



CHEMISTRY

BOOKS - S DINESH & CO CHEMISTRY (HINGLISH)

ALKALI EARTH METALS

Multiple Choice Questions

1. In alkaline earth metals, the electrons are more firmly held to the nucleus and hence

- A. atoms of alkaline earth metals are bigger than alkali metals
- B. ionization energy of alkaline earths is greater than alkali metals
- C. reactivity of alkaline earths is greater than alkali metals
- D. alkaline earths are less abundant in nature.

Answer: B



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2. Which of the following statements is correct?

A. Alkali metals are less electropositive than alkaline earth metals

B. The alkaline earth metals are denser and harder than alkali metals

C. The alkali metals are denser and harder than alkaline earth metals

D. The first ionization potential of alkali metals is more than that of alkaline earth metals

Answer: B



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3. The elements of group-2 are called alkaline earth metals because

- A. they are alkaline in nature
- B. they occur in earth's crust and form alkaline salts
- C. their oxides are alkaline and occur in earth's crust
- D. these are s-block elements

Answer: C



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4. Group 2 elements are

A. oxidising agents

B. reducing agents

C. oxidising as well as reducing agents

D. none of these

Answer: B



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5. Group 2 elements belong to

A. f-block

B. d-block

C. s-block

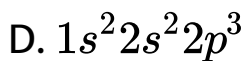
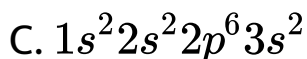
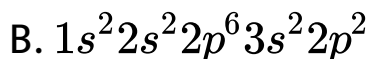
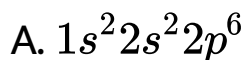
D. p-block

Answer: C



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6. which of the following configurations is correct for alkaline earth elements?



Answer: C



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7. The most abundant alkaline earth metal (in earth's crust) is

A. Radium

B. Calcium

C. Magnesium

D. Strontium

Answer: B



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8. The valence shell configuration for alkaline earth metals is represented by ns^2 where n varies from

A. 2 to 6

B. 2 to 7

C. 1 to 6

D. 1 to 7

Answer: B



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9. One of the following statements is incorrect

- A. Element of group 2 are good conductors of electricity and heat
- B. Compounds of group 2 elements are diamagnetic in nature
- C. The salts of group 2 elements are more heavily hydrated than those of elements of group 1
- D. Elements of group 2 are more electropositive than group 1 elements.

Answer: D



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10. Alkaline earth metals form dipositive ions instead of unipositive ions because

A. Dipositive ions carry more charge

B. Their second ionisation energies are not different from first ionisation energy

C. Unipositive ions do not have stable configuration

D. Dipositive ion have more hydration energy.

Answer: D



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11. Compounds of alkaline earth metals are less soluble in water than the corresponding alkali metal salts due to :

A. their high ionisation energy

B. their high lattice energy

C. their less basic character

D. their low electronegative values.

Answer: B



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12. $MgBr_2$ and MgI_2 are soluble in acetone because of

- A. their ionic nature
- B. their coordinate nature
- C. their metallic nature
- D. their covalent nature

Answer: D



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13. In case of alkaline earth metals. The oxidation state of more than two is not observed because

A. they have only two electrons in the outermost shell

B. the s-orbital can accommodate only two electrons

C. the removal of third electron involves breaking up of noble gas configuration and the energy needed for this purpose is extremely high

D. none of these

Answer: C



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14. On moving down the group, the reducing power of alkaline earth metals.

A. decreases

B. increases

C. remain unchanged

D. increases and then decreases

Answer: B



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15. The order of increasing lattice energy of the metallic compound is

A. $\text{NaCl} < \text{CaO} < \text{NaI} < \text{BaO}$

B. $\text{NaCl} \cdot \text{BaO} \cdot \text{CaO}$

C. $\text{NaCl} \cdot \text{NaCl} \cdot \text{BaO} \cdot \text{CaO}$

D. $\text{NaCl} \cdot \text{NaCl} \cdot \text{CaO} \cdot \text{BaO}$

Answer: B



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16. Which of the following is not the characteristic of Ba?

A. It emits electrons on exposure to light

B. It is a silvery white metal

C. It forms $Ba(NO_3)_2$ which is used in preparation of green fire

D. Its ionization potential is lower than radium

Answer: A



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17. A substance which gives a brick red flame and breaks down on heating to give oxygen and a brown gas is

A. Magnesium nitrate

B. Calcium carbonate

C. calcium nitrate

D. Magnesium carbonate

Answer: C



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18. Which salt gives crimson colour in flame?

A. $SrCl_2$

B. $CaCl_2$

C. NaCl

D. $MgCl_2$

Answer: A



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19. A fire work gave green light. It probably contained a salt of

A. Ca

B. Sr

C. Ba

D. Mg.

Answer: C



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20. Which of the following has exceptionally high boiling point?

A. MgO

B. NaHO

C. NaCl

D. KCl .

Answer: A



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21. Ca^{2+} ion is isoelectronic with

A. Na

B. Ar

C. Mg^{2+}

D. Sr^{2+}

Answer: B



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22. If Na^+ ion is larger than Mg^{2+} and S^{2-} ion is least soluble in water

- A. Sodium chloride
- B. sodium sulphide
- C. Magnesium chloride
- D. Magnesium sulphide

Answer: D



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23. The least abundant alkaline earth metal is

A. Barium

B. Radium

C. Strontium

D. Beryllium

Answer: B



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24. Which of the following is least reactive?

A. Calcium

B. Strontium

C. Barium

D. Radium

Answer: A



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25. Which of the following metallic chloride has covalent character?

A. Beryllium chloride

B. sodium chloride

C. Magnesium chloride

D. Barium chloride

Answer: A



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26. The atomic size of elements of group 2 compared to group 1 element are

A. equal

B. higher

C. smaller

D. none of these

Answer: C



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27. Select one of the following sets in which there are three members of group 2

A. Aluminium,sodium,potassium

B. Lithium,sodium,potassium

C. Magnesium,barium,calcium

D. Rubidium,cesium,francium

Answer: C



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28. The element with atomic no. 56 belong to

A. s-block

B. p-block

C. d-block

D. f-block

Answer: A



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29. Which of the following has the highest first ionization energy?

A. Ba

B. Mg

C. Ca

D. Be.

Answer: D



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30. The density of Ca is less than that of Mg because

- A. Nuclear charge of Ca is more than Mg
- B. Vacant 3d orbital is present in Ca
- C. Size of Ca is less than Mg
- D. none of these

Answer: B



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31. The most electronegative alkaline earth metal is

A. Be

B. Mg

C. Ca

D. Ra.

Answer: A



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32. The hydration energy of Mg^{+2} is greater than that of



Answer: B



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33. Which of the following is not an ore of magnesium?

A. Carnallite

B. Magnesite

C. Dolomite

D. Gypsum

Answer: D



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34. Important ore of magnesium is

A. Dolomite

B. Sylvine

C. Amblygonite

D. Triphylite

Answer: A



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35. Pure anhydrous $MgCl_2$ can be prepared from the hydrated salt by

- A. heating the hydrate to red heat in the atmosphere of HCl gas
- B. melting the hydrate
- C. heating the hydrate with coke
- D. heating the hydrate with Mg ribbon.

Answer: A

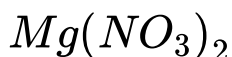


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36. Metallic magnesium is prepared by

A. Reduction of MgO by coke

B. Electrolysis of aqueous solution of



C. Displacement of Mg by iron from $MgSO_4$

solution

D. Electrolysis of molten $MgCl_2$

Answer: D



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37. Magnesium powder burns in air to give :

A. MgO

B. Mg_3N_2

C. MgCO_3

D. MgO and Mg_3N_2 both

Answer: D



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38. Epsom salt is

A. Magnesium sulphate

B. Calcium sulphate

C. Ferrous ammonium sulphate

D. Magnesium ammonium phosphate

Answer: A



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39. Anhydron, a drying agent is



D. $Mg(ClO_4)_2$

Answer: D



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40. Which of the following metal is used in the preparation of Grignard's reagent?

A. Ca

B. Mg

C. Cd

D. Both Mg.Cd.

Answer: B



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41. The sulphate of which metals given below has highest solubility in water

A. Ca

B. Ba

C. Sr

D. Mg.

Answer: D



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42. Dow metal contains

A. Mg,Al

B. Al,Mg,Zn

C. Mg,Cu,Zn

D. Mg,Cu

Answer: B



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43. Which of the following metals is present in chlorophyll?

A. Mg

B. Be

C. Ca

D. none of these

Answer: A



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44. Which of the following is used as an antacid?

A. MgO

B. $\text{Mg}(\text{OH})_2$

C. MgSO_4

D. MgCO_3

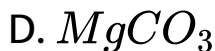
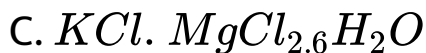
Answer: B



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45. Which of the following represent magnesite ore?

