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CHEMISTRY

BOOKS - S DINESH & CO CHEMISTRY (HINGLISH)

GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS

Occurrence

1. Most of the metals which occur in native tate

A. are very reactive

B. do not form ionic compounds

C. have low density

D. have average density

Answer: C

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2. The metal found always in the free state , is

A. Gold

B. Silver

C. Copper

D. Sodium.

Answer: A

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3. Pyrousite is an ore of

A. Magnesium

B. Maganese

C. Zinc

D. Iron

Answer: B

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4. An ore of aluminium is

A. Na_2AF_6

B. $Al_2O_3.2H_2O$

C. Al_2O_3

D. $Al_2O_3 . H_2O$.

Answer: B

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5. Which of the following elements occur free in nature

A. Iodine

B. Sulphur

C. Phosphorus

D. Magnesium

Answer: B

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6. The most abundant metal in the earth's crust is

A. Calcium

B. Aluminium

C. Iron

D. Magnesium

Answer: B

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7. The metal that occurs in the native state as well as in the combined form is

A. Silver

B. Magnesium

C. Aluminium

D. Maganese

Answer: A

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8. Sea weeds are important source of

A. Chlorine

B. Bromine

C. Iodine

D. Zinc

Answer: C

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9. The element vanadium is present in

A. Sea weeds

B. Haemoglobin

C. Chlorophyll

D. Sea cucumbers

Answer: A

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10. Living organism contains iron in

A. Chlorophyll

B. Haemoglobin

C. Eyes of animals

D. None of these

Answer: B

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11. Most abundant element in the earth's crust by weight is

A. Oxygen

B. Silicon

C. Aluminium

D. Iron

Answer: A

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12. Most abundant element in earth crust in terms of number of atoms per 100 atom is

A. Oxygen

B. Silicon

C. Aluminium

D. Hydrogen

Answer: A

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13. Element found in the sediments in the ocean floor is:

A. Iron

B. Magnesium

C. Gold

D. Iodine

Answer: C

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14. Pick up the correct statement

A. All ores are minerals

B. All minerals are ores

C. A mineral cannot be ore

D. An ore cannot be a mineral

Answer: A

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15. The element with maximum cosmic abundance is

A. Helium

B. Hydrogen

C. Nitrogen

D. Oxygen

Answer: B

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Concentration

1. The rocky and siliceous matter associated with an ore is called:

A. Silag

B. Mineral

C. Matrix or Gangue

D. Flux

Answer: C

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2. The process of removing lighter gangue particles by washing in a current if water is called:

A. Levigation

B. Liquidation.

C. Leaching

D. Cupellation

Answer: A

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3. Gravity separation method is based upon:

A. Preferential washing of ores and gangue particles

B. Differences in denstities of ore particles and impurities

C. Difference in chemical properties of ore particles and impurities

D. None of these

Answer: B

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4. Which method is used for purification of Bauxite ore ?

A. Levigation

B. Leaching

C. Electrolysis

D. Magnetic seperation

Answer: B

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5. Which of the following concentration processes will you use when the gangue is light ?

A. Gravity seperation

B. Magnetic seperation

C. Froth floatation

D. None of these

Answer: A

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6. In the froth floatation process for the purification of minerals the particles float because

A. they are light

B. they are insoluble

C. their surface is preferentially wetter by oil.

D. they bear an electrostatic charge.

Answer: C

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7. Forth floatation process for the concentration of sulphide ores is an illustration of the practical application of :

A. Adsorption

B. Absorption

C. Coagulation

D. Sedimentation

Answer: A

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8. The process by which lighter earthly particles are freed from water is called

A. Benefication

B. Levigation

C. Leaching

D. None of these

Answer: B

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9. Gravity seperation process may be used for the concentration of

A. Chalocopyrite

B. Bauxite

C. Haematite

D. Calamine

Answer: C

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10. Copper pyrites are concentrated by

A. Froth floatation process

B. Gravity seperation

C. Distillation

D. Fractionation.

Answer: A

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11. Which of the following is used as a foaming agent in froth floatation process ?

A. Pine oil

B. Amyl xanthate

C. $CuSO_4$

D. KCN

Answer: A

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12. The substance used as an activator in froth floatation process is

A. Potassium ethyl xanthate

B. Sodium cyanide

C. Copper sulphate

D. Pine oil

Answer: C

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13. Which of the following is used as a depression in froth floatation process ?

A. Amyl xanthate

B. Pine oil

C. Copper sulphate

D. Potassium cyanide

Answer: D

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14. Wolframite is seperated from cassitertite by

A. Froth floatation method

B. Levigation

C. Electromagnetic seperation method

D. Electrostatic seperation method.

Answer: C

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Roasting And Culcination

1. The process in which the ore is heated in excess of air below its metling point is known as

A. Roasting

B. Calcination

C. Reduction

D. Distillation.

Answer: A

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2. Heating pyrites in air to remove sulphur dioxide is known as

A. Calcination

B. Fluxing

C. Smeling

D. Roasting

Answer: D

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3. In roasting

A. moisture is removed

B. non metal impurties are removed

C. ore become porous

D. all the above.

Answer: D

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4. The role of calcination in metallurgical operation is

A. To remove moisture

B. To decompose carbonate

C. To derive off organic matter.

D. All the above.

Answer: D

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5. To which of the following ores , calcination process is not applicable

A. $CaCO_3$

B. $Al_2O_3.2H_2O$

C. $CaCO_3Ca(OH)_2$

D. ZnS

Answer: D

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6. Calcination and roasting are generally carried out in

A. Reverbratory furnace

B. Open hearth furnace

C. Blast furnace

D. Muffle furnace

Answer: A

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7. Roasting is carried out in case of:

A. Galena

B. Iron pyrites

C. Copper glance

D. All.

Answer: D

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8. In roasting the ores are generally converted into

A. Metal oxides

B. Hydrated metal oxides

C. Metals

D. None of these

Answer: A

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Reduction To Free Metal

1. The most common method of extraction of metals from oxide ores involve

A. Reduction with carbon

B. Electrolytic method

C. Reduction with aluminium

D. Reduction with hydrogen

Answer: A

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2. Smelling is usually carried out in

A. Blast furnace

B. Open hearth furnace

C. Muffle furnace

D. Electric furnace.

Answer: A

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3. Slag is a product of:

A. flux and coke

B. coke and metal oxide

C. flux and impurities

D. metal and flux.

Answer: C

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4. Which of the following metals cannot be extracted by smelting process ?

A. Pb

B. Fe

C. Zn

D. Al.

Answer: D

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5. An ore after levigation is found to have basic impurities. Which of the following can be used as flux during smelting ?

A. H_2SO_4

B. $CaCO_3$

C. SiO_2

D. Both CaO and SiO_2

Answer: C

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6. Which of the following processes involves smelting ?

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7. Among the following groups of oxides, the group that cannot be reduced by carbon to give the respective metals.

A. Cu_2O, Sn_2O

B. Fe_2O_3, ZnO

C. CuO, K_2O

D. PbO , Fe_3O_4

Answer: C

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8. The substance which is added to ore in order to remove impurities during smelting is called

A. Slag

B. Gangue

C. Flux

D. Matrix.

Answer: C

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9. Which of the following metal can be extracted by smelting ?

A. Al

B. Fe

C. Mg

D. All

Answer: B

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10. The reducing agent used in thermite process is

A. Magnesium

B. Chromium

C. Alluminium

D. Iron

Answer: C

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11. The process of bringing the metal or its ore into solution by the action on a suitable chemical reagent followed by extracton of the metal.

A. Electrometallurgy

B. Hydrometallurgy

C. Electro refining

D. Zone refining

Answer: B

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12. The chemical reagent used for leaching of gold and silver ores is

A. Soidum hydroxide

B. Potassium cyanide

C. Potassium cyanide

D. Sodium thiosulphate.

Answer: B

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13. Gold and silve are extracted from their respective ores by

A. Leaching

B. Smelting

C. Roasting

D. Hydro metallurgy.

Answer: D

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14. Name the metal M which is extracted based on the following equation

A. Cu

B. Au

C. Hf

D. Hg

Answer: B

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15. Which process of purification is represented by the following equation

$$\underset{\text{Impure}}{T\ i} + 2I_2 \xrightarrow{250^\circ C} Til_4 \xrightarrow{1400^\circ C} \underset{\text{pure}}{T\ i} + 2I_2$$

A. Cupellation

B. Poling

C. Van Arkel

D. Zone refining

Answer: C

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16. Van Arkel method of purification of metals involves converting the metal to

A. Volatile compound

B. Volatile unstable compound

C. Non-volatile compound

D. None of these

Answer: A

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17. Cupellation process is involved in the metallurgy of

A. Cu

B. Ag

C. Al

D. Fe

Answer: B

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18. Copper is refined by

A. Liquation

B. Cupellation

C. Bessemerisation

D. Poling

Answer: D

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Miscellaneous

1. The phenomenon of removing layers of basic oxides from metals before electroplating is called

A. Galvanising

B. Anodising

C. Poling

D. Pickling

Answer: D

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2. A naturally occurring substance from which a metal can be profitably extracted is known as

A. Mineral

B. Gangue

C. Ore

D. Flux

Answer: C

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3. The removal of impurities from an ore

A. Calcination

B. Levigation

C. Slagging

D. Refining

Answer: C

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4. One of the following metals froms a volatile corbony1 compound and this property is taken advantage of its extraction. This metal is

