



CHEMISTRY

BOOKS - NEET PREVIOUS YEAR (YEARWISE + CHAPTERWISE)

RE-NEET 2020

Others

1. Which of the following statement is not true about acid rain ?

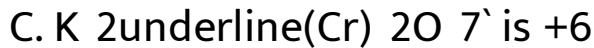
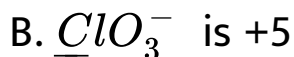
- A. It is due to reaction of SO_2 , NO_2 and CO_2 with rain water
- B. Causes no damage to monuments like Taj mahal
- C. It is harmful for plants
- D. Its pH is less than 5.6

Answer: B



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2. The oxidation number of the underlined atom in the following species. Identify the incorrect option.



Answer: A



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3. Reaction of propanamide with ethanolic sodium hydroxide and bromine will give

- A. Ethylamine
- B. Methylamine
- C. Propylamine
- D. Aniline

Answer: A



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4. A liquid compound(X) can be purified by Steam distillation only if it is

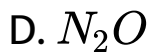
- A. Steam volatile, Immiscible with water
- B. Not steam volatile, miscible with water
- C. Steam volatile, miscible with water
- D. Not steam volatile, immiscible with water

Answer: A



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5. Among the compounds shown below which one revealed a linear structure ?

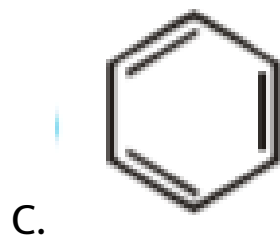
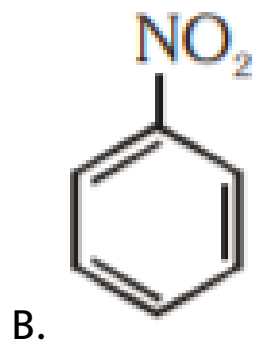
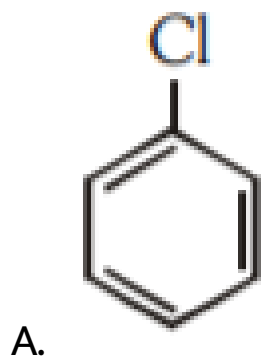


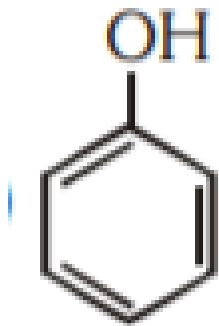
Answer: D



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6. Which of the following compounds is most reactive in electrophilic aromatic substitution?





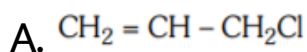
D.

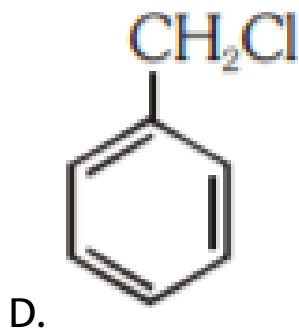
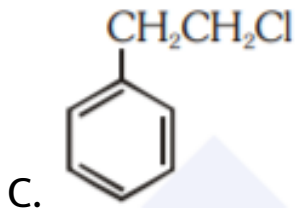
Answer: D



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7. Which of the following will not undergo S_N^1 reaction with OH^-





Answer: C



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8. Which of the following will not true about chloramphenicol ?

- A. It inhibits the growth of only grampositivebacteria
- B. It is a broad spectrum antibiotic
- C. It is not bactericidal
- D. It is bacteriostatic

Answer: A



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9. Which of the following statement is correct about Bakelite?

- A. It is a cross linked polymer
- B. It is an addition polymer
- C. It is a branched chain polymer
- D. It is a linear polymer

Answer: A



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10. If for a certain reaction $\Delta_r H$ is 30KJ mol^{-1} at 450K, the value of $\Delta_r S$ (in $\text{JK}^{-1}\text{mol}^{-1}$) for which the same reaction will be spontaneous at the same temperature is

A. 70

B. -33

C. 33

D. -70

Answer: A



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11. Match the element in column I with that in column II.

