

# CHEMISTRY

## NCERT - NCERT CHEMISTRY (GUJRATI)

### CHEMICAL BONDING

#### Problem

1. Calculation of lattice enthalpy of  $MgBr_2$  from the given data



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## Questions E Explain Briefly On The Following

1. Calculate the lattice enthalpy of  $CaCl_2$  given that the enthalpy of:

i) Sublimation of Ca in  $121.8 \text{ kJ mol}^{-1}$

ii) Dissociation of  $Cl_2$  to  $2Cl$  is  $242.8 \text{ kJ mol}^{-1}$

iii) Ionisation of  $Ca$  to  $Ca^{2+}$  is  $2422 \text{ kJ mol}^{-1}$

iv) Electron gain for

$Cl$  to  $Cl^-$  is  $-355 \text{ kJ mol}^{-1}$

v)  $\Delta H_f^{(o)}$  overall is  $-795 \text{ kJ mol}^{-1}$



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