PHYSICS

BOOKS - HC VERMA PHYSICS (HINGLISH)

LAWS OF THERMODYNAMICS

Examples

1. A gas is contained in a vessel fitted with a movable piston. The container is placed on a

stove. A total of 100 cal of heat is given to the gas and the gas does 40 J of work in the expansion resulting from heating. Calculate the increase in internal energy in the process.

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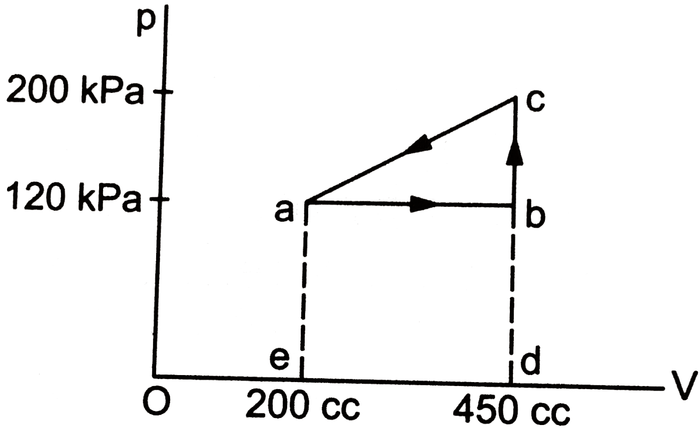
2. calculate the work done by a gas as it is taken from the state a to b, b to c and c to a as shown in



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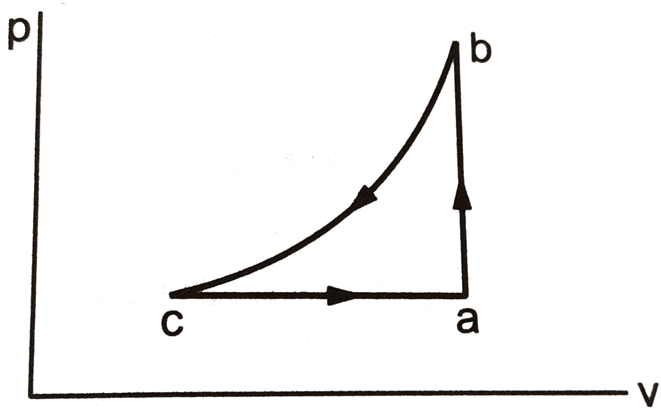
Worked Out Examples

1. A sample of an ideal gas is taken through the cyclic process abca . It absorbs 50 J of heat during the part ab, no heat during bc and rejects 70 J of heat during ca. 40 J of work is done on the gas during the part bc.(a) find the internal energy of the gas at b and c if it is 1500 J at a. (b) calculate the work done by the gas during the part ca.



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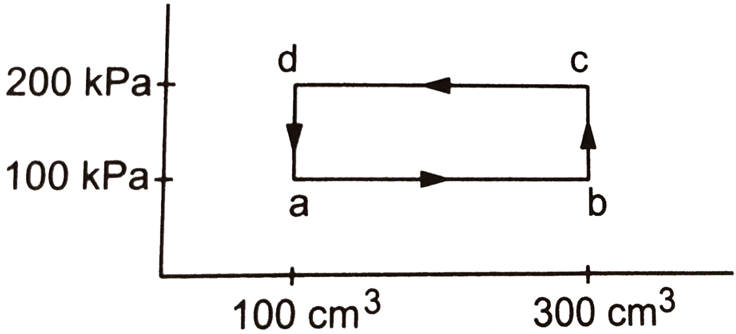
2. A thermodynamic system is taken through the cycle abcda during the parts ab, bc, cd and da. (b) find the total heat rejected by the gas during the process.



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3. calculate the increase in internal energy of 1 kg of water at $100(0)C$ when it is converted into steam at the same temperature and at 1atm (100 kPa). The density of water and steam are 1000 kg m^{-3} and 0.6 kg m^{-3} respectively. The latent heat of vaporization of water $= 2.25 \times 10^6 Jkg^1$.

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4. The internal energy of a monatomic ideal gas is 1.5 nRT. One mole of helium is kept in a cylinder of cross section $8.5cm^2$. The cylinder is closed by a light frictionless piston. The gas is heated slowly in a process during which a total of 42J heat is given to the gas. if the temperature rise through $2^{(0)}C$, find the distance moved by the piston. atmoshphere pressure `=100 kPa.

