



CHEMISTRY

BOOKS - S DINESH & CO CHEMISTRY (HINGLISH)

PRINCIPLES RELATED TO PRACTICAL CHEMISTRY

Mcq

1. During Lassaigne's test, nitrogen present in the organic compound is converted into

- A. Sodium nitride
- B. sodium azide
- C. sodium cyanide
- D. none of these

Answer: C



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2. Presence of nitrogen in which of the following compound cannot be detected by Lassaigne's test?

A. Aniline

B. hydrazine

C. Urea

D. nitrobenzene

Answer: B



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3. In Lassaigne's test, if the organic compound contains both N and S and it is fused with excess sodium metal, the sodium fusion extract will contain

A. NaCNS

B. NaCN

C. Na_2S

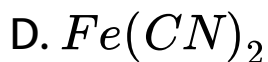
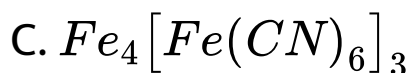
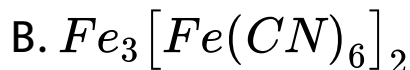
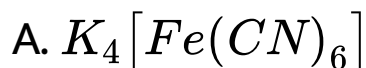
D. Both (b) and (C)

Answer: A



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4. In Lassaigne's test for the detection of nitrogen, the blue colouration is due to the formation of



Answer: C



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5. In Lassaigne's test, a blood red colouration with Fe^{3+} ions indicates the presence of

A. nitrogen

B. sulphur

C. both nitrogen and sulphur

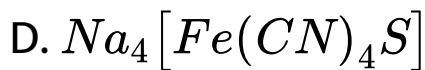
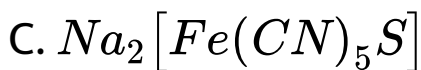
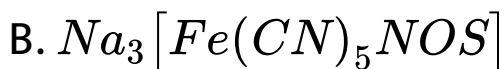
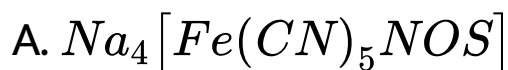
D. both nitrogen and halogen

Answer: C



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6. The formula of the compound which gives violet colour in Lassaigne's test for sulphur with sodium nitroprusside is



Answer: A



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7. The Lassaigne's extract is boiled with dil. HNO_3 before testing for halogens because

A. silver halides are soluble in HNO_3

B. Na_2S and NaCN are decomposed by HNO_3

C. Ag_2S is soluble in HNO_3

D. AgCN is soluble in HNO_3

Answer: B





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8. In Lassaigne's test, the organic compound is fused with sodium metal so as to

A. burn the compound

B. form a sodium derivative

C. convert N,S or halogen into soluble ionic compound

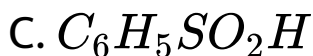
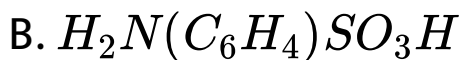
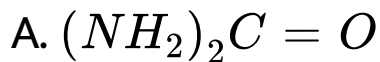
D. None of the above

Answer: C



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9. Which of the following compound will give blood red colour while doing the Lassaigne's test for N?



Answer:



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10. The function of boiling the sodium extract with concentrated nitric acid before testing halogen is

- A. to make solution clear
- B. to destroy CN^- and S^{2-} ion
- C. to make the solution acidic
- D. to bring common ion effect.

Answer:



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11. Lassaigne's test is used in the qualitative analysis to detect

A. nitrogen

B. sulphur

C. chloride

D. All of these

Answer: D



12. In Lassaigne's test for N, S and halogens, organic compound is

- A. fused with sodium
- B. dissolved with sodamide
- C. extracted with sodamide
- D. fused with calcium.

Answer:



13. In sodium fusion test of organic compounds, the nitrogen of an organic compound is converted to

- A. Sodamide
- B. sodium cyanide
- C. sodium nitrite
- D. sodium nitrate

Answer:



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14. Organic compound is fused with sodium piece in Lassaigne's test in order to

A. increase the ionisation of compound

B. increases the volume of compound

C. increases the reactivity of compound

D. converted the covalent compound to a mixture of electrovalent compounds.

Answer: A



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15. Lassaigne's test for the detection of nitrogen fails in



Answer:



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16. In Lassaigne's test for nitrogen the blue colour is due to the formation of

- A. ferric ferrocyanide
- B. potassium ferrocyanide
- C. sodium ferrocyanide
- D. sodium cyanide

Answer:



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17. Sodium extract of an organic compound gives blood red colour with $FeCl_3$. It contains

A. N

B. S

C. N and S

D. S and Cl

Answer:



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18. In Lassaigne's test when both N and S are present, blood red colour is due to the formation of

A. ferric sulphocyanide

B. ferric cyanide

C. ferric ferrocyanide

D. none

Answer: A



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19. If N and S both are present in an organic compound, during Lassaigne's test both changes to

A. Na_2S and NaSCN

B. NaSCN

C. Na_2SO_3 and NaCN

D. Na_2S and NaCNO.

Answer:



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20. The compound that does not give a blue colour in Lassaigne's test is

A. Aniline

B. glycine

C. hydrazine

D. urea

Answer:



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21. C_6H_5OH and C_2H_5OH can be distinguished by

A. sodium metal test

B. $NaHCO_3$ test

C. litmus test

D. none of these

Answer:



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22. Methyl salicylate has a smell of

A. apples

B. bitter almonds

C. bananas

D. none of these

Answer:



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23. Which of the following compounds does not interfere in the testing of alcoholic group by sodium metal test?

A. ethoxyethane

B. phenol

C. acetone

D. acetic acid

Answer:



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24. Which of the following phenolic compound does not respond to phthalein test?

A. o-Cresol

B. m-Cresol

C. p-Cresol

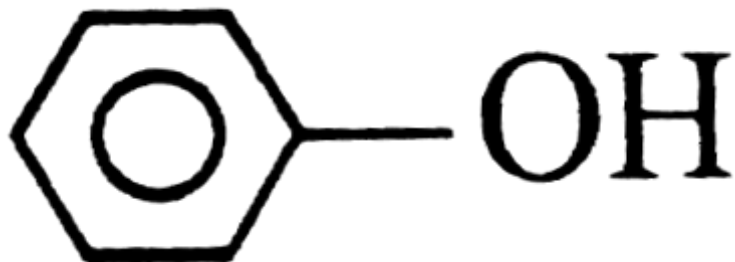
D. β -Naphthol.

Answer:

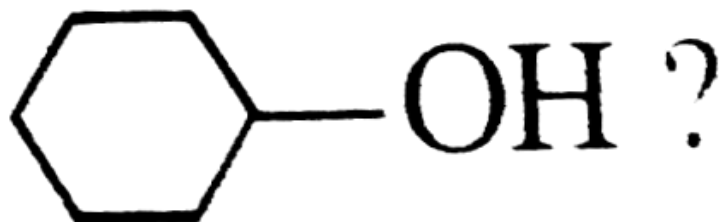


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25. Which of the following test can be used to distinguish between



and



?

- A. Br_2 water test
- B. sodium metal test
- C. $NaHCO_3$ test

D. Both (b) and (C)

Answer:



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26. Which of the following test is given both by aldehydes and ketones?

A. 2,4-DNP test

B. Schiff's reagent test

C. Tollens reagent test

D. All the three

Answer:



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27. Assertion: Aromatic aldehydes can be distinguished from aliphatic aldehydes by fehling's solution

Reason : Fehing's solution is an alkaline solution of $CuSO_4$ containing Rochelle salt.

A. Schiff's reagent test

B. Fehling's solution test

C. 2,4-DNP test

D. None of these

Answer:



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28. Ethanol and methanol can be differentiated by

A. sodium metal test

B. litmus test

C. NaHCO_3 test

D. iodoform test

Answer:



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29. Phenol and acetic acid can be differentiated by

A. litmus test

B. $NaHCO_3$ test

C. Both (A) and (B)

D. None of these

Answer:



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30. Acetic acid and formic acid can be differentiated by

A. litmus test

B. NaHCO_3 test

C. Tollen's reagent test

D. iodoform test

Answer:



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31. Benzaldehyde and acetaldehyde can be differentiated by

A. 2,4-DNP test

B. Tollen's reagent test

C. Fehling solution test

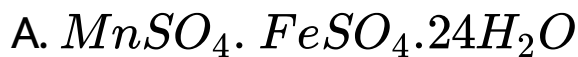
D. Schiff's reagent test

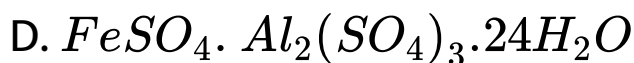
Answer:



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32. Mohr's salt is



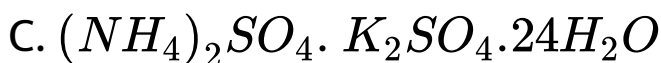
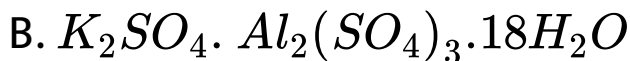
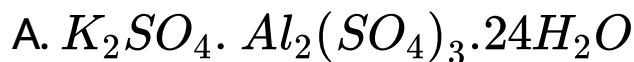


Answer:



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33. Potash alum is



D. None of these

Answer:



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34. Presence of Fe^{3+} ions in $FeSO_4$ can be checked with

A. Blue litmus paper

B. red litmus paper

C. KCNS solution

D. None of these

Answer:



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35. A few drops of conc. H_2SO_4 is added to the solution of $FeSO_4$ in water to

A. Convert Fe^{3+} ions (if any) into Fe^{2+}

ions

B. prevent salt hydrolysis

C. Both (A) and (B)

D. None of these

Answer:



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36. Coupling of benzene diazonium chloride and aniline is carried out at pH

A. 7

B. 10 – 14

C. 4 – 5

D. 6 – 8

Answer:



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37. Completion of diazotisation can be checked with

A. starch paper

B. starch -KI paper

C. litmus paper

D. $K_2Cr_2O_7$ paper.

Answer:



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38. The colour of the dye obtained by coupling benzene diazonium chloride with aniline is

A. red

B. orange

C. yellow

D. none of these

Answer:



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39. The most suitable indicator for the titration of HCOOH against NaOH is

A. methyl orange

B. phenolphthalein

C. thymol phthalein

D. Both (b) and (C)

Answer:



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40. The most suitable indicator for the titration of HBr against KOH is

A. phenolphthalein

B. methyl orange

C. bromothymol blue

D. all are correct

Answer:



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41. The most suitable indicator for the titration of NH_4OH against HNO_3 is

A. phenolphthalein

B. methyl red

C. thymol phthalein

D. none of these

Answer:



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42. If pK_{In} of an indicator is 10.5, the pH transition range for which it is most suitable is

A. 8.5 – 10.5

B. 10.5 – 12.5

C. 10.0 – 11.0

D. 9.5 – 11.5

Answer:



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43. The most suitable indicator for the titration of $Ba(OH)_2$ against HI is

A. phenolphthalein

B. methyl orange

C. Both (A) and (B)

D. none of these

Answer:



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44. In $KMnO_4$ titrations the indicator used is

A. phenolphthalein

B. methyl orange

C. Both (A) and (B)

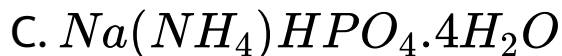
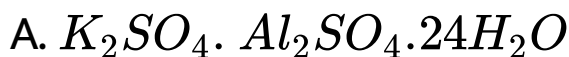
D. none of these

Answer:



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45. The salt used for performing bead test in qualitative inorganic analysis is





Answer:



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46. Which of the following cation can be detected by charcoal block test (cobalt nitrate test)?





Answer:



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47. A white salt turns yellow on heating but becomes white on cooling. It may be a salt of

A. Fe

B. Pb

C. Al

D. Zn

Answer:



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48. When borax is heated in a platinum loop, the transparent bead formed contains

A. $Na_2B_4O_7$

B. Sodium metaborate + B_2O_3

C. Water

D. CaO

Answer:



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49. A brick red colour is imparted by a

A. Ca salt

B. Sr salt

C. Na salt

D. Co salt

Answer:



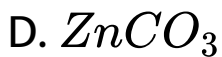
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50. A salt on reaction with dil. H_2SO_4 gives a reddish brown gas. The salt is

A. KNO_2

B. $ZnBr_2$

C. $NaNO_3$



Answer:



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51. A dark green bead in the borax bead test indicates the presence of



D. Ni^{2+}

Answer:



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52. Which is the hottest part of the flame?

A. Blue zone

B. zone of partial combustion

C. zone of complete combustion

D. zone of no combustion

Answer:



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53. In the microcosmic bead test test Cu^{2+} imparts which colour to the flame

A. green

B. yellow

C. blue

D. violet.

Answer:



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54. Flame test is not given by

A. Ba^{2+} ions

B. Ba^{2+} ions

C. Ca^{2+} ions

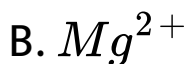
D. Na^{+} ions

Answer:



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55. A blue coloured residue obtained in cobalt nitrate charcoal cavity test is due to



Answer:



56. When a salt is heated with dil. H_2SO_4 and $KMnO_4$ solution, the pink colour of $KMnO_4$ is discharged, the mixture contains

A. sulphite

B. carbonate

C. nitrate

D. bicarbonate

Answer:



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57. A reddish brown residue in the charcoal cavity test is given by



Answer:



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58. A test tube containing a nitrate and another containing a bromide and MnO_2 are treated with conc. H_2SO_4 . The brown fumes evolved are passed through water. The water will be coloured by

- A. the nitrate
- B. the bromide
- C. both
- D. none of the two

Answer:



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59. Heating of oxalic acid with conc. H_2SO_4 gives

A. CO

B. CO_2

C. $CO + CO_2$

D. none

Answer:



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60. Bromine vapours turn starch iodine paper

A. Violet

B. Blue

C. Yellow

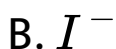
D. Red.

Answer:



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61. Solution of a salt in dilute H_2SO_4 produces deep blue colour with starch iodine solution. The salt contains

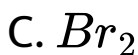


Answer:



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62. An inorganic salt when heated evolves a coloured gas which bleaches moist litmus paper. The evolved gas is



Answer:



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63. Acidified $K_2Cr_2O_7$ turns green by

A. SO_2

B. CO

C. SiO_2

D. HCl

Answer:



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64. H_2S and SO_2 can be distinguished by

A. Litmus paper

B. Lime water

C. $Pb(CH_3COO)_2$

D. HCl

Answer:



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65. Nitrates of all metals are

A. unstable

B. coloured

C. insoluble in water

D. soluble in water

Answer:



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66. NaCl, NaBr and NaI mixture on heating with conc. H_2SO_4 gives gases respectively.

A. HCl, Br_2 , I_2

B. HCl, HBr, HI

C. Cl_2 , Br_2 , I_2

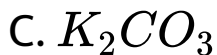
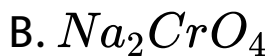
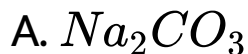
D. None

Answer:



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67. A colourless salt gives violet colour in bunsen flame, it may be



Answer:



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68. When concentrated H_2SO_4 is added to dry KNO_3 and heated, brown fumes evolve. These fumes are



Answer:



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69. Which does not give borax bead test?

A. Pb

B. Cu

C. Cr

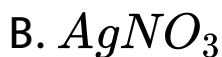
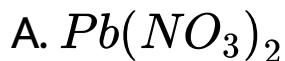
D. Fe

Answer:



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70. Of the following nitrates that one gives nitrous oxide by thermal decomposition is



Answer:



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71. The salt which is not decomposed by dil.

H_2SO_4 and conc. H_2SO_4 , will contain

A. Sulphite ion

B. chloride ion

C. nitrate ion

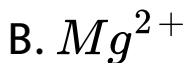
D. sulphate ion

Answer:



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72. Which one of the following cation will give a blue coloured ash when a piece of filter paper dipped in a solution containing its salt and $Co(NO_3)_2$ is burned?



Answer:



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73. Which of the following does not evolve CO_2 on treatment with dilute hydrochloric acid?

- A. Bismuth carbonate
- B. lead carbonate
- C. calcium carbonate
- D. basic lead carbonate

Answer:



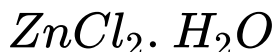


74. Fumes of hydrochloric acid obtained by heating a dry salt in a test tube indicates the presence of

A. an alkali metal chloride

B. Hg_2Cl_2

C. A hydrated chloride such as



D. An alkaline earth metal chloride.

Answer:



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75. A salt solution is acidified with dil. HCl and $BaCl_2$ solution is added. A white ppt. is formed. The salt contains



Answer:



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76. Production of a green edged flame on igniting the vapours evolved by heating a given inorganic salt with few mL of ethyl alcohol and conc. H_2SO_4 indicates the presence of a

A. Tartarate

B. oxalate

C. acetate

D. borate

Answer:



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77. Nitrate is confirmed by ring test. The brown colour of the ring test is due to the formation of

A. Ferrous nitrite

B. $FeSO_4 \cdot NO$

C. $FeSO_4 \cdot NO_2$

D. Ferrous nitrate

Answer:

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78. On heating a mixture of potassium dichromate and sodium chloride with concentrated sulphuric acid in a dry test tube, the compound formed is

A. chromiun chloride

B. chromyl chloride

C. chloric dioxide

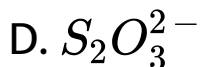
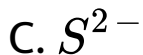
D. chromic acid

Answer:



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79. Formation of purple colour on the addition of sodium nitroprusside to sodium carbonate extract indicates the presence of



Answer:



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80. A salt solution is treated with chloroform drops. Then it is shaken with chlorine water.

Chloroform layer becomes violet, solution contains.

A. NO_3^- ion

B. NO_2^- ion

C. Br^- ion

D. I^- ion

Answer:



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81. The carbonate of which of the cation is insoluble in water

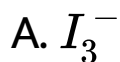


Answer:



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82. When CS_2 layer containing both Br_2 and I_2 is shaken with excess of Cl_2 water, the violet colour due to I_2 disappears and orange colour due to Br_2 appears. The disappearance of violet colour is due to the formation of



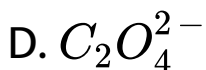
Answer:





83. To a solution of a salt, $MgSO_4$ solution is added and white ppt. appears only on heating.

The acid radical in the salt is



Answer:

84. $BaCl_2$ solution gives a white ppt. with a solution of a salt, which dissolves in dil. HCl with the evolution of a colourless, pungent smelling gas. The radical in the salt is



Answer:



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85. For the test of halides, the soda extract is acidified with

A. Dil H_2SO_4

B. Dil. HNO_3

C. Dil. HCl

D. none of the three

Answer:



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86. Sulphide ions react with $Na_2[Fe(NO)(CN)_5]$ to form purple coloured compound $Na_4[Fe(CN)_5(NOS)]$ in the reaction, the oxidation state of iron changes from

A. + 2 to +3

B. + 3 to +2

C. + 2 to +4

D. does not change.

Answer:



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87. In a mixture having nitrite and nitrate, nitrite can be destroyed by heating with

A. Na_2CO_3

B. Urea

C. Oxalic acid

D. NaCl

Answer:



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88. In a combination of NO_3^- , Br^- and I^- present in a mixture, Br^- and I^- interfere in the ring test for NO_3^- . These are removed by adding a solution of

A. $AgNO_3$

B. Ag_2SO_4

C. Ag_2CO_3

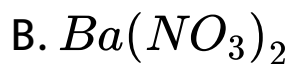
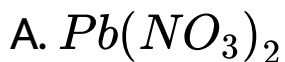
D. None of these

Answer:



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89. An aqueous solution of salt gives white precipitate with $AgNO_3$ solution as well as with dil. H_2SO_4 . It may be



Answer:



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90. Which of the following solution gives precipitate with $Pb(NO_2)_2$ but not with $Ba(NO_3)_2$?

