



CHEMISTRY

BOOKS - EVERGREEN CHEMISTRY (ENGLISH)

SAMPLE PAPER 1

Questions

1. Modern periodic table is based on which of the following property of elements ?

A. Atomic number

B. Atomic mass

C. Electronegativity

D. Electron affinity

Answer: A



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2. Which of the following is a mineral acid ?

A. Oxalic acid

B. Sulphuric acid

C. Malic acid

D. Acetic acid

Answer: B



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3. Which type of bond is formed between a metal and a non-metal ?

A. Covalent bond

B. Coordinate covalent bond

C. Ionic or electrovalent bond

D. No bonding

Answer: C



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4. What is the colour of ppt. formed when sodium hydroxide solution is added to calcium salt ?

A. Green

B. White

C. Brown

D. Red

Answer: B



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5. The correct relationship of molecular formula and empirical formula is _____

A. Empirical formula = $n \times$ molecular formula

B. Molecular formula = $n \times$ empirical formula

C. $n =$ Molecular formula \times empirical formula

D. Molecular formula = $n /$ empirical formula

Answer: B



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6. Which of the following is true for electrolysis of copper sulphate solution ?

- A. Cathode acts as electron acceptor
- B. Anode acts as electron donor
- C. Copper is deposited at cathode
- D. Sulphate ions migrate towards anode

Answer: C



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7. How many periods and groups are there in Modern Periodic Table ?

- A. 8 groups and 7 periods
- B. 7 periods and 18 groups
- C. 7 groups and 18 periods
- D. 7 groups and 15 periods

Answer: B



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8. Which of the following statement is not true regarding ionic or electrovalent bond ?

A. These are hard and brittle solids

B. They have high density

C. They are insoluble in water

D. They have high melting and boiling points

Answer: C



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9. Which type of bonding is present in ammonium ion ?

A. Covalent bond

B. Coordinate bond

C. Ionic or electrovalent bond

D. No bonding

Answer: B



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10. What is the colour of the ppt. formed when sodium hydroxide solution is added to ferric salt?

A. Dirty green

B. White

C. Brown

D. Reddish brown

Answer: D



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11. What happens when ammonium hydroxide solution is added to calcium salt?

- A. A white ppt is formed
- B. Ppt formed is soluble in excess of alkali.
- C. A colourless solution is formed
- D. No ppt is formed.

Answer: D



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12. Which formula denotes the actual number of different elements present in one molecule of compound?

- A. Molecular formula
- B. Empirical formula
- C. Empirical formula mass
- D. Percentage

Answer: A



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13. Which of the following anion can be discharged with maximum ease ?

A. Sulphate ion

B. Iodide ion

C. Hydroxyl ion

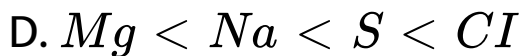
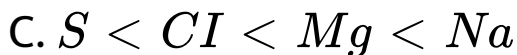
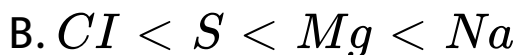
D. Nitrate ion

Answer: C



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14. What is the correct order of atomic size?



Answer: B



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15. How the ionisation energy of the elements changes on moving top to bottom in the group ?

- A. Increases
- B. decreases
- C. remains same
- D. zero

Answer: B



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16. Which of the following elements has highest ionisation energy?

A. Li

B. Na

C. K

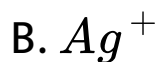
D. Cs

Answer: D



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17. Which of the ions can be discharged with maximum difficulty ?



Answer: D



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18. What is the correct relationship between vapour density and molecular mass ?

A. Relative molecular mass = Vapour density / 2

B. Vapour density = $2 \times$ Relative molecular mass

C. Relative molecular mass = $2 \times$ Vapour density

D. Vapour density = Relative molecular mass

Answer: C



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19. Which type of bond is present in chlorine molecule ?

A. Single covalent bond

B. Double covalent bond

C. Triple covalent bond

D. Ionic bond

Answer: A



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20. Identify the odd one :

A. Methane

B. HCl

C. NaCl

D. Nitrogen

Answer: C



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21. The amount of energy released when an electron is added to the outermost shell of an isolated gaseous neutral atom is called its :

A. Atomic size

B. Ionisation potential

C. Electron affinity

D. Electronegativity

Answer: C



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22. An element 'E' loses electron and becomes an ion. Name the process and ion :

A. Oxidation, cation

B. Oxidation, anion

C. Reduction, cation

D. Reduction, anion

Answer: A



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23. Which of the following is a not a strong acid ?