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5. A steam engin intakes 100g of steam at $100^{(0)}C$ per minute and cools it down to $20^{(0)}C$. Calculate the heat rejected by the steam engine per minute. Latent heart of vaporization of steam $= 540 cal g^{-1}$.

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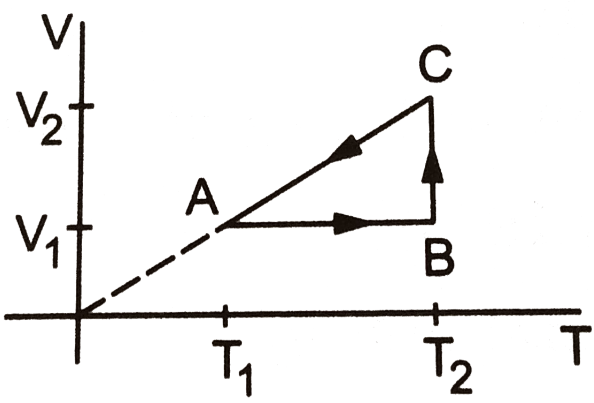
6. Shows a process ABCA performed on an ideal gas. Find the net heat given to the system during the process.



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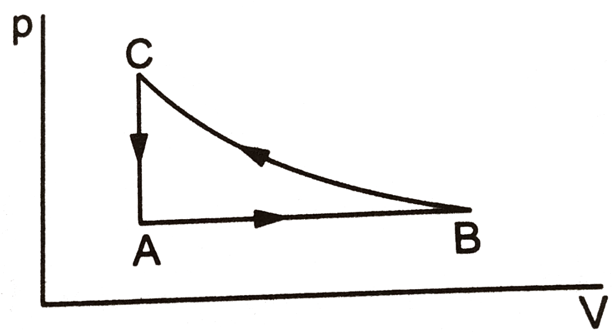
7. consider the cyclic process ABCA on a sample of 2.0 mol of an ideal gas as shown in the temperature of the gas at A and B are 300 K and 500 k respectively. A total of 1200 J heat is withdrawn from the sample in the process. Find the work done by the gas in part BC. take

$R = 8.3JK^{-1}mol^{-1}$



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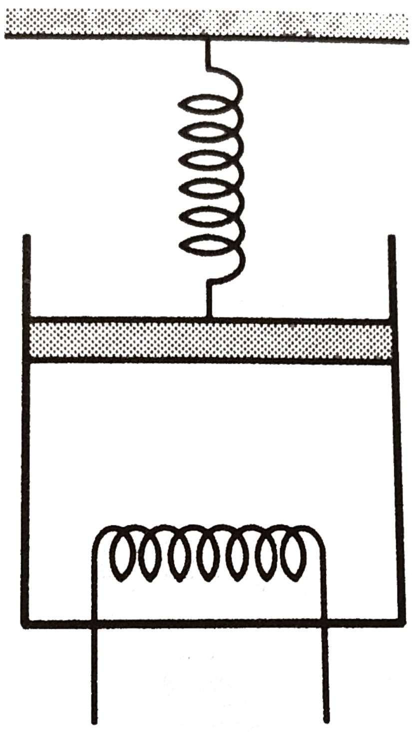
8. 2.00 mol of a monatomic ideal gas $(U = 1.5nRT)$ is enclosed in an adiabatic, fixed , vertical cylinder fitted with a smooth, light adiabatic piston.the piston is connected to a vertical spring of spring constant $200Nm^{-1}$ as shown in .the area of cross section of the cylinder is $20.0cm^2$. initially, the spring is at its natural length and the temperature of the gas is at its natural length and the temperature of the gas is 300 K. the atmosphere pressure is 100 kPa. the gas is heated sloewly for some time by means of an electric heater so as to move the poston up through 10 cm. find (a) the work done by the gas (b) the final temperature of the gas and (c) the heat supplied by the heater.



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9. A sample of an ideal gas has pressure p_0, volume V_0 and tempreture T_0. It is isothermally expanded to twice its oringinal volume.it is then compressed at constant pressure to have the original volume V_0. Finally, the gas is heated at constant volume to get the original tempreture.(a) show the process in a V-T diagram (b) calculate the heat absorbed in the process.

