



CHEMISTRY

BOOKS - CHEMISTRY

P-BLOCK ELEMENTS

P Block Elements

1. On addition of conc. H_2SO_4 to a chloride salt, colourless fumes are evolved but in case of iodide salt, violet fumes come out. This is because

A. H_2SO_4 reduces HI to I_2

B. HI is of violet colour

C. HI gets oxidised to I_2

D. HI changes to HIO_3

Answer: C

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2. In qualitative analysis when H_2S is passed through an aqueous solution of salt acidified with dil. HCl , a black precipitate is obtained. On boiling the precipitate with dil. HNO_3 , it forms a solution of blue colour. Addition of excess of aqueous solution of ammonia to this solution gives

- A. deep blue precipitate of $Cu(OH)_2$
- B. deep blue solution of $[Cu(NH_3)_4]^{2+}$
- C. deep blue solution of $Cu(NO_3)_2$
- D. deep blue solution of $Cu(OH)_2 \cdot Cu(ON_3)_2$

Answer: B

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3. In a cyclotrimetaphosphoric acid molecule, how many single and double bonds are present ?

- A. 3 double bonds, 9 single bonds
- B. 6 double bonds, 6 single bonds
- C. 3 double bonds , 12 single bonds
- D. Zero double bond, 12 single bonds

Answer: C



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4. Which of the following elements can be involved in $p\pi - d\pi$ bonding?

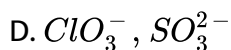
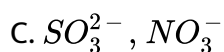
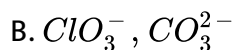
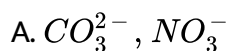
- A. Carbon
- B. Nitrogen
- C. phosphorus
- D. Boron

Answer: C



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5. Which of the following pairs of ions are isoelectronic and also isostructural ?



Answer: A



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6. Affinity for hydrogen decreases in the group from fluorine to iodine. Which of the halogen acids should have highest bond dissociation enthalpy ?



B. HCl

C. HBr

D. HI

Answer: A



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7. Bond dissociation enthalpy of E-H (E=element) bond is given below.

Compound	NH ₃	PH ₃	AsH ₃	SbH ₃
$\Delta_{\text{diss}}(E-H) / \text{kJ mol}^{-1}$	389	322	297	255

Which of the following compounds will act as strongest reducing agent?

A. NH_3

B. PH_3

C. AsH_3

D. SbH_3

Answer: D



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8. On heating with concentrated NaOH solution in an inert atmosphere of CO_2 , white phosphorus gives a gas. Which of the following statement is incorrect about the gas?

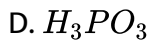
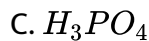
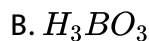
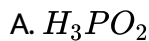
- A. It is highly poisonous and has smell like rotten fish
- B. It's solution in water decomposes in the presence of light
- C. It is more basic than NH_3
- D. It is less basic than NH_3

Answer: C



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9. Which of the following acids forms three series of salts?



Answer: C



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10. Strong reducing behaviour of H_3PO_2 is due to

A. low oxidation state of phosphorus

B. presence of two ___OH groups and one P___H bond

C. presence of one ___OH group and two P___H bonds

D. high electron gain enthalphy of phosphorus

Answer: C



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11. On heating lead nitrate forms oxides of nitrogen and lead. The oxides formed are :

A. N_2O , PbO

B. NO_2 , PbO

C. NO , PbO

D. NO , PbO_2

Answer: B



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12. Which of the following elements does not show allotropy ?

A. Nitrogen

B. Bismuth

C. Antimony

D. Arsenic

Answer: A



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13. The maximum covalency of nitrogen is

A. 3

B. 5

C. 4

D. 6

Answer: C



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14. Which of the following statements is wrong ?

- A. Single N_N bond is stronger than the single P_P bond.
- B. PH_3 can act as a ligand in the formation of coordination compound with transition elements.
- C. NO_2 is paramagnetic in nature.
- D. Covalency of nitrogen in N_2O_5 is four.

