CHEMISTRY

BOOKS - ICSE

LANGUAGE OF CHEMISTRY

Intext Questions

1. Define the term catalyst. Write one example where it is used.

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2. Name the rectants and product in the skeletal equation and balance

$Na + Cl_2 \rightarrow NaCl$

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3. Name the rectants and product in the skeletal equation and balance

$N_2 + H_2 \rightarrow NH_3$

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4. Name the rectants and product in the skeletal equation and balance

$Na + O_2 \rightarrow Na_2O$

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5. Name the rectants and product in the skeletal equation and balance

$Fe + H_2O \rightarrow Fe_3O_4 + H_2$

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6. Name the rectants and product in the skeletal equation and balance

$Na + H_2O \rightarrow NaOH + H_2$

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Exercise

1. What is a chemical reaction?

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2. What is a chemical reaction ?

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3. Why do we need to balance chemical equations?

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4. State the conditions necessary for a chemical change or reaction to take place.

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5. Differentiate between :

Reactants and products

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6. Differentiate between :

Chemical reaction and chemical equation

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7. Differentiate between :

A balanced and a skeletal chemical equation.

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8. Balance the following equation :

$KClO_3 \rightarrow KCl + O_2$

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9. Write word equations for the skeletal equation:

$Zn + HCl \rightarrow ZnCl_2 + H_2$

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10. Write word equations for the skeletal equation:

$Fe + Cl_2 \rightarrow FeCl_2$

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11. Write word equations for the skeletal equation:

$CO + O_2 \rightarrow CO_2$

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12. Write word equations for the skeletal equation:

$Ca + O_2 \rightarrow CaO$

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13. Write word equations for the skeletal equation:

$Na + O_2 \rightarrow Na_2O$

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14. Write word equations for the skeletal equation:

$NaOH + H_2SO_4 \rightarrow Na_2SO_4 + H_2O$

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15. Write word equations for the skeletal equation:

$AgBr \rightarrow Ag + Br_2$

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16. Write word equations for the skeletal equation:

$KNO_3 \rightarrow KNO_2 + O_2$

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17. Balance the chemical equation:

$FeS + HCl \rightarrow FeCl_2 + H_2S$

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18. Balance the chemical equation:

$$Na_2CO_3 + HCl \rightarrow NaCl + H_2O + CO_2$$

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19. Balance the chemical equation:

$$H_2O_2 \rightarrow H_2O$$

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20. Balance the chemical equation:

Na _(2) O + H _(2)O to NaOH`

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21. By direct combination

$H_2 + Cl_2 \rightarrow$ _____

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22. Write your observations for the chemical reactions and name the product formed :

When sugar is heated.

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23. Write your observations for the chemical reactions and name the product formed :

When manganese dioxide is added to potassium chlorate and heated.

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24. Write your observations for the chemical reactions and name the product formed :

When dilute acetic acid is poured on baking soda.

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25. Write your observations for the chemical reactions and name the product formed :

When an aqueous solution of sodium chloride

is mixed with an aqueous solution of silver nitrate.

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26. Write your observations for the chemical reactions and name the product formed :

When ammonium chloride is heated with sodium hydroxide.

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27. Write your observations for the chemical reactions and name the product formed :

When water is added to quick lime?

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28. Write skeletal equations for the following word equations.

calcium carbonate $\rightarrow$ calcium oxide + carbon dioxide

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29. Write a chemical equation for the following word equation and balance them.

Carbon + Oxygen $\rightarrow$ Carbon dioxide

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30. Write symbolic representation for the word equations and balance

Calcium oxide + Water $\rightarrow$ Calcium hydroxide

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31. Write symbolic representation for the word equations and balance

Aluminium + Chlorine $\rightarrow$ Aluminium chloride

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32. Write symbolic representation for the word equations and balance

Iron + Sulphur $\rightarrow$ Iron sulphide

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33. Write symbolic representation for the word equations and balance

Sodium carbonate + Hydrochloric acid (dil) $\rightarrow$ Sodium chloride + Water + Carbon dioxide

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34. Write symbolic representation for the word equations and balance

Barium chloride + Sodium sulphate $\rightarrow$ Barium sulphate + Sodium chloride

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35. Write symbolic representation for the word equations and balance

Iron sulphide + Hydrochloric acid (dil) $\rightarrow$ Iron (II) chloride + Hydrogen sulphide

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36. Write symbolic representation for the word equations and balance

Calcium + Water $\rightarrow$ Calcium hydroxide

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Objective Type Questions Fill In The Blanks

1. The substances which undergo chemical change are called .....

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2. The substances formed as a result of a chemical reaction are called.....

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3. During a chemical reaction, transfer of ..... Takes place.

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4. The basic condition necessary for a chemical reaction is .....

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5. In some chemical reactions, an insoluble..... Is formed when two solutions are mixed.

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Objective Type Questions True Or False

1. No new substance is formed during a chemical reaction.....

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2. When potassium iodide solution is added to lead acetate solution, a red precipitate is formed. .....

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3. A black residue is formed when sugar is heated. .....

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4. When iron and sulphur are heated together a grey mass is formed which is attracted by a magnet. .....

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Objective Type Questions Multiple Choice Questions

1. A chemical equation is a statement that describes a chemical change in terms of

A. Symbols and formulae

B. Energy

C. Number of atoms

D. Colours

Answer:

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2. Balancing a chemical equation is based on

A. Law of conservation of mass

B. Mass of reactants and products

C. Symbols and formulae

D. None of the above

Answer:

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3. Copper carbonate when heated, it turns :

A. Blue

B. Green

C. Black

D. Yellow

Answer:

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4. When lead acetate solution is added to potassium iodide solution, a precipitate is formed which is

A. Red

B. Yellow

C. White

D. Black

Answer:

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5. The gas which has a rotten egg smell is:

A. Ammonia

B. Hydrogen

C. Hydrogen sulphide

D. Oxygen

Answer:

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6. When sodium carbonate reacts with dilute hydrochloric acid, the gas evolved is

A. Carbon-dioxide

B. Nitrogen

C. Oxygen

D. Hydrogen

Answer:

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