



CHEMISTRY

BOOKS - BODY BOOKS PUBLICATION

COMPOUNDS OF NON-METALS

Example

1. Take a little ammonium chloride (NH_4Cl) in a watch glass and add a little calcium hydroxide ($Ca(OH)_2$) to it. Stir well. Can you sense any smell?

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2. Show wet blue and red litmus papers over the watch glass containing soap solution one by one. Which litmus paper shows a colour change ?



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3. Take a little ammonium chloride (NH_4Cl) in watch glass and a little calcium hydroxide ($Ca(OH)_2$) to it. Stir well. Is the gas acidic or basic?



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4. Why ammonia gas is passed over quick lime (CaO)?



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5. Ammonia gas is passed over quick lime(CaO).

What may be the reason for collecting ammonia in this manner?



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6. Ammonia gas is passed over quick lime(CaO).

What is your inference about the density of ammonia from this?



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7. What inference can be made about solubility of Ammonia in water?



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8. What inference can be made about solubility of Ammonia in water?



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9. Complete the chemical equation given below and find the product obtained when ammonia is dissolved in water.



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10. Tick ✓ which is applicable to ammonia.

Colour	Yes/No colour
Smell	Pungent smell /No smell
Property	Basic /Acidic
Solubility in water	More soluble /Less soluble
Density	Less than air /More than air



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11. When an Ammonia tanker leaks, water is sprayed to reduce its intensity. What is the reason for this?



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12. Write some uses of ammonia.



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13. Take some ammonium chloride (NH_4Cl) in a boiling tube and heat it. Don't you sense a peculiar smell?



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14.

Take some ammonium chloride (NH_4Cl) in a boiling tube and heat it. Don't you sense a peculiar smell?

Which is the gas evolved here?



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15.

Take some ammonium chloride (NH_4Cl) in a boiling tube and heat it. Don't you sense a peculiar smell?

Show a wet red litmus paper on the mouth of the boiling tube. What change can you observe?



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16.

Take some ammonium chloride (NH_4Cl) in a boiling tube and heat it. Don't you sense a peculiar smell?

Keep the litmus paper for some more time at the mouth of the boiling tube and then observe its colour change. What is the change occurred?



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17.

Take some ammonium chloride (NH_4Cl) in a boiling tube and heat it. Don't you sense a peculiar smell?

Write the chemical equation of this reaction.



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18. A glass rod dipped in conc. HCl is shown in to the jar which is filled with ammonia. What do you observe?



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19.

A glass rod dipped in conc. HCl is shown into the jar which is filled with ammonia. What do you observe?

Complete the equation and find out the product?



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20.

A glass rod dipped in conc. HCl is shown into the jar which is filled with ammonia. What do you observe?

What happens to the white powder on heating?



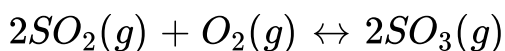
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21. Examine the chemical equation given below and write the forward and backward reaction $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g)$



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22. Examine the chemical equation given below and write the forward and backward reaction :



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23. Examine the chemical equation given below and write the forward and backward reactions : $H_2(g) + I_2(g) \leftrightarrow 2HI(g)$



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24. What happens to the rates of forward and backward reactions as time progresses?



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25. Identify the point at which the rates of both forward and backward reactions become equal?



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26. The rate of which reaction increases when the concentration of nitrogen is increased in the manufacture of Ammonia? Forward reaction/Back ward reaction.



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27. What happens if the concentration of ammonia is increased in the manufacture of Ammonia?



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28. What will be the effect of removing ammonia continuously from the system in the manufacture of Ammonia?



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29. Complete the table writing the effect of change in concentration in the system at equilibrium.

Action	Change of concentration	Change in rate
• More hydrogen is added		
• More ammonia is added		
• Ammonia is removed		
• More nitrogen is added		



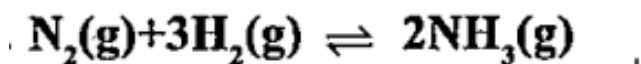
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30. $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g)$ in this equation what is the total number of moles of the reactant molecules.



