



CHEMISTRY

BOOKS - JNAN PUBLICATION

MATTER (STRUCTURE OF MATTER)

Example

1. In 1903 who had proposed the model of an atom, due to which electrons and protons were known to us?

A. Rutherford

B. Dalton

C. Goldstein

D. Joseph J Thomson

Answer:



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2. The present concept of the structure of an atom is given by which scientist?

A. Joseph J Thomson

B. Goldstein

C. Niles Bohr

D. Rutherford

Answer:



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3. Which model does not be able to explain the stability of an atom?

- A. Thomson's model
- B. Rutherford's model
- C. Bhor's model
- D. none of the above

Answer:



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4. Name an atom in which the molecules of that atom does not contain any neutrons?

A. Hydrogen

B. Oxygen

C. Sodium

D. Phosphorous

Answer:



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5. Electron was discovered by _____

A. Goldstein

B. J J Thomson

C. Niles Bohr

D. Rutherford

Answer:



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6. The correct electronic configuration of sodium is _____

A. 8,2,1 ?,

B. 2,1,8?,

C. 2,8,2 ?,

D. 2,8,1?

Answer:



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7. Who discovered neutron in 1953?

A. Nile Bohr

B. James Chadwick

C. Rutherford

D. Goldstein

Answer:



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8. The atom has _____

A. Positive charge

B. negative charge

C. some times positive and sometimes

negative

D. no charge

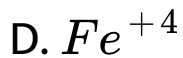
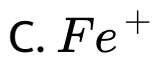
Answer:



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9. Symbo of Ferric ion is _____





Answer:



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10. Charged atoms are called _____

A. isotope

B. isobar

C. isotone

D. ion

Answer:



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11. The total number of protons and neutrons in an atom is called its _____

A. mass number

B. atomic number

C. mass

D. none of the above

Answer:



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12. State whether True or False:-

Nearly the whole mass of an atom remains concentrated at its centers.



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13. State whether True or False:-

John Dalton's experiment proved the existent of electrons in 1820.



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14. State whether True or False:-

Neutrons are electrically neutral.



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15. State whether True or False:-

An iron atom is heavier than a gold atom.



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16. State whether True or False:-

In a solid the average distance between the molecules is much larger than those in a liquid.



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17. State whether True or False:-

The kinetic energy of molecules in a liquid increases with increases of temperature.



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18. State whether True or False:-

Salt (NaCl) melts at a temperature of about $800^{\circ}C$



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19. State whether True or False:-

Ionic compounds are made of molecules.



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20. State whether True or False:-

Liquids have definite shape, but they have not definite volume.



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21. State whether True or False:-

The mass of an atom can not be measured directly by any balance.



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22. Fill in the blanks: -

In 1808 _____ published his atomic theory.



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23. Fill in the blanks: -

The smallest particle of an element that retains the properties of the element is called on_____.



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24. Fill in the blanks: -

Atoms of the same element possessing unequal number of neutrons are called_____.





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25. Fill in the blanks: -

Electrons are _____ charged.



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26. Fill in the blanks: -

The number of protons in an atom is called its

_____.



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27. Fill in the blanks: -

The total number of protons and neutrons in an atom is called its _____.



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28. Fill in the blanks: -

The _____ has mass but no charge.



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29. Fill in the blanks: -

Atoms gain or lose _____ to attain stability.



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30. Fill in the blanks: -

When an atom gains or loses an electron, an _____ is formed.



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31. Fill in the blanks: -

Solids have their own shape and _____.



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32. Fill in the blanks: -

Atoms of different elements differ in mass and _____ properties.



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33. Match the column A with column B.

Column A	Column B
a) Formula of Ammonium	i) CaCO_3
b) Niels Bohr is a	ii) roughly views like solar system
c) Formula of Calcium carbonate	iii) has many valency
d) Valency of Aluminium is	iv) has 4 protons.
e) The atomic model	v) Plasma
f) In Cuprous oxide	vi) NH_4^+
g) Iron	vii) Valency of copper is 1.
h) Nuclars of an oxygen atom	viii) 3
i) Fourth state of a substance is	ix) renowned atom scientist



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34. What does the word 'atoms' mean in Greek language?



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35. Who is the father of atomic theory?



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36. In which year atomic theory of John Dalton was published?



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37. What are an atom made up of ?



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38. Who discovered neutron?



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39. What is atomic number?



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40. What is mass number?



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41. What is Isotope?



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42. What is Isobar?



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43. What are ions?



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44. What are cations?



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45. What are anions?



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46. What is Valency of an element?



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47. How the pleasant smell of incense sticks spreads through a room?



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48. How can you prove that molecules can spread through liquids?



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49. What are the characteristics of atoms and molecules in a solid?



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50. What are the characteristics of molecules in a liquid?



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51. What are the characteristics of molecules in a gas?



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52. What were Rutherford's general conclusions about the atom?



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53. Describe Dalton's Atomic Theory.



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54. Describe the structure of electrovalent compound NaCl.



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55. What happens when the temperature of a solid is increased?



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56. What happens when the temperature of a liquid is increased?



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