



CHEMISTRY

BOOKS - NCERT EXEMPLAR CHEMISTRY (HINGLISH)

CHEMICAL REACTIONS AND EQUATIONS

Chemical Reactions And Equations

1. Which of the following is not a physical change?

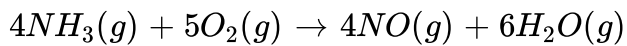
- A. Boiling of water to give water vapour
- B. Melting of ice to give water
- C. Dissolution of salt in water
- D. Combustion of Liquefied Petroleum Gas(LPG)

Answer: D



Watch Video Solution

2. The following reaction is an example of a



1. displacement reaction

2. combination reaction

3. redox reaction

4. neutralisation reaction

A. 1 and 4

B. 2 and 3

C. 1 and 3

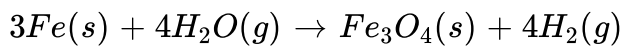
D. 3 and 4

Answer: C



Watch Video Solution

3. Which of the following statements about the given reaction are correct?



1. Iron metal is getting oxidised.
2. Water is getting reduced.
3. Water is acting as reducing agent.
4. Water is acting as oxidising agent.

A. 1, 2 and 3

B. 3 and 4

C. 1, 2 and 4

D. 2 and 4

Answer: C



Watch Video Solution

4. Which of the following are exothermic processes?

1. Reaction of water with quick lime
2. Dilution of an acid
3. Evaporation of water
4. Sublimation of camphor (crystals)

A. 1 and 2

B. 2 and 3

C. 1 and 4

D. 3 and 4

Answer: A



[Watch Video Solution](#)

5. Three beakers labelled as A, B and C each containing 25 mL of water were taken. A small amount of NaOH, anhydrous $CuSO_4$ and NaCl were added to the beakers A, B and C respectively. It was observed that there

was an increase in the temperature of the solutions contained in beakers A and B, whereas in case of beaker C, the temperature of the solution falls. Which one of the following statement (s) is (are) correct?

1. In beakers A and B, exothermic process has occurred.
2. In beakers A and B, endothermic process has occurred.
3. In beaker C, exothermic process has occurred.
4. In beaker C, endothermic process has occurred.

A. Only 1

B. Only 2

C. 1 and 4

D. 2 and 3

Answer: C



Watch Video Solution

6. A dilute ferrous sulphate solution was gradually added to the beaker containing acidified permanganate solution. The light purple colour of

the solution fades and finally disappears. Which of the following is the correct explanation for the observation?

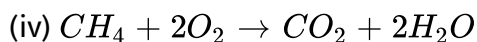
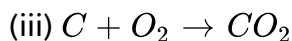
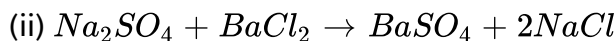
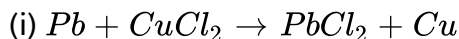
- A. $KMnO_4$ is an oxidising agent, it oxidises $FeSO_4$
- B. $FeSO_4$ acts as an oxidising agent and oxidises $KMnO_4$
- C. The colour disappears due to dilution, no reaction is involved
- D. $KMnO_4$ is an unstable compound and decomposes in the presence of $FeSO_4$ to a colourless compound

Answer: A



[Watch Video Solution](#)

7. Which among the following is (are) double displacement reaction (s)?



A. (i) and (iv)

B. Only(ii)

C. (i) and (ii)

D. (iii) and (iv)

Answer: B



Watch Video Solution

8. Which among the following statement (s) is /are true? Exposure of silver chloride to sunlight for a long duration turns grey due to

(i) the formation of silver by decomposition of silver chloride.

(ii) sublimation of silver chloride.

(iii) decomposition of chlorine gas from silver chloride.

(iv) oxidation of silver chloride.

A. (i) Only

B. (i) and (iii)

C. (ii) and (iii)

D. Only (iv)

Answer: A



Watch Video Solution

9. Solid calcium oxide reacts vigorously with water to form calcium hydroxide accompanied by liberation of heat. This process is called slaking of lime. Calcium hydroxide dissolves in water to form its solution called lime water. Which among the following are true about slaking of lime and the solution formed?

(i) It is an endothermic reaction.

(ii) It is exothermic reaction.

(iii) The pH of the resulting solution will be more than seven.