Column-I

- (a) Copper
- (b) Fluorine
- (c) Silicon
- (d) Cerium

Column-II

- (i) Non-metal
- (ii) Transition metal
- (iii) Lanthanoid
- (iv) Metalloid

Identify the correct match:

A. a-ii,b-iv,c-i,d-iii

B. a-ii,b-i,c-iv,d-iii

C. a-iv,b-iii,c-i,d-ii

D. a-i,b-ii,c-iii,d-iv

Answer: B



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12. Which of the following is a free radical substitution reaction?

A. Benzene with $B\frac{r_2}{A}lCl_3$

B. Acetylene with Hbr

C. Methane with $B\frac{r_2}{h}v$

D. Propene with $HB\frac{r}{(C_6H_5COO)_2}$

Answer: C



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13. The reaction of concentrated sulphuric acid with carbohydrates ($C_{12}H_{22}O_{11}$) is an example of

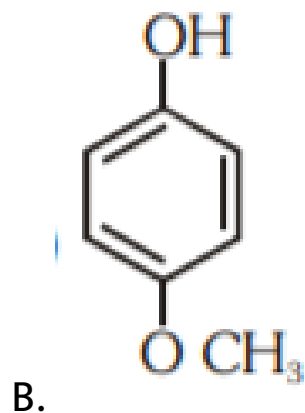
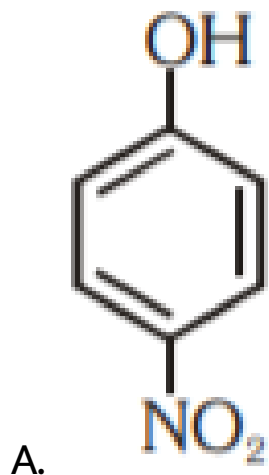
- A. Dehydration
- B. oxidation
- C. Reduction
- D. Sulphonation

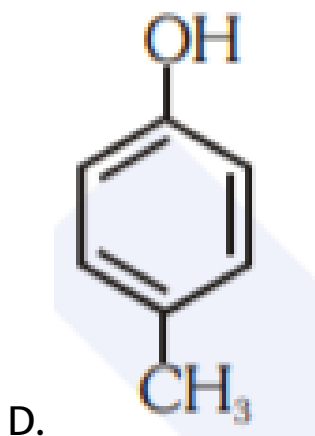
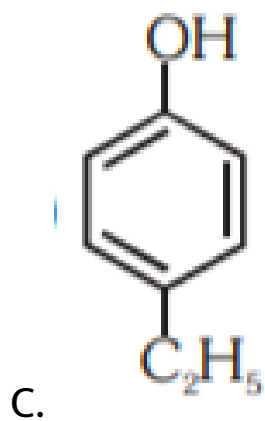
Answer: A



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14. Which of the following substituted phenols is the strongest acid?





Answer: A



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15. Match the compounds of Xe in column I with the molecular structure in column II.

Column-I

- (a) XeF_2
- (b) XeF_4
- (c) XeO_3
- (d) XeOF_4

Column-II

- (i) Square planar
- (ii) Linear
- (iii) Square pyramidal
- (iv) Pyramidal

Identify the correct match:

A. a-ii,b-i,c-iii,d-iv

B. a-ii,b-iv,c-iii,d-i

C. a-ii,b-iii,c-i,d-iv

D. a-ii,b-i,c-iv,d-iii

Answer: D



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16. The half life for a zero order reaction having 0.02M initial concentration of reactant is 100s. The rate constant (in $\text{molL}^{-1}\text{S}^{-1}$) for the reaction is

A. 1.0×10^{-4}

B. 2.0×10^{-4}

C. 2.0×10^{-3}

D. 1.0×10^{-2}

Answer: A



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17. Identify the incorrect statement from the following:

- A. Zirconium and hafnium have identical radii of 160pm and 159pm, respectively as a consequence of lanthanoid contraction.
- B. Lanthanoids reveal only +3 oxidation state
- C. The lanthanoid ions other than the f^0 type and the $1f^{(14)}$ type are all paramagnetic

D. The overall decrease in atomic and ionic radii from lanthanum to lutetium is called lanthanoid contraction

Answer: B



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18. Match the following aspects with the respective metal.