A. Iron

B. Cobalt

C. Nickel

D. Tungsten

Answer: C

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5. The process of zone refining is used for :

A. Silicon

B. Germanium

C. Gallium

D. All the above.

Answer: D

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6. The process of zone refining is based upon

A. Fractional crystallization

B. Fractional distilllation

C. Magnetic properties of impurities

D. Impurities are less fusible than metals

Answer: A

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7. Ziroconium and titanium are purified by

A. Electrolysis

B. Zone refining

C. Bessemerisation

D. Van-Arkel method

Answer: B

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8. Tin and lead can be refined by

A. Cupellation

B. Liquation

C. Poling

D. Bessemerisation

Answer: B

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9. The reason for floating of ore particles in concentration by froth floatating process is that

A. they are light

B. they are insoluble

C. they are changed

D. they are hydrophobic.

Answer: D

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Revision Questions From Competitive Exams

1. Froth floatation process is used for the concentration of

A. Oxide ores

B. Sulphide ores

C. Chloride ores

D. Amalgams

Answer: B

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2. The salt which is least likely to be found in mineral is

A. Chloride

B. Sulphate

C. Sulphide

D. Nitrate

Answer: C

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3. In metallurgical processes the flux used for removing acidic impurities is

A. Silica

B. Sodium chloride

C. Lime stone

D. Sodium carbonate.

Answer: C

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4. The oil used in the floatation method for the purification of ores is

A. Coconut oil

B. Olive oil

C. Pine oil

D. Mustard oil.

Answer: C

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5. The main function of roasting is

A. to remove the volatile matter

B. to convert the ore into oxide

C. Both (A) and (B)

D. and make slag.

Answer: C

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6. Which of the following metals is obtained by leaching its ore with dilute cyanide solution ?

A. Silver

B. Titanium

C. Vanadium

D. Zinc

Answer: A

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7. Electrolytic reduction method is used fro the extraction of

A. highly electromagnetic elements

B. highly electropositive elements

C. transition metals

D. noble metals

Answer: B

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8. In electro-refining of metal the impure me

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9. An ore of tin containing $FeCrO_4$ is concentrated by _____.

A. Magnetic seperation

B. Froth floatation

C. Electrostatic method

D. Gravity seperation

Answer: A

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10. Zone refining process is used for the

A. Concentration of an ore

B. Reduction of a metal oxide

C. Purification of metal

D. Purification of an ore.

Answer: C

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11. The ignition mixture in Alumino-thermite process contains a mixture of

A. Magnesium powder and BaO_2

Magnesium powder, aluminium and BaO_2

B. Magnesium and aluminium powders

C. Magnesium and aluminium oxide.

D.

Answer: A

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12. Which of the following is not employed for refining of metal ?

A. Poling

B. Leaching

C. Electrolysis

D. Liquation

Answer: B

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13. The underlying of blast furnace is made of

A. Graphite

B. Silica rocks

C. Fireclay bricks

D. Basic bricks

Answer: C

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14. The purpose of smelting an ore is

A. to oxidise it

B. to reduce it

C. to separate volatile impurities

D. to obtain an alloy

Answer: B

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15. Roasting is generally carried out in _____mineral.

A. Oxide ores

B. Silicate ores

C. Sulphate ores

D. Carbonate ores.

Answer: C

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16. Cupellation process is involved in the metallurgy of

A. Cu

B. Ag

C. Zn

D. Al

Answer: B

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17. The slag obtained during the extraction of copper from coper pyrites is composed mainly of

A. Au_2S

B. $FeSiO_3$

C. $CuSiO_3$

D. SiO_2

Answer: D

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18. The extraction of which of the following metals involves bessemerisation ?

A. Fe

B. Ag

C. Al

D. Cu

Answer: D

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19. Which one of the following benefication process is used for the mineral . $Al_2O_3.2H_2O$?

A. Froth floatation

B. Leaching

C. Liquation

D. Magnetic seperation

Answer: B

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20. Which of the following statements is correct regarding the slag obtained during the extraction of a metal like copper or iron ?

A. The slag is lighter and has lower metling point than the metal.

B. The slag is heavier and has lower melting point than the metal.

C. The slag is lighter and has higher melting point than the metal.

D. The slag is heavier and has higher melting point than the metal.

Answer: A

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21. Galvanishing of iron sheets is done by

A. Cu plating

B. Zn plating

C. As plating

D. Tin plating

Answer: B

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22. The natural materials from which an element can be extracted economically are called

A. Ores

B. Minerals

C. Gangue

D. None of the above

Answer: A

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23. Purification of silicon element used in semiconductors is done by

A. Zone refining

B. Heating

C. Froth floatation

D. Heating in vacuum .

Answer: A

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24. During smelting an additional substance is added which combines with impurities to foem a fusible product. It is known as

A. Slag

B. Mud

C. Gangue

D. Flux

Answer: D

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25. When a metal is to be extracted from its ore and if the gangue associated with the ore is silica, then

A. An acidic flux is needed

B. A basic flux is needed

C. Both acidic and basic fluxes are needed

D. Neither of them is needed

Answer:

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26. Which metal cannot be obtained by electrolyte sis?

A. Ca

B. Mg

C. Cr

D. Al

Answer: B::C

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27. Which is the strongest reducing agent ?

A. Rb

B. Na

C. K

D. Mg

Answer: A

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28. The most abundant metal on the surface of the earth is

A. Fe

B. Al

C. Ca

D. Hg

Answer: B

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29. The chief source of iodine, in which it is present as sodium iodate, is

A. Carnallie

B. Sea weeds

C. Caliche

D. Iodine never exists as sodium iodate.

Answer: C

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30. The process of zone refining is used in the purification of

A. Si

B. Al

C. Ag

D. Cu

Answer: A

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31. A basic lining is given to a furnace by using

A. Calcined dolomite

B. Lime stone

C. Haematite

D. Silica

Answer: A

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32. The important oxide ore of iron is

A. Sidherite

B. Haematite

C. Pyrites

D. Bauxite

Answer: B

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33. The metal extracted by cyanide process is

A. Silver

B. Copper

C. Iron

D. Sodium Aluminium .

Answer: A

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34. Malachite is an ore of

A. Iron

B. Zinc

C. Copper

D. Mercury

Answer: C

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35. Heating an ore in the absence of air below its melting point is called

A. Leaching

B. Roasting

C. Smelting

D. Calcination

Answer: D

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36. Silver containing lead as an impurity is purified by

A. Poling

B. Cupellation

C. Levigation

D. Distillation

Answer: B

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37. Which one of the following reactions is an example for calcinaion process ?

A. 2Ag + 2HCl + [O] $\rightarrow$ AgCl + H_2O

B. 2Zn + O_2 $\rightarrow$ 2ZnO

C. 2ZnS + $3O_2$ $\rightarrow$ 2ZnO + $2SO_2$

D. $MgCO_3$ $\rightarrow$ MgO + CO_2

Answer: D

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38. Which of the following metals is obtained by electrolytic reduction process ?

A. Fe

B. Cu

C. Ag

D. Al

Answer: D

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39. Cassiterite is concentrated by

A. Levigation

B. Electromagnetic seperation

C. Floatation

D. Liquefaction

Answer: B

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40. Cassiterite is an ore of

A. Mn

B. Ni

C. Sb

D. Sn

Answer: D

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41. Which of the following flux is used to remove acidic impurities in metallurgical processes ?

A. Silica

B. Lime stone

C. Sodium chloride

D. Sodium carbonate.

Answer: B

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42. Which of the following element is extracted commercially by the electrolysis of an aqueous solution of its compound ?

A. Chlorine

B. Bromine

C. Sodium

D. Aliuminium

Answer: A

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43. Galena is an ore of

A. Pb

B. Hg

C. Sn

D. Zn

Answer: A

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44. Cryolite is

A. Na_3AlF_6 and is used in the elctrolysis of alumina for decreasing elecrical conductivity

B. Na_3AlF_6 and is usual in the electrolysis of alumina for lowering the melting point of alumina

C. Na_3AlF_6 and is used in the electrolytic purification of alumina

D. Na_3AlF_6 and is used in the electrolysis of alumina .

Answer: B

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45. In the extractionof iron, slag is produced. Slag is

A. CO

B. $FeSiO_3$

C. $MgSiO_3$

D. CaSiO_3`

Answer: D

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46. Froth floatation process is used for the metallurgy of

A. Chloride ores

B. Amalgams

C. Oxide ores

D. Sulphide ores.

Answer: D

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47. Flux is used to

A. Remove all impurities from ores.

B. Reduce metal oxide

C. Remove silica

D. Remove silica and undesirable metal oxide.

Answer: D

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48. Which one of the following elements does not exist in the native form ?

