CHEMISTRY

BOOKS - EVERGREEN CHEMISTRY (ENGLISH)

SAMPLE PAPER 3

Multiple Choice Question

1. Lithium chloride is formed by transfer of electrons, which elements is getting oxidised

in the process of formation ?

A. Lithium

B. Chlorine

C. Both 1 and 2

D. None of these

Answer: A

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2. An acid which is used in soda wash and in aerated drinks is ____

A. Citric acid

B. Carbonic acid

C. Acetic acid

D. Boric acid

Answer: B

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3. If element X forms a chloride with the formula XCl_3, then X would most likely belong to the same group of the Modern Periodic Table as :

A. Na

B. Br

C. Al

D. Mg

Answer: C

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4. Valency of aluminium is:

A. 2

B. 3

C. 4

D. 5

Answer: B

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5. $Na_2CO_3.10H_2O$ is :

A. washing soda

B. baking soda

C. bleaching powder

D. tartaric acid

Answer: A

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6. The oxide and hydroxide of which metal is amphoteric:

A. Zine

B. Copper

C. Iron

D. Manganese

Answer: A

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7. Relation between vapour density and molecular weight

A. Molecular weight = 2/Vapour density

B. Molecular weight = 2 $\times$ Vapour density

C. Molecular weight $\times$ 2 = Vapour density

D. None of these

Answer: B

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8. During electrolysis of NaCl, the gas discharged at the anode is :

A. Chlorine

B. Oxygen

C. Hydrogen

D. None of these

Answer: A

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9. Magnesium hydroxide is _____

A. Monoacidic alkali

B. Diacidic alkali

C. Triacidic alkali

D. All of these

Answer: B

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10. What is the colour when methyl orange is added to Sulphuric acid ?

A. Pink

B. Red

C. Blue

D. Colourless

Answer: A

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11. Which type of bond is present in carbon tetrachloride ?

A. Ionic bond

B. Covalent bond

C. Coordinate bond

D. None of these

Answer: B

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12. An element having atomic number 19 and belongs to Alkali metals is _____

A. Li

B. F

C. K

D. Cl

Answer: C

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13. The salt which on hydrolysis forms acid is _____

A. Iron chloride

B. Aluminium acetate

C. Sodium chloride

D. All the above

Answer: C

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14. A compound which liberates reddish brown gas around the anode during electrolysis in its molten state is :

Sodium chloride

Copper (II) oxide

Copper (II) sulphate

Lead (II) bromide

A. Sodium chloride

B. Copper (II) oxide

C. Copper (II) sulphate

D. Lead (II) bromide

Answer: D

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15. The empirical formula and molecular mass of a compound are CH_2O and 180g respectively. What will be the molecular formula of the compound ?

A. $C_9H_{18}O_9$

B. CH_2O

C. $C_6H_{12}O_6$

D. $C_2H_4O_2$

Answer: C

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16. Choose the correct answer from the options given below :

Anhydrous iron(III) chloride is prepared by:

A. Direct combination

B. Simple displacement

C. Decomposition

D. Neutralization

Answer: A

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17. How many water molecules does hydrated calcium sulphate contain ?

A. 5

B. 10

C. 7

D. 2

Answer: D

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18. Identify the molecule with a single covalent bond.

A. CO_2

B. CO

C. CH_4

D. N_2

Answer: B

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19. An element having electronic configuration 2, 8, 18, 3 belongs to which group of the Modern Periodic Table ?

A. 13^{th} group

B. 3^{rd} group

C. 18^{th} group

D. 15^{th} group

Answer: A

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20. The electronic configuration of Mg is _____

A. 2,8,1

B. 2,8,7

C. 2,8,2

D. 2,8

Answer: C

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21. An element having atomic number 17 and belongs to halogens is _____

A. Li

B. F

C. K

D. Cl

Answer: D

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22. How many valence electrons are present in Mg?

A. 1

B. 2

C. 3

D. 4

Answer: B

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23. Write the name of a non-metal of group 15.

A. Nitrogen

B. Calcium

C. Phosphorous

D. None of these

Answer: A

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24. When fused lead bromide is electrolysed we observe:

A. a silver grey deposit at anode and a reddish brown deposit at cathode

B. a silver grey deposit at cathode and a reddish brown deposit at anode

C. a silver grey deposit at cathode and reddish brown fumes at anode

D. silver grey fumes at anode and reddish brown fumes 'at cathode.

Answer: C

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25. Arrange the following as per the instructions given in the brackets:

Cs, Na, Li, K, Rb (increasing order of metallic character).

