

CHEMISTRY

BOOKS - PSEB

Atoms and Molecules

Exercise

1. A 0.24 g of compound of oxygen and boron was found by analysis to contain 0.096 g of boron and 0.144g of oxygen. Calculate the

percentage composition of the compound by weight.



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2. What are polyatomic ions ? Give examples.



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3. Write the chemical formula of the following
:Magnesium chloride



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4. Write the chemical formula of the following

:Calcium oxide



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5. Write the chemical formula of the following

:Copper nitrate



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6. Write the chemical formula of the following

:Aluminium chloride



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7. Write the chemical formula of the following

:Calcium carbonate



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8. Give the names of the elements present in the following compounds :Quick lime



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9. Give the names of the elements present in the following compounds :Hydrogen bromide



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10. Give the names of the elements present in the following compounds : Baking powder



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11. Give the names of the elements present in the following compounds :Potassium sulphate



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12. Calculate the molar mass of the following substances : Ethyne C_2H_2



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13. Calculate the molar mass of the following substances : Sulphur molecule, S_8



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14. Calculate the molar mass of the following substances :Phosphorus molecule, P_4 (Atomic mass of phosphorus is 31)



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15. Calculate the molar mass of the following substances :Hydrochloric acid, HCl



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16. Calculate the molar mass of the following substances : Nitric acid, HNO_3 .



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17. What is the mass of : 1 mole of nitrogen atoms ?



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18. What is the mass of :) 4 moles of aluminium atoms (Atomic mass of aluminium = 27) ?



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19. What is the mass of : 10 moles of sodium sulphite (Na_2SO_3)?



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20. Convert into moles : 12g of oxygen gas



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21. Convert into moles : 20g of water



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22. Convert into moles : 22g of carbon dioxide



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23. What is the mass of : 0.2 mole of oxygen atoms?



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24. What is the mass of: 0.5 mole of water molecules?



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25. Calculate the number of molecules of sulphur (S_8) present in 16g of solid sulphur.



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26. Calculate the number of aluminium ions present in 0.051 g of aluminium oxide



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