A. Au

B. Pt

C. Fe

D. S

Answer: C

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49. Heating pyrites to remove sulphur is called

A. Smelting

B. Calcination

C. Liquation

D. Roasting

Answer: D

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50. Nickel is purified by thermal decomposition of its

A. Hrdride

B. Chloride

C. Azide

D. Carbonyl

Answer: D

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51. The most electropositive metals are isolated from their ores by

A. High temperture reduction with carbon

B. Self reduction

C. Thermal decomposition

D. Electrolysis of fused ionic salts

Answer: D

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52. The most abundant element in the earth's crust (by weight) is

A. Si

B. Al

C. O

D. Fe

Answer: C

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53. The process of converting hydrated alumina into anhydrous alumina is called

A. Roasting

B. Smelting

C. Dressing

D. Calcination

Answer: D

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54. Match List I with List II and select the correct answer using the codes given below the lists:

List I	List II
I. Cyanide process	(*a*) Ultrapure Ge
II. Floatation process	(*b*) Pine oil
III. Electrolytic reduction	(*c*) Extraction of Al
IV. Zone refining	(*d*) Extraction of Au

A. (I-c), (II-a), (III-d), (IV-b)

B. (I-d), (II-b), (III-c), (IV-a)

C. (I-c), (II-b), (III-d), (IV-a)

D. (I-d), (II-a), (III-c), (IV-b)

Answer: B

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55. The metal always found in the free state is

A. Ai

B. Ag

C. Cu

D. Na

Answer: A

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56. Electromagnetic seperation is used in the concentration of

A. Copper pyrites

B. Bauxite

C. Cassiterite

D. Cinnabar

Answer: C

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57. Which of the flowing metal is extracted by the electrometallurigical method?

A. Cu

B. Fe

C. Na

D. Ag

Answer: C

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58. For which ore of the metal, froth floatation method is used for concentration?

A. Horn silver

B. Bauxite

C. Cinnabar

D. Haematite

Answer: C

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59. Which of the following metal is leached by cyanide process ?

A. Ag

B. Na

C. Al

D. Cu

Answer: A

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60. Pyrolusite is a/an

A. Oxide ore

B. Sulphide ore

C. Carbide ore

D. Not an ore

Answer: A

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61. The common method of extraction of metals from oxide ores is

A. Reduction with carbon

B. Reduction with hydrogen

C. Reduction with aluminium

D. Electrolytic method.

Answer: A

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62. Which one of the following ores is not concentrated by froth floatation process ?

A. Copper pyrates

B. Pentalandite

C. Pyrolusive

D. Zinc blende

Answer: C

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63. Sulphide ores are generally concentrated by

A. Froth floatation

B. Roasting

C. Magnetic separation

D. Carbon reduction

Answer: A

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64. The metal which cannot be obtained by electrolysis of aqueous solution of its salts is

A. Ag

B. Zn

C. Cu

D. Al

Answer: D

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65. The least stable oxide at room temperature is

A. ZnO

B. CuO

C. Sb_2O_3

D. Ag_2O

Answer: D

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66. What is insoluble in water ?

A. Sodium carbonate

B. Sodium bicarbonate

C. Calcium carbonate

D. Calcium bicarbonate

Answer: C

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67. Which of the following is magnetite ?

A. Fe_2CO_3

B. Fe_2O_3

C. Fe_3O_4

D. $Fe_2O_3.3H_2O$

Answer: C

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68. Refractory materials are generally used in furnaces because

A. they can withstand high temperature

B. they are chemically insert

C. their melting point is high

D. None of these

Answer: A

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69. Which of the following is an ore of aluminium ?

A. Dolomite

B. Azurite

C. Bauxite

D. Malachite

Answer: C

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70. The method of zone relining of metal is based on the principle of $AIFMIT - 2003)$

A. greatest solubility of the impurities in the molten state than in the solid

B. greater solubility of pure metal than that of the impurity

C. higher melting point of the impurity than that of the pure metal

D. greater noble character of the solid metal than that of the impurity .

Answer: A

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71. Process of heating ore is air to remove sulphur is:

A. Calcination

B. Roasting

C. Smelting

D. None of these

Answer: B

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72. The formula of azurite is :

A. $CuCO_3 . Cu(OH)_2$

B. $2CuCO_3 . Cu(OH)_2$

C. $CuCO_2 . Cu(OH)_2$

D. $CuSO_4 . Cu(OH)_2$

Answer: B

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73. Heating of ore in presence of air to remove sulphure impurities is called

A. Calcination

B. Roasting

C. Smelting

D. None of these

Answer: B

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74. Vapour phase refining of nickel is carried out using

A. I_2

B. Cl_2

C. HCl

D. CO

Answer: D

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75. Thomas slag is

A. Calcium silicate

B. Calcium phosphate

C. Barium phosphate

D. Strontium phosphate.

Answer: B

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76. In Mc-Arthur Forrest method silver is extracted from the solution of $Na[Ag(CN)_2]$ by the use of

A. Fe

B. Na

C. Zn

D. Au

Answer: C

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77. During the extraction of iron, slag produced is

A. CO

B. $FeSiO_3$

C. $MgSiO_3$

D. $CaSiO_3$

Answer: D

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78. Silver containing lead as an impurity is not purified by

A. Poling

B. Cupellation

C. Levigation

D. Distillation

Answer: B

Watch Video Solution

79. Out of the following matals that cannot be obtained by electrolysis of the aquenous solution of their salts is

A. Ag

B. Cr

C. Cu

D. Mg

Answer: D

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80. In the froth floatation process for the facilitation of ores the ore particles float because

A. they are light.

B. their surface is not easily wetted by water

C. they bear electrostatic charge

D. they are insoluble

Answer: B

Watch Video Solution

81. Mond's process is used for

A. Ni

B. Al

C. Fe

D. Cu

Answer: A

Watch Video Solution

82. In the extraction of Fe from Fe_2O_3 the reducing agent is

A. C

B. Al

C. electrolytic reduction

D. Cu

Answer: A

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83. δG^o vs T plot in the Ellingham's diagram slopes downaward for the reaction.

A. Mg + $\frac{1}{2}O_2 \Rightarrow$ MgO

B. 2Ag + $\frac{1}{2}O_2 \Rightarrow Ag_2O$

C. $C + \frac{1}{2}O_2 \rightarrow CO$

D. $CO + \frac{1}{2}O_2 \rightarrow CO_2$

Answer: C

View Text Solution

84. Which of the following reactions taking place in the blast furnace during extraction of iron is endothermic .

A. $CaCO_3 \rightarrow CaO + CO_2$

B. $2C + O_2 \rightarrow 2CO$

C. $C + O_2 \rightarrow CO_2$

D. $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$

Answer: A

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85. Consider the following reaction at 1000° C .

(A) Zn(s) + 1/2 O_2(g) $\rightarrow$ ZnO(s)

δG^o = -360 kJ mol^{-1}

(B) C(gr) + $\frac{1}{2}O_2(g)$rarr$CO(g)$deltaG^(o) $=$

-460 kJ mol^(-1)

$Ch\infty sethec$ or $rectstatementat$1000^(@)`C .

A. zinc can be oxidized by carbon monoxide

B. zinc oxide can be reduced by graphic

C. both statements (A) and (B) are true

D. both statements (A) and (B) are false

Answer: B

View Text Solution

86. The ore having two different metal atoms is

A. haeatite

B. galena

C. magnetite

D. copper pyrites.

Answer: D

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87. Cassiterite is an ore of

A. Pb

B. Zn

C. Sn

D. Mn

Answer: C

View Text Solution

88. Blister copper is

A. Impure Cu

B. Cu alloy

C. pure Cu

D. copper having 1% impurity

Answer: D

View Text Solution

89. The magnetic ore is

A. Ferrous oxide

B. Ferric oxide

C. Ferric hydroxide

D. Ferrosoferric oxide

Answer: D

View Text Solution

90. When compared , δG^o for the formation of Al_2O_3 ,the δG^o for the formation of Cr_2O_3 is

A. same

B. unpredicted

C. higher

D. lower

Answer: C

View Text Solution

91. Which metal has a greater tendency to form metal oxide

A. Al

B. Ca

C. Cr

D. Fe

Answer: B

Watch Video Solution

92. Identify the reaction that does not take place in a blast furnace

A. $2Fe_2O_3$ + 3C $\Rightarrow$ 4Fe + CO_2

B. CO_2 + C $\rightarrow$ 2CO

C. $CaCO_3$ $\rightarrow$ CaO + CO_2

D. FeO + SiO_2 $\rightarrow$ $FeSiO_3$

Answer: D

Watch Video Solution

93. Which one of the following ores is treated by chemical leaching

A. Galena

B. Copper pyrite

C. Copper glance.

D. Argentite

Answer: D

Watch Video Solution

94. Auto reduction is employed in the metallurgy of

A. Ag

B. Al

C. Ti

D. Zn

Answer: A

Watch Video Solution

95. The ore that is concentrated by froth floatation process is

A. Zincite

B. Cinnabar

C. Bauxite

D. Malachite

Answer: B

Watch Video Solution

96. In aluminium extraction by Bayer process, alumina is extracted from bauxite by sodium hydroxide at high temperatures and pressures

$Al_2O_3(s) + 2OH^- \rightarrow 2Al_2O_2^-(aq) + H_2O$(l)

Solid impurities such as Fe_2O_3 and SiO_2 are removed and then $Al(OH)_4^-$ is reprecipitated

$2Al(OH)_4^- \rightarrow Al_2O_{3.3}H_2O$(g) + $2OH^-$ (aq)

In the industrical world :

A. Carbon dioxide is added to precipitate the alumina

B. Tempertaure and pressure are dropped and the supersaturated solution seeded

C. Both (A) and (B) are practised

D. The water is evaporated

Answer: D

Watch Video Solution

97. Bauxite is made up of Al_2O_3 + SiO_2 + TiO_2 + Fe_2O_3 . This ore is treated with conc. NaOH solution at 500 K and 35 bar pressure for a few hours and filtered hot. In the filtrate the species present are

