



# CHEMISTRY

**BOOKS - JEEVITH PUBLICATIONS**

**CHEMISTRY (KANNADA ENGLISH)**

**MODEL QUESTION PAPER 4 FOR  
PRACTICE**

**Part A**

1. Mention the sign of enthalpy of mixing for a solution showing positive deviation from Raoult's law.



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2. What happens to the solubility of a gas in a liquid on increasing temperature.



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3. Write the SI unit of molar conductance.



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4. What is the rate determining step in a multiple step reaction?



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5. Why is adsorption always exothermic?



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6. Give an example for sulphide ore of copper.



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7. Mention of noble gas not present in air.



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8. Mention any one use of iodoform.



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9. Write the IUPAC name of t-butyl bromide.



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10. Name a basic amino acid.



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**Part B**

1. Calculate the number of particles per unit cell in bcc.



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2. Give the relation between conductivity and molar conductivity.



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3. What is pseudo first order reaction? Give an example.



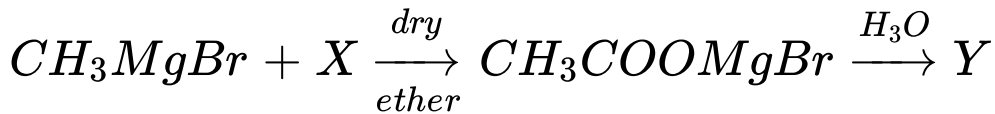
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4. What are lathanides? Give an example.



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5. Write the name of X and Y in the following reacton:



.



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6. Write the aldol condensation reaction of acetaldehyde.



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7. What are monosaccharides ? Give an example.





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8. Mention the difference between antiseptic and disinfectant.



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## Part C

1. Explain the process of obtaining "blister copper" from "copper matte" with equations.



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2. Draw the flow chart for the manufacture of sulphuric acid by Contact process. Name the catalyst used in the process.



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3. Give reason " $H_2O$ " is liquid while  $H_2S$  is gas.



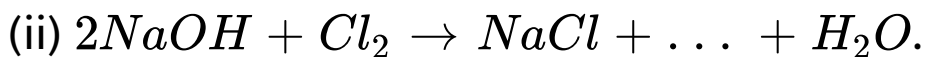
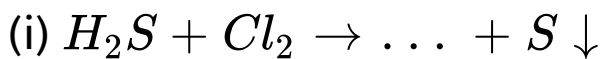
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4. Write any three anomalous properties of oxygen with respect to other element of 16th group.



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5. Complete the following equations:



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6. Give reason ''  $BrF_5$  Is more reactive than  $Br_2$  ''



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7. Calculate the magnetic moment of  $Cr^{+3}$ .



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8. Which has got smallest size

$Fe, Fe^{+2}$  &  $Fe^{3+}$  ?



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9. Write the chemical reactions in the preparation of potassium dichromate from chromite ore.



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10. Using VBT, explain the geometry and magnetic property of  $[Ni(CN)_4]^{-2}$ . (Atomic Number of Ni=28).





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11. Explain linkage isomerism with example.



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12. Give an example for homoleptic coordination compound.



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1. Calculate the packing efficiency of bcc crystal lattice.



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2. What is the Schotky defect? How does it affect the density of the crystal?



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3. A solution containing 12.5 g of non-electrolyte solution in 175g of water gave a boiling point elevation of 0.7 k. calculate the molar mass of the solute if  $K_b$  for water is  $0.52 \text{ K kg mol}^{-1}$ .



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4. State Raoult's law of relative lowering of vapour pressure of a solution.



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5. Derive an expression for the rate constant of a zero order reaction.



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6. Calculate the half-life period of a first order reaction if rate constant is  $0.45 \text{ min}^{-1}$ .



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7. Write any three differences between lyophilic and lyophobic colloids.



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8. Explain the region of delta formation on the basis of coagulation.



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9. How is SHE constructed? Mention its use.



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## Part E

1. Discuss the mechanism of  $SN^1$  reaction with suitable example.



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2. Give an example for Finkelstein reaction.



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3. Explain the mechanism of dehydration of ethanol to ethene.



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4. How is salicylic acid converted into aspirin?  
Give reaction.



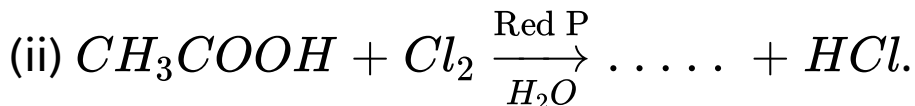
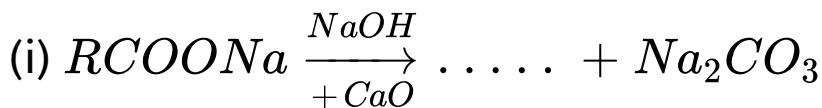
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5. Write Cannizaro's reaction of Benzaldehyde.



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6. Complete the following equations.



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7. Give reason: (i) Methyl amine is soluble in water but not in aniline.

(ii) Methyl amine is more basic than ammonia.



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8. how do you convert aniline to benzene diazonium chloride ?



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9. What is non-reducing sugar? Give an example.



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10. Draw the Haworth structure of  $\beta$ -(D) fructofuranose.



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**11.** Why is vitamin-C not stored in the body?



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**12.** What is copolymerization? Give example.



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**13.** Write any one difference between thermoplastic and thermosetting plastic.





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**14.** Identify homopolymer among Nylon-6,6 and Nylon-6.



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