Answer: A



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15. A brown ring is formed in the ring test for NO_3^- ion. It is due to the formation of

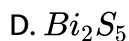
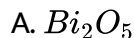
- A. $[Fe(H_2O)_5(NO)]^{2+}$
- B. $FeSO_4 \cdot NO_2$
- C. $[Fe(H_2O)_4(NO_2)]^{2+}$
- D. $FeSO_4 \cdot HNO_3$

Answer: A



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16. Elements of group 15 form compounds in +5 oxidation state. However, bismuth forms only one well characterised compound in +5 oxidation state. The compound is



Answer: B



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17. On heating ammonium dichromate and barium azide separately we get

- A. N_2 in both cases
- B. N_2 with ammonium dichromate and NO with barium azide
- C. N_2O with ammonium dichromate and N_2 with barium azide
- D. N_2O with ammonium dichromate and NO_2 with barium azide

Answer: A



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18. In the preparation of HNO_3 , we get NO gas by catalytic oxidation of ammonia. The moles of NO produced by the oxidation of two moles of NH_3 will be

- A. 2
- B. 3

C. 4

D. 6

Answer: A



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19. The oxidation state of central atom in the anion of compound NaH_2PO_2 will be

A. +3

B. +5

C. +1

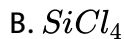
D. -3

Answer: C



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20. Which of the following is not tetrahedral in shape ?

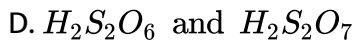
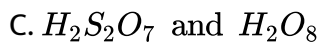
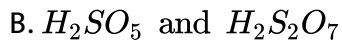
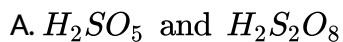


Answer: C



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21. Which of the following are peroxyacids of sulphur ?



Answer: A



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22. Hot conc. H_2SO_4 acts as moderately strong oxidising agent. It oxidises both metals and non-metals. Which of the following element is oxidised by conc. H_2SO_4 into two gaseous products. ?

A. Cu

B. S

C. C

D. Zn

Answer: C



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23. A black compound of manganese reacts with a halogen acid to give greenish yellow gas. When excess of this gas reacts with NH_3 an unstable trihalide is formed. In this process the oxidation state of nitrogen changes from

A. $-3 \rightarrow 0$

B. $-3 \rightarrow 0$

C. $-3 \rightarrow +5$

D. $0 \rightarrow -3$

Answer: A



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24. In the preparation of compounds of Xe, Bartlett has taken $\text{O}_2^+ \text{PtF}_6^-$ as a base compound. This is because

A. both O_2 and Xe have same size.

- B. both O_2 and Xe have same electron gain enthalpy.
- C. both O_2 and Xe have almost same ionisation enthalphy.
- D. both Xe and O_2 are gases.

Answer: C



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25. In solid state, PCl_5 is a.....

- A. convalent solid
- B. octahedral structure
- C. ionic solid with $[PCl_6]^+$ octahedral and $[PCl_4]^-$ tetrahedral
- D. ionic solid with $[PCl_4]^+$ tetrahedral and $[PCl_6]^-$ octahedral

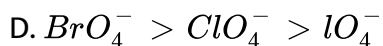
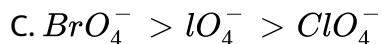
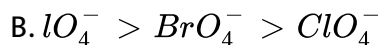
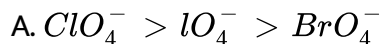
Answer: D



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26. Reduction potentials of some ions are given below. Arrange them in decreasing order of oxidising power.

Ion	ClO_4^-	IO_4^-	BrO_4^-
Reduction potential E° / V	$E^\circ = 1.19 \text{ V}$	$E^\circ = 1.65 \text{ V}$	$E^\circ = 1.74 \text{ V}$

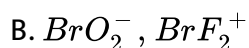
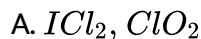


Answer: C



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27. Which of the following is isoelectronic pair ?



C. ClO_2 , BrF

D. CN^- , O_3

Answer: B



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28. If chlorine gas is passed through hot $NaOH$ solution, two changes are observed in the oxidation number of chlorine during the reaction .

These are And

A. 0 to +5

B. 0 to +3

C. 0 to -1

D. 0 to +3

Answer: A::C



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29. Which of the following options are not accordance with the property mentioned against them?

- (a) $F_2 > Cl_2 > Br_2 > I_2$ Oxodising power
- (b) $MI > MBr > MCl > MF$ Ionic character of metal halide
- (c) $F_2 > Cl_2 > Br_2 > I_2$ Bond dissociation enthalphy
- (d) $HI < HBr < HCl < HF$ Hydrogen - halogen bond strength



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30. Which of the following is correct for P_4 molecule of white phosphorus ?

