



# CHEMISTRY

## BOOKS - MAXIMUM PUBLICATION

### MODEL PAPER 1

#### Example

1. Who proposed the planetary model of atom?

A. J J Thomson

B. Rutherford

C. Bohr

D. Dalton

**Answer: B**



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2. The arrangement of eight electrons in the outermost shell of an atom is called \_\_\_\_\_



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3. The Catalyst used to decrease the rate of decomposition of hydrogen peroxide is \_\_\_\_\_



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4. The group number of noble gases in the modern periodic table is \_\_\_\_\_



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5. Mainly the mass of an atom depends up on the mass of (protons and neutrons, neutrons and electrons)



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6. The atomic number of atom is 13. Its mass number is 27.

Write the number of electron and neutrons in the atom.



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7. The atomic number of atom is 13. Its mass number is 27.

Write the electron configuration of this element.



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8. Which type of bonding exist in MgO (hint atomic no. of Mg = 120 = 8)



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**9.** Which atom loses electron? Which atom gain electron?



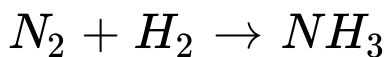
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**10.** HCl is a polar compound, Why HCl shows polar nature?



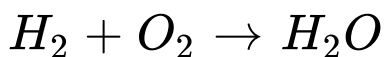
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**11.** Balance the following equations.



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**12.** Balance the following equations.



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**13.** The electronic configuration of some elements are given below.

$$12A = 2, 8, 2$$

$$6B = 2, 4$$

$$16C = 2, 9, 4$$

$$13D = 2, 8, 3$$

Which of these belong to the same group?



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**14.** The electronic configuration of some elements are given below.

$$12A = 2, 8, 2$$

$$6B = 2, 4$$

$$16C = 2, 9, 4$$

$$13D = 2, 8, 3$$

Which of these belong to same period?



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15. Match the following suitably

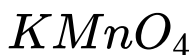
Contribution	Name of Scientist
Modern periodic law	Linus Pauling
Electronegativity	Lavosier
Law of conservation of mass	Mosely



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16. Find out the oxidation number of Mn in the following compounds

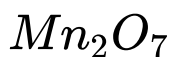
Hint: Oxidation No of  $K = +1, O = -2$



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**17.** Find out the oxidation number of Mn in the following compounds

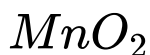
Hint: Oxidation No of  $K = +1, O = -2$



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**18.** Find out the oxidation number of Mn in the following compounds

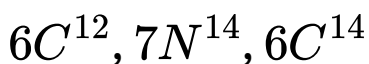
Hint: Oxidation No of  $K = +1, O = -2$





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**19.** The symbol of these atoms are given below.

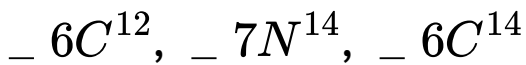


Which of these are isotopes?



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**20.** The symbol of these atoms are given below.



What is the basis of your selection?



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21. The symbol of these atoms are given below.



Write any use of an isotope



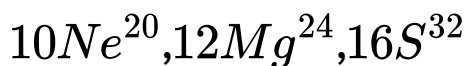
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22. Compare the properties of ionic and covalent compounds.



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23. The symbol of certain elements are given. Write their electronic configuration and find the period and group they belong

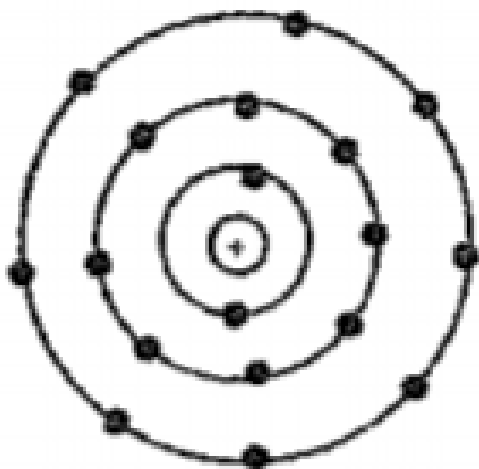


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24. The Bohr model of an atom is given below.

Analyse it and find out the following

Atomic number

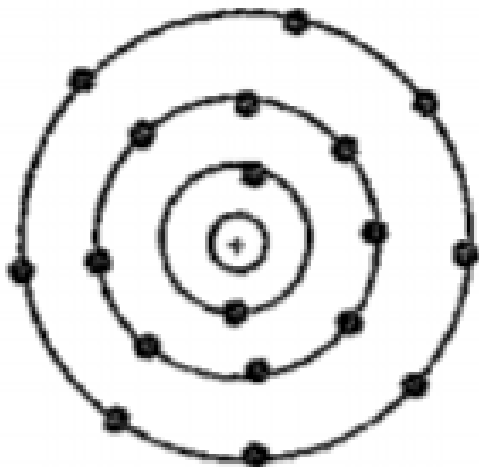


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25. The Bohr model of an atom is given below.

Analyse it and find out the following

Electron Configuration



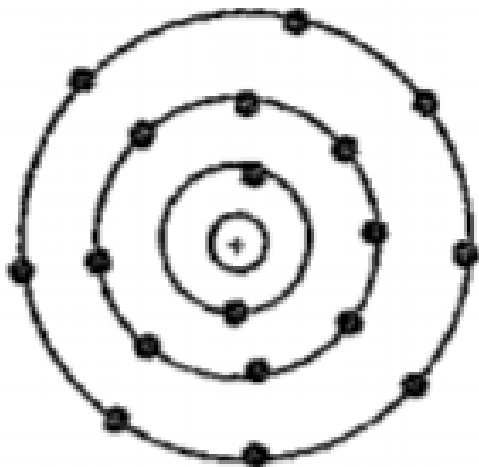
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26. The Bohr model of an atom is given below.

Analyse it and find out the following

The period and group of the element

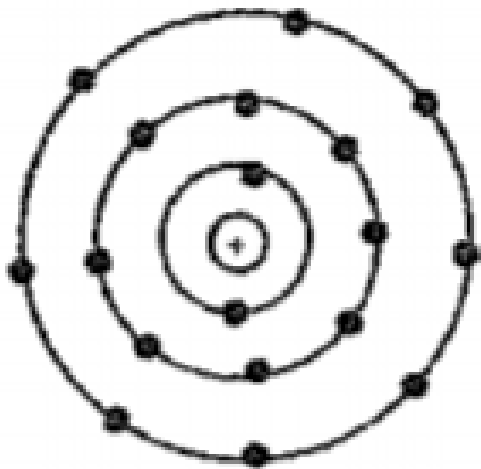


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27. The Bohr model of an atom is given below.

Analyse it and find out the following

The common name of the group.



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**28.** The electronegativity values of two elements are given below. (C=2.55 H = 2.20)

Which type of compound will be formed if these two elements combine together? Give reason.



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**29.** The electronegativity values of two elements are given below. (C=2.55 H = 2.20)

Draw the electron dot diagram of the chemical

bonding in  $CH_4$  molecule.

(At.np. C=6 H=1)



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**30.** What is the relationship between the rate of chemical reaction and temperature?



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**31.** Fill in the blanks in the following table showing how the properties varies down the

group and across the period in a periodic table.

Properties	Down the group	Across the period
Metallic nature		
Ionisation energy		
Size of the atom		
Electro-negativity		



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32. Write any two advantages and two limitations of the Mendeleev's periodic table.



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