A. $\text{CaCO}_3 \cdot \text{MgCO}_3$



Answer: D



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46. A certain metal M is used to prepare an antacid, which is used as a medicine in acidity. This metal accidentally catches fire and it was found that the fire cannot be put out by using CO_2 based extinguishers. The metal M is

A. Ca

B. C

C. Mg

D. None

Answer: C



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47. Magnesium wire burns in the atomosphere of CO_2 because

A. Magnesium acts as an oxidising agent

B. Magnesium has 2 electrons in the outermost orbit

C. Magnesium acts as a reducing agent and removes oxygen from CO_2

D. Mg forms complex with CO_2

Answer: C



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48. A certain metal is present in the soil, plants, bones, egg shells, sea shells and coral. It is also

used to remove oxygen from molten steel and its hydroxide is used to detect CO_2 . The metal is

A. Mg

B. Al

C. Ca

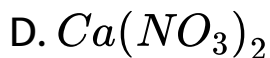
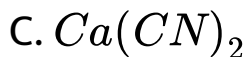
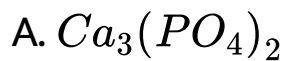
D. Na

Answer: C



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49. Nitrolim is



Answer: B



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50. Formula of gypsum salt is





Answer: B



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51. Calcium is obtained by

A. Electrolysis of molten $CaCl_2$

B. Electrolysis of a solution of $CaCl_2$ in water

C. Reduction of $CaCl_2$ with carbon

D. Roasting of lime stone.

Answer: A



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52. The metal that is extracted from sea water is

A. Fe

B. Ca

C. Mg

D. Au

Answer: C



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53. At high temperature, nitrogen combines with CaC_2 to give :

- A. Calcium cyanide
- B. Calcium cyanamide
- C. Calcium carbonate
- D. Calcium nitrate

Answer: B



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54. Slacked lime is used in the manufacture of

- A. Cement
- B. Fire bricks
- C. Pigments
- D. Medicines

Answer: A



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55. The element which is not present in asbestos, is

A. Ca

B. Ba

C. Mg

D. Si

Answer: B



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56. Calcium does not combine directly with

A. Oxygen

B. Nitrogen

C. Hydrogen

D. Carbon

Answer: B



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57. Which of the following is used as a scavenger in metallurgy?

A. Na

B. K

C. Ca

D. Zn

Answer: C



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58. which of the following is not a use of Epsom salt?

A. As a purgative

B. As a mordant in dyeing

C. As a stimulant to increase the secretion of
bile

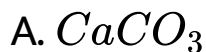
D. For removal of sulphur from petroleum

Answer: D



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59. What is X in the equation?



D. none of these

Answer: B



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60. Slacked lime is obtained from

- A. Lime stone
- B. Plaster of Paris
- C. Quik lime
- D. Nitrolim

Answer: C



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61. The milk of lime is

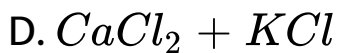
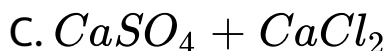
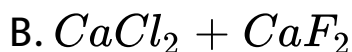
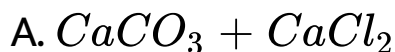
- A. Suspension of $CaCO_3$ in water
- B. Suspension of $Ca(OH)_2$ in water
- C. suspension of sand in water
- D. suspension of $CaSO_4$ in water.

Answer: B



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62. The composition of electrolyte for the manufacture of calcium by electrolytic method is



Answer: B



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63. Which of the following hydroxides is the weakest base?

A. LiOH

B. $\text{Ca}(\text{OH})_2$

C. KOH

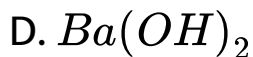
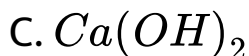
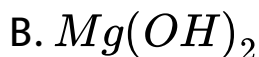
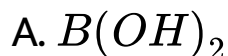
D. $\text{Sr}(\text{OH})_2$

Answer: A



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64. Which of the following hydroxides is the strongest base?

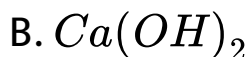
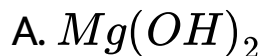


Answer: D



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65. Which of the following is the weakest base?



Answer: A



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66. A suspension of slaked lime in water is called

- A. Lime water
- B. Milk of magnesia
- C. Milk of lime
- D. None

Answer: C



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67. Which of the following statements is true about $Ca(OH)_2$?

- A. It is not used in the preparation of bleaching powder
- B. It is a light blue solid
- C. It is used in chromatography
- D. Its solution in water is called lime water

Answer: D



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68. Which of the following metal dissolves in sodium hydroxide with the evolution of hydrogen?

A. Calcium

B. Magnesium

C. Beryllium

D. Strontium

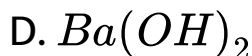
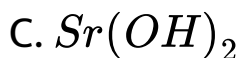
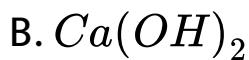
Answer: C



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69. Which of the following alkaline earth metal hydroxide is amphoteric in nature?

A. $Be(OH)_2$

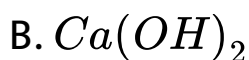
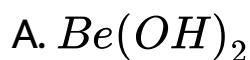


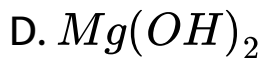
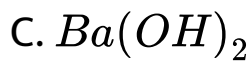
Answer: A



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70. Which out of the following will have the large value of solubility product/most soluble in water?



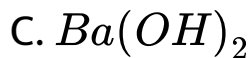


Answer: C



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71. Which out of the following represents Baryta?

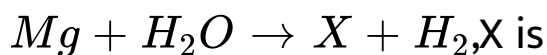


Answer: B



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72. In the reaction



A. MgO

B. $Mg(OH)_2$

C. MgH_2

D. none of these

Answer: A



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73. Superoxide of type MO_2 are formed by all except

A. Potassium

B. Beryllium

C. Strontium

D. Barium

Answer: B



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74. An amphoteric oxide dissolves in HCl to form a salt. The salt does not impart any colour to the flame and fumes in moist air. The oxide is



Answer: C



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75. Which of the following alkaline earth metal oxide is most basic?

A. BeO

B. MgO

C. CaO

D. BaO

Answer: C



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76. Which out of the following oxides will react both with acids and bases?

A. BeO

B. MgO

C. CaO

D. BaO

Answer: A



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77. which among the following does not at all show the tendency to form peroxides?

A. Lithium

B. Magnesium

C. Beryllium

D. Barium

Answer: A



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78. Which out of the following is extensively used as a refractory and in domestic heating appliances?

A. BaO

B. CaO

C. MgO

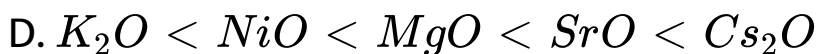
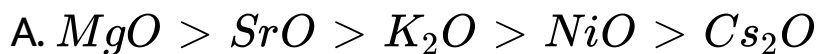
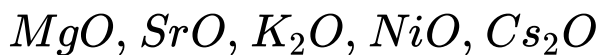
D. Na_2O

Answer: C



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79. Arrange the following in increasing order of basic strength :



Answer: C



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80. Which out of the following metal oxides is the most acidic?

A. CaO

B. MgO

C. Al_2O_3

D. Na_2O

Answer: C



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81. The nature of the oxide of radium is

A. Acidic

B. Basic

C. Neutral

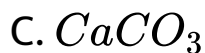
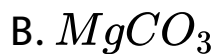
D. Amphoteric

Answer: B



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82. The carbonate that is least soluble in water



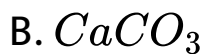
Answer: D



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83. Which out of the following does not exist in the solid state?



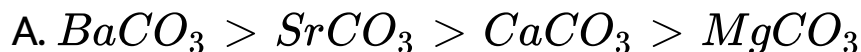


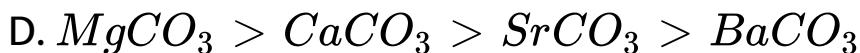
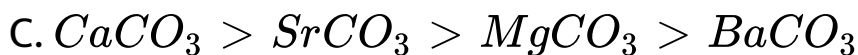
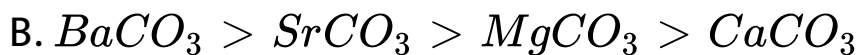
Answer: C



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84. The thermal stability of alkaline earth metal carbonates MgCO_3 , CaCO_3 , BaCO_3 and SrCO_3 decreases as:



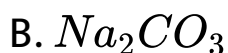


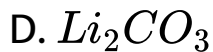
Answer: A



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85. Which of the following carbonates will not decompose on heating?



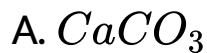


Answer: B



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86. (iii) Which is not a basic flux ?

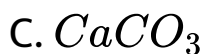
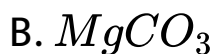


Answer: A



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87. Which of the following alkaline earth metal carbonate is thermally least stable?

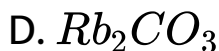
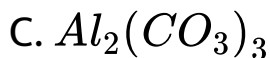
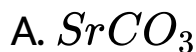


Answer: D



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88. One among the following carbonate is soluble in water

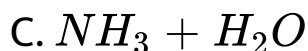
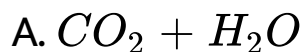


Answer: D



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89. Which of the following can dissolve limestone?