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10. A sample of 100 g water is slowly heated from $27^{(0)}C$ to $87^{(0)}C$ Calculate the water

$= 4200 J kg^{-1} K^{-1}$.

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11. A heat engine operates between a cold reservoir at tempreture $T_2 = 300K$ and a hot reservior at tempreture T_1. It takes 200 J of heat from the hot reservior and delivers 120 J of heat to the cold reservior in a cycle. What could be the minimum tempreture of the hot reservior ?

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Objective 1

1. The first law of theromodynamics is a statement of

A. conservation of heat

B. conservation of work

C. conservation of momentum

D. conservation of energy

Answer: D

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2. If heat is supplied to an ideal gas in an isothermal process.

A. the internal energy of the gas will increase

B. the gas will do positive work

C. the gas will do negative work

D. the said process is not possible.

Answer: B

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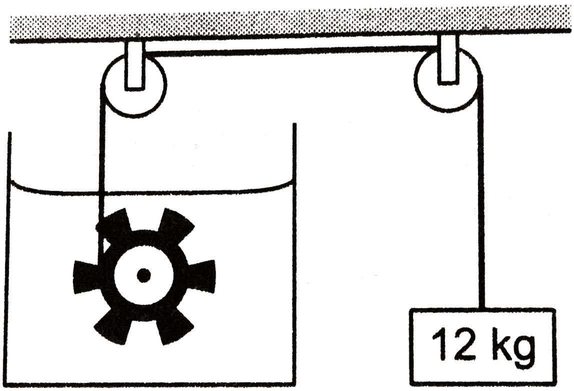
3. shows two processes A and B on a system.Let DeltaQ_(1) and DeltaQ_(2) be the heat given to the system in processes A and B respectively. Then



A. $\Delta Q_1 > \Delta Q_2$

B. $\Delta Q_1 = \Delta_2$

C. $\Delta Q_1 < \Delta Q_2$

D. $\Delta Q_1 \leq \Delta Q_2$

Answer: A

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4. Refer to let ΔU_1 and ΔU_2 be the changes in internal energy of the system in the processes A and B. then

A. $\Delta U_1 > \Delta U_2$

B. $\Delta U_1 = \Delta U_2$

C. $\Delta U_1 < \Delta U_2$

D. $\Delta U_1 \neq \Delta U_2$

Answer: B

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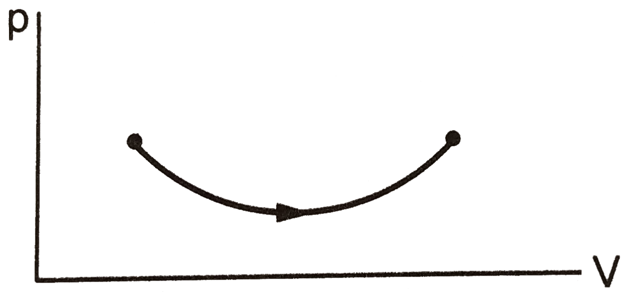
5. consider the processes on a system shown in. during the processes,the work done by the system



A. continuously increases

B. continuosly decreases

C. first increase then decreases

D. first decrease then increases then increase.

Answer: A

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6. consider the following two statements.

(A) If heat is added to a system, its temperature must increase.

(b) if positive work is done by a system in a thermodynamic process, its volume must increase.

A. both A and B are correct.

B. A is correct but B is wrong.

C. B is correct but A is wrong.

D. Both A and B are wrong.

Answer: C

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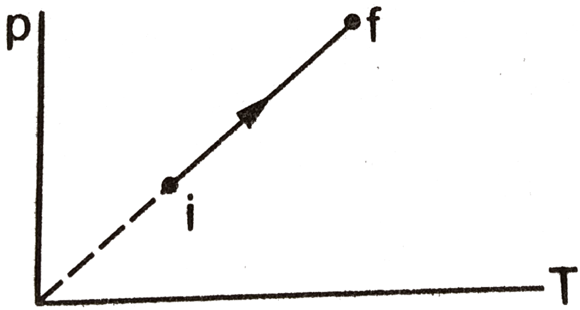
7. An ideal gas goes from the state i to the state $\int$ as shown in .the work done by the gas during the process