(iv) The pH of the resulting solution will be less than seven.

A. (i) and (ii)

B. (ii) and (iii)

C. (i) and (iv)

D. (iii) and (iv)

Answer: B



Watch Video Solution

10. Barium chloride on reacting with ammonium sulphate forms barium sulphate and ammonium chloride. Which of the following correctly represents the type of the reaction involved?

(i) Displacement reaction

(ii) Precipitation reaction

(iii) Combination reaction

(iv) Double displacement reaction

A. Only(i)

B. Only (ii)

C. Only (iv)

D. (ii) and (iv)

Answer: D

 [Watch Video Solution](#)

11. Electrolysis of water is a decomposition reaction. The mole ratio of hydrogen and oxygen gases liberated during electrolysis of water is

A. 1:1

B. 2:1

C. 4:1

D. 1:2

Answer: B

 [Watch Video Solution](#)

12. Which of the following is (are) an endothermic process(es)?

- (i) Dilution of sulphuric acid
- (ii) Sublimation of dry ice
- (iii) Condensation of water vapours
- (iv) Evaporation of water

A. Both (i) and (iii)

B. Only (ii)

C. Only (iii)

D. Both (ii) and (iv)

Answer: D



[Watch Video Solution](#)

13. In the double displacement reaction between aqueous potassium iodide and aqueous lead nitrate, a yellow precipitate of lead iodide is

formed. While performing the activity if lead nitrate is not available, which of the following can be used in place of lead nitrate?

A. Lead sulphate (insoluble)

B. Lead acetate

C. Ammonium nitrate

D. Potassium sulphate

Answer: B



Watch Video Solution

14. Which of the following gases can be used for storage of fresh sample of an oil for a long time?

A. Carbon dioxide or oxygen

B. Nitrogen or oxygen

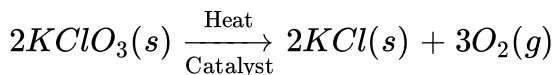
C. Carbon dioxide or helium

D. Helium or nitrogen

Answer: D

 [Watch Video Solution](#)

15. The following reaction is used for the preparation of oxygen gas in the laboratory



Which of the following statement(s) is/are correct about the reaction?

- A. It is a decomposition reaction and endothermic in nature
- B. It is a combination reaction
- C. It is a decomposition reaction and accompanied by release of heat
- D. It is a photochemical decomposition reaction and exothermic in nature

Answer: A

 [Watch Video Solution](#)

16. Which one of the following processes involve chemical reactions?

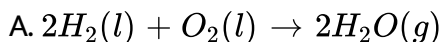
- A. Storing of oxygen gas under pressure in a gas cylinder
- B. Liquefaction of air
- C. Keeping petrol in a China dish in the open
- D. Heating copper wire in the presence of air at high temperature

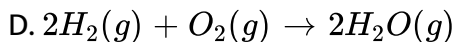
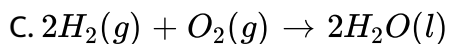
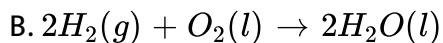
Answer: D



Watch Video Solution

17. In which of the following chemical equations, the abbreviations represent the correct states of the reactants and products involved at reaction temperature?



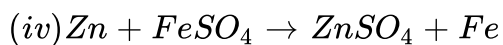
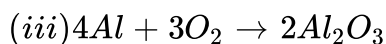
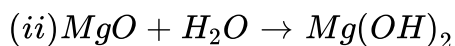
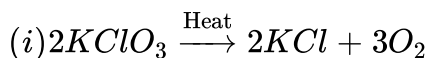


Answer: C



Watch Video Solution

18. Which of the following are combination reactions?



A. Both (i) and (iii)

B. Both (iii) and (iv)

C. Both (ii) and (iv)

D. Both (ii) and (iii)

Answer: D

 [Watch Video Solution](#)

19. Write the balanced chemical equations for the following reactions and identify the type of reaction in each case.

(a) Nitrogen gas is treated with hydrogen gas in the presence of a catalyst at 773 K to form ammonia gas.

(b) Sodium hydroxide solution is treated with acetic acid to form sodium acetate and water.

(c) Ethanol is warmed with ethanoic acid to form ethyl acetate in the presence of concentrated H_2SO_4 .