D. none of these

Answer: A



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90. Mortar is a mixture of slaked lime and sand in the ratio

A. 0.12569444444444

B. 0.04375

C. 0.08541666666667

D. 0.0444444444444444

Answer: B



View Text Solution

91. Concrete is a mixture of

A. cement, sand, gravel and water

B. cement, lime and water

C. Cement, sand and water

D. None

Answer: A



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92. By adding gypsum to cement

- A. Decrease the rate of setting of cement
- B. Bind the particles of calcium silicate
- C. Facilitate the formation of colloidal gel
- D. All the above

Answer: A



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93. The function of sand in mortar is :

- A. to decrease the hardness

- B. to prevent excessive shrinkage because of which cracks may result
- C. to increase hardness
- D. to make the mass compact

Answer: B



View Text Solution

94. Setting of cement is an

- A. Exothermic reaction
- B. Endothermic reaction

C. Neither exothermic nor endothermic

D. None

Answer: A



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95. Setting of plaster of Paris involves

A. Oxidation with atmospheric oxygen

B. Combination with atmosphere CO_2

C. Dehydrartion

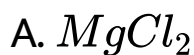
D. Hydration to yield another hydrate

Answer: D



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96. The minimum equivalent conductance in fused state is shown by-



Answer: B



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97. Ground water is hard in the region of

- A. Coal depositss
- B. Petroleum deposits
- C. Limestone deposits
- D. none of the above

Answer: C



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98. Which of the following does not react with water?

A. BeO

B. CaO

C. MgO

D. SrO

Answer: A



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99. The behavior of Be salt in pure water is

A. Basic

B. Acidic

C. Amphoteric

D. Sometimes basic and sometimes acidic
depending upon its concentration.

Answer: C



View Text Solution

100. Peroxide and super oxides of alkali metal are coloured because of

- A. Defects in crystals
- B. Presence of unpaired electrons
- C. Partially due to the defects and partially due to presence of unpaired electrons
- D. none of the above

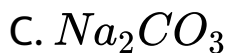
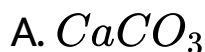
Answer: CA



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101. A solid compound ' X ' on heating gives CO_2 gas and a residue. The residue mixed with water

form 'Y'. On passing an excess of CO_2 through 'Y' in the water, a clear solution 'Z' is obtained. On boiling 'Z' compound 'X' is reformed. The compound 'X' is :

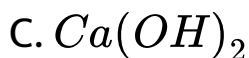


Answer: B



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102. A compound of calcium (X) is used in sugar industry for the purification of sugar. When exposed to an oxy-hydrogen flame, it becomes incandescent and starts emitting white light. On treatment with CO_2 it forms a compound, which can be decomposed to give back X at very high temperature. X is

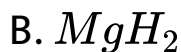


Answer: C



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103. The hydride from amongst the following that cannot be obtained directly by reaction with hydrogen is



Answer: C



Watch Video Solution

104. Hydrolith when dissolved in water liberates hydrogen. The formula of hydrolith is

A. NaH

B. CH_4

C. C_2H_6

D. CaH_2

Answer: D



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105. Which metal does not form ionic hydride-

A. Ba

B. Sr

C. Ca

D. Be.

Answer: D



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106. The solubilities of sulphates of alkaline earth metals decrease from Be to Ba because

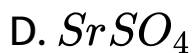
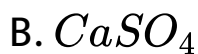
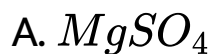
- A. Their lattice energies decrease in this order
- B. their lattice energies increase in this order
- C. lattice energies are about the same
- D. hydration energies of Be^{2+} to Ba^{2+} cations decrease in this order

Answer: D



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107. Alkaline earth sulphate that is least soluble in water



Answer: C



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108. A white salt gives a white precipitate with barium chloride solution which is insoluble in ammonia and hot water. The salt is

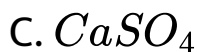
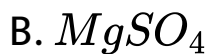
- A. Silver nitrate
- B. Magnesium sulphate
- C. Lead Acetate
- D. Potassium chromate

Answer: B



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109. which out of the following will have largest value of K_{sp} ?

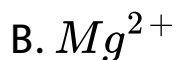


Answer: A



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110. Which of the following alkaline earth metal ions has the highest ionic mobility in aqueous solution?



Answer: D



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111. When hydrated $MgCl_2 \cdot 6H_2O$ is strongly heated,

- A. MgO is formed
- B. $Mg(OH)_2$ is formed
- C. $Mg(OH)Cl$ is formed
- D. Anhydrous $MgCl_2$ is formed

Answer: B



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112. Calcium carbide is obtained by heating

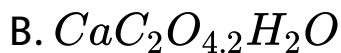
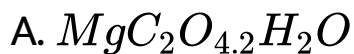
- A. Calcium carbonate
- B. A mixture of quick lime and coke
- C. A mixture of slaked lime and coke
- D. none of the above

Answer: B



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113. An oxalate of alkaline earth metal is insoluble in bases but soluble in dilute strong acid. It is also a component of most kidney stones. It is



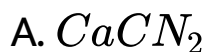
D. none

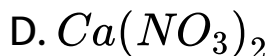
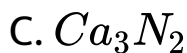
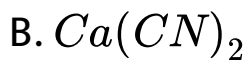
Answer: B



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114. When calcium carbide reacts with N_2 at $1000^\circ C$, the compound formed is





Answer: A



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115. Calcium metal is used to produce high vacuum because it

A. can remove water

B. can remove both O_2 and N_2

C. is a good reductant

D. is highly electropositive

Answer: B



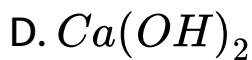
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116. calcium cyanamide on treatment with steam under pressure gives ammonia and

A. CaO

B. $\text{Ca}(\text{HCO}_3)_2$

C. CaCO_3



Answer: C



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117. Which out of the following statements is not correct for anhydrous calcium chloride ?

A. It is prepared by heating hydrated calcium chloride above 533K

B. It is used for drying alcohol and NH_3

C. It is used as a dehydrating agent to control snow and ice on highway and pavements

D. When mixed in concrete, it gives quicker initial setting and improves its strength

Answer: B



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118. Both Be and Al become passive on reaction with conc. nitric acid due to-

A. the non reactive nature of the metal

B. the non reactive nature of the acid

C. the formation of an inert layer of oxide on
the surface of the metals

D. none of these

Answer: B



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119. Which of the following statements about Be is
not correct?

A. Be has a high charge/size ratio

B. Be shows diagonal relationship with Al

C. Be forms $BeSO_4$ which is soluble in water

D. Be is unstable in air.

Answer: B



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120. BeF_2 is soluble in water whereas the fluorides of other alkaline earth metals are insoluble because of

A. ionic nature of BeF_2

B. greater hydration energy of Be^{2+} ion as compared to crystal lattice

C. covalent nature of BeF_2

D. none of these

Answer: B



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121. Which of the following is a covalent oxide?

A. CaO

B. SrO

C. MgO

D. BeO

Answer: D



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122. Pick up the wrong statement

A. Be like Al does not dissolve in alkalies

B. Oxides of both Be and Al are amphoteric

C. Beryllium chloride like Aluminium chloride is
covalent

D. Burning of limestone containing more than 10% aluminium silicates give hydraulic mortar.

Answer: A



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123. $BeSO_4$ is soluble in water but $BaSO_4$ is insoluble because

A. $BeSO_4$ is ionic while $BaSO_4$ is covalent

B. $BeSO_4$ is crystalline while $BaSO_4$ is amorphous

C. $BeSO_4$ has smaller lattice energy and high heat of hydration as compared to $BaSO_4$

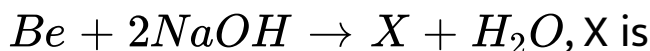
D. $BeSO_4$ is ionic while $BaSO_4$ is covalent

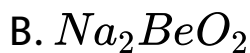
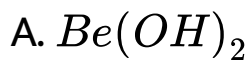
Answer: C



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124. In the reaction





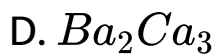
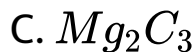
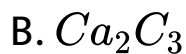
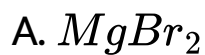
D. None of these

Answer: B



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125. An allylide on hydrolysis gives C_3H_4 (allylene). The alkaline earth metal cation of allylide dissolves in dry ether in presence of alkyl halide to form Grignard's reagent. The allylide is



Answer: C



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126. The most efficient method of extraction of beryllium is

A. reduction of beryllium halide with magnesium

B. reduction of beryllium oxide with carbon

C. electrolysis of fused beryllium chloride

D. dissociation of beryllium carbide

Answer: C



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127. Which of the following carbide can be used to prepare methane by its action with water?

A. CaC_2

B. Be_2C

C. MgC_2

D. None

Answer: B



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128. Which element of group 2 is used in treatment of cancer?

A. Magnesium

B. Radium

C. Strontium

D. Beryllium

Answer: B

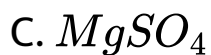


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129. Which of the following is used for taking the X – ray spectra of the digestive system:

A. $CaSO_4$

B. $BaSO_4$



Answer: B



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130. Flash bulbs contain a foil of magnesium packed in the atmosphere of

A. Nitrogen

B. Air

C. Halogens

D. SO_2

Answer: A



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131. Magnesium wire continues to burn in the atmosphere of CO_2 though $C)_2$ is not a supporter of combustion because

A. Mg acts as an oxidising agent

B. Mg has high ionization energy

C. Mg acts as a reducing agent and removes oxygen from CO_2

D. All of the above

Answer: C



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132. The mixture of $MgCl_2$ and MgO , called as Sorrel cement, is used in

A. Match industry

B. making houses

C. dental filling

D. portland cement

Answer: C



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133. Which of the following metals is present in chlorophyll ?

