



CHEMISTRY

BOOKS - EVERGREEN CHEMISTRY (ENGLISH)

QUESTION PAPER 2022 TERM 1

Multiple Choice Questions

1. In the Periodic Table, elements of Period 3 are arranged in the increasing order of

ionization potential as:

A. B, N, Cl, Ar

B. Mg, Si, S, Ar

C. Ar, Si, S, Mg

D. Si, Ar, Cl, Mg

Answer:



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2. If Relative Molecular Mass of Butane (C_4H_{10}) is 58 then its vapour density will be:

A. 58

B. 29

C. 32

D. 16

Answer:



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3. Identify one statement that holds true for electrolysis of molten lead bromide:

- A. Silver grey metal deposits at the anode
- B. Temperature is not maintained during the electrolysis
- C. Brown vapours of bromine are obtained at the anode
- D. Electrolyte contains H^+ ions along with Pb^{2+} ions

Answer:



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4. The tendency of an atom to attract shared pair of electrons to itself when forming a chemical bond is known as:

- A. Electron affinity
- B. Electronegativity
- C. Ionization potential
- D. Nuclear charge

Answer:



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5. Solid sodium chloride does not conduct electricity as:

- A. The strength of the bond is weak
- B. It contains free ions
- C. It does not contain any free ions
- D. it contains free ions as well as molecules

Answer:



6. Elements A and B have electronic configurations 8 and 13 respectively. The chemical formula formed between A and B will be :

A. AB

B. B_3A_3

C. A_2B_3

D. B_2A_3

Answer:



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7. The percentage of hydrogen present in NaOH is : (Relative Molecular Mass of NaOH = 40) (At. Wt. of H =1)

A. 2.5

B. 25

C. 0.25

D. 0.025

Answer:



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8. A salt formed by incomplete neutralization of acid by a base :

A. Basic salt

B. Acid salt

C. Normal salt

D. Complex salt

Answer:



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9. The colour of the precipitate formed after the addition of a small amount of sodium hydroxide solution to an aqueous solution of ferric chloride is:

A. gelatinous white

B. pale blue

C. reddish brown

D. dirty green

Answer:



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10. Alkaline earth metals have the same:

A. number of valence electrons

B. number of shells

C. metallic property

D. ionization potential

Answer:



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11. Which of the following compounds neither dissociate not ionise in water ?

- A. Hydrochloric acid
- B. Sodium hydroxide
- C. Potassium Nitrate
- D. Carbon tetrachloride

Answer:



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12. The table shows the electronic configuration of four elements.

element	electronic configuration
W	2, 6
X	2, 8
Y	2, 8, 1
Z	2, 8, 7

Which pair of atoms will form a covalent compound ?

A. two atoms of W

B. two atoms of X

C. an atom of W and an atom of X

D. an atom of Y and an atom of Z

Answer:



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13. Element with an atomic number 19 will:

A. accept an electron and get oxidized

B. accept an electron and get reduced

C. lose an electron and get oxidized

D. lose an electron and get reduced

Answer:



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14. Which of the following has two sets of lone pair of electrons in them ?

A. Ammonia

B. Methane

C. Water

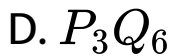
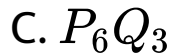
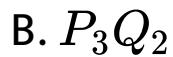
D. Ammonium ion

Answer:



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15. If the empirical mass of the formula PQ_2 is 10 and the Relative Molecular Mass is 30, then the molecular formula will be:

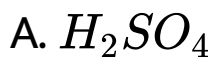


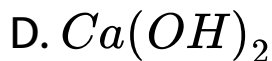
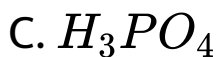
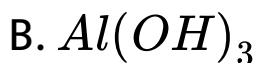
Answer:



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16. Which of the following is a tribasic acid?





Answer:



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17. If a solution of an electrolyte mixture has calcium ions, cupric ions, zinc ions and magnesium ions, which of these ions would

you see preferentially discharged at the cathode?

- A. Calcium ions
- B. Zinc ions
- C. Cupric ions
- D. Magnesium ions

Answer:



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18. Which of the following ions will readily discharge at the anode during the electrolysis of acidulated water:



Answer:



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19. If the empirical formula of a compound of CH and its vapour density is 13, then its molecular formula will be:

(At. Wt. C=12, H=1)

A. CH

B. C_2H_2

C. C_4H_4

D. C_3H_3

Answer:



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20. Aqueous solution of Cupric chloride forms a deep blue solution on addition of :

- A. dropwise sodium hydroxide
- B. excess sodium hydroxide
- C. dropwise ammonium hydroxide
- D. excess ammonium hydroxide

Answer:



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21. Which statement about conduction of electricity is correct?

A. Electricity is conducted in aqueous solution by electrons

B. Electricity is conducted in a metal wire by ions

C. Electricity is conducted in a molten electrolyte by electrons

D. Electricity is conducted in an acid solution by ions

Answer:



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22. If an element has low ionization potential, then it is likely to be a:

A. metal

B. metalloid

C. non metal

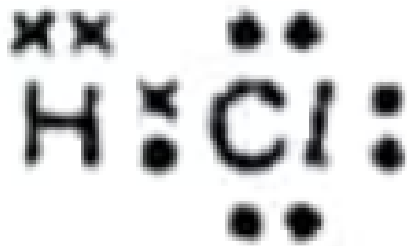
D. inert gas

Answer:

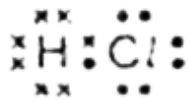


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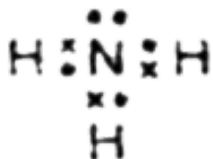
23. Which electron arrangement for the outer shell electrons in a covalent compound is correct ?