A. is positive

B. is negative

C. is zero

D. cannot be obtained from this information.

Answer: C

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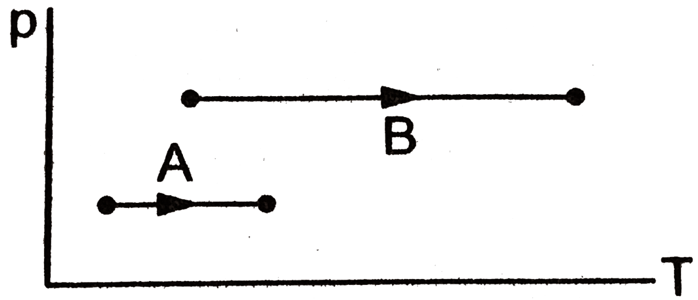
8. consider two processes on a system on a system as shown in

the volumes in the initial states are the same in the two processes and the volume in the final states are also the same.Let ΔW_1 and ΔW_2 be the work done by the sustem in the processes A and B respectively.



A. $\Delta W_1 > \Delta_2$

B. $\Delta W_1 = \Delta W_2$

C. $\Delta W_1 < \Delta W_2$

D. nothing can be said about the relation between ΔW_1 and ΔW_2.

Answer: C

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9. A gas is contained in a metallic cylinder fitted with a piston.the piston is suddenly moved in to compress the gas and is maintained at this position. As time passes the pressure of the gas in the cylinder

A. increase

B. decrease

C. remains constant

D. increase or decrease depending on the nature of the gas.

Answer: B

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Objective 2

1. The pressure p and volume V of an ideal gas both increase in a process.

A. such a process is not possible.

B. The work done by the system is positive.

C. The tempreture of the system must increase.

D. Heat supplied to the gas is equal to the change in internal energy.

Answer: B::C

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2. In a process on a system, the initial pressure and volume.

A. The initial temperature must be equal to the final temperature.

B. the initial internal energy must be equal to the final internal energy.

C. the net heat given to the system in the process must be zero.

D. the net work done by the system in the process must be zero.

Answer: A::B

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3. A system can be taken from the initial state p_1, V_1 to the final state p_{2,V_2} by two different methods, let ΔQ and ΔW represent the heat given to the system and the work done by the system. Which of the following must be the same in both the method?

A. ΔQ

B. ΔW

C. $\Delta Q + \Delta W$

D. $\Delta Q - \Delta W$.

Answer: D

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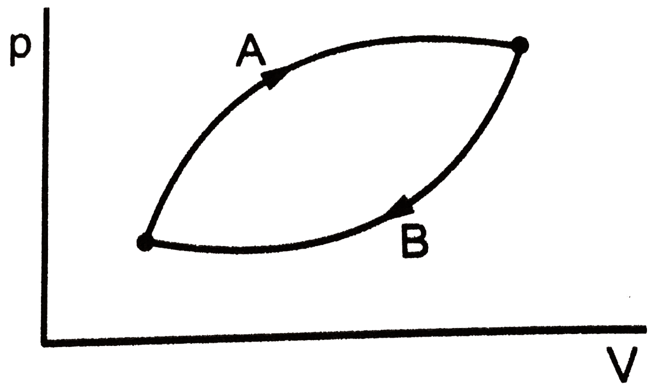
4. Refer to let ΔU_1 and ΔU_2 be the changes in internal energy in the system in process $A + B$ and ΔW be the net work done by the system in the process `A+B,



A. $\Delta U_1 + \Delta U_2 = 0.$

B. $\Delta U_1 - \Delta U_2 = 0$

C. $\Delta Q - \Delta W = 0.$

D. $\Delta Q + \Delta W = 0.$

Answer: A::C

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5. The internal enegy of an ideal gas decreases nu the same amount as the work done by the system.

A. The process must be adiabatic.

B. The process must be isothermal.

C. The process must be isobaric.

D. The temperature must decrease.

Answer: A::D

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Exercises

1. A thermally insulaated, closed copper vessel contains water at $15^0 C$. When the vessel is shaken vigorously for 15 minutes, the tempreture rises to $17^0 C$. The mass of the vessel is 100 g and that of the water is 200 g. the specific heat capacities of copper and water are $420 J kg^{-1} K^{-1}$ and $4200 J kg^{-1} K^{-1}$ respectively. negalect any thermal expansion. (a) how much heat is transferred to the liquid-vessel system ? (b)how much work has been done on this system? (c) how much is the increase in internal energy if the system ?