A. Magnesium

B. Iron

C. Sodium

D. Beryllium

Answer: A



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134. A piece of magnesium ribbon was heated to redness in an atmosphere of nitrogen and then cooled with water. The gas evolved is

A. Hydrogen

B. Oxygen

C. Nitrogen

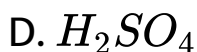
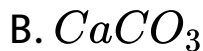
D. Ammonia

Answer: D



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135. Which out of the following drying agents will react with CO_2 and remove water vapours?



Answer: B



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136. The alloy Electron used in the construction of air crafts has the composition

A. 15% Mg, 85% Al

B. 95% Mg, 5% Zn

C. 95% Mg, 5% Al

D. 80%Mg, 20% Zn

Answer: B



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137. Which of the following is used as barium metal for getting the X -ray spectrum of the human digestive system?

A. BaO

B. $BaSO_4$

C. $BaCO_3$

D. BaS

Answer: B



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138. Bone ash contains

A. CaO

B. $\text{Ca}_3(\text{PO}_4)_2$

C. CaSO_4

D. $\text{Ca}(\text{H}_2\text{PO}_4)_2$

Answer: B



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139. The plaster of paris is hardened by

- A. liberating CO_2
- B. changing into $CaCO_3$
- C. uniting with CaO
- D. uniting with water

Answer: D



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140. Which of the following on thermal decomposition yields a basic as well as acidic oxide?

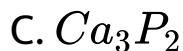


Answer: B



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141. Which out of the following is called 'Blue John' ?



Answer: B



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142. calcium cyanamide on treatment with steam under pressure gives ammonia and

- A. Calcium carbonate
- B. Calcium hydroxide
- C. Calcium oxide
- D. Calcium bicarbonate

Answer: A



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143. Which of the following is insoluble in acetic acid?

- A. Calcium oxide
- B. Calcium carbonate
- C. Calcium oxalate
- D. Calcium hydroxide

Answer: D



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144. The approximate percentage of gypsum in Portland cement is

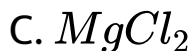
- A. 10-15%
- B. 20-30%
- C. 30-40%
- D. 2-3%

Answer: D



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145. Which of the following exists as polymeric chain in the solid state?

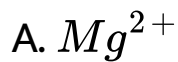


Answer: D



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146. Ca^{2+} is isoelectronic with



Answer: C



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147. which of the following elements has the highest value of second ionization energy?

A. Lithium

B. Beryllium

C. Boron

D. Magnesium

Answer: A



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148. The use of $BaSO_4$ in taking X-ray pictures of digestive tract is based on

- A. Insolubility in water
- B. great scattering of X-rays by Ba^{2+} ions
- C. Both of these
- D. None of these

Answer: C



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149. Alkaline earth metals dissolve in liquid NH_3 ,
the colour of the solution obtained is

A. Blue

B. reduction of beryllium oxide with carbon

C. purple

D. Pink

Answer: A



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1. Setting of cement is an

- A. Exothermic reaction
- B. Endothermic reaction
- C. Neither exothermic nor endothermic
- D. None

Answer: A



Watch Video Solution

2. Setting of plaster of paris is

A. Oxidation with atmospheric oxygen

B. Combination with atmosphere CO_2

C. Dehydrartion

D. Hydration to yield another hydrate

Answer: D



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3. The formula of calcium cyanamide is

A. $Ca(CN)_2$

B. CaC_2N

C. CaNCN

D. CaCHNH_2

Answer: C



Watch Video Solution

4. calcium cyanamide on treatment with steam under pressure gives ammonia and

A. Calcium carbonate

B. Calcium hydroxide

C. Calcium oxide

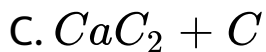
D. Calcium bicarbonate

Answer: A



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5. Nitrolim is



D. All

Answer: A



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6. Magnesium keeps on burning in

A. N_2

B. CO_2

C. N_2O

D. N_2 as well as CO_2

Answer: D



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7. Of the following the commonly used as a laboratory desiccator is

A. Anhyd. Na_2CO_3

B. Anhyd. $CaCl_2$

C. Dry NaCl

D. none of the above

Answer: B



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8. Portland cement is manufactured by using-

- A. Limestone, clay and sand
- B. Limestone, gypsum and sand
- C. Limestone, gypsum and alumina
- D. Limestone, clay and gypsum

Answer: D



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9. Drying agent which reacts with CO_2 and involves water vapour is

A. CaO

B. CaCl_2

C. CaCO_3

D. $\text{Ca}(\text{NO}_3)_2$

Answer: AC



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10. Metal which is used in Flash Bulbs?

A. Mg

B. Ba

C. Cu

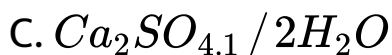
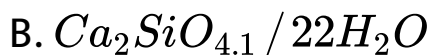
D. Ag.

Answer: C



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11. Plaster of paris is

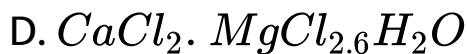
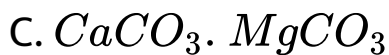
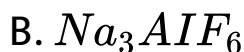


Answer: C



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12. Dolomite is:-



Answer: C



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13. Bleaching action of Bleaching powder is due to the liberation fo



Answer: C



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14. Which of the following is different from other three oxides?

A. MgO

B. SnO

C. ZnO

D. Cr_2O_3

Answer: A



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15. Plaster of Paris hardens by

- A. Giving off CO_2
- B. changing into $CaCO_3$
- C. combining with water
- D. Giving out water

Answer: C



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16. Bleaching powder is a compound having a formula

- A. $CaCO_3$

B. CaClO

C. CaOCl_2

D. CaOCl_3

Answer: C



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17. which one of the following statement is correct for alkaline earth metal?

A. They are diatomic and form ions of the type



B. They are highly electronegative elements

C. They are monoatomic and form ions of the type M^{+2}

D. They are diatomic and form ions of the type M^{2+}

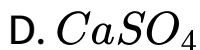
Answer: C



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18. Gypsum $CaSO_4 \cdot 2H_2O$ on heating to about $120^\circ C$ forms a compound which has the chemical

composition represented by



Answer: C

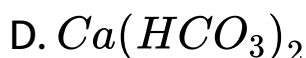
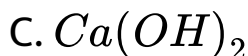
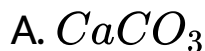


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19. Chemical A is used for water softening to remove temporary hardness. A reacts with sodium

carbonate to generate caustic soda. When CO_2 is bubbled through a solution of A, it turns cloudy.

What is the chemical formula of A ?



Answer: C



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20. Which of the following is true for magnesium?

- A. It is more electropositive than sodium.
- B. It is manufactured by electrolysis of aqueous magnesium chloride.
- C. It is a strong reducing agent.
- D. It resembles, in chemical properties, with its diagonally placed element Boron in III group of the periodic Table.

Answer: C



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21. Epsom salt is the hydrate of one of the following:

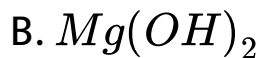
- A. Magnesium sulphate
- B. Ferrous ammonium sulphate
- C. Magnesium ammonium phosphate
- D. Calcium sulphate

Answer: A



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22. Epsom salt's chemical formula is

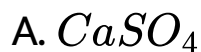


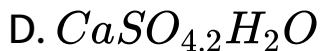
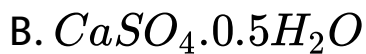
Answer: A



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23. Formula of gypsum salt is





Answer: D



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24. Which of the following compound of cement sets at the slowest rate?

A. Dicalcium silicate

B. sTricalcium silicate

C. Tricalcium aluminate

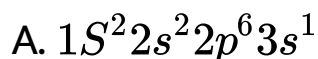
D. Tetracalcium aluminoferrite

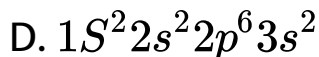
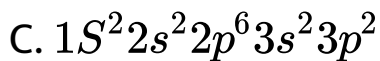
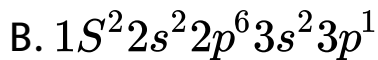
Answer: A



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25. A sudden large jump between the values of second and third ionisation energies of an element would be associated with the electronic configuration





Answer: D



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26. Which is the correct order of increasing basic strength?



C. $\text{BaOltCaOltMgOltBeO}$

D. $\text{CaOltBaOltBeOltMgO}$

Answer: B



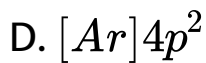
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27. Electronic configuration calcium atom can be written as

A. $[\text{Ne}]4p^2$

B. $[\text{Ar}]4s^2$

C. $[\text{Ne}]4s^2$

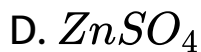
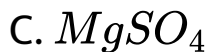
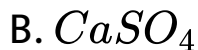


Answer: B



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28. Plaster of paris is hydrate of



Answer: B



Watch Video Solution

29. Which of the following on heating at 125°C gives Plaster of Paris?

A. Borax

B. Gypsum

C. Alum

D. Calomel

Answer: B



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30. Which is used in the laboratory for last drying of neutral gases?