A. $NaAl(OH)_4$ only

B. $Na_2Ti(OH)_6$ only

C. $NaAl(OH)_4$ and Na_2SiO_3 both

D. Na_2SiO_3 only .

Answer: C

View Text Solution

98. When copper pyrites is roasted in excess of air, a mixture of CuO + FeO is formed . FeO is present as impurities . This can be removed as slag during reduction of CuO . The flux added to form slag is

A. SiO_2 which is an acid flux

B. Lime stone which is a basic flux

C. SiO_2 which is acidic flux

D. CaO which is basic flux.

Answer: A

View Text Solution

99. Bauxite ore is made up of Al_2O_3 + SiO_2 + TiO_2 + Fe_2O_3 . This ore is treated with conc. NaOH solution at 500 K and 35 bar pressure for

a few hours and filtered hot. In the filtrate the species present are :

A. $NaAl(OH)_4$ only

B. $Na_2Ti(OH)_6$ only

C. $NaAl(OH)_4$ and Na_2SiO_3 both

D. Na_2SiO_3 only .

Answer: C

Watch Video Solution

100. The auto reduction process is not used in the metallurgy of

A. Hg

B. Cu

C. Pb

D. Fe

Answer: D

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101. The incorrect statement among the following is

A. Hydrogen is used to reduce NiO

B. Zirconium is refined by Van Arkal method

C. The sulphate ore galena is concentrated by froth floatation

D. In the metallurgy of iron, flux used in SiO_2

Answer: D

Watch Video Solution

102. Which of the following is used in thermine welding ?

A. TiO_2 + 4Na $\Rightarrow$ Ti + $2Na_2O$

B. 2Al + Fe_2O_3 $\Rightarrow$ Al_2O_3 + 2Fe

C. SnO_2 + 2C $\Rightarrow$ Sn + 2CO

D. $Cr_2O_3 + 2Al \rightarrow Al_2O_3 + 2Cr$

Answer: B

Watch Video Solution

103. Which of the following methods is used for obtaining Aluminium metal ?

A. Electrolysis of food Al_2O_3 and cryolite

B. By heating alumina with carbon

C. By heating alumina in Muffle furnace

D. By a process called pyrometallurgy

Answer: A

View Text Solution

104. According to Ellinigliam diagram, the oxidation reaction of carbon to carbon monoxide may be used to reduce which one of the following oxides at the lowest tempertaure

A. Al_2O_3

B. Cu_2O

C. MgO

D. ZnO

Answer: B

View Text Solution

105. In the equation , 4M + $8CN^-$ + $2H_2O$ + O $\rightarrow$ $4[M(CN)_2]^-$ + $4OH^-$, the metal is

A. Copper

B. Iron

C. Gold

D. Zinc

Answer: C

View Text Solution

106. Oxidation states of the metal in the mineral haematite and magnetite ,respectively are

A. II, III in haematite and III in magnetite

B. II, III in haematite and II in magnetite

C. II in haematite and II and III in magnetite

D. III is haematite and II and III in magnetite

Answer: C

View Text Solution

107. Which of the following does not occur as sulphide ore. ?

A. Zn

B. Cr

C. Ag

D. Fe.

Answer: B

View Text Solution

108. Calamine is an ore of

A. Zn

B. Cu

C. Al

D. Fe

Answer: A

View Text Solution

109. Bauxite has composition

A. Al_2O_3

B. $Al_2O_3H_2O$

C. $Al_2O_32H_2O$

D. None of these

Answer: C

View Text Solution

110. Which of the following statements is false ?

A. During roasting, moisture is removed from the ore

B. The ore is freed from all non metallic impurities

C. Calcination of ore is carried out in the absence of any blast of air

D. The concentrated zinc blende is subjected to calcination during the extraction by pyrometallurgy.

Answer: D

View Text Solution

111. In view of the sign of $\delta_n G^o$ for the following reaction :

$PbO_2 + Pb \Rightarrow 2PbO, \delta_n G^o$ lt 0

SnO_2 + Sn $\rightarrow$ 2SnO , $\delta_n G^o$ gt 0

Which oxidation states are more characteristic for lead and tin ?

A. For lead +4 , for tin +2

B. For lead +2 , for tin +2

C. For lead +2 , for tin +4

D. For lead +2, for tin +4

Answer: D

View Text Solution

112. The following reactions take place in the blast furnace in the preparation of impure iron. Identify the reaction pertaining to the formation of slag.

A.

$$Fe_2O_3(s) \rightarrow 3CO(g) \rightarrow 2Fe(l) + 3CO_2(g)$$

B. $CaCO_3(s) \rightarrow CaO(g) + CO_2(g)$

C. $CaO(s) + SiO_2(g) \rightarrow CaSiO_3(s)$

D. $2C(s) + O_2(s) \rightarrow 2CO(g)$

Answer: C

View Text Solution

113. Which of the following elements is present as the impurity to the maximum extent in pig iron ?

A. Phosphorus

B. Manganese

C. Carbon

D. Silicon

Answer: C

View Text Solution

114. Which of the following pairs of metals is purified by Van Arkel method ?

A. Ni and Fe

B. Ga and In

C. Zr and Ti

D. Ag and Au

Answer: C

View Text Solution

115. Heating mixture of $Cu_2O$and Cu_2S will give

A. Cu + SO_2

B. $Cu + SO_3$

C. Zr + Ti

D. Cu_2SO_3

Answer: A

Watch Video Solution

116. During electrolytic refining of copper some metals , present as impurity settle as anode mud. These are

A. Sn and Ag

B. Pb and Zn

C. Ag and Au

D. Fe and Ni

Answer: A

View Text Solution

117. Aluminium is extracted from Alumina (Al_2O_3) by electrolysis of a molten mixture of

A. Al_2O_3 + HF + $NaAlF_4$

B. $Al_2O_3 + CaF_2 + NaAlF_4$

C. Al_2O_3 + Na_3AlF_6 + CaF_2

D. Al_2O_3 + KF + Na_2Al_6

Answer: C

View Text Solution

118. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with

A. Copper (I) sulphide

B. Sulphur dioxide (SO_2)

C. Iron sulphide (FeS)

D. Carbon monoxide (CO).

Answer: A

View Text Solution

119. Which method of purification is represented by the following equation

$$Ti(g) + 2I_2(g) \xrightarrow{523K} Til_4(g) \xrightarrow{1700K} Ti(s) + 2I_2(g)$$

A. Cupellation

B. Poling

C. Van Arkel

D. zone refining.

Answer: C

View Text Solution

120. Identify the alloy containing a non metal as a constitunt in it

A. Invar

B. Steel

C. Bell Metal

D. Bronze

Answer: B

Watch Video Solution

121. the carbon based reduction method is not used for extraction of

A. tin from tin stone

B. iron from Fe_2O_3

C. aluminium from alumina

D. magnesium from $MgCO_3$. $CaCO_3$

Answer: C

Watch Video Solution

122. which of the following electronts is present as the impurity to the maximum extent in the pig iron?

A. Silicon

B. Phosphorus

C. Manganese

D. Carbon

Answer: C

Watch Video Solution

123. Which of the following pairs of metals uis purified by van arkel method?

A. Ag and Au

B. Ni and Fe

C. Ga and In

D. Zr and Ti

Answer: C

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124. Extraction of metal from the ore cassiterite involves

A. carbon reduction of an oxide ore

B. self-reduction of a sulphide ore

C. removal of copper impurity

D. removal of iron impurity.

Answer: A

Watch Video Solution

125. Which method of purification is represented by the following equation

`Ti(s)

A. Zone refining

B. Cupellation

C. Poling

D. Van -Arkel.

Answer: A

View Text Solution

126. Which ore of the following is a mineral of iron ?

A. Cassiterite

B. Pyrolusite

C. Magnetite

D. Malachite

Answer: C

Watch Video Solution

127. Aluminium is extracted from Alumina (Al_2O_3) by electrolysis of a molten mixture of

A. Al_2O_3 + HF + $NaAlF_4$

B. Al_2O_3 + Na_3AlF_6 +CaF_2

C. Al_2O_3 + KF + Na_3AlF_6

D. Al_2O_3 + HF + $NaAlF_4$

Answer: A

Watch Video Solution

128. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with

A. Sulphur dioxide (SO_2)

B. Iron sulphide (FeS)

C. Carbon monoxide (CO)

D. Copper(I) sulphide (Cu_2S).

Answer: C

View Text Solution

129. Identify the alloy containing a non metal as a constituent in it

A. Steel

B. Bell metal

C. Bronze

D. Invar.

Answer: C

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130. Roasting of sulphides gives the gas X as a byproduct . This is a colourless gas with choking smell of burnt sulphur and causes great damage to respiratory organs as a result of acid rain. Its aqueous solution is acidic , acts as a reducing

agent and its acid has never been isolated. The gas X is :