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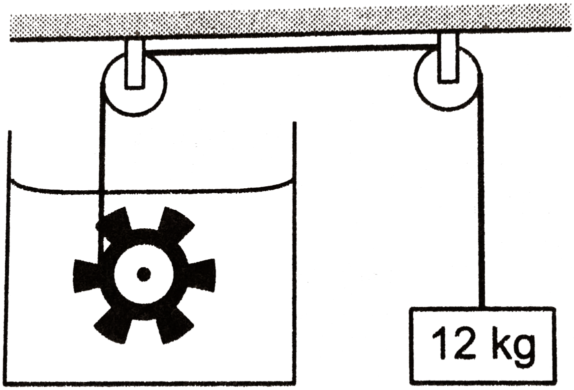
2. Shows a peddle wheel couple to a mass of 12 kg through fixed frictionless pulleus. The paddle is immersed in a liquid of heat capacity $4200 JK^{-1}$ kept in an adiabatic container. Consider a time interval in which the 12 kg block falls slowly through 70cm. (a) how much heat is given to the liquid ? (b) how much work is done on the liquid? (c) calculate the rise in the temperature of the liquid neglecting the heat capacity of the container and the paddle.



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3. A 100 kg block is started with a speed of $2.0ms^{-1}$ on a long, rough belt kept fixed in a horizontal position. The coefficient of kinetic friction between the block and the belt is 0.20.

(a) calculate the change in the internal energy of the block=-belt system as the block comes to a stop on the belt. (b) consider the sutuation from a frame of reference moving at $2.0ms^{-1}$ along the initial velocity of the block. as seen from this frame, the block is gently put on a moving belt and in due time the block starts moving with the belt at $2.0ms^{-1}$, calculate the increase in the kinetic energy of the block as it stops slipping past the belt. (c) find the work done in this frame by the external force holding the belt.

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4. calculate the change in internal energy of a gas kept in a rigid container when 100 J of heat is supplied to it.

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5. the pressure of a gas change linearly with volume from 10kPa, 200 cc to 50 kPa, 50 cc. (a) calculate the work done by the gas, (b) what is the change in the internal energy of the gas ?

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6. An ideal gas is taken from an inirial state I to a final state f in such a way that the ratio of the pressure to the absolute temperature remains constant. What will be the work done by the gas ?

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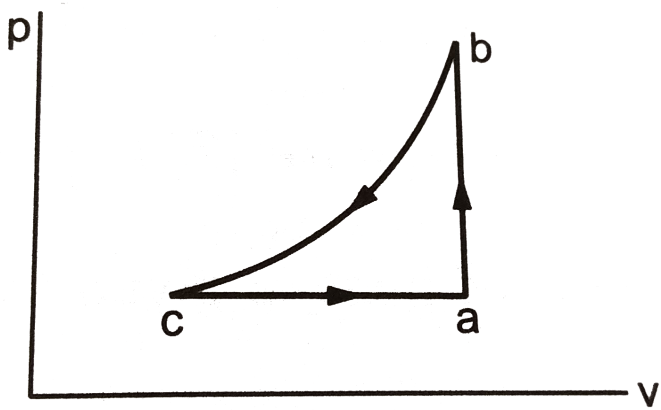
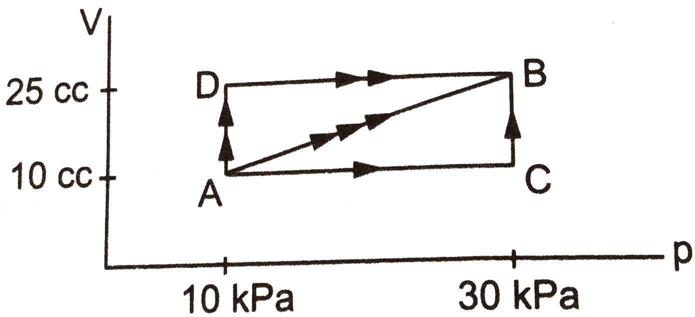
7. shows three paths through which a gas can be taken from the state A to the state B.

calculate the work done by the gas in each of the three paths.