Aspects**Metal**

- | | |
|--|-----------------|
| (a) The metal which reveals a maximum number of oxidation states | (i) Scandium |
| (b) The metal although placed in 3d block is considered not as a transition element | (ii) Copper |
| (c) The metal which does not exhibit variable oxidation states | (iii) Manganese |
| (d) The metal which in +1 oxidation state in aqueous solution undergoes disproportionation | (iv) Zinc |

A. a-i,b-iv,c-ii,d-iii

B. a-iii,b-iv,c-i,d-ii

C. a-iii,b-i,c-iv,d-ii

D. a-ii,b-iv,c-i,d-iii

Answer: B



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19. If 8g of a non-electrolyte solute is dissolved in 114g of n-octane to reduce its vapour pressure to 80 %, the molar mass of the solute is [given molar mass of n-octane is 114g mol^{-1}]

A. 40

B. 60

C. 80

D. 20

Answer: A



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20. Match the coordination number and type of hybridisation with distribution of hybrid orbitals

in space based on valence bond theory.

**Coordination
number and
type of
hybridisation**

(a) 4, sp^3

(b) 4, dsp^2

(c) 5, sp^3d

(d) 6, d^2sp^3

**Distribution
of hybrid
orbitals
in space**

(i) trigonal

bipyramidal

(ii) octahedral

(iii) tetrahedral

(iv) square planar

A. a-ii,b-iii,c-iv,d-i

B. a-iii,b-iv,c-i,d-ii

C. a-iv,b-i,c-ii,d-iii

D. a-iii,b-i,c-iv,d-ii

Answer: B



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21. The number of angular nodes and radial nodes in 3s orbital are

A. 0 and 2, respectively

B. 1 and 0, respectively

C. 3 and 0, respectively

D. 0 and 1, respectively

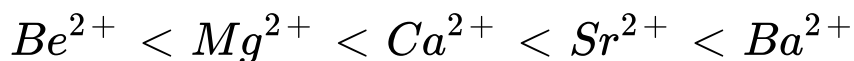
Answer: A



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22. Identify the correct statement from the following

A. The order of hydration enthalpies of alkaline earth cations



B. Lithium and magnesium show some similarities in their physical properties as they are diagonally placed in periodic table

C. Lithium is softer among all alkali metals

D. Lithium chloride is deliquescent and crystallises as a hydrate, $\text{LiCl} \cdot H_2O$

Answer: B



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23. Deficiency of which vitamin causes osteomalacia?

A. vitamin A

B. vitamin D

C. vitamin K

D. vitamin E

Answer: B



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24. Identify the wrongly matched pair.

A. PCl_5 - trigonal planar

B. SF_6 -Octahedral

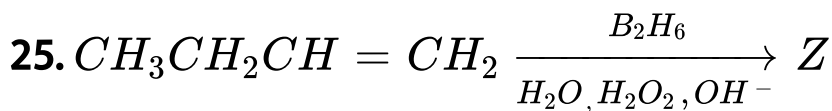
C. $BeCl_2$ -Linear

D. NH_3 -Trigonal pyramidal

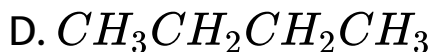
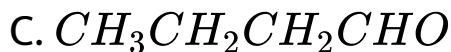
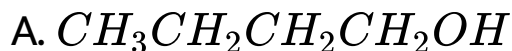
Answer: A



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What is Z?

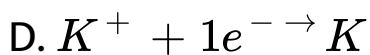
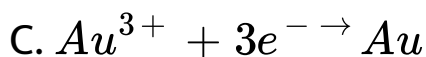
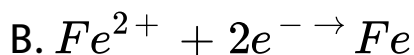
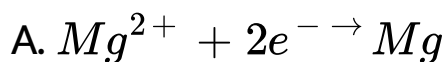


Answer: A



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26. Identify the reaction from having top position in EMF series according to their electrode potential at 298k



Answer: C



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27. Match the element in column I with methods of purification in column II.

Column I

- (a) Boron
- (b) Tin
- (c) Zirconium
- (d) Nickel

Column II

- (i) Van Arkel method
- (ii) Mond's process
- (iii) Liquation
- (iv) Zone refining

Identify the correct match:

A. a-iv,b-iii,c-i,d-ii

B. a-iv,b-iii,c-ii,d-i

C. a-ii,b-i,c-iv,d-iii

D. a-iii,b-iv,c-i,d-ii

Answer: A



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28. Which among the following salt solutions is basic in nature?