A. Phosphorus pentoxide

B. Active charcoal

C. Anhydrous calcium chloride

D. Na_3PO_4

Answer: C



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31. The atomic numbers of four elements are given below. Which one is an alkaline earth metal?

A. 10

B. 20

C. 30

D. 40

Answer: B



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32. The mixture of $MgCl_2$ and MgO is called _____.

A. Portlant cement

B. Sorrel's cement

C. Double salt

D. None

Answer: B



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33. An atom with atomic number 20 is most likely to combine chemically with the atom whose atomic number is

A. 11

B. 14

C. 16

D. 10

Answer: C



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34. Except lime, a major constituent of Portland cement is

A. silica

B. Alumina

C. Iron oxide

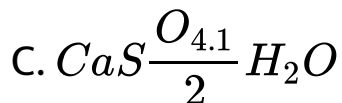
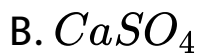
D. Magnesia

Answer: A



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35. Gypsum on heating to $390K$ gives



Answer: B

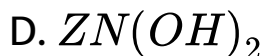
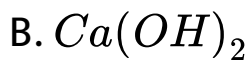


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36. Weakest base among KOH, NaOH,

$Ca(OH)_2$ and $Zn(OH)_2$ is :





Answer: D



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37. Electrolysis of fused anhydrous

$KCl \cdot MgCl_{2.6}H_2O$ gives

A. Potassium only

B. Magnesium only

C. Magnesium and chlorine

D. Potassium and Magnesium

Answer: C



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38. Beryllium is placed above magnesium in the *II* group. Beryllium dust, therefore, when added to $MgCl_2$ solution will:

A. have no effect

B. precipitate Mg metal

C. precipitate Mg

D. leads to the dissolution of beryllium metal.

Answer: A



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39. Amongst the metal Be, Mg, Ca and Sr of group 2 of the periodic table, the least ionic chloride would be formed by

A. Mg

B. Be

C. Ca

D. Sr.

Answer: B



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40. Among the following oxides, the one which is most basic is

A. Zinc oxide

B. Magnesium oxide

C. Aluminium oxide

D. Nitrogen pentoxide

Answer: B



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41. Presence of which of the following salt increases the rate of setting of Plaster of Paris.

A. NaCl

B. KCl

C. $BaSO_4$

D. $CuSO_4$

Answer: A



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42. Which of the following will liberate hydrogen by its reaction with hydrochloric acid?

A. Copper

B. phosphorus

C. Mercury

D. Magnesium

Answer: D



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43. White heavy precipitates are formed when $BaCl_2$ is added to a clear solution of compound A. Precipitates are insoluble in dilute HCl. Then, the compound A is :

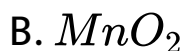
- A. a nitrate
- B. a bromide
- C. a sulphate
- D. a carbonate

Answer: C



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44. Which of the following is a true peroxide?



Answer: C



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45. Which of the following alkaline earth metals is the strongest reducing agent?

A. Ca

B. Sr

C. Ba

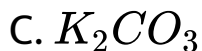
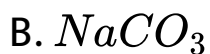
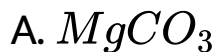
D. Mg.

Answer: C



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46. Which of the following carbonates will not decompose on heating?



D. All

Answer: A



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47. Sodium sulphate is soluble in water but barium sulphate is insoluble because

A. The hydration energy of Na_2SO_4 is more than its lattice energy

B. the lattice energy of $BaSO_4$ is more than its hydration energy

C. the lattice energy has no role to play in solubility

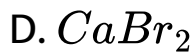
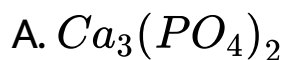
D. The lattice energy of Na_2SO_4 is less than its lattice energy. Or (E) both A and B

Answer: E



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48. White enamel of our teeth is



Answer: B



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49. Identify the correct statement:

- A. Gypsum contains a lower percentage of calcium than Plaster of Paris
- B. Gypsum is obtained by heating Plaster of Paris
- C. Plaster of Paris can be obtained by hydration of gypsum
- D. Plaster of Paris is obtained by partial oxidation of gypsum.

Answer: A



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50. Blanc fixe used in paints is

A. Finely divided $BaSO_4$

B. Paste of $Mg(OH)_2$

C. Suspension of slaked lime

D. $MgCl_{2.5}MgO.5H_2O$

Answer: A



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51. Mg burns in air to give

A. MgO

B. Mg_3N_2

C. $MgCO_3$

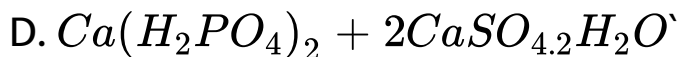
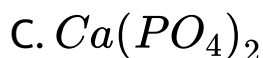
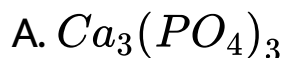
D. MgO and Mg_3N_2

Answer: D



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52. Which of the following is super phosphate of lime?



Answer: D



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53. Epsom salt's chemical formula is

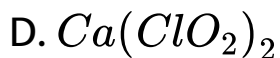
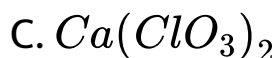
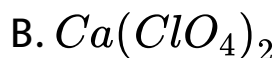


Answer: A



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54. Which of the following represent calcium chloride?



Answer: D



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55. Alkaline earth metals comes under

- A. Halogens
- B. Representative elements
- C. Transition elements
- D. Inner transition elements

Answer: B



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56. Calcium is obtained by the

A. roasting of lime stone

B. electrolysis fo solution of calcium chloride in water

C. reduction fo calcium chloride with carbon

D. electrolysis fo molten anhydrous or fused calcium chloride.

Answer: D



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57. A metal is burnt in air and the ash on moistening smells of ammonia. The metal is

A. Na

B. Ferrous ammonium sulphate

C. Mg

D. Al.

Answer: C



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58. What is X in the following reaction?

A. MgO

B. Mg

C. $\text{Mg}(\text{OH})_2$

D. $\text{Mg}(\text{OH})\text{Cl}$

Answer: A



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59. The difference of number of water molecules in gypsum and plaster of Paris is

A. $\frac{5}{2}$

B. 2

C. $\frac{1}{2}$

D. $1\frac{1}{2}$

Answer: D



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60. For two ionic solids, CaO and KI , which of the following statements is false?

- A. Lattice energy of CaO is much higher than that of KI .
- B. KI is soluble in benzene
- C. CaO has high m.p
- D. KI has high m.p

Answer: B



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61. The first ionization energies of alkaline earth metal are higher than those of the alkali metals.

This is because:

- A. there is increase in the nuclear charge of the
alkaline earth metals
- B. there is decrease in the nuclear charge of the
alkaline earth metals
- C. there is no change in the nuclear charge
- D. none of the above

Answer: A



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62. One which is not dissolved by dilute hydrochloric acid is

A. ZnS

B. MnS

C. BaCO_3

D. BaSO_4

Answer: D



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63. Which one of the salts does not impart a colour to the flame or the flame test is not shown by

A. LiCl

B. KI

C. $MgCl_2$

D. $CaCl_2$

Answer: C



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64. The one which does not show variable valency is

A. Barium

B. Titanium

C. Copper

D. Lead

Answer: A



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65. The characteristic outer electronic configuration of alkaline earth metals is

A. ns^1

B. ns^2

C. ns^2np^2

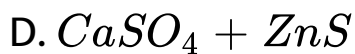
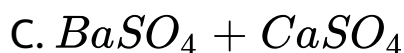
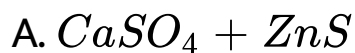
D. ns^2np^1

Answer: B



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66. Lithophone used as white pigment is a mixture of



Answer: B



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67. Alkaline earth metals form ions of the formula



Answer: A



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68. Metals belonging to the same group in the periodic table are

- A. Magnesium and sodium
- B. magnesium and copper
- C. magnesium and barium
- D. magnesium and potassium

Answer: C



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69. The compounds of alkaline earth metals have the following magnetic nature:

- A. Diamagnetic
- B. Paramagnetic
- C. Ferromagnetic
- D. Antiferromagnetic

Answer: A



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70. A chloride dissolves appreciably in cold water.

When placed on platinum wire in Bunsen flame, no distinctive colour is noticed, the cation would be



Answer: A



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71. Which of the following is incorrect?

A. Mg burns in air releasing dazzling light rich in UV rays

B. $\text{CaCl}_{2.6}\text{H}_2\text{O}$ when mixed with ice gives, freezing mixture

C. Mg cannot form complexes

D. Be can form complexes due to its very small size

Answer: C



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72. Halides of alkaline earth metals form hydrates such _____ as $MgCl_2 \cdot 6H_2O$, $CaCl_2 \cdot 6H_2O$, $BaCl_2 \cdot 2H_2O$ and $SrCl_2 \cdot 2H_2O$. This shows that halides of group 2 elements :

- A. are hygroscopic in nature
- B. act as dehydrating agent
- C. can absorb moisture from air
- D. All of the above

Answer: D



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73. Which of the following substance can be used for drying neutral or basic gases?