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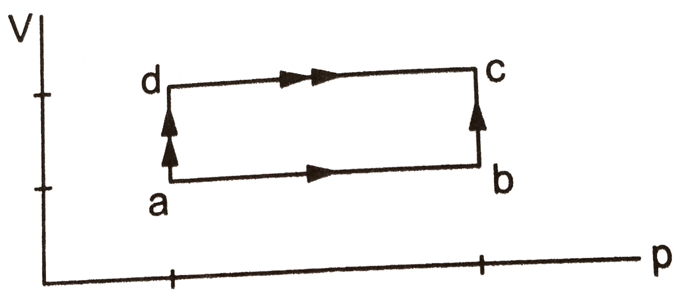
8. when a system is taken through the process abc shown in 80 J of heat is absorbed by the system and 30 J of work is done by it. If the system does 10 J work during the process adc, how much heat flows into it during the process ?



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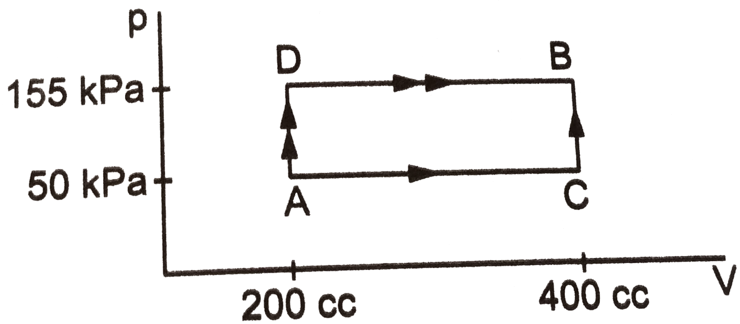
9. 50 eal of heat should be supplied to take a system from the state A to the state B through the path ABC as shown in ,find the quantity of heat to be supplied to take it from A to B via ADB.



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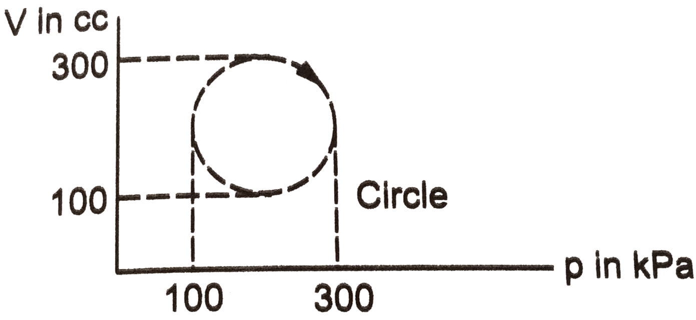
10. calculate the heat absorbed by a system in going through the cyclic process shown in ,



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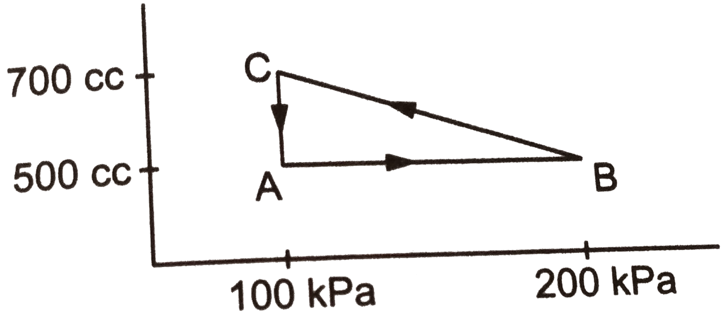
11. A gas is taken through a cyclic process ABCA as shown in, if 2.4 cal of heat is given in the process, what is the value of J ?



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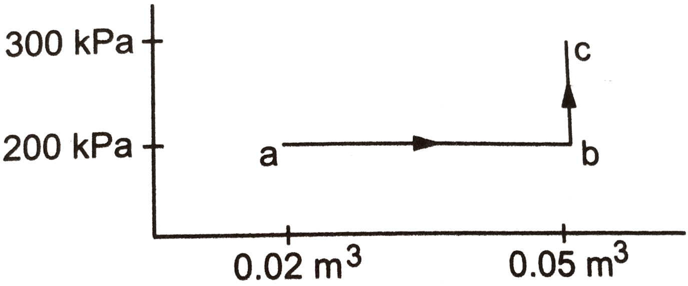
12. A substance is taken through the process abc as shown in, if the internal energy of the substance increase by 5000 J and a heat of 2625 cal is given to the system, calculate the value of J.



