



CHEMISTRY

BOOKS - S CHAND CHEMISTRY (HINGLISH)

ACID, BASES AND SALTS

Solved Examples

1. You have been provided with three test tubes. One of them contains distilled water and the other two contain an acidic solution and a basic solution, respectively. If you are given only red litmus paper, how will you identify the contents of each test tube?

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2. Metal compound A reacts with dilute hydrochloric acid to produce effervescence. The gas evolved extinguishes a burning candle. Write a balanced chemical equation for the reaction if one of the compounds formed is calcium chloride.

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3. A solution reacts with crushed egg-shells to give a gas that turns lime-water milky. The solution contains

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4. Equal lengths of magnesium ribbons are taken in test tubes A and B. Hydrochloric acid (HCl) is added to test tube A, while acetic acid (CH_3COOH) is added to test tube B. Amount and concentration taken for both the acids are same. In which test tube will the fizzing occur more vigorously and why?

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5. You have two solutions, A and B. The pH of solution A is 6 and pH of solution B is 8. Which solution has more hydrogen ion concentration? Which of this is acidic and which one is basic?

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6. Five solutions A,B,C,D and E when tested with universal indicator showed pH as 4,1,11,7 and 9, respectively. Which solution is

- (a) neutral?
- (b) strongly alkaline?
- (c) strongly acidic?
- (d) weakly acidic?
- (e) weakly alkaline?

Arrange the pH in increasing order of hydrogen-ion concentration.

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1. What colour do the following indicators turn when added to a base or alkali (such as sodium hydroxide)?

(a). Methyl orange

(b). Litmus

(c). Red cabbage extract.



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2. What colours do the following indicators turn when added to an acid (such as hydrochloric acid)?

(a). Litmus

(b). Methyl orange



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3. Name an indicator which is red in acid solution but turns blue in basic solution.

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4. Name an indicator which is pink in alkaline solution but turns colourless in acidic solution.

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5. When a solution is added to a cloth strip treated with onion extract, then the smell of onion cannot be detected. State whether the given solution contains an acid or a base.

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6. When a solution is added to vanilla extract, then the characteristic smell of vanilla cannot be detected. State whether the given solution is an acid or a base.

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7. Which gas is usually liberated when an acid reacts with a metal? Illustrate with an example. How will you test for the presence of this gas?

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8. Name the gas evolved when dilute HCl reacts with sodium hydrogencarbonate. How is it recognised?

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9. Give the names and formulae of two strong acids and two weak acids.



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10. Name one natural source of each of the following acids.

(a). Citric acid

(b). Oxalic acid

(c). Lactic acid

(d). Tataric acid.



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11. Name one animal and one plant whose stings contain formic acid (or methanoic acid).



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12. How is the concentration of hydronium ions (H_3O^+) affected when a solution of an acid is diluted?



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13. Write word equations and balanced equations for the reactions taking place when:

- (a). Dilute sulphuric acid reacts with zinc granules.
- (b). Dilute hydrochloric acid reacts with magnesium ribbon.
- (c). Dilute sulphuric acid reacts with aluminium powder.
- (d). dilute hydrochloric acid reacts with iron filings.



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14. Complete and balance the following chemical equations

- (a). $Zn(s) + HCl(aq) \rightarrow$
- (b). $Na_2CO_3(s) + HCl(aq) \rightarrow$
- (c). $NaHCO_3(s) + HCl(aq) \rightarrow$
- (d). $NaOH(aq) + HCl(aq) \rightarrow$
- (e). $CuO(s) + HCl(aq) \rightarrow .$



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15. Fill in the blanks in the following sentences:

- (a). Acids have a _____ taste and they turn _____ litmus to _____
- (b). Substances do not show their acidic properties without _____
- (c). Acids produce _____ ions on dissolving in water.
- (d). Those substances whose smell (or odour) changes in acidic or basic solution are called _____ indicators.
- (e). Onion and vanilla extract are _____ indicators.



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16. (a). What is an indicator? Name three common indicators.

(b). Name the acid-base indicator extracted from lichen.

(c). What colour does the turmeric paper turn when put in an alkaline solution.?



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17. What is an olfactory indicator? Name two olfactory indicators. What is the effect of adding sodium hydroxide solution to these olfactory indicators?

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18. Which gas is usually liberated when an acid reacts with a metal? Illustrate with an example. How will you test for the presence of this gas?

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19. While diluting an acid, why is it recommended that the acid should be added to water and not water to the acid?

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20. What happens when an acid reacts with a metal hydrogencarbonate?

Write equation of the reaction which takes place.

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21. (a). What happens when dilute hydrochloric acid is added to sodium carbonate? Write a balanced chemical equation of the reaction involved.

(b). Which gas is liberated when dilute hydrochloric acid reacts with sodium carbonate? How will you test for the presence of this gas?

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22. What happens when an acid reacts with a base? Explain by taking the example of hydrochloric acid and sodium hydroxide. Give equation of the chemical reaction which takes place. What is the special name of such a reaction?

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23. What happens when an acid reacts with a metal oxide? Explain with the help of an example. Write a balanced equation for the reaction involved.