- A. It has 6 lone pairs of electrons
- B. It has six P___P single bonds
- C. It has three P___P single bonds
- D. It has lone pair of electrosn

Answer: B::D



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31. Which of the following statements are correct?

- I. Among halogens, radius ratio between iodine and fluorine is maximum.
- II. Leaving F-F bond, all halogens have weaker X-X bond than X-X' bond in interhalogens.
- III. Among interhalogen compounds, maximum number of atoms are present in iodine fluoride.
- IV. Interhalogen compounds are more reactive than halogen compounds.

The correct option is



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32. Which of the following statements are correct for SO_2 gas?

- (a) It acts as bleaching agent in moist conditions.
- (b) Its molecule has linear geometry.

(c) Its dilute solution is used as disinfectant.

(d) It can be prepared by the reaction of dilute H_2SO_4 with metal sulphide.



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33. Which of the following statements are correct ?

(a) All the three N___O bond lengths in HNO_3 are equal.

(b) All P___Cl bond lengths in PCl_5 molecule in gaseous state are equal

(c) P_4 molecule in white phosphorus has angular strain therefore white phosphorus is very reactive

(d) PCl_5 is ionic in solid state in which cation is tetrahedral and anion is octahedral. And anion is octahedral.



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34. Which of the following orders are correct as per the properties mentioned against ?

(a) $As_2O_3 < SiO_2 < P_2O_3SO_2$ Acid strength.

(b) $AsH_3 < PH_3 < NH_3$ Enthalpy of vaporisation.

(c) $S < O < Cl < F$ More negative electron gain enthalpy.

(d) $H_2O > H_2S > H_2Se > H_2Te$ Thermal stability.



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35. Which of the following statements are correct?

(a) S___S bond is present in $H_2S_2O_6$

(b) In peroxosulphuric acid (H_2SO_5) sulphur is in + 6 oxidation state

(c) Iron powder along with Al_2O_3 and K_2O

is used as a catalyst in the preparation of NH_3

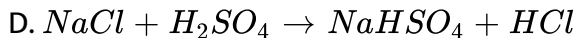
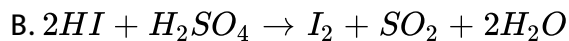
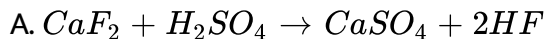
by Haber's process (d) ΔH is positive for the preparation of

SO_3 by catalytic oxidation of SO_2



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36. In which of the following reactions conc. H_2SO_4 is used as an oxidising reagent?



Answer: B::C



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37. Which of the following statements are true?

- (a) Only type of interactions between particles of noble gases are due to weak dispersion forces.
- (b) Ionisation enthalpy of molecular oxygen is very close that of xenon.
- (c) Hydrolysis of XeF_6 is a redox reaction.
- (d) Xenon fluorides are not reactive.



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38. In the preparation of H_2SO_4 by Contact process, why is SO_3 not absorbed directly in water to form H_2SO_4 ?



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39. Write a balanced chemical equation for the reaction showing catalytic oxidation of NH_3 by atmospheric oxygen.



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40. Write the structure of pyrophosphoric acid.



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41. PH_3 forms bubbles when passed slowly in water but NH_3 dissolves. Explain why ? In PCl_5 Phosphorus is in sp^3d hybridised state but all its five bonds are not equivalent. Justify your answer with reason.



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42. In PCl_5 phosphorus is in $sp^3 d$ hybridised state but all its five bonds are not equivalent. Justify your answer with reason.



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43. Why is nitric oxide paramagnetic in gaseous state but the solid obtained on cooling it is diamagnetic ?



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44. Give one reason to explain why ClF_3 exists but FCl_3 does not exist?



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45. Out of H_2O which one has higher bond angle and why?



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46. SF_6 is known but SC_6 is not. Why?

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47. On reaction with Cl_2 phosphorus forms two types of halides 'A' and 'B'.

Halide 'A' is yellowish- white powder but halide 'B' is colourless oily liquid.

Identify A and B and write the formulae of their hydrolysis products.

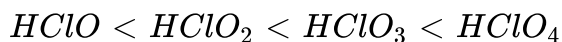
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48. In the ring test of NO_3^- ion, Fe^{2+} ion reduces nitrate ion to nitric oxide, which combines with Fe^{2+} (aq) ion to form brown complex .

Write the reactions involved in the formation of brown ring.