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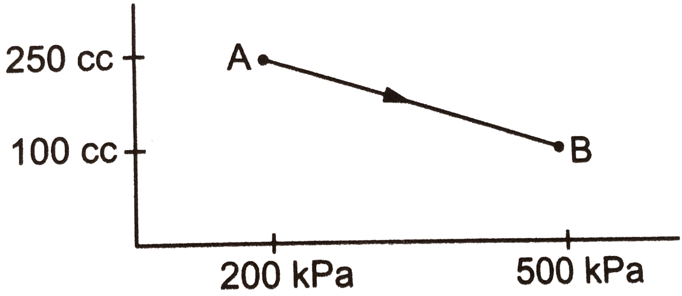
13. A gas is taken along the path AB as shown in , if 70 cal of heat is extracted from the gas in the process, calculate the change in the internal energy of the system.



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14. The internal energy of a gas is given by $U = 1.5pV$. It expands from $100cm^2$ to $200cm^3$ against a constant pressure of $1.0 \times 10^5 Pa$. Calculate the heat absorbed by the gas in the process.

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15. A gas is enclosed in a cylindrical vessel fitted with a frictionless piston.the gas is slowly heated for some time. During the process, 10 J of heat is suopplied and the piston is found to move out 10cm. Find the increase in the internal energy of the gas , the area of cross section of the gas, the area of cross section of the cylinder=4 cm^(4) and the atmospheric pressure $= 100kPa$.

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16. A gas is initinaly at a pressure of 100 kPa and its volume is $2.0m^3$. Its pressure is kept constant and the volume is changed from $2.0m^3 \rightarrow 2.5^3$. Its volume is now kept constant and the pressure is increased from 100 lPa to 200 kPa. The gas is brought back to its initial state, the pressure varying linearly with its volume. (a) whether the heat is supplied to or exterted from the gas in the complete cycle ? (b) how much heat was supplied or extracted ?

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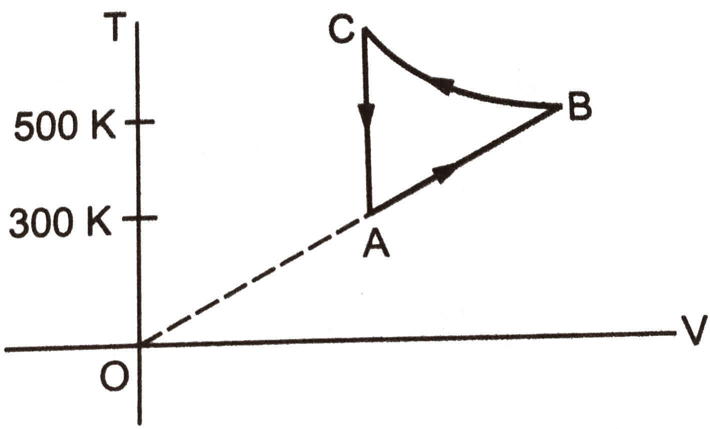
17. Consider the cyclic process ABCA, shown in, performed on a sample of 2.0 mol if an ideal gas. A total of 1200 J if heat is withdrawn from the sample in the process. Find the work done by the gas during the part BC.



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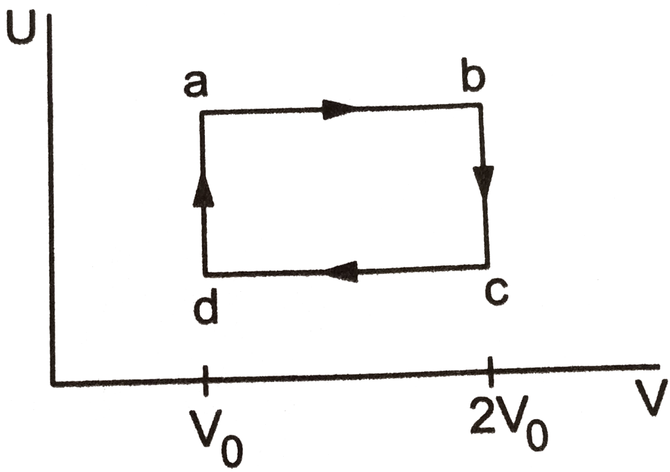
18. shows the variation in the internal energy U with the volume V of 2.0 mol of an ideal gas in a cyclic process abcda. The temperatures of the gas at b and c are 500 K and 300 K respectively. Calculate the heat absorbed by the gas during the process.