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24. (a). What are organic acids and mineral acids?

(b). Give two examples each of organic acids and mineral acids.

(c). State some of the uses of mineral acids in industry.

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25. What is meant by strong acids and weak acids? Classify the following into strong acids and weak acids:

HCl , HC_3COOH , H_2SO_4 , HCO_3 , H_2CO_3 , H_2SO_3 .

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26. Why do HCl , HNO_3 , etc., show acidic characters in aqueous solutions while solutions of compounds like alcohol and glucose do not show acidic character?

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27. What happens when an acid reacts with a base? Explain by taking the example of hydrochloric acid and sodium hydroxide. Give equation of the chemical reaction which takes place. What is the special name of such a reaction?

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28. Why should curd and sour substances not be kept in brass and copper vessels?

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29. (a). What is produced if an acid is added to a base?

(b). Why does dry HCl gas not change the colour of dry litmus paper?

(c). What colour does phenolphthalein indicator turn when added to an alkali (such as sodium hydroxide)?

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30. (a). Why do acids not show acidic behaviour in the absence of water?

(b). Why does an aqueous solution of an acid conduct electricity?

(c). Why does distilled water not conduct electricity whereas rain water does?

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31. (a). What happens when an acid reacts with a metal carbonate?

Explain with the help of an example. Write chemical equation of the reaction involved.

(b). What happens when carbon dioxide gas is passed through lime

water:

(i). For a short time?

(ii). for a considerable time?

Write equations of the reactions involved.

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32. Compounds such as alcohols and glucose also contain hydrogen but are not categorised as acids. Describe an Activity to prove it.

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33. 10 mL of a solution of NaOH is found to be completely neutralised by 8 mL of a given solution of HCl. If we take 20 mL of the same solution of NaOH, the amount HCl solution (the same solution as before) required to neutralise it will be

A. 4 mL

B. 8 mL

C. 12 mL

D. 16 mL

Answer: D



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34. Which one of the following types of medicines is used for treating indigestion?

A. antibiotic

B. analgesic

C. antacid

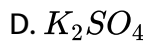
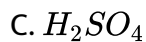
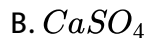
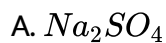
D. antiseptic

Answer: C



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35. A solution reacts with marble chips to produce a gas which turns lime water milky. The solution contains:



Answer: C



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36. One of the following is not an organic acid. This is :

A. ethanoic acid

B. formic acid

C. citric acid

D. carbonic acid

Answer: D



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37. The property which is not shown by acids is:

- A. they have sour taste
- B. they feel soapy
- C. they turn blue litmus red
- D. their pH is less than seven

Answer: B



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38. The indicators which turn red in acid solution are:

- A. turmeric and red litmus

B. phenolphthalein and methyl orange

C. blue litmus and methyl orange

D. phenolphthalein and red litmus

Answer: C



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39. The discomfort caused by indigestion due to overeating can be cured by taking:

A. vinegar

B. lemon juice

C. baking soda

D. caustic soda

Answer: C



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40. The property which is common between vinegar and curd is that they

:

- A. have sweet taste
- B. have bitter taste
- C. are tasteless
- D. have sour taste

Answer: D



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41. The indicator which produces a pink colour in an alkaline solution is:

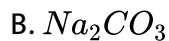
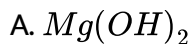
- A. methyl orange
- B. turmeric
- C. litmus paper

D. phenolphthalein

Answer: D

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42. A solution reacts with zinc granules to give a gas which burns with a 'pop' sound. The solution contains:



Answer: D

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43. When a piece of limestone reacts with dilute HCl , a gas X is produced. When gas X is passed through lime water then a white precipitate Y is formed. On passing excess of gas X, the white precipitate dissolves forming a soluble compound Z.

(a). What are X, Y and Z?

(b). Write equations for the reactions which take place:

(i). When limestone reacts with dilute HCl

(ii). When gas X reacts with lime water to form white precipitate Y

(iii). When excess of gas X dissolves white precipitate Y to form a soluble compound Z.

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44. If some one is suffering from acidity in stomach, which of the following would you suggest as remedy. Orange juice, Coka Cola, Baking soda solution ?

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45. On adding dilute hydrochloric acid to copper oxide powder, the solution formed is blue-green.

(a). Predict the new compound formed which imparts a blue-green colour to solution.

(b). Write a balanced chemical equation of the reaction which takes place.

(c). On the basis of the above reaction, what can you say about the nature of copper oxide?



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46. A white shirt has a yellow stain of curry. When soap is rubbed on this shirt during washing, the yellow stain turns reddish-brown. On rinsing the shirt with plenty of water, the reddish-brown stain turns yellow again.

(a). Name the natural indicator present in curry stain.

(b). Explain the changes in colour of this indicator which take place during washing and rinsing the shirt.

(c). What is the nature of soap (acidic/basic) as shown by the indicator present in curry stain?

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47. You have been provided with three test tubes. One of them contains distilled water and the other two contain an acidic solution and a basic solution, respectively. If you are given only red litmus paper, how will you identify the contents of each test tube?