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31.



In this equation what is the total number of moles of the reactant molecules?

What about the products?



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32. In the manufacture of ammonia, the reaction in which direction result in the decrease in the number of molecules?



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33. What happens when the pressure of system is decreased?



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34. In manufacturing of ammonia, What if the pressure of the system is decreased?

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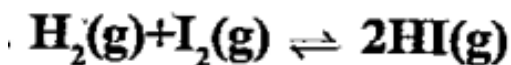
35. In the manufacture of ammonia, why is a pressure of 150-300 atm used?

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36. $H_2(g) + I_2(g) \leftrightarrow 2HI(g)$ What is the total number of moles of reactants?

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37.



What is the total number of moles of reactants?

What about the products?



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38. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g}) + \text{Heat}$ Which is the endothermic reaction in this? Forward reaction/Bakward reaction.

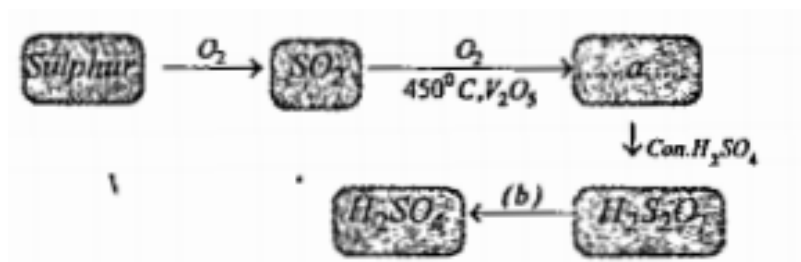


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39. Sulphuric acid is formed also by the direct dissolution of sulphur trioxide in water. Still, sulphur trioxide is not directly dissolved in water. Why?

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40. Complete the flow chart.



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41. Take 5mL water in a test tube and slowly add concentrated sulphuric acid to it. Touch the bottom of the test tube. What

do you feel? Is the reaction exothermic or endothermic?



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42. What are the constituent elements of sugar?



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43. In constituent of sugar, Which is the black substance in the product formed?



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44. What is the ratio of hydrogen and oxy-gen is sugar?



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45. Which is the substance that absorbed hydrogen and oxygen from sugar in the ratio as in water?

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46. Complete the table by involving the activities given below.

No.	Activity	Observation
1.	Dropping Con. H_2SO_4 on a cotton cloth.	
2.	Adding Con. H_2SO_4 to glucose taken in a small beaker.	
3.	Adding Con. H_2SO_4 to a watch glass in which CuSO_4 crystals are taken.	

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47. Why is concentrated sulphuric acid not used as a drying agent in the preparation of ammonia?



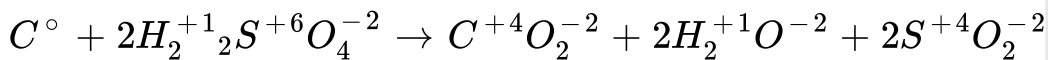
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48. Add concentrated sulphuric acid to a test tube containing a small quantity of carbon. Heat it. What do you observe?



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49. Analyse the chemical equation and find the reason for your observation.



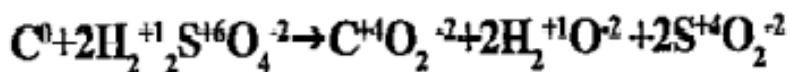
What is the oxidation state of elemental carbon?



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50.

Analyse the chemical equation and find the reason for your observation.



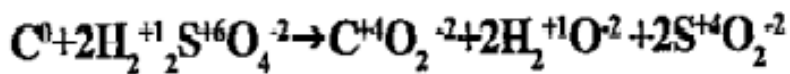
What is the oxidation state of the carbon in carbon dioxide?



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51.

Analyse the chemical equation and find the reason for your observation.



