



CHEMISTRY

BOOKS - ICSE

ELEMENTS AND COMPOUNDS

Check Your Progress Fill In The Blanks

1. Assertion: Metals are good conductors of electricity

Reason:Metals conduct electricity in solid as well as in molten state



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2. Elements that do not react chemically are called _____



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3. Elements are represented by _____



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4. The symbol of sodium is _____



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5. The elements present in common salt are _____ and _____



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Exercises Tick The Most Appropriate Answer

1. Which of the following is a metal?

A. sulphur

B. chlorine

C. gold

D. bromine

Answer:



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2. Which of the following is a non-metal ?

A. silver

B. iodine

C. copper

D. iron

Answer:



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3. Which of the following is a metalloid?

A. copper

B. silver

C. oxygen

D. silicon

Answer:



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4. What is the symbol for Iron →

A. I

B. Fe

C. In

D. Na

Answer:



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5. Atoms of different elements combine to form a molecule of

- A. an element
- B. a compound
- C. both a and b
- D. cannot say

Answer:



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Exercises Fill In The Blanks

1. Pure substances are classified into ____ and ____



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2. The _____ of an element is the abbreviation of its full name.



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3. Natrium is the Latin name of _____



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4. The symbol of cobalt is _____



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5. The molecular formula of water is _____



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Exercises Matching

1. Match the columns

- | | |
|-----------------------------------|--------------|
| 1. Smallest unit of a compound | a. molecule |
| 2. Sulphur | b. noble gas |
| 3. Helium | c. non-metal |
| 4. A diatomic molecule | d. carbon |
| 5. An element having valency four | e. hydrogen |
| | f. metalloid |



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Exercises Write True Or False Correct The False Statements

1. Which of the following statements are true for pure substances?

(i) Pure substances contain only one kind of particles.

(ii) Pure substances may be compounds or mixtures.

(iii) Pure substances have the same composition throughout.

(iv) Pure substance can be exemplified by all elements other than nickel.



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2. Some metals exist as gases.



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3. Sulphur is a metalloid.



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4. Compounds are substances made up of two or more elements.



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5. A molecule of hydrogen is made up of three atoms of hydrogen.



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Exercises Name The Following

1. Elements that do not react chemically are called _____



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2. Elements showing properties of both metals and non-metals



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3. The smallest particle of an element or a compound which has independent existence



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4. The number of atoms present in one molecule of an element is called its_____



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5. The combining capacity of an element



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Exercises Write The Symbols

1. Helium



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2. Boron



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3. Carbon



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4. Neon



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5. Sodium



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6. Chlorine



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7. Silver



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8. Gold



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9. Oxygen



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10. Potassium



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Exercises Symbols Of Some Elements Are Given Below Name The Elements

1. Be



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2. F



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3. Al



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4. Si



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5. Ar is symbol of



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6. Ca



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Exercises Name The Constituents Of The Following Compounds

1. Carbon dioxide is a compound.



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2. Iron sulphide



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3. Common salt is



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4. Ammonia



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Exercises Write The Chemical Formulae Of The Following Compounds

1. Water has the formula



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2. Carbon dioxide is a compound.



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3. Formula of Zinc sulphide



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4. Calcium oxide is



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Exercises Answer The Following In Short

1. What are the two types of pure substances?



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2. What is a compound?



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3. What is a molecule?





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4. Which can exist independently - atoms or molecules?



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5. What do you understand by the term 'triatomic molecule'? Give an example of the same.



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6. What is a molecular formula?



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7. What are the things required for writing the molecular formula of a compound?



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Exercises Answer The Following In Detail

1. What are the four groups into which all the elements are classified? Name two of each group



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2. What is meant by the symbol of an element?
What is its significance?



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3. State three characteristics of a compound.



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4. State two differences between an element and a compound.



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5. State three characteristics of a molecule.



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Think And Answer

1. The atoms of gold are different from the atoms of silver. True or false? Give reason.



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2. Why are the noble gases monoatomic?



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3. What is the importance of a chemical formula? State two points.



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4. The chemical formulae of two compounds can be same. Do you agree?



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