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48. A substance X which is used as an antacid reacts with dilute hydrochloric acid to produce a gas Y which is used in one type of fire-extinguisher. Name the substance X and gas Y. Write a balanced equation for the chemical reaction which takes place.

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49. How is the neutralisation of a carbonate with an acid different from the neutralisation of an oxide or a hydroxide?



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50. What happens to (a) the H^+ ions, and (b) temperature of the solution, when an acid is neutralised?



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51. Name the gas evolved when zinc granules are treated/heated with:

(a). Hydrochloric acid solution

(b). Sodium hydroxide solution



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52. What is the common name of water soluble bases?



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53. What is common in all the water soluble bases (or alkalis)?

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54. Why does tooth decay start when the pH of mouth is lower than 5.5?

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55. What is the pH of a neutral solution?

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56. Which is more acidic: a solution of pH=2 or a solution of pH=6?

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57. Which is more basic (or more alkaline) : a solution of H^+ concentration of 10^{-8} or a solution of $\text{pH}=11$?

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58. Name the scientist who developed the pH scale.

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59. Name the indicator which can give us an idea of how strong or weak an acid or base is.

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60. The pH of soil A is 7.5 while that of soil B is 4.5. Which of the two soils, A or B, should be treated powdered chalk to adjust its pH and why?

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61. What is the name of the indicator which can be used for testing the pH of a solution?

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62. What colour will universal indicator show if you add it to the following substances?

(a). Potassium hydroxide pH=12

(b). Soda water, pH=5

(c). Sulphuric acid, pH=2.

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63. A beaker of concentrated hydrochloric acid has a pH of 1. What colour will full range universal indicator turn if it is added to this beaker? Is it a strong or a weak acid?

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64. Two solutions X and Y are tested with universal indicator. Solution X turns orange whereas solution Y turns red. Which of the solutions is a stronger acid?

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65. Two solutions A and B have pH values of 3.0 and 9.5 respectively. Which of these will turn litmus solution from blue to red and which will turn phenolphthalein from colourless to pink?

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66. Two drinks P and Q gave acidic and alkaline reactions, respectively one has a pH value of 9?

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67. Two solutions X and Y have $\text{pH}=4$ and $\text{pH}=8$, respectively. Which solution will give alkaline reaction and which one acidic?

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68. Fill in the following blanks with suitable words

(a). Acids have a pH ___ than 7.

(b). Alkalis have a pH ___ than 7.

(c). Neutral substances have a pH of ____

(d) The more acidic a solution, the ___ the pH .

(e). The more alkaline a solution, the _____ the pH .

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69. Fresh milk has a pH of 6. How do you think the pH will change as it turns into curd? Explain your answer.

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70. (a). What is a universal indicator? For what purpose is it used?

(b). How does a universal indicator work?

(c). Water is a neutral substance. What colour will you get when you add a few drops of universal indicator to a test-tube containing water?

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71. Which chemical is injected into the skin of a person.

(a). During an ant's sting?

(b). During the nettle leaf hair sting?

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72. (a). Explain the pH change as the cause of tooth decay. How can tooth decay caused by pH change be prevented?

(b). Explain how pH change in the lake water can endanger the lives of aquatic animals (like fish). What can be done to lessen the danger to the lives of aquatic animals in the lake?



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73. (a). What happens during a bee sting? What is its remedy?

(b). What happens during a wasp sting? What is its remedy?



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74. (a). Why is it wrong to treat a bee sting with vinegar?

(b). Why is it wrong to treat a wasp sting with baking soda solution?



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75. (a). What does the pH of a solution signify? Three solutions A, B and C have pH values of 6, 4 and 10 respectively. Which of the solutions is highly acidic?

(b). A farmer has found that the pH of soil in his fields is 4.2. Name any two chemical materials which he can mix with the soil to adjust its pH.



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76. (a). The pH values of six solutions A to F are given below:

A=0, B=11, C=6, D=3, E=13, F=8

Which of the above solutions are (i) acids (ii) alkalis?

(b). Name the acids or alkalis used to make (i) can batteries (ii) explosives (iii) soaps (iv) fertilisers.

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77. (a). The pH of a cold drink is 5. What will be its action on blue and red litmus solutions?

(b). The pH values of three acids A, B and C having equal molar concentrations are 5.0, 2.8 and 3.5 respectively. Arrange these acids in order of the increasing acid strengths.

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78. Under what soil conditions do you think a farmer would treat the soil of his fields with quicklime (calcium oxide), or slaked lime (calcium hydroxide) or chalk (calcium carbonate)?

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79. Which acid is produced in our stomach ? What happens if there is an excess of acid in the stomach? How can its effect be cured?

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80. The soil in a field is highly acidic. Name two materials which can be added to this soil to reduce its acidity. Give the reason for your choice.

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81. What is meant by strong bases and weak bases? Classify the following into strong bass and weak bases:

NH_4OH , $Ca(OH)_2$, $NaOH$, KOH , $Mg(OH)_2$.

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82. What ions are present in the solutions of following substances? (write the symbols only)

(i) Hydrochloric acid

(ii). Nitric acid

(iii) Sulphuric acid

(iv). Sodium hydroxide

(v). Potassium hydroxide

(vi). Magnesium hydroxide.