A. Sodium chloride

B. sodium sulphate

C. sodium nitrate

D. sodium hydrogen phosphate

Answer:



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91. Chromyl chloride vapours are dissolved in water and acetic acid and lead acetate solution is added, then

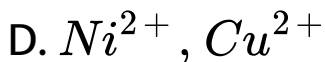
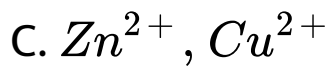
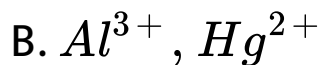
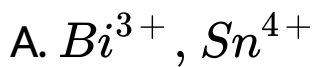
- A. The solution will remain colourless
- B. the solution will become dark green
- C. A yellow solution will be obtained
- D. A yellow ppt. will be obtained

Answer:



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92. Which one among the following pairs of ions cannot be separated by H_2S in dilute hydrochloric acid?



Answer:



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93. Sometimes yellow turbidity appears while passing H_2S gas even in the absence of II group radicals. This is because

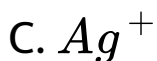
- A. Sulphur is present in the mixture as impurity
- B. IV group radicals are precipitated as sulphides
- C. of the oxidation of H_2S gas by some acid radical
- D. III group radicals are precipitated as hydroxide.

Answer:



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94. The ion that cannot be precipitated by both HCl and H_2S is

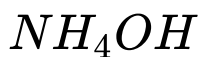


Answer:



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95. Mark the compound which turns black with



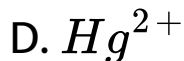
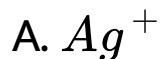
- A. lead chloride
- B. mercurous chloride
- C. mercuric chloride
- D. silver chloride

Answer:



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96. Which of the following radicals does not belong to group I?



Answer:



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97. Mg is not precipitated in group V because

- A. $MgCO_3$ is soluble in water
- B. $MgCO_3$ is soluble in NH_4Cl
- C. $MgCO_3$ is soluble in NH_4OH
- D. None

Answer:



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98. On passing H_2S into saturated solution of $BaCl_2$, white ppt. obtained is of

- A. hydrogen chloride
- B. formation of a complex
- C. barium chloride
- D. barium sulphide.

Answer:



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99. Concentrated nitric acid is added before proceeding to test for group III members. This is to

A. oxidise any remaining H_2S

B. Convert ferrous ions to ferric ions

C. form nitrates which gives granular precipitate

D. increases ionisation of ammonium hydroxide

Answer:



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100. To a solution of a substance, gradual addition of ammonium hydroxide results in a brown ppt. which does not dissolve in excess of NH_4OH . However, when HCl is added to the original solution, a white ppt. is formed. The solution contained

A. Lead salt

B. silver salt

C. mercurous salt

D. copper salt

Answer:



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101. To avoid the precipitation of hydroxides of Ni^{2+} , Co^{2+} , Mn^{2+} along with those of the third group cations, the solutions should be

A. heated with few drops of conc. HNO_3

B. heated with excess of ammonium chloride

C. concentrated to small volume

D. none of these

Answer:



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102. Silver, mercury (ous) and lead are grouped together in the scheme of qualitative analysis because they form

A. soluble nitrates

B. carbonates which dissolve in dilute nitric
acid

C. insoluble chlorides

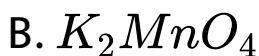
D. All of above

Answer:



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103. In the fourth group, $Mn(OH)_2$ on heating with PbO_2 and conc. HNO_3 gives pink colour due to the formation of

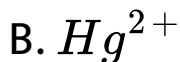


Answer:



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104. An inorganic salt solution on treatment with HCl gives a white ppt. Which of the following metal ions is possible?



Answer:



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105. H_2S in the presence of HCl precipitate group II but not group IV because

A. HCl activates H_2S

B. HCl increases the concentration of Cl^-

C. HCl decreases the concentration of S^{2-}
ions

D. HCl lowers the solubility of H_2S in the
solution

Answer:

106. Distinguishing reagent between silver and lead salts is

A. H_2S gas

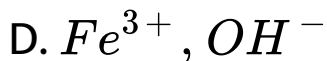
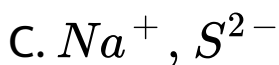
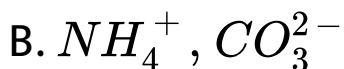
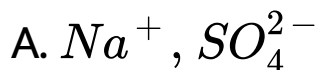
B. Dil. HCl solution

C. $NH_4Cl(Solid) + NH_4OH$ solution

D. $NH_4Cl(solid) + (NH_4)_2CO_3$ solution

Answer:

107. Which of the following pairs of ions would be expected to form precipitate when dilute solutions are mixed?



Answer:



108. A metal chloride solution on mixing with K_2CrO_4 solution gives a yellow precipitate, insoluble in acetic acid. The metal may be

- A. Mercury
- B. Zinc
- C. silver
- D. Lead

Answer:



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109. A black sulphide is formed by the action of

H_2S on

- A. Cupric chloride
- B. Cadmium chloride
- C. Zinc chloride
- D. Sodium chloride.

Answer:



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110. H_2S will precipitate the sulphide of all the metals from the solution of chlorides of Cu, Zn and Cd if

- A. The solution is aqueous
- B. the solution is acidic
- C. the solution is dilute acidic
- D. none of the above solution is present

Answer:



111. When H_2S is passed through an ammoniacal salt solution X, a white precipitate is obtained. Then X can be a

- A. cobalt salt
- B. nickel salt
- C. maganese salt
- D. zinc salt

Answer:





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112. Lead has been placed in the group I and II because

- A. it shows the valency of one and two
- B. it forms insoluble $PbCl_2$
- C. it forms lead sulphide
- D. it is partially soluble in water

Answer:



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113. Potassium ferrocyanide is used in the detection of

A. Cu^{2+} ions

B. Fe^{3+} ions

C. both

D. none

Answer:



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114. Which of the following is insoluble in dil. HNO_2 but soluble in aqua regia?

A. HgS

B. PbS

C. Bi_2S_3

D. CuS

Answer:



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115. Which of the following precipitates K^+ from its solution?

- A. Sodium ferrocyanide
- B. sodium cobaltinitrite
- C. sodium argento cyanide
- D. sodium bicarbonate

Answer:



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116. NH_4CNS can be used to test one or more out of Fe^{3+} , Co^{2+} , Cu^{2+}

A. Fe^{3+} only

B. Co^{2+} , Cu^{2+}

C. Fe^{3+} , Cu^{2+}

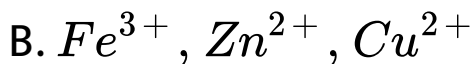
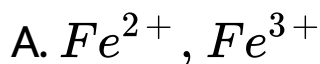
D. All

Answer:



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117. $K_4[Fe(CN)_6]$ can be used to detect one or more out of Fe^{2+} , Fe^{3+} , Zn^{2+} , Cu^{2+} , Cd^{2+}



Answer:



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118. For dissolution of an ionic solid in water

$$\Delta_{\text{sol}}H^{\ominus} =$$

A. $\Delta_{\text{lattice}}H^{\ominus} + \Delta_f H^{\ominus}$

B. $\Delta_f H^{\ominus} + \Delta_{\text{hyd}}H^{\ominus}$

C. $\Delta_a H^{\ominus} + \Delta_f H^{\ominus}$

D. none of these

Answer:



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119. For dissolution of a solid in a liquid, ΔS is generally

A. $+ve$

B. $+ve$

C. zero

D. Both (A) and (B)

Answer:



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120. If enthalpy of neutralisation of dil HCl and NaOH(aq) is x then enthalpy of neutralisation of dil. HI and $Ba(OH)_2$ (aq) is

A. $x / 2$

B. $2x$

C. x

D. slightly less than x in magnitude

Answer:



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121. Which of the following can be enthalpy of neutralisation of a weak acid and a strong base?

A. -57.1kJ

B. 58.1kJ

C. -56.1kJ

D. 114.2kJ

Answer:



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122. Suspended impurities from a colloidal solution of gum in water can be removed by

- A. filtration
- B. dialysis
- C. electrodialysis
- D. ultrafiltration

Answer:



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123. Excess H^+ and Cl^+ ions from a colloidal solution of $Fe(OH)_3$ can be removed by

A. filtration

B. dialysis

C. sublimation

D. chromatography

Answer:



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124. Which of the following sols cannot be prepared by boiling the dispersed phase with dispersion medium (water)?

A. Egg albumin sol in water

B. Gum sol in water

C. Starch sol in water

D. none of these

Answer:



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125. Which of the following sols can only be prepared by boiling the dispersed phase with dispersion medium?

A. Egg albumin sol in water

B. Gum sol in water

C. $Fe(OH)_3$ sol in water

D. Starch sol in water

Answer:



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126. Which of the following sol will be destabilised with ionic impurities - if present?

- A. Gum sol in water
- B. Starch sol in water
- C. Egg albumin sol in water
- D. $Fe(OH)_3$ sol in water

Answer:



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127. Which of the following sol will not be destabilised with ionic impurities - if present?

A. $Fe(OH)_3$ sol in water

B. Gold sol in water

C. Silver sol in water

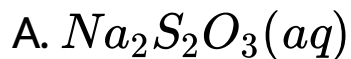
D. Starch sol in water

Answer:



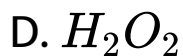
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128. Iodine can react with



B. Starch

C. Both (A) and (B)

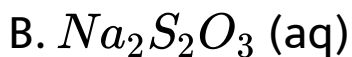
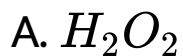


Answer:



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129. Iodine does not react with



C. Starch

D. all the three

Answer:



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130. Iodine reacts with $Na_2S_2O_3$ (aq) to give products which are

A. pale yellow coloured

B. dark blue coloured

C. violet coloured

D. colourless.

Answer:



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1. Brown ring is made for



Answer:



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2. In the precipitation of the iron group in qualitative analysis, ammonium chloride is added before adding ammonium hydroxide is

A. Decrease concentration of OH^- ions

B. Prevent interference by phosphate ions

C. increases concentration of Cl^- ions

D. increases in the concentration of NH_4^+

ions.