A. SO_3

B. H_2S

C. SO_2

D. CO_2

Answer: D

View Text Solution

131. Sulphide ores are common for the metals

A. Ag , Cu and Pb

B. Ag , Cu and Sn

C. Ag , Mg and Pb

D. Al , Cu and Pb

Answer: C

Watch Video Solution

132. Vapour phase refining for the purification of

A. Hg

B. Ti

C. Zn

D. Ni.

Answer: B::D

Watch Video Solution

133. Aluminium is used in the extraction of

A. Cr

B. Fe

C. Ca

D. Mn

Answer: A::D

View Text Solution

Selected Straight Objective Type Mcqs

1. Common impurities present in bauxite are

A. CuO

B. ZnO

C. Fe_2O_3

D. SiO_2

Answer: C::D

View Text Solution

2. Auto-reduction process is used for the extraction of :

A. Hg

B. Fe

C. Cu

D. Pb

Answer: A::C::D

Watch Video Solution

3. Which of the following metal/metals can be used for the extraction of Ti from TiO_2

A. Na

B. Al

C. Mg

D. Zn

Answer: A::C

View Text Solution

4. Hydrogen is used as a reducing agent in the extraction of

A. Mo

B. Ca

C. W

D. Na

Answer: A::C

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5. In nature , metals are found in the

A. native state

B. oxidised form

C. reduced form

D. form of alloys.

Answer: A::B

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6. In nature, non metals are found in the

A. native state

B. reduced form

C. oxidised form

D. form of alloys with metals.

Answer: A::B::C

View Text Solution

7. Leaching is used for the concentration of

A. red bauxite

B. haematite

C. gold ore

D. silver ore.

Answer: A::C::D

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8. Pick out the correct statements

A. All minerals are ores

B. All ores are minerals

C. All minerals are not ores

D. All ores are not minerals.

Answer: B::C

Watch Video Solution

Watch Video Solution

9. Of the following that cannot be obtained by electrolysis of the aqueous solution of their states is /are

A. Ag

B. Mg

C. Cu

D. Al

Answer: B::D

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Watch Video Solution

10. The element which exists in the liquid state at room tempertaure are

A. Na

B. Br

C. Hg

D. Ga

Answer: B::C

Watch Video Solution

11. The major role of fluorpar (CaF_2) which added in small quantities in the electrolyte reduction of alumina dissolved in fused cryolite (N_3AlF_6) is

A. as a catalyst

B. to make the fused mixture very conducting

to lower the temperature of the melt

C. to lower the temperature of the melt

D. to decrease the rate of oxidation of carbon at anode.

Answer: B::C

Watch Video Solution

12. Which of the following oxides is/are amphoteric ?

A. Na_2O

B. CaO

C. Al_2O_3

D. SnO_2

Answer: C::D

Watch Video Solution

13. Extraction of metal from the ore cassiterite involves

A. Carbon reduction of an oxide ore

B. Self reduction of a sulphide ore

C. Removal of copper impuruty

D. Removal of iron impurity.

Answer: A::C::D

Watch Video Solution

14. The carbon-based reduction method is not used in the extraction of

A. tin from SnO_2

B. iron from Fe_2O_3

C. aluminium from aluminimum

D. magnesium from $MgCO_3$. $CaCO_3$

Answer: C::D

Watch Video Solution

15. In the metallurgical processes, aluminium acts as

A. an oxidising agent

B. a reducing agent

C. a flus

D. a solder.

Answer: B

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16. Hydrogen will not reduce

A. heated cupric oxide

B. heated ferric oxide

C. heated stannic oxide

D. heated aluminium oxide.

Answer: D

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17. Froth floatation process may be used to increase the concentration of the mineral in

A. Chalopyrites

B. Bauxite

C. Haematite

D. Calamine

Answer: A

Watch Video Solution

18. In which of the following mineral aluminium is not present ?

A. Cryolite

B. Mica

C. Feldspar

D. Fluorspar

Answer: D

View Text Solution

19. An important ore of magensium is

A. Malachite

B. Cassiterite

C. Carnallite

D. Galena

Answer: C

Watch Video Solution

20. Among the following statements the incoorect one is

A. Calamine and siderte are oxides

B. Argentite and cuprite are oxides

C. Zinc blende and iron pyrites are sulphides.

D. Malachite and azurite are ores of copper.

Answer: B

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21. In the commercial electrochemical process for aluminium extraction, the electrolyte used is

A. $Al(OH)_3$ in NaOH solution

B. An aqueous solution of $Al_2(SO_4)_3$

C. A molten mixture of Al_2O_3 and Na_3AlF_6

D. A motlen mixture of Al_2O_3 and $Al(OH)_3$

Answer: C

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22. Which of the following process is used in the extractive metallurgy of magnesium ?

A. Fused salt electrolysis

B. Self reduction

C. Aqueous solution electrolysis

D. Thermite reduction .

Answer: A

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23. Pb and Sn are extracted from their chief ore by

A. carbon reductioin and self reduction

B. self reduction

C. electrolysis and self reduction

D. self reduction and electrolysis.

Answer: B

Watch Video Solution

24. Which of the following ores is best concentrated by froth floatation method

A. Malachite

B. Cassiterite

C. Galena

D. Magnetite

Answer: C

View Text Solution

25. During electrolytic refining of copper, some metals present as impurity settle as anode mud. These are

A. Sn and Ag

B. Pb and Zn

C. Ag and Au

D. Fe and Ni

Answer: A

View Text Solution

26. Heating mixture of Cu_2O and Cu_2S will give

A. Cu + SO_2

B. Cu + SO_3

C. CuO + CuS

D. Cu_2SO_3

Answer: C

View Text Solution

27. Sulphide ores of metals are usually concentrated by froth floatation process. Which one of the following sulphide others offers an exception and is concentrated by chemical leaching ?

A. Sphalerite

B. Argentite

C. Galena

D. Copper pyrite.

Answer: B

View Text Solution

28. Which of the following statements, about the advantage of roasting of sulphide ore before reduction is not true

A. Roasting of the sulphide to the oxide is thermodynamics feasible

B. Carbon and hydrogen are suitable reducing agents for metal sulphides

C. The $\delta(fG)^{o0}$ of the sulphide is greater than those for CS_2 and H_2S

D. The $\delta(fG)^o$ is negative for roasting of sulphide ore to oxide.

Answer: B

View Text Solution

29. Extraction of zinc blende is achieved by

A. Electrolytic reduction

B. Roasting followed by reduction with carbon

C. Roasting followed by reduction with other metals

D. Roasting followed by self reduction.

Answer: B

View Text Solution

30. Which of the following factors is of no significance for roasting sulphide ores to oxides and not subjecting the sulphide ores to carbon reduction directly ?

A. CO_2 is more volatile than CS_2

B. Metal sulphides are thermodynamically more stable than CS_2

C. CO_2 is thermodynamically more stable than CS_2

D. Metal sulphides are less stable than the corresponding oxides.

Answer: B

View Text Solution

31. Native silver metal forms a water soluble complex with a dilute solution of NaCN in the presence of

A. nitrogen

B. oxygen

C. carbon dioxide

D. argon

Answer: B

View Text Solution

Reason Assertion Type Mcqs

1. Assertioin: Silicate minerals are seldon used for hte extraction of a metal.

Reason: Silicate minerals are very rare.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: D

View Text Solution

2. Assertion: Carbon is not used as a reducing agent for the extraction of group 2 metals,

Reason: Carbon react with group 2 metals to give carbides.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: A

View Text Solution

3. Assertioin: Platinum and gold occur in native state.

Reason: Platinum and gold are expensive metals.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: C

View Text Solution

4. Assertion: Al is used a reducing agent in the manufacture of Fe.

Reason: Al can easily extract oxygen from iron oxide.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: D

View Text Solution

5. Assertion: Wolframite impurity is separated from tin stone (SnO_2) by magnetic separation.

Reason: Tin stone is ferromagnetic and is attracted by the magnet.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: C

View Text Solution

6. Assertion: Pb, Sn and Bi are purified by liquation,

Reason: Pb, Sn and Bi have lower m.pt then the impurities present .

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: A

View Text Solution

7. Assertion: Elements Fe, Ni, Co, Pd and Rb are called siderophiles.

Reason: Sideros in Greek means iron.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: A

View Text Solution

8. Assertion: Elements O, N and Ar are called atmophiles

Reason: Atoms in Greek means which cannot be subdivided.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: C

View Text Solution

9. Assertion: Nitrate ores are very rare.

Reason: Atmos in Greek means which cannot be subdivided.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: B

View Text Solution

10. Assertion: Roasting ore with NaCl is called chloridising roasting.

Reason: Chlorides are easily soluble in water.

A. Both A and R ar true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

Answer: C

View Text Solution

Linked Comprehnsion Types

1. Aluminium is quite reactive and hence does not occur in nativie forms. Aluminium is the third most abundnat element in the earth's crust. The important ores of aluminium are

i) Bauxite $Al_2O_32H_2O$

ii) Cryolite Na_3AlF_6

iii) Feldspar $KAISi_3O_8$

iv) Alunite $K_2SO_4Al_2(SO)_4Al(OH)_3$

Aluminium is extracted from Bauxite. Bauxite is generally associated with iron oxide and silica as impurity. Bauxite ore containing excess of ferric oxide (Red bauxite) is powdered and digested with a strong solution of caustic soda in an auto clave when alumina $(Al_2O_3$ passes into solution as sodium aluminate while ferric oxide remains undissolved and is filtered out-

$$Al_2O_3 + 2Na \rightarrow 2NaAlO_2 + H_2O$$

The solutioin is diluted with water and a small amount of freshly precipated aluminium hydroxidel is added as seeding agent. The precipated $Al(OH)_3$ is filtered, washed and

ignited to get pure alumina.

$$2Al(OH)_3 \xrightarrow{\Delta} \underset{\text{Alumina}}{A} l_2O_3 + 3H_2O$$

Pure alumina is dissolved in molten cryolite and is electrolyted using graphite electrodes. Aluminium obtained after this step is $99\,\%$ pure. It can further be refined by Hoope's process.

