is soap not suitable for washing clothes when the water is hard?





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5. Name the compound of sodium which loses its water of crystallization when exposed to air.



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6. Name the substance which on treatment with chlorine yields bleaching powder.



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7. Which compound of calcium produces limelight?



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8. Name the compound of sodium which is used to make borax and glass.



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9. Name any two substances which do not contain water of crystallization.



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10. The chemical formula of soda ash is



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11. Give one important use of plaster of paris.



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Short Answer Questions

1. Which of the following substances can be obtained by using common salt as the raw material ?



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2. How is washing soda different from sodium bicarbonate ?



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3. How is sodium carbonate converted into sodium bicarbonate ?



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4. Name the chemical substance which is used as a flux in the extraction of metals.



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5. Are the crystalline salts really dry ? Give your answer with reasons .



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6. A white compound of sodium is used to remove hardness of water and also as a reagent in the laboratory . Identify the compound and mention two its uses.



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7. Explain why baking soda is used as an antacid .



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8. Name the compound of calcium which gets hardened when treated with water. Give the equation for the reaction involved.



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9. Write the chemical formula of gypsum. How is it converted into plaster of Paris ?



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10. Why is plaster of paris stored in a moistureproof vessel ?



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11. What is the difference between slaked lime and lime water ?



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12. Why is sodium chloride required in our body?



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13. What happens when a sample of sodium hydrogen carbonate is heated ? Write the equation of the reaction involved.



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14. What happens when bleaching powder is heated with dilute H_2SO_4 ? Give equation of the reaction.



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15. How many molecules of water of crystallization are present in the following substance ?

(i) Crystal of copper sulphate.

(ii) Gypsum

(iii) Washing soda



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16. (a). What is bleaching powder? How is bleaching powder prepared? Write chemical equation of the reaction involved in the

preparation of bleaching powder.

(b) What happens when bleaching powder reacts with dilute sulphuric acid? Give equation of the reaction involved.

(c) State two important uses of bleaching powder.



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17. Describe an activity to show that blue copper sulphate contain water of crystallization .



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18. Why is the formula of plaster of paris written as $CaSO_4 \cdot \frac{1}{2}H_2O$?



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19. Give the important properties of baking soda.



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20. Give two important uses of washing soda and baking soda each.



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21. Describe the manufacture of bleaching powder why does its effectiveness decrease after prolonged storage ?



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22. What is the function of common salt in the manufacture of soap ?



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Long Answer Questions

1. How is sodium chloride obtained from sea - water ? State two uses of sodium chloride .



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2. A chemical compound having smell of chlorine is used to remove yellowness of white clothes in laundries. Name the compound and write the chemical equation involved in its preparation.



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3. A baker found that the cake he had prepared was hard and small in size. What ingredient had he forgotten to add that would

have made the cake fluffy ? Explain with reasons .



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4. What is baking soda ? How is it prepared from sodium chloride ? How does it differ from baking powder ?



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5. Give the principles involved in the manufacture of caustic soda by chlor - alkali process. Write chemical equations for the reactions that occur during the process. What are the by-products of this process ?



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6. (i) Name the chemical substance used in hospitals for setting fractured bones .

(ii) How will you prepare slaked lime ?



7. A pellet of a basic substance absorbs moisture from the air gives a soapy touch. The substance is also produced during chlor-alkali process.

(i) Identify the substance .

(ii) What happens when the substance is treated with CO_2 ? ,

(iii) Write the balanced chemical equation for the reaction of the substance with CO_2