Answer:



3. A salt is heated with dilute H_2SO_4 and then with conc. H_2SO_4 . No reaction takes place. It may be

A. Nitrate

B. Sulphide

C. Oxalate

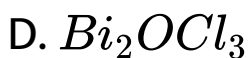
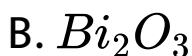
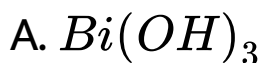
D. sulphate

Answer:



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4. When bismuth chloride is poured into a large volume of water then white precipitate produced is



Answer:



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5. The brown ring test for nitrate employs

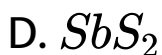
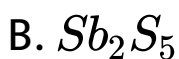
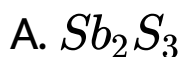
- A. Barium chloride
- B. Ferrous sulphates
- C. Nitric acid
- D. none of the above

Answer:



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6. A precipitate ofwould be obtained on adding HCl to a solution of Sb_2S_3 in yellow ammonium sulphide.



Answer:



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7. The alkaline earth metal that imparts apple green colour to the bunsen flame when introduced in it in the form of its chloride is

A. Barium

B. Strontium

C. Calcium

D. Magnesium

Answer:



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8. A yellow precipitate obtained in II group of the qualitative analysis was soluble in aqueous NaOH and insoluble in dil. HNO_3 . This shows the presence of

A. Tin

B. Antimony

C. Arsenic

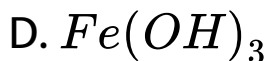
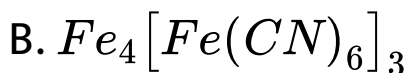
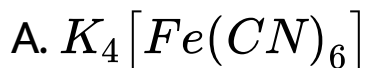
D. Cadmium

Answer:



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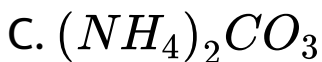
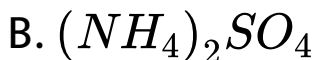
9. Ferric ion forms a prussian blue coloured ppt. due to



Answer:



10. Which one of the following can be used in place of NH_4Cl for the identification of the third group radicals?



Answer:



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11. In qualitative analysis NH_4Cl is added before NH_4OH

A. The dissociation of NH_4OH increases

B. The concentration of OH^- increases

C. The concentration of both OH^- and

NH_4^+ increases

D. The concentration of OH^- decreases

Answer:



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12. A green mass is formed in the charcoal cavity test when a colourless salt (X) is fused with cobalt nitrate. X contains

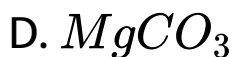
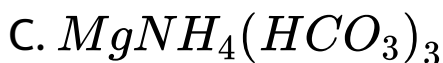
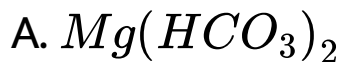


Answer:



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13. The presence of magnesium is confirmed in the qualitative analysis of the formation of a white crystalline ppt. which is due to



Answer:



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14. In India at the occasion of marriages, the fire works are used, which of the following gives green flame?

A. Ba

B. K

C. Be

D. Na

Answer:



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15. Nitrate is confirmed by ring test. The brown colour of the ring is due to the formation of

A. Ferrous nitrite

B. $FeSO_4 \cdot NO$

C. $FeSO_4 \cdot NO_2$

D. Ferrous nitrate

Answer:



16. $Fe(OH)_3$ can be separated from $Al(OH)_3$ by the addition of

- A. Dil.HCl
- B. NaCl solution
- C. NaOH solution
- D. $NH_4Cl + NH_4OH$

Answer:



17. A light greenish coloured salt was soluble in water. On passing H_2S into the solution, a black ppt. was obtained which dissolved readily in HCl. The metal ion present is



Answer:



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18. Colour of cobalt chloride solution is

A. Pink

B. Black

C. Colourless

D. Green

Answer:



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19. If NaOH is added to an aqueous solution of zinc ions a white precipitate appears and on adding excess of NaOH, the precipitate dissolves. In this solution, zinc exists in the

A. Cationic part

B. Anionic part

C. Both in the cationic and anionic parts

D. there is no zinc ion in the solution

Answer:



20. Group reagent for the precipitation of group II basic radicals for the qualitative analysis is

A. Dil. $\text{HCl} + \text{H}_2\text{S}$

B. $\text{NH}_4\text{OH} + \text{H}_2\text{S}$

C. only H_2S

D. none of these

Answer:



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21. In the fifth group, $(NH_4)_2CO_3$ is added to precipitate out the carbonates, we do not add Na_2CO_3 because

A. $CaCO_3$ is soluble in Na_2CO_3

B. Na_2CO_3 increases the solubility of fifth group carbonates

C. $MgCO_3$ will be precipitated out in fifth group

D. none of these

Answer:



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22. Concentrated sodium hydroxide can separate a mixture of

A. Al^{3+} and Cr^{3+}

B. Cr^{3+} and Fe^{3+}

C. Al^{3+} and Zn^{2+}

D. Zn^{2+} and Pb^{2+}

Answer:



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23. A metal sulphide which is soluble in water and white in colour is

A. CuS

B. Na_2S

C. PbS

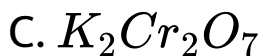
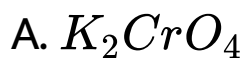
D. ZnS

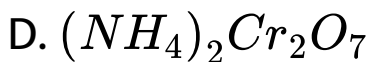
Answer:



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24. In the chromyl chloride test the reagent used is





Answer:



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25. Addition of a solution of oxalate to an aqueous solution of mixture of Ba^{2+} , Sr^{2+} and Ca^{2+} will precipitate



C. Ba^{2+} and Sr^{3+}

D. All the above

Answer:

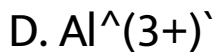
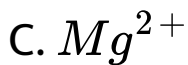


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26. Which of the following cations is detected by the flame test?

A. NH_4^+

B. K^+



Answer:



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27. A salt which on heating with conc. H_2SO_4 gives violet vapours is

A. sulphate

B. bromide

C. iodide

D. nitrate

Answer:



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28. Which does not give borax bead test?

A. Cr^{3+}

B. Cu^{2+}

C. Mn^{2+}

D. Pb^{2+}

Answer:



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29. The reagent silver sulphate solution is used to separate

A. Nitrate and bromide

B. nitrate and chlorate

C. Bromide and iodide

D. Nitrate and nitrite

Answer:



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30. In qualitative inorganic analysis, phosphate, if present, is to be eliminated in the appropriate group in order to detect the radical

A. Pb^{2+}



Answer:



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31. Which of the following radicals will not be precipitated by passing H_2S in concentrated acid solution?

A. Copper

B. Antimony

C. Arsenic

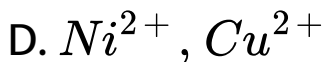
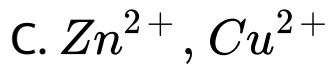
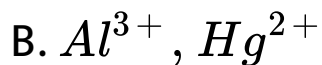
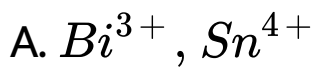
D. Cadmium

Answer:



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32. Which one among the following pairs of ions cannot be separated by H_2S in dilute hydrochloric acid?



Answer:



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33. Which one of the following metals will give blue ash when its salt is heated with Na_2O_3 solid and $Co(NO_3)_2$ on a charcoal piece?

A. Cu

B. Mg

C. Al

D. Zn

Answer:



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34. Yellow ammonium sulphide solution is a suitable reagent for the separation of

A. HgS and PbS

B. PbS and Bi_2S_3

C. Bi_2S_3 and CuS

D. CdS and As_2S_3

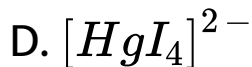
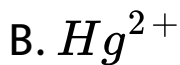
Answer:



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35. In Nessler's reagent, the active ion is

A. Hg^+



Answer:



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36. Which of the following forms a hydroxide highly soluble in water?





Answer:



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37. A metal salt solution gives a yellow precipitate with silver nitrate. The precipitate dissolves in dilute nitric acid as well as in ammonium hydroxide. The solution contains

A. bromide

B. iodide

C. phosphate

D. chromate

Answer:



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38. A metal salt solution forms a yellow precipitate with potassium chromate in acetic acid, a white precipitate with dilute H_2SO_4

but gives no precipitate with sodium chloride or iodide, it is

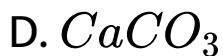
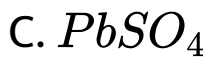
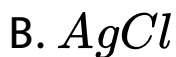
- A. Lead carbonate
- B. Basic lead carbonate
- C. Barium carbonate
- D. strontium carbonate

Answer:



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39. Which is soluble in NH_4OH ?

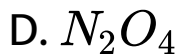
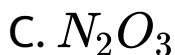
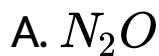


Answer:



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40. Which of the following combines with Fe(II) ions to form a brown complex?



Answer:



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41. Which gives violet colour with borax?

A. Fe

B. Ni

C. Co

D. Mn

Answer:



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42. When conc. H_2SO_4 is added to dry KNO_3

brown fumes evolve. These fumes are of

A. SO_2

B. SO_3

C. NO

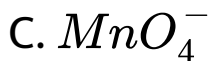
D. NO_2

Answer:



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43. Nessler's reagent is used to detect



Answer:



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44. The compound insoluble in acetic acid is

- A. Calcium oxide
- B. calcium dioxide
- C. calcium oxalate
- D. calcium hydroxide

Answer:

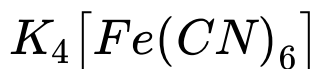


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45. Prussian blue is formed when

- A. Ferrous sulphates reacts with $FeCl_3$

B. Ferric sulphate reacts with



C. Ferrous ammonium sulphate reacts with



D. Ammonium sulphates reacts with $FeCl_3$

Answer:



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46. What product is formed by mixing the solution of $K_4[Fe(CN)_6]$ with the solution of $FeCl_2$?

- A. Ferro ferricyanide
- B. Ferric ferrocyanide
- C. Ferriferricyanide
- D. none.