Principle ore of aluminium is

A. Bauxite

B. Corrundum

C. Cryolite

D. Feldspar.

Answer: A

View Text Solution

2. Aluminium is quite reactive and hence does not occur in nativie forms. Aluminium is the third most abundnat element in the earth's crust. The important ores of aluminium are

i) Bauxite $Al_2O_32H_2O$

ii) Cryolite Na_3AlF_6

iii) Feldspar $KAISi_3O_8$

iv) Alunite $K_2SO_4Al_2(SO)_4Al(OH)_3$

Aluminium is extracted from Bauxite. Bauxite is generally associated with iron oxide and silica as impurity. Bauxite ore containing excess of ferric oxide (Red bauxite) is powdered and digested with a strong solution of caustic soda in an auto clave when alumina $(Al_2O_3$ passes into solution as sodium aluminate while ferric oxide remains undissolved and is filtered out-

$$Al_2O_3 + 2Na \rightarrow 2NaAlO_2 + H_2O$$

The solutioin is diluted with water and a small amount of freshly precipated aluminium hydroxidel is added as seeding agent. The precipated $Al(OH)_3$ is filtered, washed and

ignited to get pure alumina.

$$2Al(OH)_3 \xrightarrow{\Delta} \underset{\text{Alumina}}{A} l_2O_3 + 3H_2O$$

Pure alumina is dissolved in molten cryolite and is electrolyted using graphite electrodes. Aluminium obtained after this step is $99\,\%$ pure. It can further be refined by Hoope's process.

Which mineral has been named correctly

A. Bauxite $(Al_2O_32H_2O$

B. Corrundum Al_2O_3

C. Cryolite Na_3AlF_6

D. Feldspar $Be_3Al_2Si_6O_{13}$

Answer: A

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3. Aluminium is quite reactive and hence does not occur in nativie forms. Aluminium is the third most abundnat element in the earth's crust. The important ores of aluminium are

i) Bauxite $Al_2O_32H_2O$

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Which of the following metals are extracted by electrolytic reaction

A. Copper

B. Aluminium

C. Zinc

D. Silver

Answer: B

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Pure alumina is dissolved in molten cryolite and is electrolyted using graphite electrodes. Aluminium obtained after this step is $99\,\%$ pure. It can further be refined by Hoope's process.

Which one of the following benefication processes is used for the mineral $Al_2O_32H_2O$

A. Froth floatation

B. Leaching

C. Liquation

D. Magnetic separation

Answer: B

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5. Aluminium is quite reactive and hence does not occur in nativie forms. Aluminium is the third most abundnat element in the earth's crust. The important ores of aluminium are

i) Bauxite $Al_2O_3 2H_2O$

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In electrolysis of Al_2O_3 by Hall and Heroult process

A. Cryolite $(Na_3AlF_6$ lowers the melting point of alumina (Al_2O_3) and increases its eletrical conductivity.

B. Al is obtained at cathode and CO_2 at anode.

C. Both A and B are correct

D. None of A, B are correct.

Answer: A

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6. Copper is the most noble of first row transition metals and occurs in small deposits in several countries. Ores of copper include chalanthrite. $(CuSO_4.5H_2O)$, atacanrite $\left[Cu_2Cl(OH)_3\right]$ cuprite (Cu_2O), copper glance (Cu_2S) nad malachite $\left[Cu_2(OH)_2CO_3\right]$. However, $80\,\%$ of the world copper production comes from the ore chalcopyrite $(CuFeS_2)$. The extraction of ore chalcopyrite $(CuFeS_2)$. The extraction of copper from chalcopyrite involves partial roasting, removal of iron nad self reduction.

Partial roasting of chalcopyrite produces

A. Cu_2S and FeO

B. Cu_2O and FeO

C. CuS and Fe_2S_3

D. Cu_2O and Fe_2O_3

Answer: A

View Text Solution

7. Copper is the most noble of first row transition metals and occurs in small deposits in several countries. Ores of copper include

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Iron is removed from chalcopyrite as

A. FeO

B. FeS

C. Fe_2O_3

D. $FeSiO_3$

Answer: D

View Text Solution

8. Copper is the most noble of first row transition metals and occurs in small deposits in several countries. Ores of copper include chalanthrite. $(CuSO_4.5H_2O)$, atacanrite $\left[Cu_2Cl(OH)_3\right]$ cuprite (Cu_2O), copper glance (Cu_2S) nad malachite $\left[Cu_2(OH)_2CO_3\right]$.

However, 80% of the world copper production comes from the ore chalcopyrite $(CuFeS_2)$. The extraction of ore chalcopyrite $(CuFeS_2)$. The extraction of copper from chalcopyrite involves partial roasting, removal of iron nad self reduction.

In self reduction, hte reducing spent is

A. S

B. O^{2-}

C. S^{2-}

D. SiO_2

Answer: C

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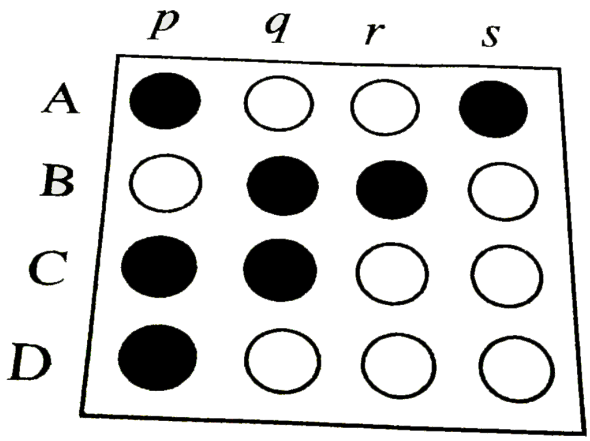
Matrix Match Type Mcqs

1. Here each question contains statements given in two columns which have to the matched. Sttements in column I are labelled as A, B, C and D whereas statement in column II are labelled as p,q,r and s. The answers of these questions are to appropriatly bubbled as illustrated with the

following example.

If the correct matches are A-p, A-r, B-q, B-r, C-p, C-q and D-p, then correctly bubbled 4×4 matrix should look like the following.



	Column I		Column II
A	$PbS \longrightarrow PbO$	p	Roasting
B	$CaCO_3 \longrightarrow CaO$	q	Calcination
C	$ZnS \longrightarrow Zn$	r	Carbon reduction
D	$Cu_2S \longrightarrow Cu$	s	Self reduction

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2. Here each question contains statements given in two columns which have to the matched. Sttements in column I are labelled as A, B, C and D whereas statement in column II are labelled as p,q,r and s. The answers of these questions are to appropriatly bubbled as illustrated with the following example.

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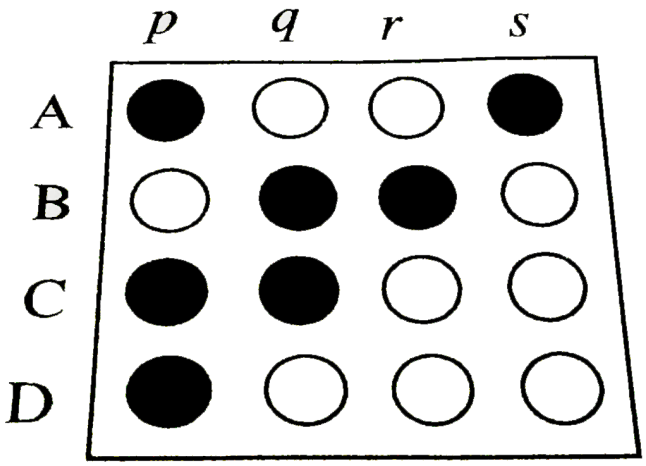

	Column I		Column II
A	Self reduction	p	lead
B	Carbon reduction	q	silver
C	Complex formation and displacement by by metal	r	copper
D	Decomposition of iodide	s	Boron

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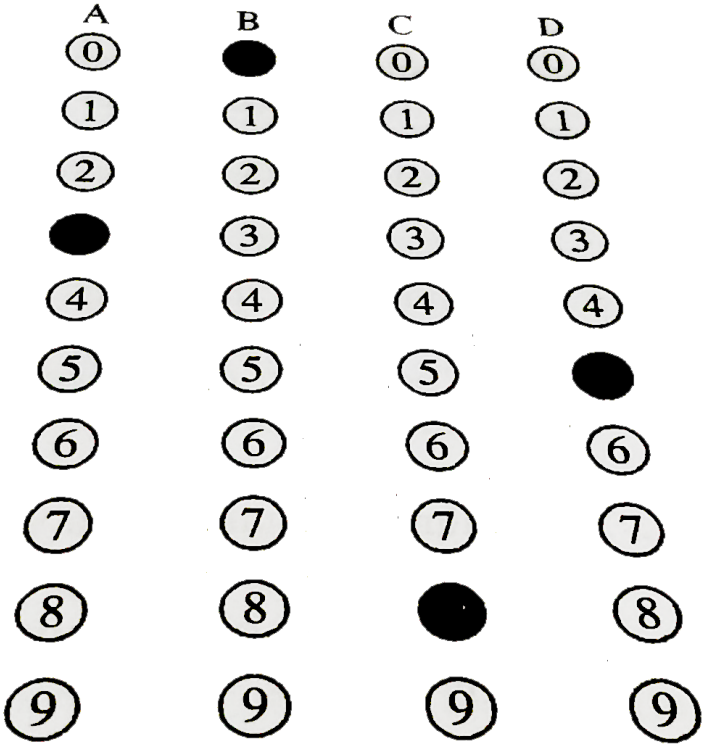
Integer Type Questions

1. The answer to each of the following questions is a single digit integer from 0 to 9. If the correct answer to the question number A,B, C and D say 3,0,8,5 respectively, then the correct darkening of bubbles should be as follow.



