



CHEMISTRY

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CHEMISTRY (GUJRATI ENGLISH)

QUESTION ASKED IN NEET - 2018

Mcq

1. A mixture of 2.3 g formic acid and 4.5 g oxalic acid is treated with conc. H_2SO_4 . The

evolved gaseous mixture is passed through KOH pellets. Weight (in g) of the remaining product at STP will be

A. 4.4

B. 1.4

C. 2.8

D. 3

Answer: C



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2. In which case is the number of molecules of water maximum ?

A. 10^{-3} mole of water

B. 18 ml of water

C. 0.00224 L of water vapours at 1 atm and
273 K

D. 0.18 g of water

Answer: B



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3. Among CaH_2 , BeH_2 , BaH_2 the order of ionic character is :



Answer: B



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4. Which one is a wrong statement ?

A. The value of m for d_z^2 is zero.

B. Total orbital angular momentum of electron in 's' orbital is equal to zero.

C. The electronic configuration of N atom is



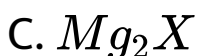
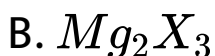
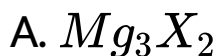
D. An orbital is designated by three quantum numbers while an electron in an atom is designated by four quantum numbers.

Answer: C



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5. Magnesium reacts with an element (X) to form an ionic compound. If the ground state electronic configuration of (X) is $1s^2 2s^2 2p^3$, the simplest formula for this compound is



D. MgX_2

Answer: A



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6. Which one of the following elements is unable to form MF_6^{3-} ion ?

A. In

B. Ga

C. B

D. Al

Answer: C



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7. The correct order of atomic radii in group 13 elements is

A. B It Ga It Al It In It Tl

B. B It Al It In It Ga It Tl

C. B It Ga It Al It Tl It In

D. B It Al It Ga It In It Tl

Answer: A::D



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8. Consider the following species :

CN^+ , CN^- , NO and CN

Which one of these will have the highest bond order?

A. CN

B. NO

C. CN^+

D. CN^-

Answer: D



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9. In the structure of CIF_3 , the number of lone pairs of electrons on central atom 'Cl' is

A. three

B. one

C. four

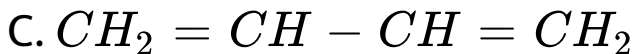
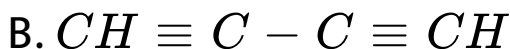
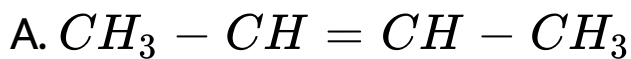
D. two

Answer: D



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10. Which of the following molecules represents the order of hybridisation sp^2 , sp^2 , sp , sp from left to right atoms?



Answer: D



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11. Given van der Waals constant for NH_3 , H_2 , O_2 and CO_2 are respectively 4.17,

0.244, 1.36 and 3.59, which one of the following gases is most easily liquefied ?

A. CO_2

B. NH_3

C. O_2

D. H_2

Answer: B



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12. The correction factor 'a' to the ideal gas equation corresponds to

A. forces of attraction between the gas molecules.

B. density of the gas molecules.

C. electric field present between the gas molecules.

D. volume of the gas molecules.

Answer: A

13. The bond dissociation energies of X_2 , Y_2 and XY are in the ratio of 1: 0.5 : 1. ΔH for the formation of XY is -200kJmol^{-1} . The bond dissociation energy of X_2 will be

A. 400kJmol^{-1}

B. 200kJmol^{-1}

C. 800kJmol^{-1}

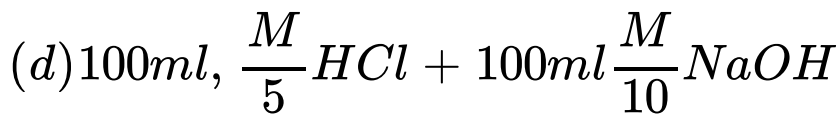
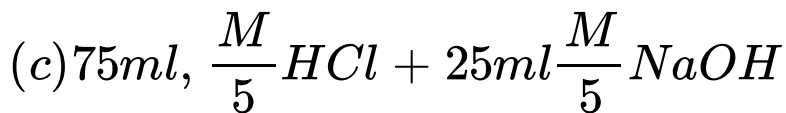
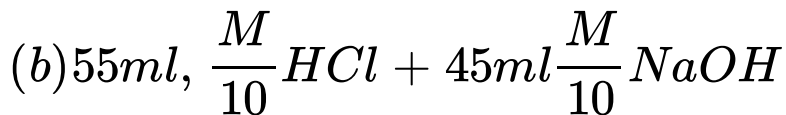
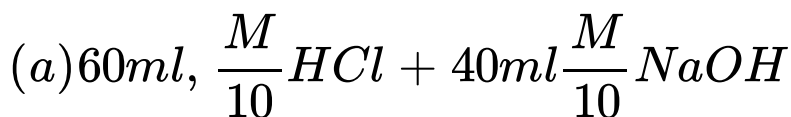
D. 100kJmol^{-1}

Answer: C



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14. Following solutions were prepared by mixing different volumes of NaOH and HCl of different concentrations:



pH of which one of them will be equal to 1?

A. c

B. b

C. d

D. a

Answer: A



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15. The solubility of $BaSO_4$ in water is $2.42 \times 10^{-3} gL^{-1}$ at 298 K. The value of its

solubility product (K_{sp}) will be (Given molar mass of $BaSO_4 = 233 \text{ g mol}^{-1}$)

A. $1.08 \times 10^{-8} \text{ mol}^2 \text{ L}^{-2}$

B. $1.08 \times 10^{-10} \text{ mol}^2 \text{ L}^{-2}$

C. $1.08 \times 10^{-14} \text{ mol}^2 \text{ L}^{-2}$

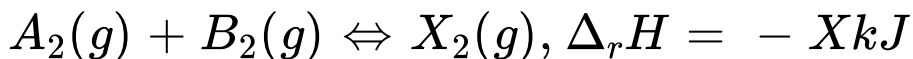
D. $1.08 \times 10^{-12} \text{ mol}^2 \text{ L}^{-2}$

Answer: B



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16. Which one of the following conditions will favour maximum formation of the product in the reaction,



- A. High temperature and low pressure
- B. Low temperature and high pressure
- C. High temperature and high pressure
- D. Low temperature and low pressure

Answer: B



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