Answer:

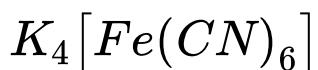


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47. A blue colouration is not obtained when

A. Ammonium hydroxide dissolves in copper sulphate

B. copper sulphate solution reacts with



C. Ferric chloride reacts with sodium

ferrocyanide

D. Anhydrous white $CuSO_4$ is dissolved in

water

Answer:



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48. Which of the following will not give positive chromyl chloride test?

A. Copper chloride, $CuCl_2$

B. mercuric chloride, $HgCl_2$

C. Zinc chloride, $ZnCl_2$

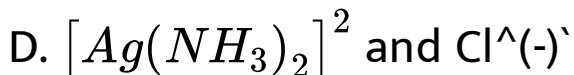
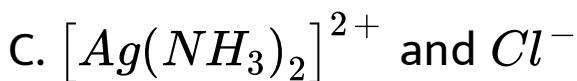
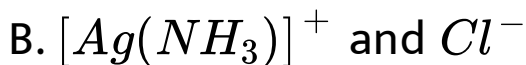
D. Anilinium chloride $C_6H_5NH_3Cl$

Answer:



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49. AgCl dissolves in ammonia solution giving



Answer:



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50. Nitrate of all metals are

A. coloured

B. unstable

C. soluble in water

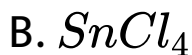
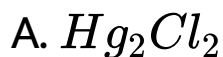
D. insoluble in water.

Answer:



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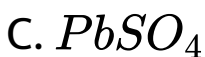
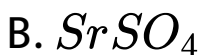
51. When excess of SnCl_2 is added to a solution of HgCl_2 , a white precipitate turning grey is obtained. The grey colour is due to the formation of



Answer:



52. A white crystalline substance dissolves in water. On passing H_2S in this solution, a black ppt. is obtained. The black ppt. dissolves completely in hot HNO_3 . On adding few drops of conc. H_2SO_4 , a white ppt. is obtained. The ppt. is that of



D. $CdSO_4$

Answer:



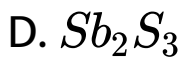
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53. Of the following sulphides which one is insoluble in dil. Acids but soluble in alkalies

A. PbS

B. CdS

C. FeS



Answer:



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54. When chlorine water is added to an aqueous solution of potassium halide in the presence of chloroform, a violet colour is obtained on adding more of chlorine water, the violet colour disappears, and a colourless

solution is obtained. This test confirms the presence of the following in aqueous solution

A. Iodide

B. Bromide

C. Chloride

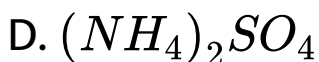
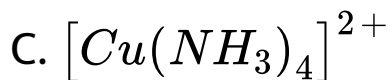
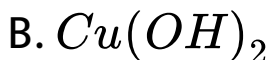
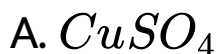
D. Iodide and bromide

Answer:



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55. When excess of dilute NH_4OH is added to an aqueous solution of copper sulphate, an intense blue colour is obtained. This is due to the presence of

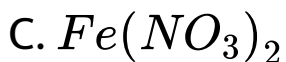


Answer:



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56. Composition of brown ring is



Answer:



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57. Chloride of which element is coloured?

A. Ag

B. Hg

C. Zn

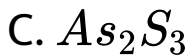
D. Co.

Answer:



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58. The composition of golden spangles is



Answer:



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59. Chromyl chloride test is performed for the confirmation of the presence of the following in a mixture

A. sulphate

B. chromium

C. chloride

D. chromium and chloride

Answer:



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60. An aqueous solution of colourless metal sulphate M gives a white precipitate with NH_4OH . This was soluble in excess of

NH_4OH . On passing H_2S through this solution a white ppt. is formed. The metal M in the salt is

A. Ca

B. Ba

C. Al

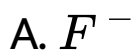
D. Zn

Answer:



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61. Which of the following will not produce a precipitate with $AgNO_3$ solution?

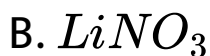
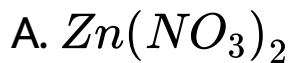


Answer:



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62. The aqueous solution of the following salts will be coloured



D. potash alum

Answer:



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63. When $AgNO_3$ is strongly heated, the products formed are

A. NO and NO_2

B. NO_2 and O_2

C. NO_2 and N_2O

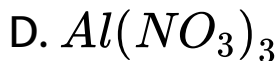
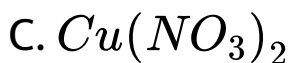
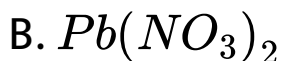
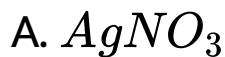
D. NO_2 and O_2

Answer:



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64. Which of the nitrates on strong heating leaves the metal as the residue?



Answer:



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65. Which will give borax bead test with blue bead?

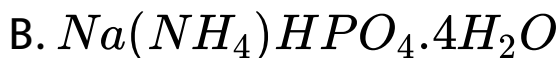


Answer:



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66. Microcosmic salt is



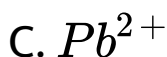
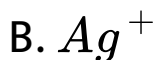
D. none of the above

Answer:



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67. The ion which is not precipitated by H_2S in the presence of HCl is

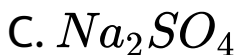
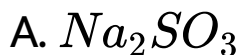


Answer:



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68. Which of the following salt would give SO_2 with hot and dil. H_2SO_4 and also decolourise Br_2 water?



Answer:



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69. A salt which gives CO_2 with hot H_2SO_4 and also decolourizes acidified $KMnO_4$ on warming is

A. HCO_3^-

B. CO_3^{2-}

C. Oxalate

D. Acetate

Answer:



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70. Which of the following precipitates does not dissolve even in large excesses of NH_4OH ?

A. AgCl

B. AgBr

C. AgI

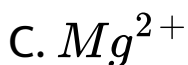
D. None of these

Answer:



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71. The reagent NH_4Cl and aqueous NH_3 will precipitate



Answer:



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72. Cu^{2+} ions will be reduced to Cu^+ ions by the addition of an aqueous solution of

A. KF

B. KCl

C. KI

D. KOH

Answer:



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73. AgCl is soluble in

A. aqua regia

B. H_2SO_4

C. HCl

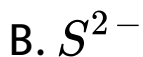
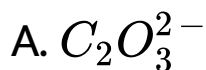
D. NH_4OH

Answer:



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74. A substance on treatment with dil. H_2SO_4 liberates a colourless gas which produces (i) turbidity with baryta water and (ii) turns acidified dichromate solution green. The reaction indicates the presence of



Answer:



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75. In which of the following solvents, AgBr will have the highest solubility?

A. 10^{-3} M NaBr

B. 10^{-3} M NH_4OH

C. pure water

D. 10^{-3} M HBr

Answer:



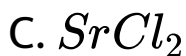
76. When copper nitrate is strongly heated, it is converted into

- A. Cu metal
- B. Cupric oxide
- C. Cuprous oxide
- D. Copper nitrate

Answer:



77. The salt insoluble in cold water but soluble in boiling water is



Answer:



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78. On the addition of a solution containing CrO_4^{2-} ions to the solution of Ba^{2+} , Sr^{2+} and Ca^{2+} ions, the ppt obtained first will be of



D. A mixture of all the three

Answer:



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79. A pale green crystalline metal salt of M dissolves freely in water. It gives a brown precipitate on addition of aqueous NaOH. The metal salt solution also gives a black precipitate on bubbling H_2S in aqueous medium. An aqueous solution of the metal salt decolourizes the pink colour of the permanganate solution. The metal in the metal salt solution is

A. copper

B. Aluminium

C. Lead

D. Iron

Answer:



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80. Turn bull's blue is a compound

A. Ferricyanide

B. Ferrous ferricyanide

C. Ferrous cyanide

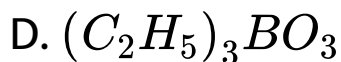
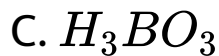
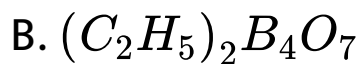
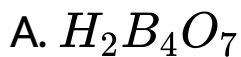
D. Ferriferrocyanide

Answer:



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81. Sodium borate on reaction with conc. H_2SO_4 and C_2H_5OH gives a compound A which burns with a green edged flame. The compound A is



Answer:



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82. On addition of aqueous NaOH to a salt solution, a white gelatinous precipitate is

formed, which dissolves in excess of alkali. The salt solution contains

- A. Chromous ions
- B. Aluminium ions
- C. Barium ions
- D. Iron ions.

Answer:



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83. Brown ring test is used to detect

A. Iodide

B. Nitrate

C. Iron

D. Bromide

Answer:



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84. Metal halide insoluble in water is

A. AgI

B. $CaCl_2$

C. KBr

D. AgF

Answer:



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85. When $K_2Cr_2O_7$ crystals are heated with conc. HCl, the gas evolved is

A. O_2

B. Cl_2

C. CrO_2Cl_2

D. HCl

Answer:



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86. Which is most soluble in water?

A. AgCl

B. AgBr

C. AgI

D. AgF

Answer:



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87. On passing H_2S gas in II group sometimes the solution turns milky. It indicates the presence of

A. Oxidising agent

B. acidic salt

C. Thiosulphate

D. Reducing agent

Answer:



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88. Dimethyl glyoxime in a suitable solvent was refluxed for 10 minutes with pure pieces of nickel sheet, it will result in

A. Red ppt.

B. Blue ppt.

C. Yellow ppt.

D. no ppt.

Answer:



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89. A mixture of chlorides of copper, cadmium, chromium, iron and aluminium was dissolved in water acidified with HCl and hydrogen sulphide gas was passed for sufficient time. It was filtered, boiled and a few drops of nitric acid were added while boiling. To this solution ammonium chloride and sodium hydroxide were added and filtered. The filtrate shall give test for

A. sodium and iron

B. sodium and aluminium

C. aluminium and iron

D. sodium, iron , cadmium and aluminium

Answer:



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90. A metal is burnt in air and the ash on moistening smells of ammonia. The metal is

A. Na

B. Fe

C. Mg

D. Al

Answer:



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91. Which of the following pairs is not distinguished by passing H_2S ?