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19. Find the change in the internal energy of 2 kg of water as it heated from $0^0C \rightarrow 4^0C$. The specific heat capacity of water is $4200 J kg^{-1}K^{-1}$ and its densities at 0^0C and 4^0C are $999.9 kgm^{-3}$ and $1000 kgm^{-3}$ respectively. atmospheric pressure $= 10^5$ Pa.

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20. Calculate the increase in the internal energy of 10 g of water when it is heated from $0^0 C \rightarrow 100^0 C$ and converted into steam at 100 kPa. The density of steam $= 0.6 kgm^{-3}$ specific heat capacity of water $= 4200 Jkg^{-1^0 C^{-3}}$ latent heat of vaporization of water $= 2.25 \times 106 Jkg^{-1}$

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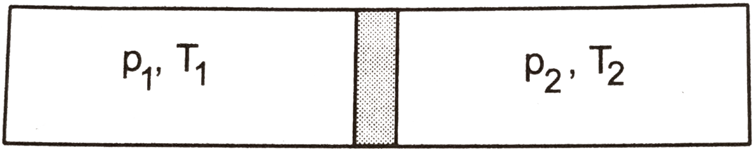
21. shows a cylindrial tube of volume V with adiabatic walls containing an ideal gas. The internal energy of this ideal gas is given by 1.5nRT. The tube is divided into two equal parts by a fixed diathermic wall. Initially, the pressure and the tempreture are p_1, T_1 on the left and p_2, T_2 on the right. the system is left for sufficient time so that the tempreture becomes equal on the two sides. (a) how much work has been done by the gas on the left part ? (b) find the final pressures on the two sides. (c) find the final equilibrium tempreture. (d) how much heat has flown from the gas on the right to the gas on the left ?



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22. An adiabatic vessel of total volume V is divided into two equal parts by a conducting separtor. The is fixed in this position. The part on the left contains one mole of an ideal gas $(U = 1.5nRT)$ and the part on the right contains two moles of the same gas. initially, the pressure on each side is p. the system is left for sufficiant time so that a steady state is reached. find (a) the work done by the gas in the left part during the process, (b) the temperature on the two sides in the beginning, (c) the final common temperature reached by the gases, (d) the heat given to the gas in the right part and (e) the increase in the internal energy of the gas in the left part.

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Short Answer

1. Should the internal energy of a system necessarily increase if heat is added to it?

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2. Should the internal energy of a system necessarily increase if its tempreture is increased?

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3. A cylinder containing a gas is lifted from the first floor to the second floor. What is the amount of work done on the gas? What is the amount of work done by the gas? Is the internal energy of the gas increased ? Is the tempreture of the gas increased?

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4. A force F is applied on a block of mass M. The block is displaced through a distance d in the direction of the force. What is the work done by the force on the block? Does the internal energy change because of this work?

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5. The outer surface of a cylinder containing a gas is rubbed vigorously by a polishing machine. The cylinder and its gas become warm.is the energy transferred to the gas heat or work?

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6. When we rub our hands they become warm. Have we supplied heat to the hands?

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7. A closed bottle contains some liquid. The bottle is shaken vigorously for 5 minutes. It is found that the tempreture of the liquide is increased. Is heat transferred of the liquid? Is work done on the liquid? Neglect expansion on heating.

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8. The final volume of a system is equal to the initial volume in a certain process. Is the work done by the system necessarily zero? Is it necessarily nonzero?

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9. Can work be done by a system without changing its volume?

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10. An ideal gas is pumped into a rigid container having diathermic walls so that the tempreture remains constant. In a certain time interval energy of the contains of the container also doubled in the interval?

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11. When a tyre bursts, the air coming out is cooler than the surrounding air.Explain.

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12. When we heat an object, it expands. Is work done by the object in this process? Is heat given to the object equal to the increase in its internal energy?

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13. When we stir a liquid vigorously, it becomes warm. Is it a reversible process?

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14. What should be the condition for the efficiency of a carnot engine to be equal to 1?

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15. When an object cools down, heat is withdrawn from it. Does the entropy of the object decrease in this process? If yes, is it a violation of the second law of thermodynamics stated in terms of increased in entropy?

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