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49. Explain why the stability of oxoacids of chlorine increases in the order given below.



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50. Explain why ozone is thermodynamically less stable than oxygen ?



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51. P_4O_6 reacts with water according to equation $\text{P}_4\text{O}_6 \rightarrow 4\text{H}_3\text{PO}_3$.

Calculate the volume of 0.1MNaOH solution required to neutralise the acid formed by dissolving 1.1g of P_4O_6 in H_2O .



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52. White phosphorus reacts with chlorine and the product hydrolyses in the presence of water. Calculate the mass of HCl obtained by the hydrolysis of the product formed by the reaction of 62 g of white phosphorus with chlorine in the presence of water.



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53. Name three oxoacids of nitrogen. Write the disproportionation reaction of that oxoacid of nitrogen is in +3 oxidation state.



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54. Nitric acid forms an oxide of nitrogen on reaction with P_4O_{10} . Write the reaction involved. Also write the resonating structures of the oxide of nitrogen formed.



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55. (i) White phosphorus (ii) red phosphorus and (iii) black phosphorus.

Write the difference between white red and black phosphorus on the basis of their structure and reactivity.

Phosphorus has three allotropic forms __



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56. Given an example to show the effect of concentration of nitric acid on the formation of oxidation product.



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57. PCl_5 reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH_3 solution. Write the reactions involved to explain what happens.



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58. Phosphorus forms a number of oxoacids. Out of these oxoacids, phosphinic acid has strong reducing property. Write its structure and also write a reaction its reducing behaviour.



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59. Match the compounds given in Column I with the hybridisation and shape given in Column II and mark the correct option.



A.

<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
1	3	4	2

B.

<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
1	2	4	3

C.

<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
4	3	1	2

D.

<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
4	1	2	3

Answer:



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60. Match the formulas of oxides given in Column I with the type of oxide given in Column II and mark the correct option.

Column I	Column II
A. Pb_3O_4	1. Neutral oxide
B. N_2O	2. Acidic oxide
C. Mn_2O_7	3. Basic oxide
D. Bi_2O_3	4. Mixed oxide

A. $\begin{matrix} A & B & C & D \\ 1 & 2 & 3 & 4 \end{matrix}$

B. $\begin{matrix} A & B & C & D \\ 4 & 1 & 2 & 3 \end{matrix}$

C. $\begin{matrix} A & B & C & D \\ 3 & 2 & 4 & 1 \end{matrix}$

D. $\begin{matrix} A & B & C & D \\ 4 & 3 & 1 & 2 \end{matrix}$

Answer:



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61. Match the items of Columns I and II and mark the correct option.

Column I	Column II
A. H_2SO_4	1. Highest electron gain enthalpy
B. CCl_3NO_2	2. Chalcogen
C. Cl_2	3. Tear gas
D. Sulphur	4. Storage batteries

A. $\begin{matrix} A & B & C & D \\ 4 & 3 & 1 & 2 \end{matrix}$

B. $\begin{matrix} A & B & C & D \\ 3 & 4 & 1 & 2 \end{matrix}$

C. $\begin{matrix} A & B & C & D \\ 4 & 1 & 4 & 3 \end{matrix}$

D. $\begin{matrix} A & B & C & D \\ 2 & 1 & 3 & 4 \end{matrix}$

Answer:



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62. Match the species given in Column I with the shape given in Column II and mark the correct option.

Column I	Column II
A. SF_4	1. Tetrahedral
B. BrF_3	2. Pyramidal
C. BrO_3^-	3. Sea-saw shaped
D. NH_4^+	4. Bent T-shaped

A. $\begin{matrix} A & B & C & D \\ 3 & 2 & 1 & 4 \end{matrix}$

B. $\begin{matrix} A & B & C & D \\ 3 & 4 & 2 & 1 \end{matrix}$

C. $\begin{matrix} A & B & C & D \\ 1 & 2 & 3 & 4 \end{matrix}$

D. $\begin{matrix} A & B & C & D \\ 1 & 4 & 3 & 2 \end{matrix}$

Answer:



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63. Match the items of Columns I and II and mark the correct option.

Column I	Column II
A. Its partial hydrolysis does not change oxidation state of central atom.	1. He
B. It is used in modern diving apparatus.	2. XeF_6
C. It is used to provide inert atmosphere for filling electrical bulbs.	3. XeF_4
D. Its central atom is in sp^3d^2 hybridisation.	4. Ar

- A.