A. Hg,Pb

B. Cd,Pb

C. As,Cd

D. Zn,Mn

Answer:



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92. In qualitative analysis NH_4Cl is added before NH_4OH

A. To decreases OH^- concentration

B. To increases OH^- concentration

C. for making HCl

D. Statement is wrong.

Answer:



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93. Which compound does not dissolve in hot dil. HNO_3 ?

A. HgS

B. PbS

C. CuS

D. CdS

Answer:



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94. An aqueous solution $FeSO_4 \cdot Al_2(SO_4)_3$ and chrome alum is heated with excess of Na_2O_2 and filtered. The materials obtained are

- A. A colourless filtrate and a green residue
- B. A yellow filtrate and a green residue
- C. A yellow filtrate and a brown residue
- D. a green filtrate and a brown residue.

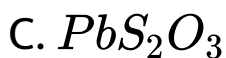
Answer:



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95. A salt on treatment with dil. HCl gives a pungent smelling gas and a yellow precipitate. The salt gives green flame when tested. The

solution gives a yellow ppt. with potassium chromate. The salt is

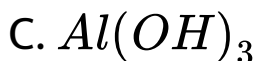
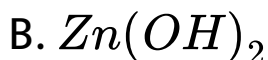
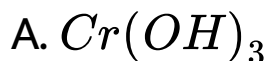


Answer:



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96. Which of the following compound on reaction with NaOH and Na_2O_2 gives yellow colour?



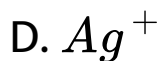
D. none of these

Answer:



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97. Which of the following cannot give iodometric titration?

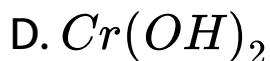
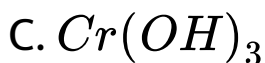
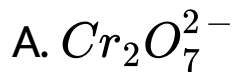


Answer:



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98. CrO_3 dissolves in aqueous NaOH to give



Answer:



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99. The only cations present in a slightly acidic solution are Fe^{3+} , Zn^{2+} and Cu^{2+} . The reagent that when added in excess to this solution would identify and separate Fe^{3+} in one step is

A. 2M HCl

B. 6M NH_3

C. 6 M NaOH

D. H_2S gas

Answer:



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100. When HNO_3 is dropped into the palm and washed with water, it turns yellow. It shows the presence of

A. NO_2

B. N_2O

C. NO

D. N_2O_3

Answer:



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101. Which one of the following is not efflorescent

A. Hydrated Na_2CO_3

B. hydrated $CuSO_4$

C. NaOH

D. All of these

Answer:



102. $K_2[HgI_4]$ detects ion/group

A. NH_2

B. NO

C. NH_4^+

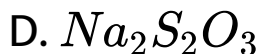
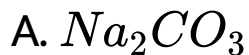
D. Cl^-

Answer:



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103. Which of the following does not react with AgCl ?

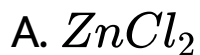


Answer:



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104. Which one of the following does not produce metallic sulphide with H_2S ?



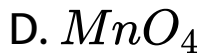
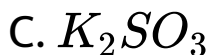
Answer:



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105. By passing H_2S gas in acidified $KMnO_4$,

we get



Answer:



Watch Video Solution

106. Which metal salt gives a violet coloured bead in the borax bead test ?



Answer:



Watch Video Solution

107. Which of the following gives a precipitate with $Pb(NO_3)_2$ but not with $Ba(NO_3)_2$?

A. Sodium chloride

B. sodium acetate

C. sodium nitrate

D. sodium hydrogen phosphate

Answer:



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108. Whci of the following is soluble in yellow ammonium sulphide ?

A. CuS

B. CdS

C. SnS

D. PbS

Answer:



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109. Which of the following statement is correct?

A. Fe^{2+} gives brown colour with ammionium thiocyanate

B. Fe^{2+} gives blue ppt. with potassium ferricyanide

C. Fe^{3+} gives brown colour with potassium ferrocyanide

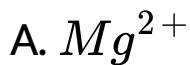
D. Fe^{3+} gives red colour with potassium
ferrocyanide

Answer:



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110. A chloride dissolves appreciably in cold water. When placed on platinum wire in Bunsen flame, no distinctive colour is noticed, the cation would be



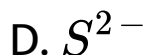
Answer:



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111. A white sodium salt dissolves in water to give a solution which is neutral to litmus. When silver nitrate solution is added to the

solution, a white ppt. is obtained which does not dissolve in dil. HNO_3 . The anion is



Answer:



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112. A mixture of two salts is not water soluble but dissolves completely in dil. HCl to form a colourless solution. The mixture could be

A. $AgNO_3$ and KBr

B. $BaCO_3$ and ZnS

C. $FeCl_3$ and $CaCO_3$

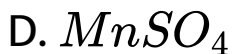
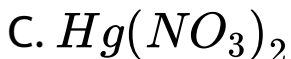
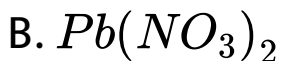
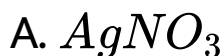
D. $Mn(NO_3)_2$ and $MgSO_4$

Answer:



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113. Three separate samples of a solution of a single salt gave these results. One formed a white ppt. with excess ammonia solution, one formed a white ppt. with dil. NaCl solution and one formed a black ppt. with H_2S . The salt could be

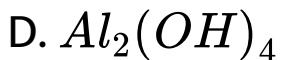
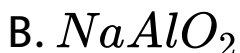
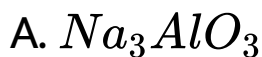


Answer:



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114. Action of caustic soda on $Al(OH)_3$ gives a compound having formula



Answer:



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115. When a substance A reacts with water, it produces a combustible gas B and a solution of substance C in water. When another substance D reacts with this solution of C . It also produces gas B on reaction with dilute sulphuric acid at room temperature. A imparts a deep golden yellow colour to the

smokless flame of bunsen flame A,B , C and D
are respectively

A. $Na, H_2, NaOH, Zn$

B. K, H_2, KOH, Al

C. $Ca, H_2, Ca(OH)_2, Sn$

D. $CaC_2, C_2H_2, Ca(OH)_2, Fe$

Answer:



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116. One litre flask is full of brown bromine fumes. The intensity of brown colour of vapour will not decrease appreciably on adding to the flask some

A. Pieces of marble

B. Animal charcoal powder

C. CCl_4

D. CS_2

Answer:



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117. Which one of the following ionic species will impart colour to an aqueous solution?



Answer:



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118. Which is not dissolved in dil. HCl?

A. ZnS

B. MnS

C. $BaSO_3$

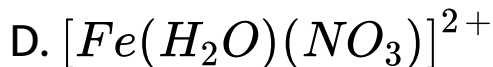
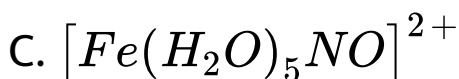
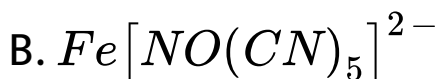
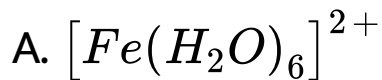
D. $BaSO_4$

Answer:



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119. The brown ring test for NO_3^- is due to the formation of the complex ion with formula

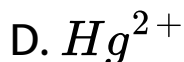
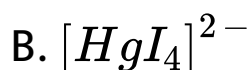
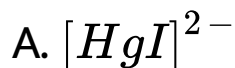


Answer:



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120. In Nessler's reagent, the ion present is



Answer:



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121. When I_2 is passed through KCl, KF and KBr

A. Cl_2 and Br_2 are evolved

B. Cl_2 is evolved

C. Cl_2 , F_2 and Br_2 are evolved

D. None of these

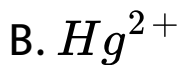
Answer:



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122. Mercurous ion is represent as

A. Hg_2^{2+}

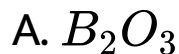


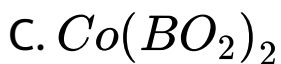
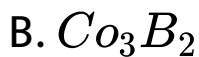
Answer: B



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123. In the borax test of Co^{2+} , the blue colour of bead is due to the formation of





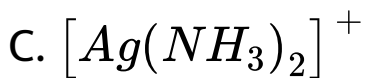
Answer: B::C::D



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124. AgCl is soluble in NH_4OH solution. The solubility is due to the formation of





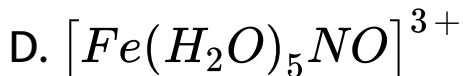
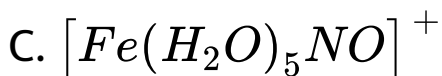
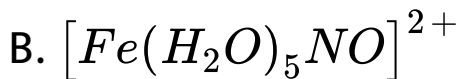
Answer: A::C



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125. Correct formula of the complex formed in the brown ring test for nitrates is





Answer: B::D



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126. Which of the following gives blood red colour with KCNS?



B. Fe^{3+} ions

C. Al^{3+}

D. Zn^{2+}

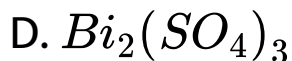
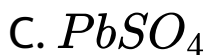
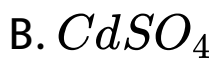
Answer:



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127. Which of the following sulphates is insoluble in water ?

A. $CuSO_4$



Answer:



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128. Which of the following is not precipitated as sulphide by passing H_2S in the presence of conc. HCl ?

A. Copper

B. Arsenic

C. Cadmium

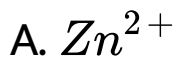
D. Lead

Answer:



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129. The metal ion which is precipitated when H_2S is passed with HCl is



Answer:



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130. The gas that turns lime water milky is



B. SO_2

C. Both of these

D. None of these

Answer:



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131. Which of the following is not a preliminary test used to detect ions ?

A. Borax bead test

B. Flame test

C. Brown ring test

D. Permanganate test

Answer:



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132. Which of the following metal sulphides has maximum solubility in water ?