A. $LI < Na < K < Rb < Cs$

B. $Li < Na < Cs < K < Rb$

C. $K < Rb < Cs < Li < Na$

D. $Cs < Li < Na < K < Rb$

Answer: A

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26. A chloride which forms a precipitate that is soluble in excess of ammonium hydroxide is :

A. Calcium chloride

B. Ferrous chloride

C. Ferric chloride

D. Copper chloride

Answer: D

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27. Sodium carbonate is a basic salt because it is a salt of a:

A. strong acid and strong base

B. weak acid and weak base

C. strong acid and weak base

D. weak acid and strong base

Answer: D

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28. A polar covalent bond will be formed in which one of these pair of atoms:

A. HF

B. H_2

C. Cl_2

D. O_2

Answer: A

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29. Sodium carbonate is a basic salt because it is a salt of a:

A. C_5H_4

B. C_6H_4

C. C_5H_{10}

D. C_5H_2

Answer: A

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30. The vessel in which electrolysis of lead bromide is carried out is:

A. Clay crucible

B. Glass vessel

C. Silica crucible

D. Aluminium vessel

Answer: C

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31. Arrange the following as per instructions given in the brackets:

Cl, F, Br, I (increasing order of electron affinity)

A. $Br < I < F < Cl$

B. $I < Br < Cl < F$

C. $Br < Cl < I < F$

D. $F < Cl < Br < I$

Answer: B

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32. A solution of the compound which gives a dirty green precipitate with sodium hydroxide.

A. Ammonium sulphate

B. Lead carbonate

C. Ferrous sulphate

D. Chlorine

Answer: C

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33. Alkalis are :

A. acids, which are soluble in water

B. acids, which are insoluble in water

C. bases, which are insoluble in water

D. bases, which are soluble in water

Answer: D

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34. Aluminium has a tendency to loose ..... electrons.

A. 2 electrons

B. 1 electron

C. 4 electrons

D. 3 electrons

Answer: D

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35. An organic compound contains carbon , hydrogen and oxygen . Its elemental analysis gave C ,38.41% and $H, 9.67\%$. The empirical formula of the compound would be

A. CH_3O

B. CH_2O

C. CHO

D. CH_4O

Answer: A

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36. Which soln. becomes a deep/inky blue colour when excess of ammonium hydroxide is added to it.

A. Copper nitrate

B. Iron (II) sulphate

C. Iron (III) chloride

D. Lead nitrate

Answer: A

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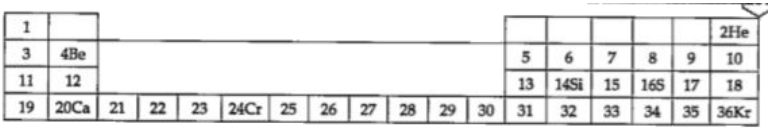
37. The diagram given below is a part of Periodic Table. Study the table and answer the questions given below the table :



1																2He
3	4Be											5	6	7	8	9	10
11	12											13	14Si	15	16S	17	18
19	20Ca	21	22	23	24Cr	25	26	27	28	29	30	31	32	33	34	35	36Kr

Name two elements in same group of Periodic Table

A. Oxygen and Uranium

B. Oxygen and sulphur

C. Calcium and Hydrogen

D. Cromium and Rubidium

Answer: B

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1																	2He
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19	20Ca	21	22	23	24Cr	25	26	27	28	29	30	31	32	33	34	35	36Kr

Name the transition metal.

A. Chromium

B. Sulphur

C. Calcium

D. Oxygen

Answer: A

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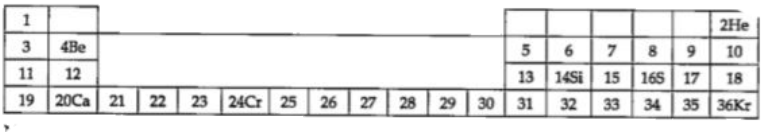

1																2He
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Name an element, which reacts vigorously with water.

A. Argon

B. Boron

C. Calcium

D. Uranium

Answer: C

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1																	2He
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19	20Ca	21	22	23	24Cr	25	26	27	28	29	30	31	32	33	34	35	36Kr

Which element forms very corrosive acid?

A. Chromic acid produced by chromium

B. Oxalic acid produced by Oxygen

C. Calcium carbonate acid produced by calcium

D. Ferric oxide produced by Iron

Answer: A

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