	<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
1	4	2	3	
- B.

	<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
1	2	3	4	
- C.

	<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
2	1	4	3	
- D.

	<i>A</i>	<i>B</i>	<i>C</i>	<i>D</i>
1	3	2	4	

Answer:



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64. Assertion (A) N_2 is less reactive than P_4 .

Reason (R) Nitrogen has more electron gain enthalpy then phosphorus.

- A. Both Assertion and Reason are correct statements , and Reason is the correct explation of the Assertion.
- B. Both Assertion and Reason are correct statements , and Reason is not the correct explanation of the Assertion.
- C. Assertion is correct, but Reason is wrong statement.

D. Assertion is wrong but Reason is correct statement.

Answer:



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65. Assertion (A) HNO_3 makes iron passive.

Reason (R) HNO_3 forms a protective layer of ferric nitrate on the surface of iron.

A. Both Assertion and Reason are correct statements , and Reason is the correct explanation of the Assertion.

B. Both Assertion and Reason are correct statements , and Reason is not the correct explanation of the Assertion.

C. Assertion is correct, but Reason is wrong statement.

D. Assertion is wrong but Reason is correct statement.

Answer:



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66. Assertion (A) HI cannot be prepared by the reaction of KI with concentrated H_2SO_4 .

Reason (R) HI has lowest H___X bond strenth among halogen acids.

- A. Both Assertion and Reason are correct statements , and Reason is the correct explanation of the Assertion.
- B. Both Assertion and Reason are correct statements , and Reason is not the correct explanation of the Assertion.
- C. Assertion is correct, but Reason is wrong statement.
- D. Assertion is wrong but Reason is correct statement.

Answer:



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67. Assertion (A) Both rhombic and monoclinic sulphur exist as S_8 but oxygen exists as O_2 .

Reason (R) Oxygen forms $p\pi - p\pi$ multiple bond due to small size and small bond length but $p\pi - p\pi$ bonding is not possible in sulphur.

- A. Both Assertion and Reason are correct statements , and Reason is the correct explanation of the Assertion.
- B. Both Assertion and Reason are correct statements , and Reason is not the correct explanation of the Assertion.
- C. Assertion is correct, but Reason is wrong statement.
- D. Assertion is wrong but Reason is correct statement.

Answer:



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68. Assertion (A) $NaCl$ reacts with concentrated H_2SO_4 to give colourless fumes with pungent smell. But on adding MnO_2 the fumes become greenish yellow.

Reason (R) MnO_2 oxidises HCl to chlorine gas which is greenish yellow.

A. Both Assertion and Reason are correct statements , and Reason is the correct explanation of the Assertion.

B. Both Assertion and Reason are correct statements , and Reason is not the correct explanation of the Assertion.

C. Assertion is correct, but Reason is wrong statement.

D. Assertion is wrong but Reason is correct statement.

Answer:



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69. Assertion:- SF_6 cannot be hydrolysed but SF_4 can be.

Reason:- Six F atoms in SF_6 prevent the attack of H_2O on sulphur atom of SF_6

- A. Both Assertion and Reason are correct statements , and Reason is the correct explanation of the Assertion.
- B. Both Assertion and Reason are correct statements , and Reason is not the correct explanation of the Assertion.
- C. Assertion is correct, but Reason is wrong statement.
- D. Assertion is wrong but Reason is correct statement.

Answer:



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70. An amorphous solid A burns in air to form a gas B which turns lime water milky. The gas is also produced as a by - product during roasting of

sulphide ore. This gas decolourises acidified aqueous $KMnO_4$ solution and reduces Fe^{3+} to Fe^{2+} . Identify the solid A and the gas B and write the reactions involved.



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71. On heating lead (II) nitrate gives a brown gas "A". The gas "A" on cooling changes to colourless solid "B". Solid B on heating with NO changes to a blue solid 'C'. Identify 'A', 'B' and 'C' and also write reactions involved and draw the structures of 'B' and 'C'.



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72. On heating compound (A) gives a gas (B) which is a constituent of air. This gas when treated with 3 moles of hydrogen (H_2) in the presence of a catalyst gives another gas (C) which is basic in nature. Gas C on further oxidation in moist condition gives a compound (D) which is a part of acid rain. Identify compounds (A) to (D) and also give necessary equations of all the steps involved.



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