(d) Ethane is burnt in the presence of oxygen to form carbon dioxide, water and releases heat and light.

 [Watch Video Solution](#)

20. Write the balanced chemical equations for the following reactions and identify the type of reaction in each case.

(a) Iron (III) oxide reacts with aluminium and gives molten iron and aluminium oxide.

(b) Magnesium ribbon is burnt in an atmosphere of nitrogen gas to form solid magnesium nitride.

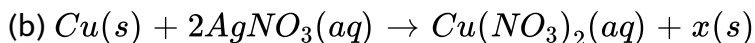
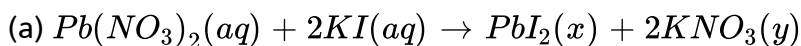
(c) Chlorine gas is passed in an aqueous potassium iodide solution to form potassium chloride solution and solid iodine.

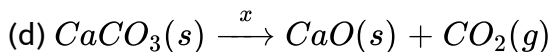
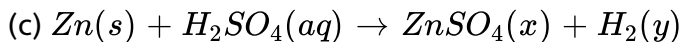
(d) Ethanol is burnt in air to form carbon dioxide, water and releases heat.



[Watch Video Solution](#)

21. Complete the missing components/variables given as x and y in the following reactions





 [Watch Video Solution](#)

22. Which among the following changes are exothermic or endothermic in nature?

(a) Decomposition of ferrous sulphate

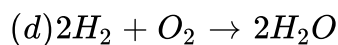
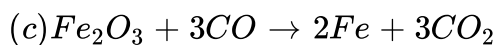
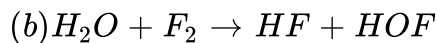
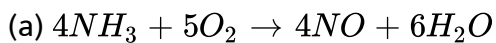
(b) Dilution of sulphuric acid

(c) Dissolution of sodium hydroxide in water

(d) Dissolution of ammonium chloride in water

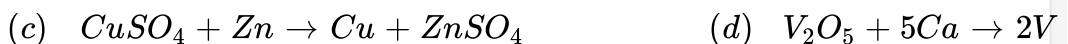
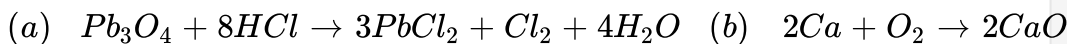
 [Watch Video Solution](#)

23. Identify the reducing agent in the following reactions.



 [Watch Video Solution](#)

24. Identify the oxidising agent (oxidant) in the following reactions



Watch Video Solution

25. Write the balanced chemical equations for the following reaction.

(a) Sodium carbonate on reaction with hydrochloric acid in equal molar concentrations gives sodium chloride and sodium hydrogen carbonate.

(b) Sodium hydrogen carbonate on reaction with hydrochloric acid gives sodium chloride, water and liberates carbon dioxide.

(c) Copper sulphate on treatment with potassium iodide precipitates cuprous iodide iodide (Cu_2I_2), liberates iodine gas and also forms potassium sulphate.



Watch Video Solution

26. A solution of potassium chloride when mixed with silver nitrate solution, an insoluble white substance is formed. Write the chemical reaction involved and also mention the type of the chemical reaction.

 [Watch Video Solution](#)

27. Ferrous sulphate decomposes with the evolution of a gas having a characteristic odour of burning sulphur. Write the chemical reaction involved and identify the type of reaction.

 [Watch Video Solution](#)

28. Why do fire flies glow at night?

 [Watch Video Solution](#)

29. Grapes hanging on the plant do not ferment but after being plucked from the plant can be fermented. Under what conditions do these grapes ferment? Is it a chemical or a physical change?

 [Watch Video Solution](#)

30. Which among the following are physical or chemical changes?

- (a) Evaporation of petrol
- (b) Burning of Liquefied Petroleum Gas (LPG)
- (c) Heating of an iron rod to red hot
- (d) Curdling of milk
- (e) Sublimation of solid ammonium chloride

 [Watch Video Solution](#)

31. During the reaction of some metals with dilute hydrochloric acid, following observations were made.

- (a) Silver metal does not show any change

(b) The temperature of the reaction mixture rises when aluminium (Al) is added

(c) The reaction of sodium metal is found to be highly explosive

(d) Some bubbles of a gas are seen when lead (Pb) is reacted with the acid

Explain these observations giving suitable reasons.