How many of the following metals can be refined by the vapour phase refining?

Zr, Zn, Cd, Hg, Ni, Co, Pt, Fe, Ti

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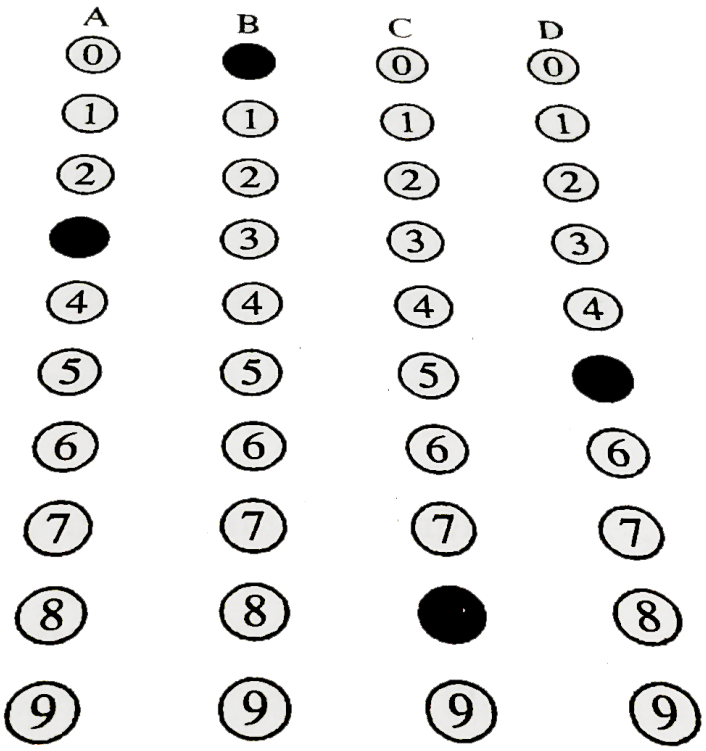
2. The answer to each of the following questions is a single digit integer from 0 to 9. If the correct answer to the question number A,B, C and D say 3,0,8,5 respectively, then the correct darkening of bubbles should be as follow.



Amonst the following, the total number of ores, which can be concentrated by from floatation process haematite, bauxite, galena, copper pyrites, sphalerite caneterite, calamine, argentite, chalcorite.

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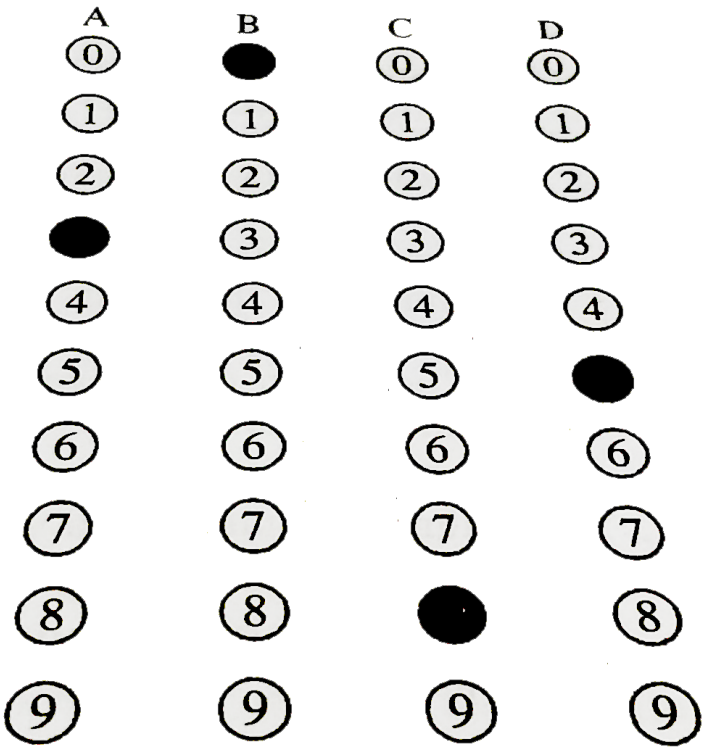
3. The answer to each of the following questions is a single digit integer from 0 to 9. If the correct answer to the question number A,B, C and D say 3,0,8,5 respectively, then the correct darkening of bubbles should be as follow.



Among the following, total number of metals which occur inthe native in the earth's crust are.

Fe,Zn,Na,Au,Ni,Sb,Sn,Pt,Hg

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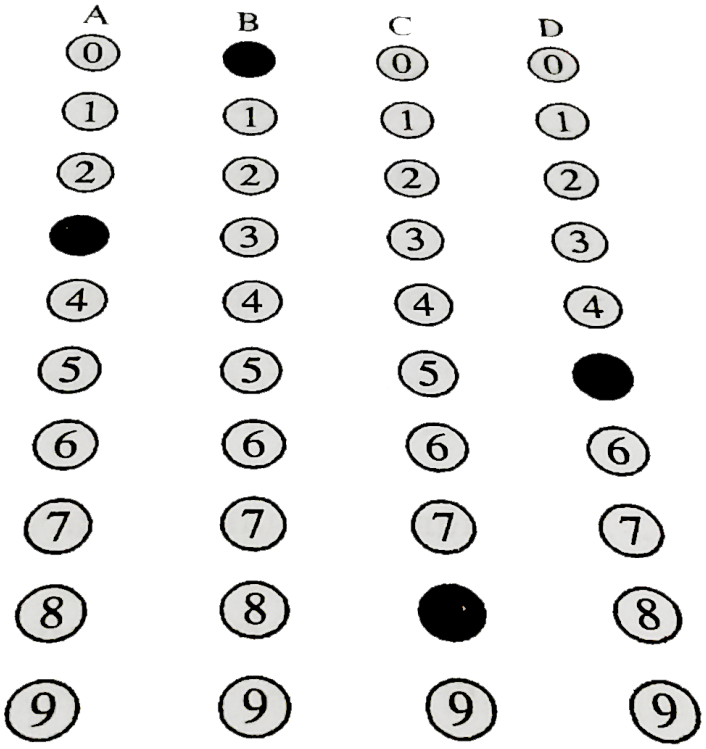
4. The answer to each of the following questions is a single digit integer from 0 to 9. If the correct answer to the question number A,B, C and D say 3,0,8,5 respectively, then the correct darkening of bubbles should be as follow.



Amonst the following, the ores which are roasted to convert them into their corresponding metal oxides are: Alumina, zinc, blende, iron pyrites, copper pyrites, galena, chalcorite

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Ultimate Preparatory Package

1. Which of the following is used for the seperation of ZnS and PbS during the froth floatation process ?

A. KCl

B. KCl

C. NH_4NO_3

D. None of these.

Answer: (B)

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2. The element generally extracted from sea-weed ashes is

A. Cl_2

B. Na

C. I_2

D. K

Answer: C

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3. A vey high cosmic abundance of Fe is due to

A. its ferro-magnetic nature

B. its high nuclear binding energy

C. its high reactivity towards O_2

D. None of these.

Answer: B

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4. The pine oil used in froth floatation process an example of

A. froth stabilizer

B. coagulant .

C. depressent

D. coagulant .

Answer: A

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5. Iron pyrites is generally roasted to obtain Fe_3O_4 first. $Fe-(3)O_4$ can be concentrated by

A. gravity seperation

B. froth floatation

C. leaching

D. magnetic seperation.

Answer: D

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6. Mond process is for the purification of

A. Mo

B. Mn

C. Ni

D. None of these.

Answer: C

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7. Silica is added to roasted copper pyrites during smelting in order to remove

A. cuprous oxide

B. cuprous sulphide

C. ferous oxide

D. ferous sulphide

Answer: C

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8. The non metal which is generally occur in the oxidised form in nature is

A. C

B. N

C. S

D. None of these.

Answer: D

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9. Direct reduction of Fe_2O_3 by CO is possible at temperature.

A. less than 1123 K

B. equal to or less than 1123 K

C. more than 1123 K

D. equal to or more than 1123 K.

Answer: B

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10. Direct reduction of Fe_2O_3 by Carbon is possible at temperature.

