



CHEMISTRY

BOOKS - OSWAAL PUBLICATION CHEMISTRY (KANNADA ENGLISH)

Sample Paper 2

Exercise

1. What is the effect of rise in temperature on the solubility of gases in liquids?



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2. Define osmotic pressure.

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3. Mention the concentration of H^+ ions in the solution used in SHE.

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4. $2SO_2(g) + O_2(g) \xrightarrow{NO_2(g)}$ Is this reaction an example for Homogeneous or Heterogeneous catalysis.

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5. Name the depressant used in separation of ZnS from PbS by froth floatation process.

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6. Which noble gas does not occur in nature?

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7. $R - X + NaI \xrightarrow{\text{Dry Acetone}} RI + NaX$. This reaction is known as.....

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8. Give reason : Acetic acid is soluble in water.

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9. Among the following which is a fat soluble vitamin

Vitamin - B_{12} , Vitamin - C, Vitamin - D.

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10. Give two differences between p-type & n-type semiconductors.

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11. What is limiting molar conductivity? Represent graphically the variation in molar conductivity with concentration for acetic acid.

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12. Rate constant of a first order reaction is $6.93 \times 10^{-3} \text{min}^{-1}$. Calculate the half-life period.

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13. (i) What is actinide contraction?

(ii) Which is the common oxidation state exhibited by actinides?

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15. How do you prepare diethyl ether by dehydration of ethanol.

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16. How do you convert benzamide to benzoic acid?

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17. Give one example each for (a) Tranquilizer (b) Antiseptic.

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18. What is saponification with an example.

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19. How copper is refined by electrolytic method?

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20. (i) Write the structure and mention basicity of hypophosphorous acid.

(ii) Which gas is liberated when zinc reacts with dil. HNO_3 ?

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22. Draw the flow chart for the manufacture of sulphuric acid by Contact process. Name the catalyst used in the process.

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23. (a) Give two reasons for the anomalous behaviour of fluorine.
(b) Give one example of interhalogen compounds.

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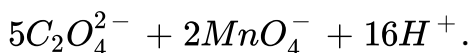
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25. Write any two characteristics of interstitial compounds.

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26. (i) Write the two chemical equations to show the inter conversion of chromates & dichromates in aqueous solution.

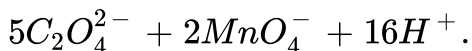
(ii) Complete the equation:



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27. (i) Write the two chemical equations to show the inter conversion of chromates & dichromates in aqueous solution.

(ii) Complete the equation:



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28. With the help of valence bond theory account for the geometry and magnetic property of $(\text{Co}(\text{NH}_3)_6)^{3+}$.

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29. What is an ambidentate ligand ?

Name the type of structural isomerism that arises in the coordination compound containing such a ligand.

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30. Give the IUPAC name of $K_2[Zn(OH)_4]$.

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31. (a) Calculate the packing efficiency in c ccp crystal lattice.

(b) What is the number of particles per unit cell of a simple cube.

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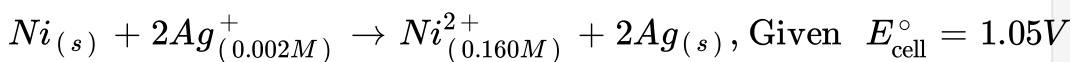
33. Calculate the osmotic pressure of 0.05% urea solution in water at 20°C . Given $R = 0.0821 \text{ atm mol}^{-1}\text{K}^{-1}$. Molar mass of urea = 60 g mol^{-1} .

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34. Give two general characteristics of an ideal solution of two liquids.

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35. Calculate the e.m.f. of the cell in which the following reaction takes place.



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36. A galvanic cell after use is recharged by passing current through it. What type of cell is it? Give an example.

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37. Rate constant of reaction at 300 K and 400 K are $0.0345S^{-1}$ and $0.1365S^{-1}$ respectively. Calculate the activation energy for the reaction.

[Given : $R = 8.314JK^{-1}mol^{-1}$]

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38. Derive an expression for half-life of zero order reaction.

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39. (a) Mention two applications of adsorption.

(b) What are emulsions? Give an example for O/W emulsion.

(c) What is the cause for Brownian movement?

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42. Explain $S_N - 2$ reaction mechanism ?

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43. Name the product formed when chloromethane reacts with aqueous KOH and Alcoholic AgCN.

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44. Give an example of polyhalogen compound.

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45. Explain esterification reaction between acetic acid and ethyl alcohol as example.

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46. Boiling point of alcohol is greater than the boiling point of hydrocarbons of comparable molar masses. Why?

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47. What is the effect of the group on the acidity of phenols?

– NO_2

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48. Explain Etards reaction.

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49. Name the reagent used in the decarboxylation of carboxylic acid.

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50. How do you convert benzene diazonium chloride into chlorobenzene. Name the reaction.

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51. Explain Hofmann's bromamide reaction with an example.

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52. Write Haworth structure for maltose.

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53. What are hormones? Give one biological function of insulin.

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54. What are nucleosides?

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55. Name the monomers of Nylon-6,6.

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56. How is neoprene prepared ?

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57. Give an example of thermoplastic polymer.

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