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- 83.** (a). What would you expect the pH of pure water to be?
- (b). What colour would the universal indicator show in an aqueous solution of sugar? Why?
- (c). A sample of rain water turned universal indicator paper yellow. What would you expect its pH to be? Is it a strong or a weak acid?

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- 84.** (a). What do you think will be the pH in the stomach of a person suffering from indigestion: Less than 7 or more than 7?
- (b). What do you think will be the pH of an antacid solution: less than 7 or more than?
- (c). How does an antacid work?
- (d). Name two common antacids.

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85. Separate the following into substances having pH values above and below 7. How do these influence litmus paper?

(i) Lemon juice

(ii) Solution of washing soda

(iii) toothpaste

(iv). Vinegar

(v). Stomach juices.



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86. (a). Do basic solutions also have H^+ (aq) ions? If yes, then why are they basic?

(b). When a solution becomes more acidic, does the pH get higher or lower?



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87. (a). What happens when zinc granules are heated with sodium hydroxide solution? Write equation of the reaction which takes place.

(b). What happens when bases react with non-metal oxides? Explain with the help of an example. What does this reaction tell us about the nature of non-metal oxides?



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88. (a). What effect does the concentration of H^+ (aq) ions have on the nature of a solution?

(b). What effect does the concentration of OH^- ions have on the nature of a solution?

(c). Someone put some universal indicator paper into vinegar. The pH is 3. What does this tell you about the vinegar?

(d). Someone put some universal indicator paper onto wet soap. the pH is 8. What does this tell you about the soap?

(e). State whether a solution is acidic, alkali or neutral if its pH is:

(i) 9

(ii) 4

(iii) 7

(iv) 1

(v) 10

(vi) 3.



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89. One of the following is a medicine for indigestion. This is:

A. sodium hydroxide

B. manganese hydroxide

C. magnesium hydroxide

D. potassium hydroxide

Answer: C



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90. Bee sting contains:

- A. an acidic liquid
- B. a salt solution
- C. an alkaline liquid
- D. an alcohol

Answer: A



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91. Wasp sting contains:

- A. a sugar solution
- B. an acidic liquid
- C. a salt solution
- D. an alkaline liquid

Answer: D



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92. One of the following does not inject an acidic liquid into the skin through its sting. This is:

A. honey bee

B. ant

C. wasp

D. nettle leaf hair

Answer: C



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93. A solution turns red litmus blue, its pH is likely to be

A. 1

B. 4

C. 5

D. 10

Answer: D

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94. A solution turns blue litmus red. Its pH is likely to be:

A. 7

B. 5

C. 8

D. 14

Answer: B

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95. A solution turns phenolphthalein indicator pink. The most likely pH of this solution will be:

A. 6

B. 4

C. 9

D. 7

Answer: C



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96. The colour of methyl orange indicator in a solution is yellow. The pH of this solution is likely to be:

A. 7

B. less than 7

C. 0

D. more than 7

Answer: D



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97. Bee stings can be treated with:

A. vinegar

B. sodium hydrogencarbonate

C. potassium hydroxide

D. lemon juice

Answer: B



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98. Wasp stings can be treated with :

A. baking soda

B. vinegar

C. washing soda

D. milk of magnesia

Answer: B



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99. it has been found that rubbing vinegar on the stung area of the skin of a person given him relief. Ther person has been stung by:

A. wasp

B. ant

C. honey bee

D. nettle leaf hair

Answer: A



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100. Fresh milk has a pH of 6. When milk changes into curd, the pH value will:

- A. become 7
- B. become less than 6
- C. become more than 7
- D. remain unchanged.

Answer: B



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101. Which acid is produced in our stomach ? What happens if there is an excess of acid in the stomach? How can its effect be cured?

- A. acetic acid
- B. citric acid
- C. hydrochloric acid
- D. sulphuric acid

Answer: C

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102. The daffodil plants grow best in a soil having a pH range of 6.0 to 6.5. If the soil in a garden has a pH of 4.5, which substance needs to be added to the soil in order to grow daffodils?

- A. salt
- B. lime
- C. sand
- D. compost

Answer: B

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103. A milkman adds a very small amount of baking soda to fresh milk.

(a) Why does he shift the pH of the fresh milk from 6 to slightly alkaline?

(b) Why does this milk take a long time to set as curd?

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104. Which of the following elements would form oxides which would indicate pH values less than seven, using moist pH paper ?

Magnesium, carbon, sulphur, Hydrogen, copper.

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105. The pH values of five solutions, A, B, C, D and E are given below:

A 1

B 5

C 7

D 11

E 13 Itbr. Which solution is (i) weakly alkaline (ii) neutral (iii) strongly acidic (iv) strongly alkaline, and (v) weakly acidic?



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106. Potatoes grow well on Anhad's farm which has soil with a pH of 5.5. Anhad decided to add lot of lime to soil so that he can grow broccoli in the same farm:

- (a). Do potatoes grow better in acidic or alkaline soil?
- (b). Does broccoli grow better in acidic or alkaline soil?



