CHEMISTRY

BOOKS - MAXIMUM PUBLICATION

MODEL PAPER 2

Example

1. Which scientist proposed the law of electrolysis?

A. John Dalton

B. Michel Faraday

C. Humphry Davy

D. William Crookes

Answer: B

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2. What Is the maximum number of electrons that can be accommodate.In the 'M' shell of an atom?

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3. Which of the following is a polar molecule?

A. MgO

B. NaCl

C. HCl

D. CH_4

Answer: C

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4. Which elements are called rare earths?

A. Lanthanoids

B. Actinoids

C. Halogens

D. Noble gases

Answer: A

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5. The valency of two elements are given below.

Al-3,O-2

The chemical formula of aluminium oxide is _____

A. AlO

B. Al_2O_3

C. Al_3O_2

D. AlO_2

Answer: B

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6. There are 11 electrons and 12 neutrons In an atom.

Find out the atomic number and mass number.

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7. There are 11 electrons and 12 neutrons In an atom.

Write the electron configuration of the element.

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8. What are isotopes?

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9. Which are the Isotopes of Hydrogen.

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10. The atomic numbers of certain elements are given below.

N=7, O=8, Cl=17

Identify the nature of covalent bonds In the molecules given below N_2

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11. The atomic numbers of certain elements are given below.

N=7, O=8, Cl=17

Identify the nature of covalent bonds In the molecules given below O_2

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12. The atomic numbers of certain elements are given below.

N=7, O=8, Cl=17

Identify the nature of covalent bonds In the molecules given below Cl_2

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13. What are the characteristics of transition elements?

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14. What are redox reactions?

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15. What is the relation between surface area and rate of chemical reaction?

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16. The atomic number of two elements are given below.

Mg = 12, O = 8

Write the electron configuration of these elements.

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17. The atomic number of two elements are given below.Mg = 12,O = 8

Which type of compound we be formed If these two elements combine together?

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18. The atomic number of two elements are given below.Mg = 12,O = 8

Draw the electron dot diagram of the compound formation.

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19. Analyse the table given below.

Fill up the blanks

Element	Atomic number	Valency
A	11	1
B	8	.....
C	17	.....

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20. Analyse the table given below.

Write the chemical formula of the compound

formed by A and B

Element	Atomic number	Valency
A	11	1
B	8	.....
C	17	.....

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21. Which of the following chemical equation are balanced?

$H_2 + O_2 \rightarrow H_2O$

$CO_2 + C \rightarrow 2CO$

$N_2 + H_2 \rightarrow NH_3$

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22. Balance the unbalanced chemical equations.

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23. There are 4 electrons in the M shell of an element.

Write the electron configuration of the element.

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24. What Is atomic number?

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25. Find out the period and group of the element in the periodic table with atomic number 14.

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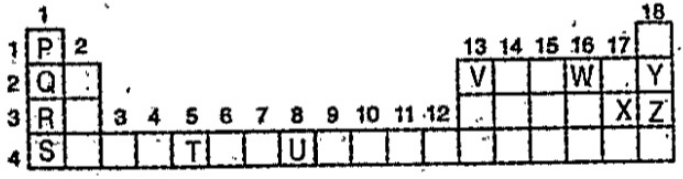
26. An incomplete form of the periodic table Is given below.(symbols are not real)

Which are transition elements?



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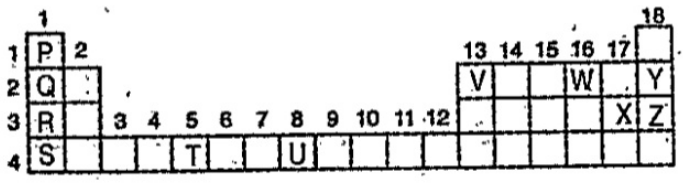
27. An incomplete form of the periodic table Is given below.(symbols are not real)

Which element has the highest Ionisation

energy?



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28. Find out the oxidation number of 'S' in the following compounds. (Oxidation number H = +1 O = -2)

H_2SO_4

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29. Find out the oxidation number of 'S' in thefollowing compounds. (Oxidation number H = +1 O = -2)

SO_2

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30. Analyse the following chemical equation and find out the element undergone oxidation, reduction, oxidising agent and reducing agent.

$S + O_2 \rightarrow SO_2$

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31. Sodium chloride Is an ionic compound and wax is a covalent compound. How can you compare the properties of these two?

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32. Write the main postulates of the Bohr model of atom.

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