- A. Ammonium chloride
- B. Ammonium sulphate
- C. Ammonium nitrate
- D. sodium acetate

Answer: D



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29. In which of the sols, the colloidal particles are with negative charge?

A. TiO_2

B. Haemoglobin

C. Starch

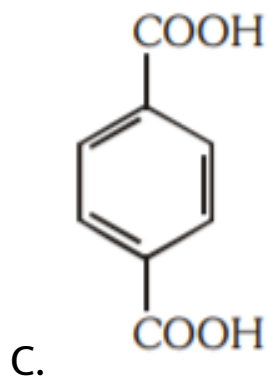
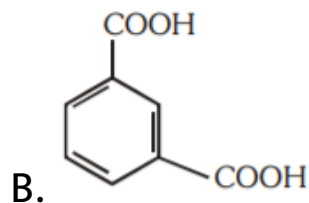
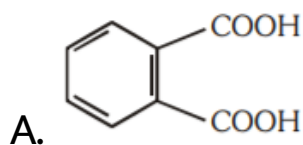
D. Hydrated Al_2O_3

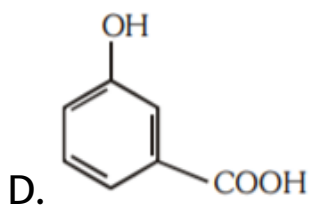
Answer: C



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30. Which of the following acid will form an (a)
Anhydride on heating and
(b) Acid amide on strong heating with ammonia?

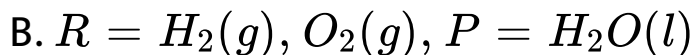
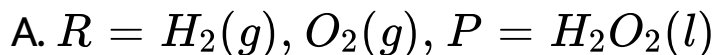




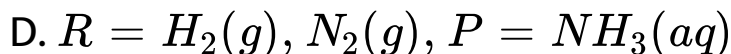
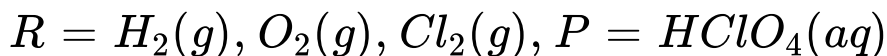
Answer: A

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31. In a typical fuel cell, the reactants(R) and product (P) are:



C.



Answer: B



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32. In collision theory of chemical reaction, Z_{AB} represents

A. the fraction of molecules with energies greater than E_a

B. the collision frequency of reactants, A and B

C. steric factor

D. the fraction of molecules with energies equal to E_a

Answer: B



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33. Which of the following statement is not true about glucose?

A. It is an aldohexose

B. It contains five hydroxyl groups

C. It is a reducing sugar

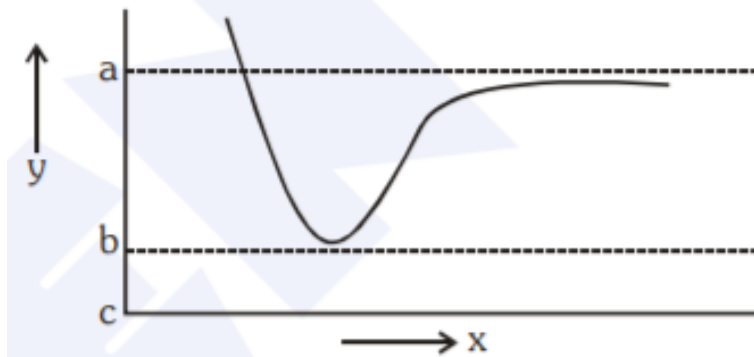
D. It is an aldopentose

Answer: D



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34. The potential energy(y) curve for H_2 formation as a function of internuclear distance(x) of the H atoms is shown below



The bond energy of H_2 is :

A. $b - a$

B. $\frac{c - a}{2}$

C. $\frac{b - a}{2}$

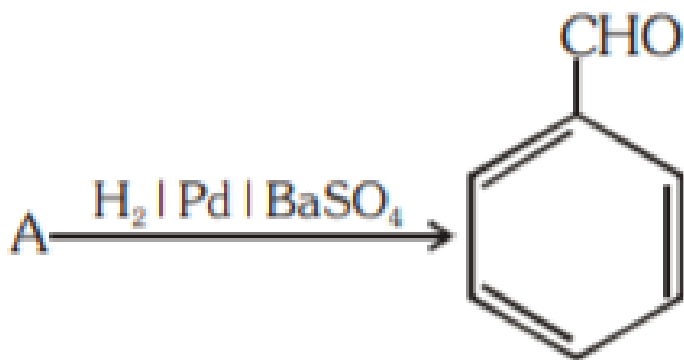
D. $c - a$

Answer: A



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35. Identify compound (A) in the following reaction:



A. Benzoyl chloride

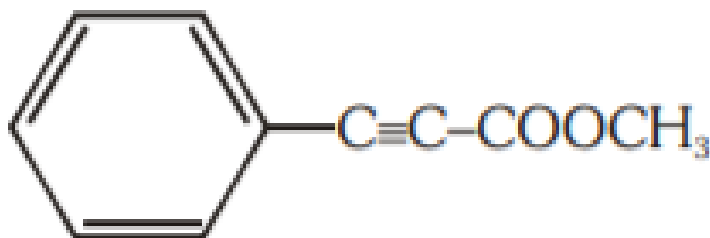
B. Toulene

C. Acetophenone

D. Benzoic acid

Answer: A

36. How many sp^2 hybridised carbon atoms and π bonds respectively are present in the following compound?



A. 7,5

B. 8,6

C. 7,6

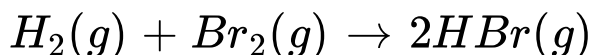
D. 8,5

Answer: C



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37. At standard conditions, if the change in the enthalpy for the following reaction is -109KJmol^{-1}



Given that bond energy of H_2 and Br_2 is 435kJmol^{-1} and 192kJmol^{-1} , respectively. What is the bond energy of HBr ?

A. 368

B. 736

C. 518

D. 259

Answer: A



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38. The minimum pressure required to compress $600dm^3$ of a gas at 1bar to $150dm^3$ at $40^\circ C$ is

A. 4.0^-

B. 0.2^-

C. 1.0^-

D. 2.5^-

Answer: A



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39. What is the role of gypsum, $CaSO_{4.2}H_2O$ in setting of cement? Identify the correct option from the following:

A. to fasten the setting process

B. to provide water molecules for hydration process

C. to help to remove water molecules

D. to slow down the setting process

Answer: D



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40. Which of the following oxide is amphoteric in nature?

A. SnO_2

B. SiO_2

C. GeO_2

D. CO_2

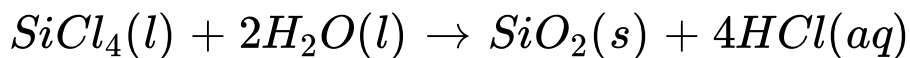
Answer: A



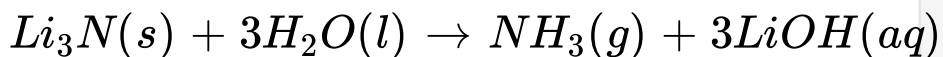
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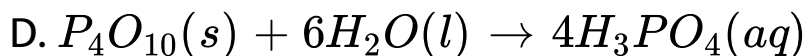
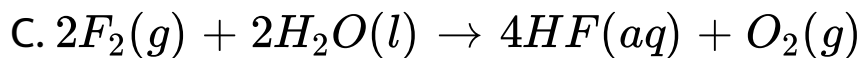
41. Which one of the following reactions does not come under hydrolysis type reaction?

A.



B.





Answer: C



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42. Which one of the following compounds shows both, Frenkel as well as Schottky defects?

A. AgBr

B. AgI

C. NaCl

D. ZnS

Answer: A



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43. One mole of carbon atom weighs 12g, the number of atoms in it is equal to, (Mass of carbon-12 is $1.9926 \times 10^{-23}g$)

A. 1.2×10^{23}

B. 6.022×10^{22}

C. 12×10^{22}

D. 6.022×10^{23}

Answer: D



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44. Isotonic solutions have same

- A. vapour pressure
- B. freezing temperature
- C. osmotic pressure
- D. boiling temperature

Answer: C



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45. The solubility product for a salt of the type AB is 4×10^{-8} . What is the molarity of its standard solution?

A. $2 \times 10^{-4} \text{ mol} / L$

B. $16 \times 10^{-16} \text{ mol} / L$

C. $2 \times 10^{-16} \text{ mol} / L$

D. $4 \times 10^{-4} \text{ mol} / L$

Answer: A



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