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107. Here are some results of solutions tested with universal indicator paper:

Sulphuric acid: Red

Metal polish : Dark blue

Washing-up liquid : Yellow

Milk of magnesia : Light blue

Oven cleaner : Purple

Car battery acid : Pink

Arrange the solutions in order of their increasing pH values (starting with the one with the lowest pH).



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108. Solution A turns universal indicator blue to purple whereas solution B turns universal indicator orange to red.

- (a). What will be the action of solution A on litmus?
- (b). What will be action of solution B on litmus?
- (c) Name any two substances which can give solutions like A.
- (d). Name any two substances which can give solutions like B.
- (e). What sort of reaction takes place when solution A reacts with solution B?



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109. A first-aid manual suggests that vinegar should be used to treat wasp stings and baking soda for bee stings.

What does this information tell you about the chemical nature of:

(a). Wasp stings?

(b). Bee stings?



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110. Hydrochloric acid reacts with a metal X to form a gas Y which burns with a pop sound. Sodium hydroxide solution also reacts with the same metal X (on heating) to form the same gas Y.

(a) Name X and Y

(b). Write the chemical equation of the reaction of metal X with (i) hydrochloric acid, and (ii) sodium hydroxide solution.



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111. What is the chemical formula of (a) baking soda, and (b) washing soda?

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112. Write the chemical formula of (i) soda ash, and (ii) sodium carbonate decahydrate.

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113. State whether the following statement is true or false:

Copper sulphate crystals are always wet due to the presence of water of crystallisation in them.

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114. Which of the following salt has a blue colour and why?

$CuSO_4 \cdot 5H_2O$ or $CuSO_4$

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115. What would be the colour of litmus in a solution of sodium carbonate?

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116. State the common and chemical names of the compound formed when plaster of paris is mixed with water.

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117. With which substance should chlorine be treated to get bleaching powder?

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118. What is the commercial name of calcium sulphate hemihydrate?

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119. Name the product formed when Cl_2 and H_2 produced during the electrolysis of brine are made to combine.

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120. Name a calcium compound which hardens on wetting with water.

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121. Name a sodium compound which is a constituent of many dry soap powders.

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122. Name a metal carbonate which is soluble in water.

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123. Name an acid which is present in baking powder.

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124. Name the metal whose carbonate is known as washing soda.

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125. Which compound is used as an actacid in medicine: $NaHCO_3$ or Na_2CO_3 ?

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126. What is the common name of (a) $NaHCO_3$ and (b) $Na_2CO_3 \cdot 10H_2O$?

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127. Write the chemical name and formula of (a) common salt, and (b) caustic soda.

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128. What are the two main ways in which common salt (sodium chloride) occurs in nature?

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129. Name the major salt present in sea-water?

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130. (a) Name the process by which common salt is obtained from seawater

(b) Name the process by which common salt is purified.

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131. Why is sodium chloride required in our body?

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132. Name three chemicals made from common salt (or sodium chloride).

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133. Give any two uses of common salt (sodium chloride).

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134. What name is given to the common salt which is mined from underground deposits? How was this salt formed?

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135. Name the salt which is used as a preservative in pickles, and in curing meat and fish.

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136. Name the raw material used for the production of caustic soda.

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137. The electrolysis of an aqueous solution of sodium chloride gives us three products. Name them.

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138. During the electrolysis of a saturated solution of sodium chloride, where is:

- (a). Chlorine formed?
- (b) Hydrogen formed?
- (c) Sodium hydroxide formed?

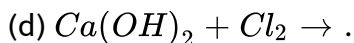
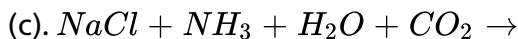
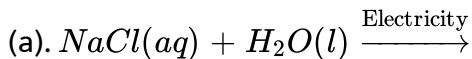
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139. Fill in the following blanks:

- (a) Common salt is obtained from sea-water by the process of _____
- (b) Rock salt is mined just like _____
- (c) Chemical formula of washing soda is _____
- (d) Sodium hydrogencarbonate is _____ soda whereas sodium carbonate is _____ soda.
- (e). The chemical formula of plaster of Paris is _____

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140. Complete and balance the following chemical equations:



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141. What is washing soda? State two properties and two uses of washing soda.



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142. Write the formulae of sodium chloride and sodium carbonate. Explain why an aqueous solution of sodium chloride is neutral but an aqueous solution of sodium carbonate is basic (or alkaline). Write chemical equations of the reactions involved.



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143. Write the chemical formula of ammonium chloride. Explain why an aqueous solution of ammonium chloride is acidic in nature? Illustrate your answer with the help of a chemical equation.

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144. What is baking soda? Write the chemical name of baking soda. Give the important uses of baking soda. How does baking soda differ chemically from washing soda?

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145. Describe how sodium hydrogencarbonate (baking soda) is produced on a large scale. Write equation of the reaction involved.

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146. What happens when a cold and concentrated solution of sodium chloride reacts with ammonia and carbon dioxide? Write the chemical equation of the reaction which takes place.