A. Hydrochloric acid

B. Sulphuric acid

C. Nitric acid

D. Acetic acid

Answer: D



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24. Bases furnishions in solutions.

A. H^+

B. OH^-

C. Both (1) and (2)

D. None of these

Answer: B



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25. Choose the hydroxide which is insoluble in excess of ammonium hydroxide ?

A. Calcium hydroxide

B. Ferric hydroxide

C. Copper hydroxide

D. Zinc hydroxide

Answer: B



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26. Which formula denotes the simplest positive integer ratio of atoms present in a compound?

A. Molecular formula

B. Empirical formula

C. Empirical formula mass

D. Percentage

Answer: B



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27. The kind of particles present in carbonic acid are _____ .

A. Ions

B. Molecules

C. Both (1) and (2)

D. None of these

Answer: C



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28. Which of the following contains molecules only?

A. Sugar

B. NaOH

C. HCl

D. Carbonic acid

Answer: A



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29. Which of the following is not an oxyacid ?

A. Hydrochloric acid

B. Sulphuric acid

C. Nitric acid

D. Acetic acid

Answer: A



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30. An example of hydracid is _____ .

A. Hydrochloric acid

B. Sulphuric acid

C. Nitric acid

D. Acetic acid

Answer: A



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31. Which of the following statement is not true regarding acids ?

- A. They have a sour taste.
- B. They are good conductors of electricity
- C. They turn red litmus to blue.
- D. They are corrosive in nature.

Answer: C



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32. Which of the following is a coloured ion ?

A. Nickel

B. Lead

C. Zinc

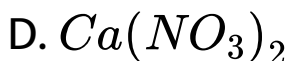
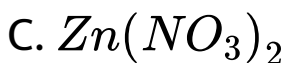
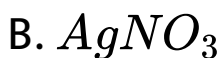
D. Chloride

Answer: A



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33. Four solutions are taken. A copper strip is added to these? The colour of which of the following will turn blue ?



Answer: B



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34. What happens when ammonia solution is added first drop-wise and then in excess to zinc sulphate?

A. A reddish brown ppt is formed which is insoluble in excess of alkali.

B. Dirty green ppt is formed which is insoluble in excess of alkali.

C. A gelatinous white ppt is formed which is soluble in excess of alkali.

D. No ppt will be formed even in excess of alkali.

Answer: C



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35. If the vapour density of a gas is 15, its relative molecular mass will be _____

A. 30

B. 7.5

C. zero

D. can't be determined

Answer: A



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36. The salt which on hydrolysis forms acid is

- A. Iron chloride
- B. Aluminium acetate
- C. Sodium chloride
- D. All of the above

Answer: A



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37. Some elements are given below in the form of a table : The elements are mentioned with their own symbol.

Li	Be	B	C	O	F	Ne
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To which period these elements belong?

A. First

B. Second

C. Third

D. Fourth

Answer: B



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38. Some elements are given below in the form of a table : The elements are mentioned with their own symbol.

Li	Be	B	C	O	F	Ne
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Which element is missing in given series?

- A. Oxygen
- B. Nitrogen
- C. Phosphorous
- D. Sulphur

Answer: B



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39. Some elements are given below in the form of a table : The elements are mentioned with their own symbol.

Li	Be	B	C	O	F	Ne
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To which group does missing elements belongs ?

A. 15

B. 16

C. 17

D. 18

Answer: A



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40. Some elements are given below in the form of a table : The elements are mentioned with their own symbol.

Li	Be	B	C	O	F	Ne
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Which element belongs to group 18 and what is that group elements called ?

A. O, chalcogens

B. Li, alkali metals

C. F, halogens

D. Ne, Inert gases

Answer: D



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41. Some elements are given below in the form of a table : The elements are mentioned with their own symbol.

Li	Be	B	C	O	F	Ne
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What is valency of group 18 elements ?

A. Zero

B. One

C. Two

D. Three

Answer: A



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