Answer:



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133. The phenomenon in which white transparent crystal change into white powder is called

A. Deliquescence

B. Efflorescence

C. Allotropy

D. Sublimation

Answer:



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134. The compound formed in the borax bead test of Cu^{2+} ion in oxidising flame is

A. Cu

B. $CuBO_3$

C. $Cu(BO_2)_2$

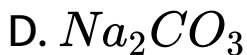
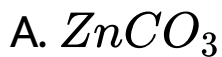
D. None of these

Answer:



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135. Which of the following does not give CO_2 on heating ?



Answer:



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136. When H_2S gas is passed through the HCl containing aqueous solution of

$CuCl_2$, $HgCl_2$, $BiCl_3$ and $CoCl_2$, it does not precipitated out

A. CuS

B. HgS

C. Bi_2S_3

D. CoS

Answer:



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137. Mark the correct statement

A. I group basic radicals precipitates as chlorides

B. IV group basic radicals precipitated as sulphides

C. V groups basic radicals precipitates as carbonates

D. All the above statements are correct

Answer:



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138. Potassium chromate solution is added to an aqueous solution of a metal chloride. The precipitate thus obtained are insoluble in acetic acid. These are subjected to flame test, the colour of the flame is

A. Lilac

B. Apple green

C. Crimson red

D. Golden yellow

Answer:



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139. In borax bead test which compound is formed ?

A. Orthoborate

B. Metaborate

C. Double oxide

D. Tetraborate

Answer:



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140. Consider the following observation:

ItbRgt $M^{n+} + HCl \rightarrow$ white precipitate

$\xrightarrow{\Delta}$ water soluble.

The metal ion M^{n+} will be

A. Hg^{2+}



Answer:

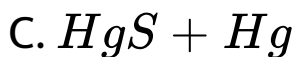
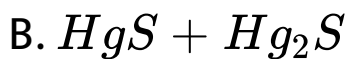


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141. When H_2S is passed through Hg_2^{2+} , we

get





Answer:



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142. How do we differentiate between Fe^{3+} and Cr^{3+} in group III ?

A. By adding excess of NH_4OH

B. by increasing NH_4^+ ions concentration

C. by decreasing OH^- ion concentration

D. Both (b) and (C)

Answer:



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143. Which compound does not dissolve in hot dilute HNO_3 ?

A. HgS

B. CuS

C. PbS

D. CdS

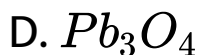
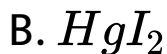
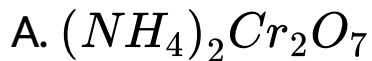
Answer:



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144. A red solid is insoluble in water. However, it becomes soluble if some KI is added to water. Heating rod solid in a test tube produces violet coloured fumes and droplets

of metal appear on the cooler parts of test tube. The red solid is



Answer:



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145. Which of the following nitrates will leave behind a metal on strong heating ?

- A. Ferric nitrate
- B. copper nitrate
- C. maganese nitrate
- D. silver nitrate

Answer:



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146. Which one of the following statements is correct ?

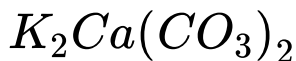
A. Manganese salt give violet borax bead test in the reducing flame

B. from a mixed precipitate of AgCl and AgI , ammonia solution dissolves only AgCl

C. Ferric ions give a deep green precipitate on adding potassium ferrocyanide solution

D. On boiling a solution having K^+ , Ca^{2+}

and HCO_3^- ions we get a precipitate of



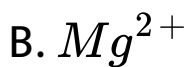
Answer:



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147. In Nessler's reagent for the detection of ammonia the active species is





Answer: B



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148. In qualitative analysis ,in order to detect second group basic radical, H_2S gas is passed in the presence of dil. HCl to

A. Increases the dissociation of H_2S

B. Decreases the dissociation of the salt
solution

C. decreases the dissociation of H_2S

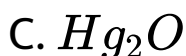
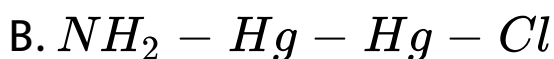
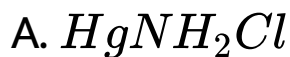
D. increases the dissociation of salt
solution

Answer:



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149. Calomel (Hg_2Cl_2) on reaction with ammonium hydroxide gives

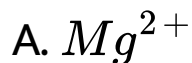


Answer:



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150. The radical can be confirmed by Borax bead test is



Answer:



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151. Which gives (s) yellow precipitate with

K_2CrO_4 ?

A. Ba^{2+} ions

B. Cd^{2+}

C. Ca^{2+}

D. Sr^{2+}

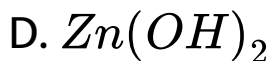
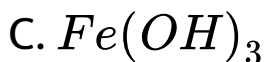
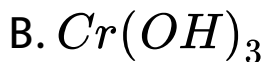
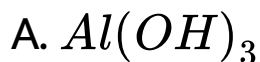
Answer:



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Selected Straight Objective

1. Which of the following is (are) soluble in excess of NaOH ?



Answer:



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2. KI solution identifies



Answer:



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3. Diphenylamine reagent gives a deep blue colour with a solution. It contains

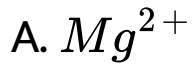


Answer:



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4. Flame test is not given by



Answer:



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5. Borax bead test is not given by

A. Copper salts

B. Aluminium salts

C. Nickel salts

D. magnesium salts

Answer:



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6. Potassium ferrocyanide is used in the detection of

A. Fe^{2+} ion

B. Fe^{3+} ions

C. Cu^{2+} ion

D. Cd^{2+} ion

Answer:



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7. Potassium cyanide is used for separating



Answer:



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8. A solution giving yellow ppt. with ammonium molybdate contains



Answer:



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9. Which is not decomposed by dil. H_2SO_4 ?

A. Chloride

B. carbonate

C. Nitrate

D. Acetic acid

Answer:



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10. A solution of salt in HCl when diluted with water turns milky. It indicates the presence of

A. Al

B. Bi

C. Sb

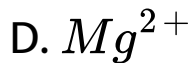
D. Zn

Answer:



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11. The reagent NH_4Cl and aqueous NH_3 will precipitate



Answer: B



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12. Which of the following statement (s) is (are) correct when a mixture of NaCl and $K_2Cr_2O_7$ is gently warmed with conc. H_2SO_4 ?

A. A deep red vapour is evolved

B. the vapour when passed into NaOH solution gives a yellow solution of Na_2CrO_4

C. Chlorine gas is evolved

D. Chromyl chloride is formed

Answer:



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13. Which of the following statements (s) is (are) correct with reference to ferrous and ferric ions

A. Fe^{3+} gives brown colour with potassium ferricyanide

B. Fe^{3+} gives blue precipitate with potassium ferricyanide

C. Fe^{3+} gives red colour with potassium thiocyanate

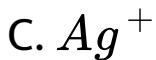
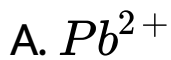
D. Fe^{3+} gives brown colour with ammonium thiocyanate

Answer:



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14. The ion that cannot be precipitated by both HCl and H_2S is

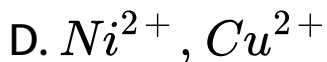
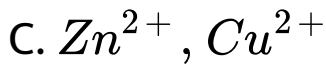
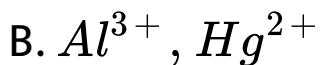
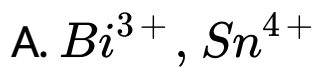


Answer:



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15. Which one among the following pairs of ions cannot be separated by H_2S in dilute hydrochloric acid ?

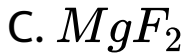
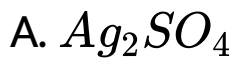


Answer:



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16. Which of the following compounds is expected to be coloured ?



Answer:



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17. Sodium nitrate decomposes above $800^\circ C$
to give

A. N_2

B. O_2

C. NO_2

D. Na_2O

Answer:



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18. On heating ammonium dichromate the gas evolved is

A. oxygen

B. ammonia

C. nitrous oxide

D. nitrogen

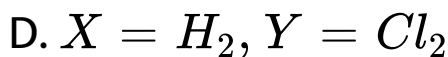
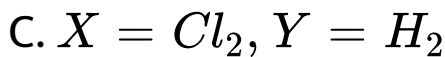
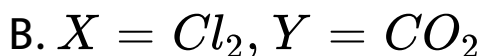
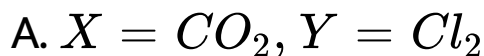
Answer:



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19. A gas X is passed through water to form a saturated solution. The aqueous solution on treatment with silver nitrate give a white

precipitate. The saturated aqueous solution dissolve magnesium ribbon with evolution of a colourless gas 'Y'. Identify X and Y

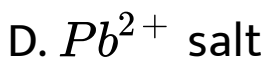
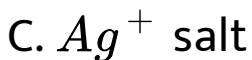
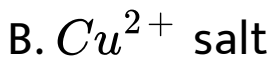
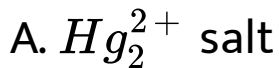


Answer:



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20. An aqueous solution of a substance give a white precipitate on treatment with dilute hydrochloric acid, which dissolve on heating. When hydrogen sulphide is passed through the hot acidic solution, a black precipitate is observed. the substance is a

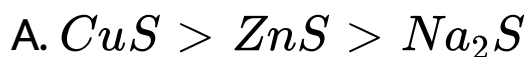


Answer:



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21. Identify the correct order of solubility of Na_2S , CuS and ZnS in aqueous solution



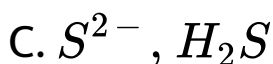
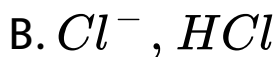
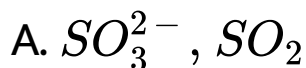
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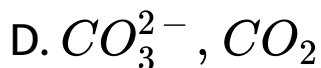


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22. $[X] + H_2SO_4 \rightarrow [Y]$ colourless with irritating smell. $[Y] + K_2Cr_2O_7 + H_2SO_4 \rightarrow$ green solution

$[X]$ and $[Y]$ are





Answer:



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23. A sodium salt on treatment with $MgCl_2$ gives white precipitate only on heating. The anion of the sodium salt is :





Answer:



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24. A metal nitrate on reaction with KI gives black precipitate and with excess KI gives orange solution. The metal ion is ,





Answer:

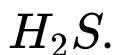


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Assertion And Reason

1. CuS will give H_2S in dilute acid test.

! All sulphide react with dil. H_2SO_4 to give



A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



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2. $PbCl_2$ will give HCl in conc. H_2SO_4 test.