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147. (a). What is the common name of sodium hydrogencarbonate? It Brgt
(b). What happens when a solution of sodium hydrogencarbonate is heated? Write equation of the reaction involved.
(c) Explain why, sodium hydrogencarbonate is used as an antacid.

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148. (a). What will happen if heating is not controlled while preparing plaster of paris?
(b) Write an equation to show the reaction between plaster of paris and water.

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149. (a) What happens when copper sulphate crystals are heated strongly? Explain with the help of an equation.

(b) What happens when a few drops of water are added to anhydrous copper sulphate? Explain with the help of an equation.

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150. (a). What is the chemical name of bleaching powder.?

(b) What is the chemical formula of bleaching powder?

(c) What are the materials used for the preparation of bleaching powder?

(d) State one use of bleaching powder (other than bleaching).

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151. (a) Name a sodium compound used for softening hard water.

(b) Which compound of calcium is used for disinfecting drinking water

supply?

(c) Name a metal compound which has detergent properties (cleansing properties).

(d) Name one compound of calcium which is used for removing the colour of a coloured cloth.

(e) State a peculiar (or remarkable) property of plaster of paris.

(f) Name the substance obtained by the action of chlorine on solid (dry) slaked lime.



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152. (a). What is gypsum? What happens when gypsum is heated to $100^{\circ}C$ (373K)?

(b) Name a sodium compound which is used for making borax and glass.

(c) Name the compound which is used in hospitals for setting fractured bones.

(d) Which is the real bleaching agent present in bleaching powder?



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153. (a) What is baking powder? How does it make the cake soft and spongy?

(b). In addition to sodium hydrogen carbonate, baking powders contain a substance X. Name the substance X what is the role of substance X in the baking powder?

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154. State two uses each of the following compounds:

(a) Sodium hydroxide

(b) Chlorine

(c) Hydrogen

(d) Hydrochloric acid.

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155. (a) What is the common name of the compound $CaOCl_2$?

(b) Name the raw material used for the preparation of plaster of paris.

(c) Which property of plaster of paris is utilised in making casts of broken limbs in hospitals?

(d) Explain why chlorine is used for sterilising drinking water supply.

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156. (a). What happens when concentrated solution of sodium chloride (brine) is electrolysed? Write the equation of the reaction involved.

(b) Why is the electrolysis of a concentrated solution of sodium chloride known as chlor-alkali process?

(c) Name three products of the chlor-alkali process. State two uses of each of these products.

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157. (a). What is bleaching powder? How is bleaching powder prepared? Write chemical equation of the reaction involved in the preparation of bleaching powder.

(b) What happens when bleaching powder reacts with dilute sulphuric

acid? Give equation of the reaction involved.

(c) State two important uses of bleaching powder.

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158. (a). What is plaster of paris? Write the chemical formula of plaster of paris.

(b) How is plaster of paris prepared? Write chemical equation of the reaction involved.

(c) Explain why plaster of paris should be stored in a moisture-proof container.

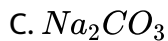
(d) State two important uses of plaster of paris.

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159. The salt which will give an acidic solution on dissolving in water is:

A. KCl

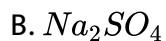
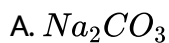
B. NH_4Cl



Answer: B

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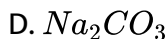
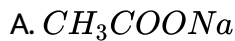
160. One of the following salts will give an alkaline solution on dissolving in water. This is:



Answer:

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161. The salt which will give a neutral solution on dissolving in water will be :

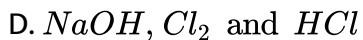
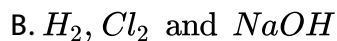
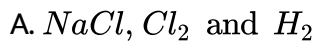


Answer: C



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162. The products of chlor-alkali process are :



Answer: B

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163. The number of molecules of water of crystallisation present in washing soda crystals is:

- A. five
- B. two
- C. ten
- D. seven

Answer: C

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164. The salt whose aqueous solution will turn blue litmus to red is:

A. ammonium sulphate

B. sodium acetate

C. sodium chloride

D. potassium carbonate

Answer: A

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165. The aqueous solution of one of the following salts will turn red litmus to blue. This salt is :

A. potassium sulphate

B. sodium sulphat

C. sodium chloride

D. potassium carbonate

Answer: D

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166. The salt whose aqueous solution will have no effect on either red litmus or blue litmus is

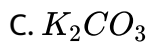
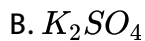
- A. potassium sulphate
- B. sodium carbonate
- C. ammonium sulphate
- D. sodium acetate

Answer: A

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167. The aqueous solution of one of the following salts will turn phenolphthalein indicator pink. This salt is :

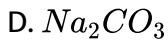
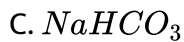
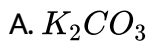
- A. *KCl*



Answer: C

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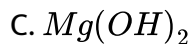
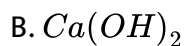
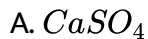
168. The formula of baking soda is:



Answer: C

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169. Which of the following is treated with chlorine to obtain bleaching powder ?

