



CHEMISTRY

BOOKS - EVERGREEN CHEMISTRY (ENGLISH)

SAMPLE PAPER 05

Question

1. The value of second ionisation energy is than first ionisation energy.

A. more

B. less

C. same

D. None of these

Answer: A



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2. Which of the following is present in pure water ?

A. Ions

B. Molecules

C. Atoms

D. None of these

Answer: B



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3. The vapour density of an element is 14.

Calculate its molecular mass.

A. 14

B. 7

C. 28

D. Infinite

Answer: C



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4. What is the colour of aluminium salts ?

A. Red

B. Green

C. White

D. Colourless

Answer: D



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5. If a solution changes red litmus blue, what can you say about its pH value, less than 7 or greater than 7?

A. Lower than 7

B. More than 7

C. Equal to 7

D. Zero

Answer: B



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6. A solid has properties like high boiling point, high melting point, hard, soluble in polar solvents. Which type of solid is this?

A. Ionic solid

B. Covalent solid

C. Coordinate bond

D. All Question

Answer: A



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7. Answer the following questions in one word or one sentence:

What is the colour of phenolphthalein in a solution of NaOH?

A. Red

B. Blue

C. Green

D. Pink

Answer: D



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8. Electrode at which reduction takes place is known as

A. Anode

B. Cathode

C. Both 1 and 2

D. None of these

Answer: B



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9. Which of the following is not true for electrolysis of NaCl ?

A. Oxidation occurs in Na

B. Reduction occurs at Cl

C. Oxidation takes place at cathode

D. It is a redox reaction.

Answer: A::B::C



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10. A compound having molecular formula $C_6H_{14}O_6$ and empirical formula $C_3H_7O_3$, value of n is _____

A. 1

B. 2

C. 3

D. 4

Answer: B



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11. What will be the colour of precipitate obtained when NaOH solution is added to iron (II) sulphate solution ?

A. Red

B. Green

C. White

D. Colourless

Answer: B



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12. What can you say about the pH of a solution that liberates carbon dioxide from sodium carbonate?

A. Lower than 7

B. More than 7

C. Equal to 7

D. Zero

Answer: A



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13. How many electrons are gained or lost by calcium to attain nearest noble gas configuration ?

A. 2 electrons gain

B. 2 electrons lost

C. 3 electrons gain

D. 3 electrons lost

Answer: B



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14. Which of the conditions are necessary for the formation of ionic bond ?

A. Low ionisation energy

B. High electron affinity

C. High lattice energy

D. All the above

Answer: D



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15. Which of the following is equal to number of electrons in the valence shell ?

A. Atomic mass

B. Group number

C. Period number

D. Atomic volume

Answer: B



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16. In which period of the periodic table, an element with atomic number 14 is placed ?

A. 4

B. 3

C. 2

D. 1

Answer: B



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17. What will be the empirical formula of a compound if the percentage of elements present is A = 39.56 %, B = 7.74%, C = 52.70% ?

(Atomic mass of A = 12, B = 1, C = 16)

A. ABC

B. $A_3B_7C_3$

C. $A_6B_7C_3$

D. $A_3B_3C_6$

Answer: B



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18. Name :

a yellow monoxide that dissolves in hot and concentrated caustic alkali.

A. Cu_2O

B. ZnO

C. PbO

D. none of these

Answer: C



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19. How pH of a neutral solution be increased ?

A. By adding acid

B. By adding alkali

C. Both 1 and 2

D. None of these

Answer: B



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20. An example of salt formed by a base and an acid is...

- A. Sodium chloride
- B. Calcium sulphate
- C. Lithium chloride
- D. Sodium chlorate

Answer: A



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21. Which of the following is an example of covalent bond ?

A. Calcium oxide

B. Water

C. Ammonia

D. Sodium chloride

Answer: B



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22. Which process occur in electropositive ion ?

A. Oxidation

B. Reduction

C. Both 1 and 2

D. None of these

Answer: A



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23. In which period of the periodic table, an element with atomic number 14 is placed ?

A. 4

B. 3

C. 2

D. 1

Answer: A



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24. Which of the following is least reactive:

A. F

B. Br

C. Cl

D. I

Answer: D



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25. What happens to pale blue ppt. obtained when ammonia solution is added to copper sulphate solution on addition of excess of alkali ?

- A. Ppt. is soluble
- B. Ppt. is insoluble
- C. No effect on ppt
- D. Sparingly soluble

Answer: A



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26. Which of the following is true when lead nitrate is treated with sodium hydroxide solution, and ppt. obtained is dissolved in excess of alkali ?

- A. White ppt. soluble in excess of alkali
- B. White ppt, insoluble in excess of alkali
- C. Green ppt. soluble in excess of alkali
- D. Green ppt. insoluble in excess of alkali

Answer: A



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27. An example of salt formed by a metal and an acid is _____

- A. Sodium chloride
- B. Calcium sulphate
- C. Lithium chloride
- D. Sodium chlorate

Answer: B



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28. Which process occur in electropositive ion ?

- A. Oxidation
- B. Reduction
- C. Both 1 and 2
- D. None of these

Answer: A



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29. Hydrogen chloride molecule contains:

- A. Polar covalent bond
- B. Double bond
- C. Coordinate bond
- D. Electrovalent bond

Answer: A



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30. The gas liberated when Sodium Sulphite reacts with dilute sulphuric acid is

- A. Carbon dioxide
- B. Hydrogen
- C. Hydrogen sulphide
- D. Sulphur dioxide

Answer: D



31. Choose the correct answer from the options given below :

Hydroxide of this metal is soluble in sodium hydroxide solution.

A. Magnesium

B. Lead

C. Silver

D. Copper

Answer: B



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32. Vapour density of a gas is 22. What is its molecular mass ?

A. 23

B. 22

C. 44

D. 11

Answer: C



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33. _____ the chemical change that occurs at this electrode is called _____.

- A. Anode, oxidation
- B. Anode, reduction
- C. Cathode, oxidation
- D. Cathode, reduction

Answer: D



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34. The cathode production of the electrolysis of zinc iodide is :

A. Iodine

B. Zinc

C. Zinc oxide

D. Chloride

Answer: B



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35. The electrolysis of acidified water is an example of:

A. reduction

B. oxidation

C. redox reaction

D. catalytic reaction

Answer: D



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36. The ratio of the mass of a certain volume of gas to the mass of an equal volume of hydrogen under the same conditions of temperature and pressure is known as _____

- A. Vapour density
- B. Empirical formula
- C. Molecular formula

D. Percentage composition

Answer: A



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37. The electronic configuration of three elements X, Y and Z is given below: X = 2 Y=2,6
Z=2,8, 2

Which element belongs to the second period ?

A. X

B. Y

C. Z

D. none of these

Answer: B



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38. The electronic configuration of three elements X, Y and Z is given below: X = 2 Y=2,6
7=2,8, 2

A. X

B. Y

C. Z

D. none of these

Answer: A



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39. The electronic configuration of three elements X, Y and Z is given below: X = 2 Y=2,6
7=2,8, 2

A. 1

B. 2

C. 3

D. 4

Answer: B



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40. The electronic configuration of three elements X, Y and Z is given below: X = 2 Y=2,6
7=2,8, 2

A. X

B. Y

C. Z

D. All of these

Answer: C



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