A. less than 1123 K

B. equal to or less than 1123 K

C. more than 1123 K

D. equal to or more than 1123 K.

Answer: B

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11. Elements of group 1 generally form chloride ores. The only other element which form chloride ores is

A. Cu

B. Ag

C. Pb

D. Hg

Answer: B

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12. Malachite is basic copper carbonate. The only other elelment forming basic carbonate ore is

A. Bi

B. Ba

C. Mn

D. Mo

Answer: B

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13. The element not extracted from silicate ores is

A. Th

B. U

C. Sc

D. Zr

Answer: B

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14. Matte in metallurgy is,

A. artificially produced oxides

B. artificially produced sulphides

C. natural sulphides

D. None of these.

Answer: B

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Brain Teasers 15

1. High purtiy can be obtained by using

A. thermite process

B. carbon reduction

C. electrolytic reduction

D. hydrogen reduction.

Answer: C

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Brain Teasers 16

1. Concentration of pentladite is carried out by using

A. gravity process

B. froth floatation process

C. electromagnetic process

D. Any of these.

Answer: B

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Brain Teasers 17

1. Dross is

A. infusible impurity in an easily fusible metal

B. non-volatile impurity in a volatile metal

C. non metallic impurity in a metal

D. None of these.

Answer: A

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Brain Teasers 18

1. Metal which cannot be refined by electrolytic refining process is

A. Cu

B. Cr

C. Ni

D. Al

Answer: B

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Brain Teasers 19

1. Metal which obtained by reduction with carbon is

A. Fe

B. Zn

C. Pb

D. Cu

Answer: D

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Brain Teasers 20

1. Metal which cannot be extracted by smelting process is

A. Pb

B. Fe

C. Zn

D. Al.

Answer: D

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Brain Teasers 21

1. Which of the following oxide cannot be reduced to metal by carbon ?

A. CuO

B. HgO

C. Al_2O_3

D. Fe_2O_3

Answer: C

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Brain Teasers 22

1. In our solar system

A. density of larger planets is more than smaller planets

B. density of smaller planets is more than larger planets

C. density of a planet does not depend upon its size density of all the planets is the same.

D. density of larger planets is more than smaller planets

Answer: B

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Brain Teasers 23

1. The only metal which exist in nature as well as combined state is

A. Al

B. Au

C. Pt

D. Ag

Answer: D

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Brain Teasers 24

1. The metal found in sediments in ocean floor is

A. Ag

B. Mg

C. Mn

D. Pt.

Answer: C

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Brain Teasers 25

1. Gold can easily dissolve in

A. conc.HNO_3

B. conc.HCl

C. KCN(aq) + O_2

D. None of these.

Answer: C

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Brain Teasers 26

1. Pick out the correct statement

A. Inner core of the earth is made up of almost pure solid iron

B. Inner core of the earth is made up of almost pure molten iron

C. Inner core of the earth is made up of molten rock

D. Inner core of the earth is made up of solid rock.

Answer: A

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Brain Teasers 27

1. The most abundant metal (by wkt) on earth's crust is

A. Na

B. Fe

C. Al

D. Ca

Answer: C

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Brain Teasers 28

1. The sulphide ore which is red in colour is

A. zinc blende

B. iron pyrites

C. galena

D. cinnabar.

Answer: D

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Brain Teasers 29

1. If the impurities present in an ore or the products obtained from them are basic, the refractory lining of the furnace should preferrably be

A. acidic

B. basic

C. neutral

D. amphoteric.

Answer: B

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Brain Teasers 30

1. Carbonates ores are mostly of

A. group 1 elements

B. group 2 elements

C. transition elements

D. group 13 elements

Answer: B

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Brain Teasers 31

1. The non metal which occur only in the oxidised form in nature is

A. carbon

B. nitrogen

C. chlorine

D. None of these.

Answer: A

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Brain Teasers 32

1. Artificially produced sulphides during metallurgical operations is known as

A. gangue

B. flux

C. regulus

D. None of these.

Answer: A

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Brain Teasers 33

1. Which is not an example of basic refractory material ?

A. Calcined limestone

B. Calcined dolamine

C. Magnetie

D. Sand stone.

Answer: A

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Brain Teasers 34

1. Which is not an example of acidic refractory ?

A. Silica

B. Quartz

C. Sand stone

D. Chromite

Answer: A

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Brain Teasers 35

1. Refractory material of good thermal conductivity is used in

A. blast furnace

B. muffle furnace

C. reverberatory furnace

D. open hearth furnace.

Answer: A

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Brain Teasers 36

1. Pick out the wrong statement. Roasting is done to

A. make the ore lumpy

B. make the ore porous

C. remove volatile impurities

D. reduce the oxide to metal.

Answer: A

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Unit Test 10

1. An ore after levigation is found to have acidic impurities. Which of the following can be used as flux during smelting operation.

A. H_2SO_4

B. $CaCO_3$

C. SiO_2

D. Both $CaCO_3$ and SiO_2

Answer: (B)

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Unit Test 11

1. Extraction of silver from Ag_2S by the use of sodium cyanide is example of

A. roasting

B. hydrometallurgy

C. electrometallurgy

D. Smelting

Answer: (B)

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Unit Test 12

1. Which of the following metal can be extracted by smelting ?

A. aluminium

B. magnesium

C. iron

D. None of these.

Answer: (C)

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Unit Test 13

1. Froth floatation process is used for the concentration of

A. oxide ores

B. sulphide ores

C. chlorine ores

D. amalgams

Answer: (B)

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Unit Test 14

1. Slag is formed by the reaction between

A. impurities and coke

B. impurities and ore

C. impurities and flux

D. flux and coke .

Answer: (C)

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Unit Test 15

1. The substance which added to the ore in order to remove impurities during smelting is called

A. slag

B. gangue

C. flux

D. matrix

Answer: (C)

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Unit Test 16

1. The reason for floating of ore particles in concentration by froth floatation is that

A. they are light

B. they are insoluble

C. they are charged

D. they are hydrophobic.

Answer: (D)

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Unit Test 17

1. Heating pyrites in air to remove sulphur is known as

A. Calcination

B. Roasting

C. Smelting

D. Bessimerisatiion

Answer: (B)

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Unit Test 18

1. Alkali metals can be extracted form their salts by

A. reduction with carbon

B. electrolysis of their fused halides

C. electrolysis of aqueous solution of their halides

D. reduction with aluminium.

Answer: (B)

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Unit Test 19

1. Zone refining process is used for the

A. Concentration of an ore.

B. Reduction of a metal oxide

C. Purification of metals

D. Purification of an ore.

Answer: (C)

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Unit Test 20

1. Which of the following statements is correct regrading the slag obtained during the extraction of a metal like copper or iron?

A. The slag is lighter and has lower melting point than the metal.

B. The slag is heavier and has lower melting point than the metal

C. The slag is lighter and has higher melting point than the metal

D. The slag is heavier and has higher melting point than the metal.

Answer: (A)

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Unit Test 21

1. In the purification of aluminium by Hoope's process, impurities of silicon and copper are added to molten aluminium in order to

A. make the melt conducting

B. lower the melting point of the melt

C. smooth deposit of aluminium

D. make the melt heavier.

Answer: (B)

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Unit Test 22

1. The slag consists of molten impurities, generally , in the form of

A. Metal carbonate

B. Metal silicate

C. Metal oxide

D. Metal nitrate

Answer: (B)

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Unit Test 23

1. In the metallurgical processes the flux used for removing acidic impurities is

A. Silica

B. Sodium chloride

C. Lime stone

D. Sodium carbonate

Answer: (C)

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Unit Test 24

1. Electrolyte reduction method is used in the extraction of

A. Highly electromagnetic elements

B. Highly electropositve elements

C. transition elements

D. Noble metals.

Answer: (B)

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Unit Test 25

1. Gravity seperation process may be used for the concentration of

A. Chalcopyrites

B. Bauxite

C. Haemattie

D. Calamine.

Answer: (C)

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Unit Test 26

1. Which method of purification is represented by the following equation?

$$Ti(s) + 2I_2(g) \xrightarrow{523K} Til_4(g) \xrightarrow{1700K} Ti(s) + 2I_2(g)$$

A. Cupellation

B. Poling

C. Van Arkel

D. Zone refining.

Answer: (C)

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Unit Test 27

1. Ellingham diagram represents:

A. Change of ΔG with temperature

B. Change of ΔH with temperature

C. Change ΔG with pressure

D. Change of $(\Delta G - T\Delta S)$ with temperature.

Answer: A

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Unit Test 28

1. To carry out a reduction process, select a temperature so as to make

A. δG negative

B. δG positive

C. δH negative

D. δH positive.

Answer: (A)

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Unit Test 29

1. The reduction of an oxide by aluminium is called

A. Ellingham process

B. Goldschmidt's alumino thermite process

C. Kroll's process

D. Van Arkel process.

Answer: (B)

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Unit Test 30

1. Select the correct statement

A. In the decomposition of an oxide into oxygen and gaseous metal, entropy increases.

B. Decomposition of an oxide is an endothermic change

C. To make δG^o negative , temperature should be high enough so that TδS^o gt δH^o

D. All are correct statements.

Answer: (D)

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Unit Test 31

1. Which does not represent correct method

A. $TiCl_2$ + Mg $\rightarrow$ Ti + MgCl : Kroll

B. $Ni(CO)_4$ $\rightarrow$ Ni + 4CO : Mond

C. Ag_2CO_3 $\rightarrow$ 2Ag + CO_2 + $\frac{1}{2}O_2$: Van Arkel

D. ZrI_4 $\rightarrow$ Zr + $2I_2$: Van Arkel

Answer: (C)

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Unit Test 32

1. Softening of lead means

A. melting pure lead at high temperature

B. removal of impurities , except silver, present in commercial lead

C. formation of lead only

D. formation of 100 % pure lead.

Answer: (B)

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Unit Test 33

1. Corondum is .....mineral of Al.

A. silicate

B. oxide

C. double salt

D. sulphate.

Answer: (B)

Watch Video Solution

Unit Test 34

1. In zone-refining methode the molten zone

A. consists of impurities alone

B. contains more impurities than the original metal

C. contains the purified metal only

D. move the either side.

Answer: (B)

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