- A. Calcium carbonate
- B. Sodium carbonate
- C. Sodium bicarbonate
- D. Calcium oxide

Answer: D



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74. Bleaching powder loses its power on keeping for a long time because

- A. It change into calcium hypochlorate
- B. It changes into $CaCl_2$ and $Ca(OH)_2$
- C. It absorbs moisture
- D. It changes into calcium chloride and calcium chlorate.

Answer: D



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75. Among the alkaline earth metals, the element forming predominantly covalent compound is

A. Barium

B. Strontium

C. calcium

D. Beryllium

Answer: D



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76. Plaster of Paris is used

- A. as a plaster for walls
- B. in dentistry and surgery
- C. in metallurgical process
- D. as a drying agent

Answer: B



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77. Magnesium is an important component of which biomolecule occurring extensively in living world?

A. Haemoglobin

B. Chlorophyll

C. Florigen

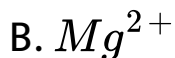
D. ATP.

Answer: B



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78. Which of the following is the smallest cation?



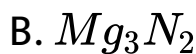
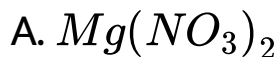
Answer: D



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79. A metal X on heating in nitrogen gas gives Y , Y on treatment with H_2O gives a colourless

gas which when passed through $CuSO_4$ solution gives a blue colour. Y is:

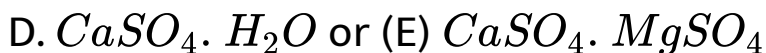
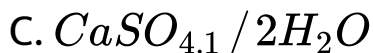


Answer: B



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80. Formula of plaster of Paris is



Answer: C



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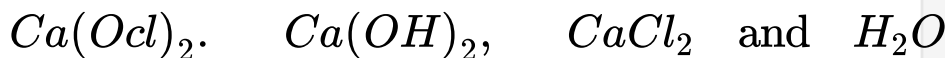
81. The reaction of slaked lime with Cl_2 gas gives



C. A

mixture

of



D. Quick lime or (e) Baryta water.

Answer: C



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82. The nitride salt of Ca when treated with H_2O gives

A. N_2

B. CaO



Answer: D



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83. The name of $CaS\frac{O_{4.1}}{2}H_2O$ is

A. Gypsum

B. Epsom salt

C. Plaster of Paris

D. Dolomite

Answer: C



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84. When one mole of bleaching powder is completely decomposed, then the mass of chlorine gas that is liberated will be

A. 35.45g

B. 70.90g

C. 17.72g

D. 88.60g

Answer: B



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85. In the presence of cobalt chloride ($CoCl_2$), bleaching powder decomposes to form

A. $CaCO_3$ and O_3

B. ClO_2 and CaO

C. Cl_2O and CaO

D. $CaCl_2$ and O_2

Answer: D



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86. Limestone is not used in which of the following manufacturing processes?

- A. phosphorus from phosphorite
- B. Ordinary (soda lime) glass
- C. Iron from haematite
- D. Solvay process of sodium carbonate

Answer: A



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87. On heating quicklime with coke in an electric furnace, we get

A. Ca and CO_2

B. $CaCO_3$

C. CaO

D. CaC_2

Answer: D



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88. Magnesium can be obtained by

- A. Reducing MgO with coke
- B. Reducing magnesium salt solution with Fe
- C. Electrolysis of fused magnesium salt
- D. Electrolysis of $Mg(NO_3)_2$ solution

Answer: C



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89. The metal present in chlorophyll is _____

A. Chromium

B. Cobalt

C. Magnesium

D. Iron

Answer: C



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90. A solution of $MgCl_2$ in water has pH

A. $\text{lt}7$

B. $\text{gt}7$

C. 7

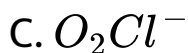
D. 14.2

Answer: A



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91. Bleaching action of Bleaching powder is due to the liberation of



D. O^{2-} or (E) Cl^{-}

Answer: B



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92. 30g Mg and 30g O_2 are reacted and the residual mixture contains:

A. 60g of MgO only

B. 40g of MgO and 20g of O_2

C. 45g of MgO and 15g of O_2

D. 50g of MgO and 10g of O_2

Answer: D



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93. Which of the following salt will give a green colour in fire works?

A. Sodium

B. potassium

C. Barium

D. Calcium

Answer: C



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94. Bleaching powder is obtained by the interaction of Cl_2 and

A. Dry slaked lime

B. hot $Ca(OH)_2$

C. cold $Ca(OH)_2$

D. conc. $CaCl_2$

Answer: A



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95. Carnallite on electrolysis gives

A. Ca and Cl_2

B. Na and CO_2

C. Al and Cl_2

D. Mg and Cl_2 or (E) K and CO_2

Answer: D



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96. The solubilities of carbonates decreases down the magnesium group due to a decrease in

A. entropy of solution formation

B. lattice energies of solids

C. hydration energies of cation

D. Inter ionic attraction

Answer: C



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97. The substance not likely to contain $CaCO_3$ is:

- A. dolomite
- B. a marble statue
- C. calcined gypsum
- D. sea shells.

Answer: C



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98. Several blocks of magnesium are fixed to the bottom of a ship to

- A. prevent puncturing by under sea rock
- B. Keep away the sharks
- C. make the ship lighter
- D. to prevent the action of water and salt.

Answer: D



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99. In curing cement plasters, water is sprinkled from time to time. This helps in

- A. converting sand into silicic acid
- B. keeping it cool
- C. developing interlocking needle like crystals of hydrated silicates.
- D. hydrating sand and gravel mixed with cement

Answer: C



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100. The correct order of solubility of the sulphates of alkaline earth metals in water is

A. $\text{BaSO}_4 < \text{CaSO}_4 < \text{MgSO}_4 < \text{SrSO}_4$

B. $\text{MgSO}_4 < \text{BaSO}_4 < \text{CaSO}_4 < \text{SrSO}_4$

C. $\text{BaSO}_4 < \text{MgSO}_4 < \text{CaSO}_4 < \text{SrSO}_4$

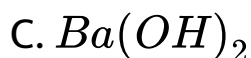
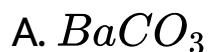
D. $\text{MgSO}_4 < \text{CaSO}_4 < \text{BaSO}_4 < \text{SrSO}_4$

Answer: C



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101. The product obtained on fusion of $BaSO_4$ and Na_2CO_3 is

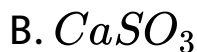


Answer: A



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102. Dead burnt plaster is



Answer: D



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103. Which of the following is not an ore of magnesium?

A. gypsum

B. Magnesite

C. Dolomite

D. Carnallite

Answer: C



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104. Pick out statement (s) which is/are not true about diagonal relationship of Li and Mg:

A. Polarising powers of Li^{\oplus} and Mg^{2+} ions are almost the same.

B. Like Li, Mg decomposes water very fast.

C. LiCl and $MgCl_2$ are deliquescent.

D. Like Li, Mg readily reacts with liquid bromine at ordinary temperature.

A. 1 and 4

B. 2 and 3

C. only 2

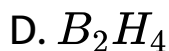
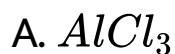
D. only A or (E) 2 and 4

Answer: E



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105. Which of the following exists in polymeric form

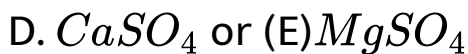
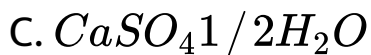


Answer: B



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106. Dead burnt plaster is

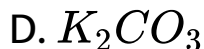
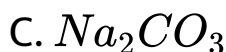
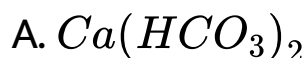


Answer: D



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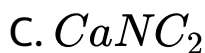
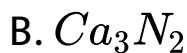
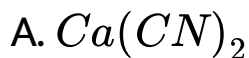
107. A solid compound ' X ' on heating gives CO_2 gas and a residue. The residue mixed with water forms ' Y '. On passing an excess of CO_2 through ' Y ' in water, a clear solution, ' Z ' is obtained. On boiling ' Z ', compound ' X ' is reformed. The compound ' X ' is



Answer: B



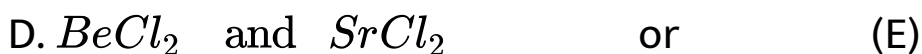
108. When calcium carbide is heated in atmosphere of nitrogen in an electric furnace, the compound formed is



Answer: D



109. Which pair of the following chlorides does not impart color to the flame ?



Answer: B



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110. Which of the following pairs of substance would give same gaseous product on reaction with water?

A. Na and Na_2O_2

B. Ca and CaH_2

C. Ca and CaO

D. Ba and BaO_2 or (E) Ca and CaC_2

Answer: B



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111. Which of the following does not contain magnesium

A. Magnetite

B. Asbestos

C. Magnesite

D. Carnallite

Answer: A



View Text Solution

112. Property of the alkaline earth metals that increases with their atomic number is

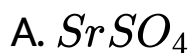
- A. solubility of their hydroxides in water
- B. solubility of their sulphates in water
- C. ionisation energy
- D. electronegativity

Answer: A



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113. Which one of the following alkaline earth metal sulphates has its hydration enthalpy greater than its lattice enthalpy?