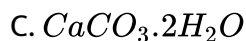
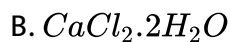
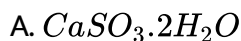


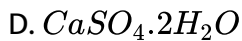
Answer: B



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170. Plaster of paris is prepared by heating one of the following to a temperature of $100^\circ C$ this is:



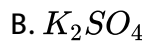
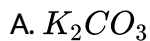


Answer: D



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171. A salt whose aqueous solution will have a pH of more than 7 will be:

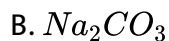


Answer: A



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172. A salt is dissolved in water and the pH of this salt solution is measured with a universal indicator paper. If the pH of solution is less than 7, the salt is most likely to :



Answer:



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173. Which of the following salts will give an aqueous solution having pH of almost 7?



C. CaCl_2

D. KCl

Answer:

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174. P and Q are aqueous solutions of sodium chloride and sodium hydroxide, respectively. Which of these will turn:

(a) Blue litmus red?

(b) red litmus blue?

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175. The metal salt A is blue in colour. When salt A is heated strongly over a burner, then a substance B is eliminated and a white powder C is left behind. When a few drops of a liquid D are added to powder C, it becomes blue again. What could be A, B, C and D?



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176. When the concentrated aqueous solution of substance X is electrolysed, then NaOH , Cl_2 and H_2 are produced. Name the substance X. What is the special name of the process.?



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177. Consider the following substances:

NaCl , Ca(OH)_2 , NaHCO_3 , NH_3 , Na_2CO_3 , H_2O , Cl_2 , CO_2 , $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$

- Which two substances combine to form bleaching powder?
- Which four substances are utilised in the production of washing soda?
- Which compound represents plaster of Paris?
- Which compound is a part of baking powder?
- Which compound is used as an antacid?



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178. Give one example each of a salt which gives an aqueous solution having:

(a). pH less than 7

(b) pH equal to 7 (b) pH equal to 7 (c) pH more than 7.

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179. A compound X which is prepared from gypsum has the property of hardening when mixed with a proper quantity of water.

(a) Identify the compound X

(b) Write the chemical equation for its preparation

(c) For what purpose is it used in hospitals?

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180. Consider the following salts:

Na_2CO_3 , $NaCl$, NH_4Cl , CH_3COONa , K_2SO_4 , $(NH_4)_2SO_4$

Which of these salts will give.

(a) Acidic solutions?

(b) neutral solutions?

(c) Basic solutions (or alkaline solutions)?



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181. A white powdery substance having strong smell of chlorine is used for disinfecting drinking water supply at waterworks. Identify the substance. Give its chemical name and write the chemical reaction for its preparation.



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182. A salt X when dissolved in distilled water gives a clear solution which turns red litmus blue. Explain the phenomenon.



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183. A person found that the cake prepared by him is hard and small in size. Which ingredient has he forgotten to add that would have caused to cake to rise and become light. Explain your answer.

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184. A white chemical compound becomes hard on mixing with proper quantity of water. It is also used in surgery to maintain joints in a fixed position. Name the chemical compound.

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185. When chlorine and sodium hydroxide being produced during the electrolysis of brine are allowed to mix, a new chemical is formed. Name this chemical and write its uses.

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186. Write the name and formula of one salt each which contains.

(a) Two molecules of water of crystallisation. (b) five molecules of water of crystallisation.

(c). Ten molecules of water of crystallisation.

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187. How many molecules of water of crystallisation (per formula unit) are present in:

(a) Copper sulphate crystals?

(b) washing soda?

(c) gypsum?

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188. You have been provided with three test tubes. One of them contains distilled water and the other two contain an acidic solution and a basic

solution, respectively. If you are given only red litmus paper, how will you identify the contents of each test tube?

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189. Why should curd and sour substances not be kept in brass and copper vessels?

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190. Which gas is usually liberated when an acid reacts with a metal? Illustrate with an example. How will you test for the presence of this gas?

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191. Metal compound A reacts with dilute hydrochloric acid to produce effervescence. The gas evolved extinguishes a burning candle. Write a

balanced chemical equation for the reaction if one of the compounds formed is calcium chloride.

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192. Why do HCl, HNO_3 , etc., show acidic characters in aqueous solutions while solutions of compounds like alcohol and glucose do not show acidic character?

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193. Why does an aqueous solution of an acid conduct electricity?

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194. Why does dry HCl gas not change the colour of the dry litmus paper?

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195. While diluting an acid, why is it recommended that the acid should be added to water and not water to the acid?

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196. How is the concentration of hydronium ions (H_3O^+) affected when a solution of an acid is diluted?

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197. How is the concentration of hydroxide ions (OH^-) affected when excess base is dissolved in a solution of sodium hydroxide?

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198. You have two solutions, A and B. The pH of solution A is 6 and pH of solution B is 8. Which solution has more hydrogen ion concentration?

Which of this is acidic and which one is basic?

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199. What effect does the concentration of $H^+(aq)$ ions have on the nature of the solution?

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200. Do basic solutions also have $H^+(aq)$ ions? If yes, then why are these basic?

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201. Under what soil conditions do you think a farmer would treat the soil of his fields with quicklime (calcium oxide), or slaked lime (calcium hydroxide) or chalk (calcium carbonate)?