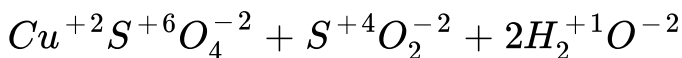
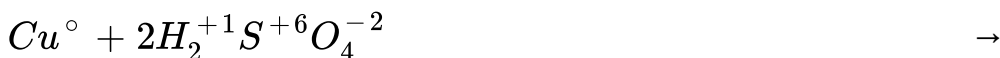
Was carbon oxidized or reduced in this reaction?

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52. What is oxidising agent?

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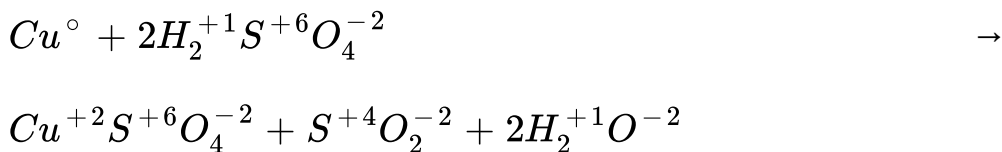
53. See the reaction between concentrated sulphuric acid and copper.



Is copper oxidized or reduce in this case?

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54. See the reaction between concentrated sulphuric acid and copper.



Which is the oxidizing agent in this reaction? Which is the reducing agent?



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55. Analyse the given chemical equation.

$Na_2SO_4 + BaCl_2 \rightarrow BaSO_4 + 2NaCl$ Which substance is soluble in water among the products?



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56. When 1 ml Barium Chloride solution is added to the solution given below:

No.	Solution	By adding BaCl ₂ solution	When dilute HCl is added to this
1.	MgSO ₄		
2.	ZnSO ₄		

Which substance is the white pre-cipitate?



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57. When 1 ml Barium Chloride solution is added to the solution given below:

No.	Solution	By adding BaCl ₂ solution	When dilute HCl is added to this
1.	MgSO ₄		
2.	ZnSO ₄		

Does the white precipitate dis-solve when dilute hydrochloric acid is added to it?

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58. Write down the observation in the table given below, when 1mL Barium chlo-ride solution is added to the solutions given in the table.

No.	Solution	By adding BaCl ₂ solution	When dilute HCl is added to this
1.	MgSO ₄		
2.	ZnSO ₄		

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59. In which of the following rever-sible reactions does change in pressure not influence equili- brium? What is the reason?

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60. What is the use of applying high pressure during the formation of ammonia from nitrogen and hydrogen?



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61. $C(s) + H_2O(g) \xrightleftharpoons{\text{heat}} CO(g) + H_2(g)$ Identify the reactants and products.



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62. $C(s) + H_2O(g) \rightleftharpoons CO(g) + H_2(g)$

a) Identify the reactants and products.

b) Products are frequently removed from the system. What happen to the system? Explain the reason.



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63. $2NO(g) + O_2(g) \leftrightarrow 2NO_2(g) + \text{heat}$ In this reaction how do the following changes influence the amount of the product: Decrease in temperature



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64. $2NO(g) + O_2(g) \leftrightarrow 2NO_2(g) + \text{heat}$ In this reaction how do the following changes influence the amount of the product: Increase in pressure



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65. $2NO(g) + O_2(g) \leftrightarrow 2NO_2(g) + \text{heat}$ In this reaction how do the following changes influence the amount of the product: Increase in concentration of oxygen.



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66. $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$: What change in pressure is required for the maximum yield of the product?



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67. $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$: What is the change in concentration required for increasing the rate of the forward reaction?



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68. The chemical equation of one of the different stages of manufacturing sulphuric acid by contact process is given below. Find out the influence of the following factors in the reaction given below. $2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g) + \text{heat}$:
Increase the amount of oxygen.