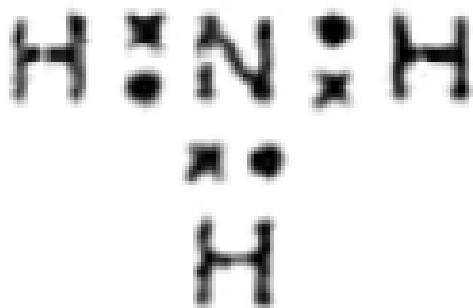
A.



B.



C.



D.

Answer:



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24. The products formed when an acid reacts with a base is :

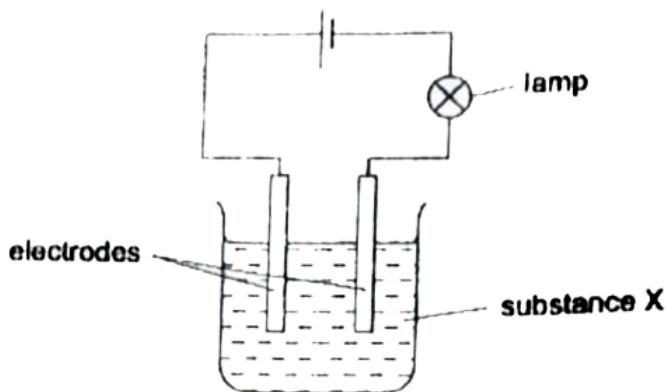
- A. salt and hydrogen
- B. salt and oxygen
- C. salt and water
- D. salt and carbon dioxide

Answer:



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25. In the circuit below, the lamp lights up



What could X be ?

A. a solution of alcohol in water

B. a solution of sodium chloride in water

C. sugar solution

D. solid potassium chloride

Answer:



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26. Which one of the following is a non metallic cation ?

A. K^+



Answer:



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27. Type of bonding present in hydrogen chloride :

A. metallic

B. ionic

C. covalent

D. coordinate

Answer:



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28. How does the metallic and non-metallic character vary on moving from left to right in a period?

A. increases

B. decreases

C. remains same

D. first increases and then decreases

Answer:



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29. The aqueous solution of the following compounds which contains both ions and molecules is

A. sulphuric acid

B. nitric acid

C. acetic acid

D. hydrochloric acid

Answer:



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30. The basic oxide which is an alkali :

A. Copper oxide

B. Sodium oxide

C. Ferric oxide

D. Zinc oxide

Answer:



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31. If the pH of solution is '2' , then the solution is a

A. strong acid

B. strong alkali

C. weak acid

D. weak alkali

Answer:



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32. The acidity of aluminium hydroxide is :

A. 3

B. 1

C. 4

D. 2

Answer:



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33. Hydracids are those acids which contain :

A. Hydrogen with any metal

B. Hydrogen , a non-metal and oxygen

C. Hydrogen and a non-metal other than oxygen

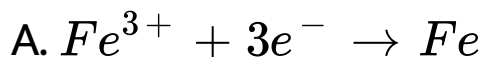
D. Hydrogen and oxygen only

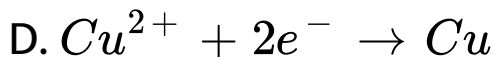
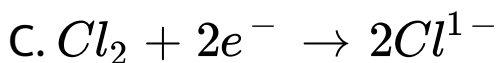
Answer:



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34. The oxidation reaction among the following is :





Answer:



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35. A student added excess of sodium hydroxide solution to each of the salt solution . An insoluble precipitate formed was observed in :

A. Calcium nitrate

B. Zinc nitrate

C. Lead nitrate

D. Sodium nitrate

Answer:



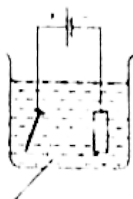
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36. Which apparatus could be used to electroplate an iron nail with copper ?

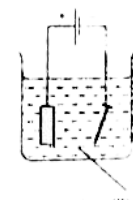
A.



B. aqueous copper(II) sulphate

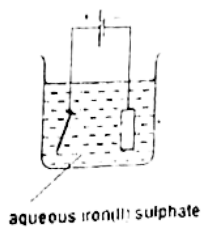


C.



aqueous iron(II) sulphate

D.



aqueous iron(III) sulphate

Answer:



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37. The table below shows the electronic arrangements of six atoms, A to F .

atom	A	B	C	D	E	F
electronic configuration	2,5	2	2,6	2,8,6	2,8,8	2,8,3

With respect to the table select the following :

Two atoms from the same group of the periodic table :

A. D and E

B. C and D

C. E and F

D. C and E

Answer:



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38. The table below shows the electronic arrangements of six atoms, A to F .

atom	A	B	C	D	E	F
electronic configuration	2,5	2	2,6	2,8,6	2,8,8	2,8,3

With respect to the table select the following :

Two noble gases :

A. A and B

B. E and F

C. B and E

D. D and E

Answer:



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39. The table below shows the electronic arrangements of six atoms, A to F .

atom	A	B	C	D	E	F
electronic configuration	2,5	2	2,6	2,8,6	2,8,8	2,8,3

With respect to the table select the following :

The atom which is the most electronegative :

A. A

B. B

C. C

D. F

Answer:



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40. The table below shows the electronic arrangements of six atoms, A to F .

atom	A	B	C	D	E	F
electronic configuration	2,5	2	2,6	2,8,6	2,8,8	2,8,3

With respect to the table select the following :

The atom which has the highest ionization potential :

A. A

B. B

C. E

D. F

Answer:



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