



CHEMISTRY

BOOKS - MAXIMUM PUBLICATION

MODEL PAPER 3

Example

1. Identify the relation and fill up the blank.

Law of electrolysis : _____



Watch Video Solution

2. Identity the relation and fill up the blank.

Law of conservation of mass : _____



[Watch Video Solution](#)

3. Which one of the following does not belong to the group.

(Mendeleev, Newlands, Dobreiner, Niels Bohr)



[Watch Video Solution](#)

4. Which of the noble gas does not have octet arrangement in the outermost shell?



[Watch Video Solution](#)

5. Which elements show similarity in chemical properties in groups and periods in the periodic table?

(Noble gases, Representative elements
transition elements, Lanthanoids)



[Watch Video Solution](#)

6. Conducts electricity in molten state and when in solutions. This statement is applicable to what type of compounds?



[Watch Video Solution](#)

7. Name the elements included in the first period of the periodic table.



[Watch Video Solution](#)

8. Analyse the following statements and correct them if there is a mistake.

The mass of an atom mainly depends upon the mass of the electrons and neutrons present in it.



Watch Video Solution

9. Analyse the following statements and correct them if there is a mistake.

Down the group in .a periodic table ionisation energy decreases electronegativity increases.



[Watch Video Solution](#)

10. What are oxidation and reduction ?



[Watch Video Solution](#)

11. From the following statements

Choose those applicable to covalent compounds.

- Dissolves in organic compounds
- Dissolves in water
- High melting and boiling point.
- Found in solid liquid and gaseous states.
- Conducts electricity.



[Watch Video Solution](#)

12. The valency of iron (Fe) is 3 and oxygen (O) is 2.

Write the chemical formula of iron oxide



[Watch Video Solution](#)

13. The chemical formula of aluminium chloride is $AlCl_3$. Write the valencies of Al and Cl.



Watch Video Solution

14. Write the equation to find out the maximum number of electrons that can be accommodated in a shell.



Watch Video Solution

15. Find out the maximum number of electrons that can be accommodated in the third shell of an atom.



Watch Video Solution

16. The atomic number of magnesium is 12 and Fluorine is 9.

Write the electron configuration of these elements.



Watch Video Solution

17. The atomic number of magnesium is 12 and Fluorine is 9.

Write the electron dot formula of the compound formed by these two elements.



Watch Video Solution

18. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55, O=3.44 Cl = 3.6 H=2.20, Na = 0.93)





[Watch Video Solution](#)

19. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55, O=3.44 Cl = 3.6 H=2.20, Na = 0.93)

NaH



[Watch Video Solution](#)

20. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55,

O=3.44 Cl = 3.6 H=2.20, Na = 0.93

HCl



[Watch Video Solution](#)

21. Classify the following compounds as ionic and covalent. (Hint:- electronegativity C = 2.55,

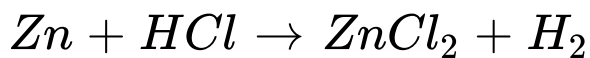
O=3.44 Cl = 3.6 H=2.20, Na = 0.93

NaCl



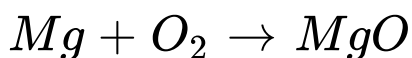
[Watch Video Solution](#)

22. Balance the following chemical equations.



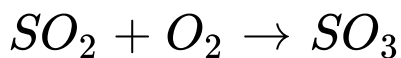
Watch Video Solution

23. Balance the following chemical equations.



Watch Video Solution

24. Balance the following chemical equations.



Watch Video Solution

25. What is a catalyst?



Watch Video Solution

26. Write an experiment to prove the influence of a catalyst in a chemical reaction.



Watch Video Solution

27. What are representative elements?



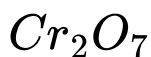
Watch Video Solution

28. What are the general characteristics of these elements.



View Text Solution

29. Find out the oxidation number of 'Cr' in the following compounds ?(Hint : oxidation number $k = + 1 , 0 = -2$)



Watch Video Solution

30. Find out the oxidation number of 'Cr' in the following compounds ?(Hint : oxidation number $k = + 1 , 0 = -2$)





[Watch Video Solution](#)

31. Draw the electron dot diagram of the formation of CCl_4 , molecule (Hint at . no. C =6
Cl =17)



[Watch Video Solution](#)

32. The mass number of an atom is 32. There are 6 electrons in the 'M' shell.

Write the electron configuration of this element.



[Watch Video Solution](#)

33. The mass number of an atom is 32. There are 6 electrons in the 'M' shell.

What is the number of neutrons?



[Watch Video Solution](#)

34. The mass number of an atom is 32. There are 6 electrons in the 'M' shell.

Draw the Bohr model of the atom.



Watch Video Solution