! All chlorides react with conc. H_2SO_4 , on heating to give HCl

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

3. $ZnCO_3$ will not give any gas when treated with conc. H_2SO_4

! CO_3^{2-} can only be detected in dil. Acid test

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. Both A and are false

Answer:



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4. \$ CdS is yellow in colour

! Cd^{2+} salts are yellow in colour.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



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5. \$ A brown gas which intensifies on adding copper turnings in conc. H_2SO_4 test is NO_2

! Copper reacts with conc. HNO_3 to give NO_2

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true.

Answer:



Watch Video Solution

6. \$ CuS is blue in colour.

! All Cu^{2+} salts are blue in colour.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



7. \$ Acidified $K_2Cr_2O_7$ is turned green when SO_2 is passed through it.

! In this reaction SO_2 acts as a reducing agent.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

8. \$ White ppt. of AgCl is soluble in NH_4OH

! It is due to the formation of soluble complex.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



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9. \$ All soluble sulphides give white ppt. with $BaCl_2$ solution.

! BaS is insoluble in water

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true.

Answer:



Watch Video Solution

10. \$ A solution of $BiCl_3$ in conc. HCl when diluted with water gives white ppt.

! $BiCl_3$ is insoluble in dil. HCl

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

11. \$ Borax bead test is applicable to coloured salts

! In borax bead test, coloured salts are decomposed to give coloured metal metaborates.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

12. S^{2-} Phosphates are identified by the yellow precipitate obtained on adding ammonium molybdate solution

! Ammonium phosphomolybdate is a yellow compound.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

13. Cu^{2+} and Cd^{2+} are separated by first adding KCN solution and then passing H_2S gas.

! KCN reduces Cu^{2+} to Cu^+ and forms a complex with it.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

14. \$ Oxalate gives white ppt. with calcium chloride in the presence of HCl

! Calcium chloride is insoluble in water.

- A. Both A and R are true and R is the correct explanation of A
- B. Both A and R are true but R is not a correct explanation of A
- C. A is true but R is false
- D. A is false but R is true.

Answer:



Watch Video Solution

15. \$ V group basic radicals are precipitated as their carbonates in presence of NH_4OH

! NH_4OH maintains the pH of the solution basic.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



Watch Video Solution

16. \$ Sb(III) is not precipitated as sulphide when in its alkaline solution H_2S is passed.

! The concentration of S^{2+} ions in alkaline medium is inadequate for precipitation.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



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17. \$ Addition of NH_4OH to an aqueous solution of $BaCl_2$ in the presence of NH_4Cl

(excess) precipitates $Ba(OH)_2$.

! $Ba(OH)_2$ is insoluble in water.

A. Both A and R are true and R is the correct explanation of A

B. Both A and R are true but R is not a correct explanation of A

C. A is true but R is false

D. A is false but R is true.

Answer:



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Ultimate Preparatory Package

1. Colour of $KMnO_4$ is decolourised without evolution of any gas. The radical present may be



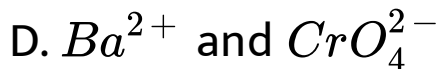
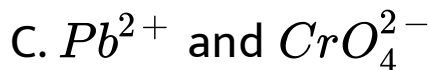
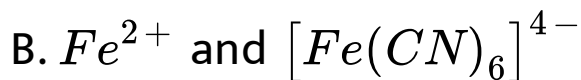
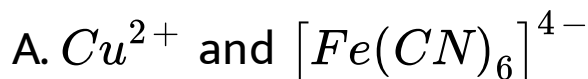
D. Both (b) and (C)

Answer:



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2. Reddish brown (chocolate) precipitated is formed with



Answer:



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3. An aqueous solution Hg^{2+} , Hg_2^{2+} , Pb^{2+} and Cd^{2+} . The addition of 6N HCl will precipitate

- A. Hg_2Cl_2 only
- B. $PbCl_2$ only
- C. $PbCl_2$ and Hg_2Cl_2
- D. $PbCl_2$ and $HgCl_2$

Answer:



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4. Nitric acid is generally not used for the preparation of original solution in analysis of basic radicals because it

- A. is an oxidising agent
- B. is reducing agent
- C. forms insoluble nitrates
- D. forms soluble nitrates.

Answer:



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5. Sulphuric acid is not for the preparation of original solution in analysing basic radicals because

- A. it is a reducing agent
- B. it forms insoluble sulphate
- C. it forms a soluble complex
- D. it is viscous in nature

Answer:



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6. Some salts, although contain two different metal elements, give test for only one of them in solution, such salts are

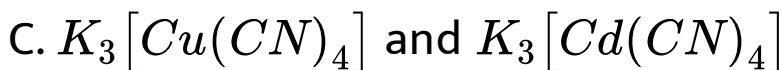
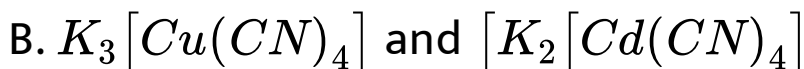
- A. normal salts
- B. double salts
- C. complex salts
- D. basic salts

Answer:



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7. The complex compound formed with KCN solution is added to solution containing both Cu^{2+} and Cd^{2+} ions are



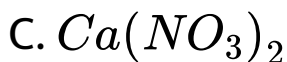
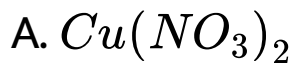
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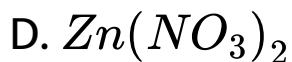


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Brain Teasers 23

1. Which of the following nitrates does not give NO_2 on heating ?





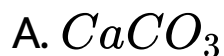
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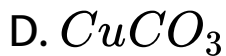
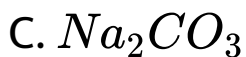
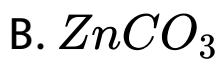


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Brain Teasers 24

1. Which of the following carbonate does not give CO_2 on heating ?





Answer:



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Brain Teasers 25

1. Which of the following carbonate cannot be easily detected by dilute acid test ?

A. $ZnCO_3$

B. Na_2CO_3

C. $PbCO_3$

D. $FeCO_3$

Answer:



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Brain Teasers 26

1. A pink violet salt on heating change to blue.

It may be due to the presence of

A. $MnCl_2$

B. $CoCl_2$

C. $MnSO_4$

D. none of these

Answer:



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1. White ppt. of $PbSO_4$ is soluble in

A. Conc. HCl on heating

B. Conc. HNO_3

C. $(NH_4)_2CO_3$

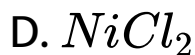
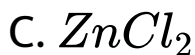
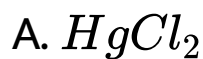
D. CH_3COONH_4

Answer:



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1. which of the following will not respond to chromyl chloride test ?



Answer:



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Brain Teasers 29

1. The chloride soluble in hot water is

A. AgCl

B. PbCl_2

C. Hg_2Cl_2

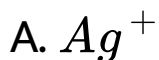
D. none of these

Answer:



Brain Teasers 30

1. A solution of a colourless salt in conc. HCl on dilution with water gives white ppt. it is due to



D. none of these

Answer:



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Brain Teasers 31

1. What is correct about the cations of group V?

A. These cations should be tested in the sequence Ba,Sr,Ca

B. These cations should be tested in the sequence Ca,Sr,Ba

C. These cations should be tested in the sequence Sr,Ca,Ba

D. These cations can be detected in any sequence

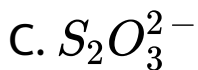
Answer:



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1. A solution of a colourless salt when acidified with dil. HCl, slowly turns milky with a pale yellow tint. When $AgNO_3$ solution is added to the solution of the salt, white ppt. slowly changing to yellow, orange, brown and finally to black is obtained. The anion present in the salt is





Answer:



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Brain Teasers 33

1. Ni^{2+} with dimethylglyoxime reagent in alkaline solution (NH_4OH) gives

A. Green ppt

B. Blue ppt.

C. white ppt

D. brilliant red ppt.

Answer:



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Brain Teasers 34

1. Cu^{2+} ions with NH_4OH solution gives

- A. Blue ppt.
- B. Blue solution
- C. Violet ppt.
- D. None of these

Answer:



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1. Cu^{2+} salts with $K_4[Fe(CN)_6]$ gives

A. Reddish brown ppt.

B. Deep blue solution

C. Blue ppt.

D. None of these

Answer:



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1. Cu^{2+} salts are

- A. deep blue in colour
- B. light blue in colour
- C. light green in colour
- D. colourless.

Answer:



Watch Video Solution

1. CdS is

- A. white in colour
- B. yellow in colour
- C. black in colour
- D. pink in colour

Answer:



Watch Video Solution

1. HgS is

A. red in colour

B. black in colour

C. Both (A) and (B)

D. none.

Answer:



Watch Video Solution

1. CuS is

- A. blue in colour
- B. black in colour
- C. dirty white in colour
- D. none of these

Answer:



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Brain Teasers 40

1. If a salt solution containing Fe^{2+} is not treated with a few drops of HNO_3 before starting for group III the ppt. obtained in group III will be

A. redish brown in colour

B. dirty green in colour

C. black

D. white

Answer:



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Brain Teasers 41

1. I^- can be detected by violet colour in CS_2 layer obtained on adding chlorine water. However, on adding excess of Cl_2 , the violet colour disappears. It is due to

- A. the formation of interhalogen compound Icl
- B. removal of I_2 from CS_2 layer by Cl_2
- C. precipitation of I_2 from solution
- D. none of these

Answer:



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1. Green edged flame test for BO_3^{3-} with alcohol and conc. H_2SO_4 is generally carried out in a china dish. However, the test should be carried out in a test tube if one of the cation present in solution is



Answer:



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