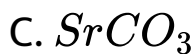
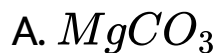


Answer: C



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114. Which of the following carbonates is the most stable

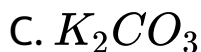
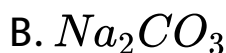
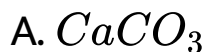


Answer: D



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115. The compound A on heating gives a colourless gas and a residue that is dissolved in water to obtain B. Excess of CO_2 is bubbled through aqueous solution of B, C is formed which is recovered in the solid form. Solid C on gentle heating gives back A. The compound is:-



Answer: A



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116. Which of the following compounds has the lowest melting point ?

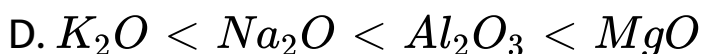
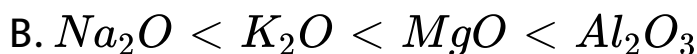
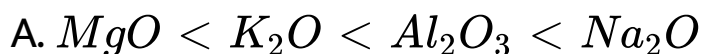


Answer: D



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117. Which one of the following order represents the correct sequence of the increasing basic nature of the given oxides?

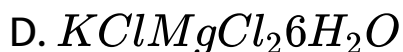
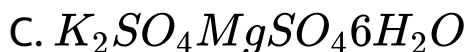


Answer: C



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118. which one of the following represents the composition of carnallite mineral?

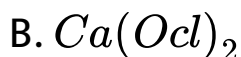


Answer: D



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119. Which one of the following is present as an active ingredient in bleaching powder for bleaching action?



Answer: B



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120. Match list I with list II for the composition of substance and select the correct answer using the code given below

- A. (a) (b) (c) (d)
 (iii) (iv) (i) (ii)
- B. (a) (b) (c) (d)
 (ii) (iii) (iv) (i)
- C. (a) (b) (c) (d)
 (i) (ii) (iii) (iv)
- D. (a) (b) (c) (d)
 (iv) (iii) (ii) (i)

Answer: B



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121. The alkaline earth metal which has the least density is

A. Mg

B. Be

C. Sr

D. Ca or (E) Ba

Answer: D



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122. Choose the incorrect statement in the following

A. BeO is almost insoluble but BeSO_4 is soluble in water

B. BaO is soluble but BaSO_4 is insoluble in water

C. LiI is more soluble than KI in alcohol

D. Both Li and Mg form solid hydrogen carbonates

Answer: D



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123. Which of the following on thermal decomposition yields a basic as well as an acidic oxide?

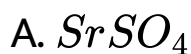


Answer: B



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124. Which one of the following alkaline earth metal sulphates has its hydration enthalpy greater than its lattice enthalpy?

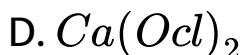
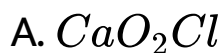


Answer: C



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125. Which one of the following is present as an active ingredient in bleaching powder for bleaching action?

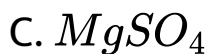
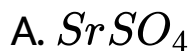


Answer: D



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126. In which of the following is the hydration energy higher than the lattice energy?

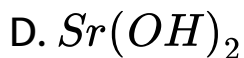
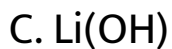
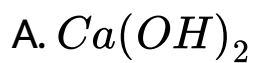


Answer: C



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127. Which of the following is the weakest base:-



Answer: A



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Selected Straight Objective Type Mcqs

1. Alkaline earth metal (s) which does (do) not impart colour to the flame is (are)

A. Be

B. Mg

C. Ca

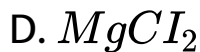
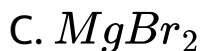
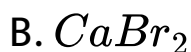
D. Sr.

Answer: A,B



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2. Halides of alkaline earth metal (s) soluble in organic solvent (e.g., acetone) is (are)

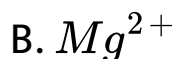


Answer: A,C,D



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3. Halides (Cl' , Br' and I') of alkaline earth metal (s) which cannot be dehydrated by heating is (are)

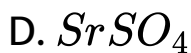
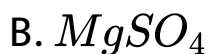


Answer: B,D



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4. Alkaline earth metal sulphates which crystallises in hydrated form are



Answer: A,B,C



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5. Oxalates of alkaline earth metal (s) sparingly soluble in water (is) are

A. Be

B. Ba

C. Ca

D. Sr.

Answer: B,C,D



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6. Substance(s) which reduces/reduce the rate of setting of Plaster of Paris is (are)

A. NaCl

B. Alum

C. Borax

D. $CaSO_4$

Answer: B,C



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7. Setting of cement is an

- A. expansion
- B. contraction
- C. evolution of heat
- D. absorption of heat

Answer: A,C



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8. Calcium oxide can be used for drying of

A. Cl_2

B. H_2

C. NH_3

D. CO_2

Answer: B,C



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9. Magnesium is used as a reducing agent for the production of

A. Mn

B. Ti

C. Zr

D. Hf

Answer: B,C,D



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10. When zeolite, which is hydrated sodium aluminium silicate, is treated with hard water, the sodium ions are exchanged with

A. H^{+} ions

B. Ca^{2+} ions

C. SO_4^{2-} ions

D. Mg^{2+} ions or (e) OH^- ions

Answer: B,D



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11. In Epsom salt, the number of molecules of water directly coordinated with Mg^{2+} ion are

A. 2

B. 4

C. 6

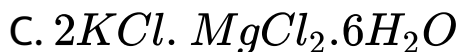
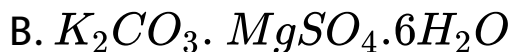
D. 7

Answer: C



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12. Potash magnesia is



D. none of these

Answer: B



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13. Lightest construction metal in industry is

A. Lithium

B. Calcium

C. Aluminium oxide

D. Magnesium

Answer: D



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14. Alkaline earth metal oxide not having rock salt structure is

A. BeO

B. MgO

C. CaO

D. SrO

Answer: A



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15. Lightest alkaline earth metal is

A. Be

B. Mg

C. Ca

D. Sr

Answer: C



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16. The compound insoluble in acetic acid is

- A. Calcium oxide
- B. Calcium carbonate
- C. Calcium oxalate
- D. Calcium hydroxide

Answer: C



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17. An important ore of magnesium is

- A. Malchite
- B. Cassiterite

C. Carnallite

D. Galena

Answer: C



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18. The decreasing order of the second ionisation potential of K , Ca and Ba is

(At. No : K = 19 , Ca = 20 , Ba = 56)

A. KgtCagtBa

B. CagtBagtK

C. BagtKgtCa

D. KgtBagtCa

Answer: A



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19. Which of the following process is used in the extractive metallurgy of magnesium ?

A. Fused salt electrolysis

B. Self reduction

C. Aqueous solution electrolysis

D. Thermite reduction

Answer: A



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20. One mole of magnesium nitride on reaction with an excess of water gives

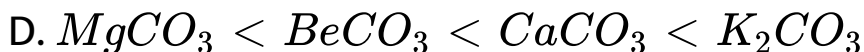
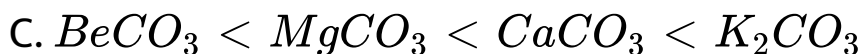
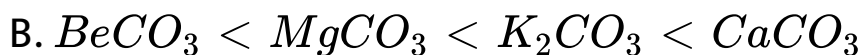
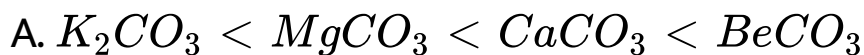
- A. one mole of ammonia
- B. one mole of nitric acid
- C. two moles of ammonia
- D. two moles of nitric acid

Answer: C



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21. The correct order of increasing thermal stability of K_2CO_3 , $MgCO_3$, $CaCO_3$, and $BeCO_3$ is

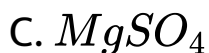
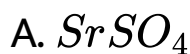


Answer: C



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22. In which of the following is the hydration energy higher than the lattice energy?

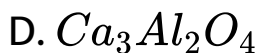
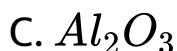
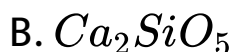
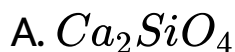


Answer: C



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23. Among the following compounds of cement which is present in the highest amount



Answer: B



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1. Assertion Alkaline earth metals dissolve in liquid ammonia to give a blue-black coloured solution.

The blue-black colour of the solution is due to the presence of ammoniated electrons.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: D



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2. Assertion Be shows similarities with Al.

Reason Be^{2+} and Al^{3+} have same ionic radii.

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: B



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3. Assertion Anhydrous $CaCl_2$ cannot be used for drying NH_3 gas

Reason NH_3 gas is basic in nature.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: B



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4. Assertion Gypsum is added to cement to increase its rate of setting.

Reason Gypsum is calcium sulphate hemihydrate.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: C



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5. Assertion Setting of cement is an exothermic process.

Reason It involves dehydration of calcium aluminates and calcium silicates.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: C



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6. Assertion Anhydrous $MgCl_2$ cannot be obtained by heating $MgCl_{2.6}H_2O$

$MgCl_{2.6}H_2O$ is a highly stable compound, not affected by heat.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: D



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7. Assertion Bleaching powder is a mixed salt.

Reason In the presence of $CoCl_2$ bleaching powder decomposes to give $CaCl_2$ and O_2

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true
- or (E) Both A and R are false.

