CHEMISTRY

BOOKS - MAXIMUM PUBLICATION

MODEL PAPER 3

Example

1. Identity the relation and fill up the blank.

Law of electrolysis : _____

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2. Identity the relation and fill up the blank.

Law of conservation of mass : _____

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3. Which one of the following does not belong to the group.

(Mendeleev, Newlands, Dobreiner, Niels Bohr)

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4. Which of the noble gas does not have octet arrangement in the outermost shell?

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5. Which elements show similarity in chemical properties in groups and periods in the periodic table?

(Noble gases, Representative elements transition elements, Lanthanoids)

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6. Conducts electricity in molten state and when in solutions. This statement is applicable to what type of compounds?

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7. Name the elements included in the first period of the periodic table.

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8. Analyse the following statements and correct them if there is a mistake.

The mass of an atom mainly depends upon the mass of the electrons and neutrons present in it.

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9. Analyse the following statements and correct them if there is a mistake.

Down the group in .a periodic table ionisation energy decreases electronegativity increases.

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10. What are oxidation and reduction ?

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11. From the following statements

Choose those applicable to covalent compounds.

- Dissolves in organic compounds
- Dissolves in water
- High melting and boiling point.
- Found in solid liquid and gaseous states.
- Conducts electricity.

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12. The valency of iron (Fe) is 3 and oxygen (O) is 2.

Write the chemical formula of iron oxide

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13. The chemical formula of aluminium chloride is $AlCl_3$. Write the valencies of Al and Cl.

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14. Write the equation to find out the maximum number of electrons that can be accommodated in a shell.

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15. Find out the maximum number of electrons that can be accommodated in the third shell of an atom.

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16. The atomic number of magnesium is 12 and Fluo rine is 9.

Write the election configuration of these elements.

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17. The atomic number of magnesium is 12 and Fluo rine is 9.

Write the electron dote formula of the compound formed by these two elements.

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18. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55, 0=3..44 CI = 3.6 H=2.20, Na = 0.93

CO_2

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19. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55, 0=3..44 CI = 3.6 H=2.20, Na = 0.93

NaH

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20. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55,

0=3..44 Cl = 3.6 H=2.20, Na = 0.93

HCl

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21. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55, 0=3..44 Cl = 3.6 H=2.20, Na = 0.93

NaCl

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22. Balance the following chemical equations.

$Zn + HCl \rightarrow ZnCl_2 + H_2$

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23. Balance the following chemical equations.

$Mg + O_2 \rightarrow MgO$

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24. Balance the following chemical equations.

$SO_2 + O_2 \rightarrow SO_3$

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25. What is a catalyst?

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26. Write an experiment to prove the influence of a catalyst in a chemical reaction.

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27. What are representative elements?

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28. What are the general characteristics of these el ements.

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29. Find out the oxidation number of 'Cr' in the following compounds ?(Hint : oxidation number k = + 1 , 0 = -2)

Cr_2O_7

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30. Find out the oxidation number of 'Cr' in the following compounds ?(Hint : oxidation number k = + 1 , 0 = -2)

$K_2Cr_2O_7$

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31. Draw the electron dot diagram of the formation of CCl_4, molecule (Hint at . no. C =6 CI =17)

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32. The mass number of an atom is 32. There are 6 elec trons in the 'M' shell.

Write the electron configuration of this element.

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33. The mass number of an atom is 32. There are 6 elec trons in the 'M' shell.

What is the number of neutrons?

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34. The mass number of an atom is 32. There are 6 elec trons in the 'M' shell.

Draw the Bothh model of the atom.

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