

## **CHEMISTRY**

## **BOOKS - GRB CHEMISTRY (HINGLISH)**

## **IONIC EQUILIBRIUM**

Others

1. The following equilibrium is established when hydrogen

chloride is dissolved in acetic acid

$$HCl(aq) + CH_3COOH(aq) \Leftrightarrow Cl^-(aq) + CH_3COOH_2^+(aq)$$

The set that characterises the conjugate acid-base pairs is:

A.  $(HCl, CH_3COOH)$  and  $(CH_3COOH_2^+, Cl^-)$ 

- B.  $(HCl, CH_3COOH_2^+)$  and  $(CH_3COOH, Cl^-)$
- C.  $(CH_2COOH_2^+, HCl)$  and  $(Cl^-, CH_3COOH)$
- D.  $(HCl, Cl^{-})$  and  $(CH_3COOH_2^{+}, CH_3COOH)$

## **Answer: D**



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