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202. What is the common name of the compound $CaOCl_2$?

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203. Name the substance which on treatment with chlorine yields bleaching powder.

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204. Name the sodium compound which is used for softening hard water.

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205. What will happen if a solution of sodium hydrocarbonate is heated?

Give the equation of the reaction involved.

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206. Write an equation to show the reaction between Plaster of Paris and water

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207. A solution turns blue litmus red. Its pH is likely to be:

- A. 1
- B. 4
- C. 5
- D. 10

Answer: 10

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208. A solution reacts with crushed egg-shells to give a gas that turns lime-water milky. The solution contains

A. $NaCl$

B. HCl solution

C. $LiCl$

D. KCl

Answer: HCl



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209. 10 mL of a solution of NaOH is found to be completely neutralised by 8 mL of a given solution of HCl. If we take 20 mL of the same solution of NaOH, the amount HCl solution (the same solution as before) required to neutralise it will be

A. $4mL$

B. $4mL$

C. $12mL$

D. $16mL$

Answer: 16 mL



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210. Which one of the following types of medicines is used for treating indigestion?

A. Antibiotic

B. Analgesic

C. Antacid

D. Antiseptic

Answer: Antacid



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211. Write word equations and then balanced equations for the reaction taking place when

- A. dilute sulphuric acid reacts with zinc granules
- B. dilute hydrochloric acid reacts with magnesium ribbon
- C. dilute sulphuric acid reacts with aluminium powder
- D. dilute hydrochloric acid reacts with iron filings

Answer:

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212. Compounds such as alcohols and glucose also contain hydrogen but are not categorized as acids. Describe an activity to prove it.

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213. Why does distilled water not conduct electricity, whereas rain water does?

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214. Why do acids not show acidic behaviour in the absence of water?

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215. Five solutions A,B,C,D and E when tested with universal indicator showed pH as 4,1,11,7 and 9, respectively. Which solution is

- (a) neutral?
- (b) strongly alkaline?
- (c) strongly acidic?
- (d) weakly acidic?
- (e) weakly alkaline?

Arrange the pH in increasing order of hydrogen-ion concentration.

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216. Equal lengths of magnesium ribbons are taken in test tubes A and B. Hydrochloric acid (HCl) is added to test tube A, while acetic acid (CH_3COOH) is added to test tube B. Amount and concentration taken for both the acids are same. In which test tube will the fizzing occur more vigorously and why?

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217. Fresh milk has a pH of 6. How do you think the pH will change as it turns into curd? Explain your answer.

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218. A milkman adds a very small amount of baking soda to fresh milk.

(a) Why does he shift the pH of the fresh milk from 6 to slightly alkaline?

(b) Why does this milk take a long time to set as curd?



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219. Plaster of Paris should be stored in a moisture-proof container.

Explain why?

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220. What is a neutralisation reaction? Give two examples.

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221. Give two important uses of washing soda and baking soda.

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Short Answer Type Question

1. With the help of labelled diagrams, describe an activity to show that acids produce ions only in aqueous solution.

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2. Explain how anhydrous copper sulphate can be used to detect the presence of moisture (water) in a liquid.

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3. What does a soda-acid type of fire extinguisher contain? How does it work? Explain the working of a soda-acid fire extinguisher with the help of a labelled diagram.

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Long Answer Type Questions

1. (a). Define an acid and a base. Give two examples of each.
- (b). Give the names and formulae of two strong bases and two weak bases.
- (c). What type of ions are formed:
 - (i) When an acid is dissolved in water?
 - (ii). When a base (or alkali) is dissolved in water?
- (d). Write the neutralisation reaction between acids and bases in terms of the ions involved.
- (e). Write any two important uses of bases.



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2. (a). Describe how washing soda is produced starting from sodium chloride (common salt). Write equations of all the reactions involved.
- (b) State whether an aqueous solution of washing soda is acidic or alkaline? Give reason for your answer.
- (c) What is meant by saying that washing soda has detergent properties?
- (d) Give two important uses of washing soda (or sodium carbonate).



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3. (a). What is a salt? Give the names and formulae of any two salts. Also name the acids and bases from which these salts may be obtained.

(b) What is meant by 'a family of salts' ? Explain with examples.

(c) What is meant by 'hydrated' and 'anhydrous' salts? Explain with examples.

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Questions Based On High Ordre Thinking Skills Hots

1. A group of students measured the pH of some substances they found in their homes. Their results are given in the following table:

<i>Substance</i>	<i>pH</i>	<i>Substance</i>	<i>pH</i>
Apples	3.0	Salt	7.0
Baking soda	8.5	Sugar	7.0
Black coffee	5.0	Toothpaste	9.0
Household ammonia	12.0	Vinegar	3.0
Lemon juice	2.5	Washing soda	11.5
Milk	6.5		

(a). What would the student have used to measure the pH?

- (b). Which solution is the most acidic?
- (c). Which solution is the most alkaline?
- (d). Which solutions are neutral?
- (e) Which solution can be used to treat wasp stings?
- (f) Which solution can be used to treat bee stings?



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