Answer: A



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8. Assertion A solution of $MgCl_2$ in water is basic in nature.

Reason $MgCl_2$ when dissolved in water undergoes hydrolysis to give free OH^- ions.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: A



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9. Assertion When SO_2 is passed through lime water, it first turns milky, but becomes clear when excess of SO_2 is passed through it.

Reason The milkiness is due to the formation of insoluble $CaSO_3$ and with excess of SO_2 , $Ca(HSO_3)_2$ is formed which is soluble in water.

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true
- or (E) Both A and R are false.

Answer: D



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10. Assertion Calcium acetylide is CaC_2

Reason CaC_2 is a calcium salt of acetylene.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: A



11. Assertion (A): magnesium is not present in enamel of human teeth.

Reason (R): Magnesium is an essential elements for biological functions of human beings.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: A



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12. Assertion: Barium is not required for normal biological function in human.

Reason: Barium does not show variable oxidation state.

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true
- or (E) Both A and R are false.

Answer: A



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13. Assertion (A): $BaCO_3$ is more soluble in HNO_3 than in water.

Reason (R): Carbonate is a weak base and reacts with H^{\oplus} ions to form strong acid causing barium salt to dissociate.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: A



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14. Assertion (A): magnesium is not present in enamel of human teeth.

Reason (R): Magnesium is an essential elements for biological functions of human beings.

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true
- or (E) Both A and R are false.

Answer: A



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15. Assertion (A): Addition of NH_4OH to an aqueous solution of $BaCl_2$ in the presence of excess of NH_4Cl precipitates $Ba(OH)_2$.

Reason (R): $Ba(OH)_2$ is insoluble in water.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true

or (E) Both A and R are false.

Answer: C



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Ultimate Preparatory Package

1. BeCl_2 is not soluble in

A. acetone

B. water

C. ether

D. carbon tetrachloride

Answer:



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2. Ethyl alcohol can be dried with

A. $CaCl_2$

B. Ca

C. Na

D. none of these

Answer:



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3. Anhydrous CaCl_2 can be used for drying

A. alcohols

B. ammonia

C. H_2S

D. methane

Answer:



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4. Alkaline earth metal not having rock salt structure is

A. BeO

B. CaO

C. BaO

D. MgO

Answer:



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5. Hydrated chloride of the alkaline earth metal.

Which cannot be dehydrated on heating is

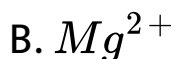


Answer:



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6. Sulphate of alkaline earth metal which crystallises without water of hydration is that of



Answer:



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7. Oxalates of alkaline earth metal, highly soluble in water is that of

A. Be

B. Sr

C. Ba

D. Ca

Answer:



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8. A solution of Ca in liquid ammonia when evaporated to dryness gives as residue

A. pure Ca

B. $[Ca(NH_3)_x]^{2+}$

C. $[Ca(NH_3)_6]^{2+}$

D. $Ca(NH_3)_6$

Answer:



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9. $MgCl_2 \cdot NH_4Cl \cdot 6H_2O$ on heating in air gives as residue

A. $Mg(OH)Cl$

B. $MgCl_2$

C. $Mg(OH)_2$

D. MgO

Answer:



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10. Calcium bicarbonate is

- A. a white crystalline solid
- B. a white powder
- C. a colourless crystalline solid
- D. none of these

Answer:



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11. The metal with highest specific heat is

A. Be

B. Al

C. Ag

D. Au

Answer:



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12. The alkaline earth metal with least density is

A. Be

B. Mg

C. Ca

D. Sr.

Answer:



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13. The alkaline earth metal harder than glass

A. Be

B. Mg

C. Ca

D. none of these

Answer:



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Brain Teasers 17

1. Which of the following is used to remove lignin from wood up?

- A. Unslaked lime
- B. Slaked lime
- C. Calcium bisulphate

D. Calcium bisulphite

Answer:



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2. India has no deposits of sulphur. However, large deposits of gypsum occur in india. How can you obtain SO_2 from gypsum quantitative?

A. by heating anhydrous gypsum to $1200^{\circ}C$ alone

- B. by heating anhydrous gypsum with coke to 1200°C
- C. by treating gypsum with conc. HNO_3 at 700°C
- D. none of these

Answer:



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3. Which of the following is best suited to dry crystals of ammine salts (complexes in which NH_3

is a ligand e.g., $[Ni(NH_3)_6]Br_2$)

A. Anhy. $CaCl_2$

B. Anhy. CaO

C. A mixture of CaO and NH_4Cl

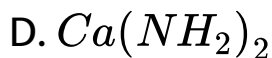
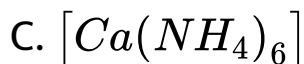
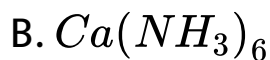
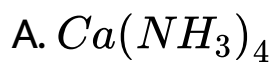
D. None of these

Answer:



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4. A solution of calcium in liquid ammonia on evaporation produces

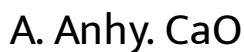


Answer:



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5. What is 'anhydrone'?



C. Anhy. $MgCl_2$

D. Anhy. $MgClO_4$

Answer:



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6. BeO is least soluble in

A. pure water

B. aqueous solution of $BeCl_2$

C. dilute HCl

D. dilute NaOH solution

Answer:



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7. A burning magnesium ribbon will continue to burn in

A. N_2 and steam

B. CO_2 and SO_2

C. N_2 , CO_2 and SO_2

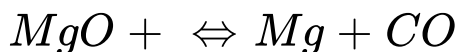
D. N_2 , CO_2 , SO_2 and steam

Answer:



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8. Consider the reaction



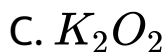
- A. The forward reaction is not possible as magnesium has more affinity for oxygen.
- B. The forward reaction is possible at a temperature of above 2000°C
- C. The forward reaction is possible at 2000°C if the products of reaction are cooled rapidly.
- D. none of these

Answer:



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9. Which of the following statements is a yellow solid in pure water?



Answer:



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10. Which of the following statements is not correct?

- A. K is more reactive than Na
- B. KOH is a stronger base than NaOH
- C. KOH is less hygroscopic than NaOH
- D. KOH in moist air does not produce a crust of carbonate like NaOH.

Answer:





11. The lesser solubility of $NaHCO_3$ than Na_2CO_3

in water is mainly due to

A. higher lattice energy due to smaller size of

HCO_3^- ion

B. smaller hydration energy for larger HCO_3^-

ion

C. intermolecular H-bonding

D. intermolecular H-bonding between

$NaHCO_3$ molecules and water molecules of

crystallisation.

Answer:



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12. The yellow light used for street and road illumination is from

A. mercury vapour lamp

B. Na

C. NaClO

D. $NaClO_3$

Answer:



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13. Which is manufactured by electrolysis of fused NaCl?

A. NaOH

B. K

C. Ca

D. Zn

Answer:



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14. Which of the following is used as scavenger in metallurgy?

A. Na

B. K

C. Ca

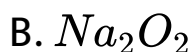
D. Zn

Answer:



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15. Which of the following substances acts as an oxidising as well as a reducing agent?



Answer:



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16. $CaCO_3$ exist in

- A. only one crystalline form
- B. only two crystalline forms
- C. three crystalline forms
- D. four crystalline forms

Answer:



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17. Gypsum is not chemically similar to

- A. Selenite
- B. alabaster
- C. anhydrite
- D. fluorspar

Answer:



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18. The function of sand in mortar is :

- A. to decrease the hardness
- B. to make the mass compact

C. to decrease the plasticity of the mass

D. to prevent cracks when the mortar sets to a solid mass

Answer:



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19. The blue coloured mineral 'Lapis Lazuli' which is used as a semi-precious stone is a mineral of the following class

A. Sodium aluminosilicates

B. Zinc cobaltate

C. Basic copper carbonate

D. prussian blue

Answer:



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20. A common food preservative in squashes and fruit juices is

A. Sodium metabisulphite

B. sodium sulphide

C. sodium sulphate

D. none of these

Answer:



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Unit Test 12

1. K^+ and Ca^{2+} are isoelectronic but differ in

A. electronic configuration

B. ionic size

C. both (A) and (B)

D. none of the above

Answer:



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2. Which causes nerve signals in animals?

A. Electrical potential gradient due to transfer

K^{+} ions

B. Electrical potential gradient due to transfer

of Na^{+} ions in $(Na^{+} - K^{+})$ pump

C. Electrical potential gradient set up due to transfer of Ca^{2+} ions

D. No nerve signal exists in animals

Answer:



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3. Alkali metals readily dissolve in liquid ammonia to give blue coloured solutions. The blue colour is believed to be due to

A. ammoniated cations

B. ammoniated anions

C. ammoniated electrons

D. ammoniated cations and electrons

Answer:



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4. Na_2O_2

A. is diamagnetic in nature

B. is a salt of dibasic and H_2O_2

C. oxidizes Cr^{3+} (green) to CO_4^{2-} (yellow)

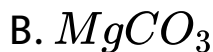
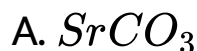
D. all the correct properties of Na_2O_2

Answer:



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5. Which of the following compound decompose decomposes at highest temperature-



Answer:



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