[Watch Video Solution](#)

32. A substance X, which is an oxide of a group 2 element, is used intensively in the cement industry. This element is present in bones also.

On treatment with water it forms a solution which turns red litmus blue.

Identify X and also write the chemical reactions involved.



[Watch Video Solution](#)

33. Write a balanced chemical equation for each of the following reactions and also classify them.

(a) Lead acetate solution is treated with dilute hydrochloric acid to form lead chloride and acetic acid solution.

(b) A piece of sodium metal is added to absolute ethanol to form sodium ethoxide and hydrogen gas.

(c) Iron (III) oxide on heating with carbon monoxide gas reacts to form solid iron and liberates carbon dioxide gas.

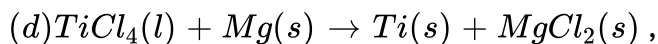
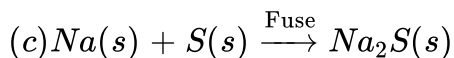
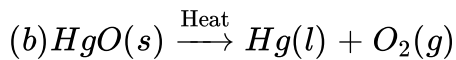
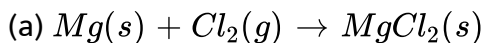
(d) Hydrogen sulphide gas reacts with oxygen gas to form solid sulphur and liquid water.

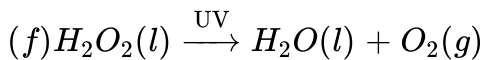
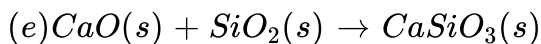
 [Watch Video Solution](#)

34. Why do we store silver chloride in dark coloured bottles?

 [Watch Video Solution](#)

35. Balance the following chemical equation and identify the type of chemical reaction.





 [Watch Video Solution](#)

36. A magnesium ribbon is burnt in oxygen to give a white compound X accompanied by emission of light. If the burning ribbon is now placed in an atmosphere of nitrogen, it continues to burn and forms a compound Y.

(a) Write the chemical formulae of X and Y.

(b) Write a balanced chemical equation, when X is dissolved in water.

 [Watch Video Solution](#)

37. Zinc liberates hydrogen gas when reacted with dilute hydrochloric acid, whereas copper does not. Explain, why?

 [Watch Video Solution](#)

38. A silver article generally turns black when kept in the open for a few days. The article when rubbed with toothpaste again starts shining.

(a) Why do silver articles turn black when kept in the open for a few days?

Name the phenomenon involved.

(b) Name the black substance formed and give its chemical formula.



[Watch Video Solution](#)

39. On heating blue coloured powder of copper (II) nitrate in a boiling tube, copper oxide (Black), oxygen gas and a brown gas X is formed

(a) Write a balanced chemical equation of the reaction.

(b) Identify the brown gas X evolved.

(c) Identify the type of reaction.

(d) What could be the pH range of aqueous solution of the gas X ?



[Watch Video Solution](#)

40. What happens when piece of

- (a) zinc metal is added to copper sulphate solution ?
- (b) aluminium metal is added to dilute hydrochloric acid ?
- (c) silver metal is added to copper sulphate solution ?

Also write the balanced chemical equation if the reaction occurs.



[Watch Video Solution](#)

41. What happens when zinc granules are treated with dilute solution of H_2SO_4 , HCl , HNO_3 , $NaCl$ and $NaOH$? Also write the chemical equations if reaction occurs.



[Watch Video Solution](#)

42. On adding a drop of barium chloride solution to an aqueous solution of sodium sulphite, white precipitate is obtained :

- (a) Write a balanced chemical equation for the reaction involved.
- (b) What other name can be given to this precipitation reaction

(c) On adding dilute hydrochloric acid to the reaction mixture, white precipitate disappears. Why ?

 [Watch Video Solution](#)

43. You are provided with two containers made up of copper and aluminium.

You are also provided with solutions of dilute HCl, dilute HNO_3 , $ZnCl_2$ and H_2O . In which of the containers these solutions can be kept?

 [Watch Video Solution](#)

Long Answer Type Questions

1. Give the characteristic tests for the following gases : (a) CO_2 (b) SO_2
(c) O_2 (d) H_2

 [Watch Video Solution](#)