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69. The chemical equation of one of the different stages of manufacturing sulphuric acid by contact process is given below. Find out the influence of the following factors in the reaction given below. $2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g) + \text{heat}$:
Pressure is increased



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70. The chemical equation of one of the different stages of manufacturing sulphuric acid by contact process is given below. Find out the influence of the following factors in the reaction given below. $2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g) + \text{heat}$: Catalyst vanadium pentoxide (V_2O_5) is added.

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71. The chemical equation of one of the different stages of manufacturing sulphuric acid by contact process is given below. Find out the influence of the following factors in the reaction given below. $2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g) + \text{heat}$: SO_3 is removed.

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72. Calcium oxide (CaO) is used as drying agent in the preparation of Ammonia in laboratory. Can concentrated H_2SO_4 be used as drying agent instead of CaO ? Justify your answer.



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73. Which property of sulphuric acid is shown in the following situation.

- a) During the preparation of chlorine the gas is passed through concentrated H_2SO_4
- b) Wooden cupboards appeared to be burnt, when concentrated sulphuric acid happened to fall on it.



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74. Which property of sulphuric acid is shown in the following situation.

a) During the preparation of chlorine the gas is passed through concentrated H_2SO_4

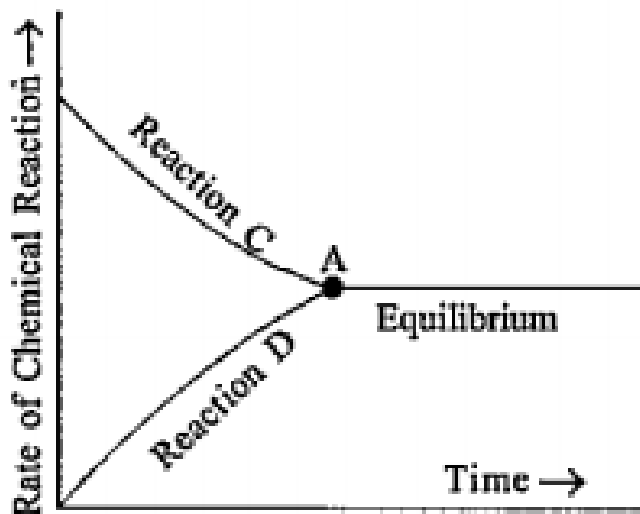
b) Wooden cupboards appeared to be burnt, when concentrated sulphuric acid happened to fall on it.



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75. The graph for the reaction

$N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$ is given below.



Identify

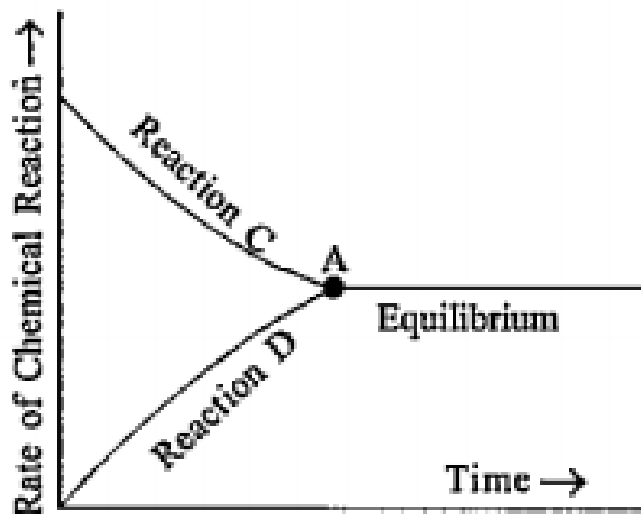
and write the reactions C and D



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76. The graph for the reaction

$N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$ is given below.



What

happens to the position of point A in the graph when a catalyst is used? Redraw the graph.

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77. It is often said that the production of sulphuric acid is a bench mark of the industrial development of country. Prepare a note based on the various uses of sulphuric acid.

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78. Fill half of a beaker of capacity 50 mL with sugar. Add concentrated sulphuric acid so that the sugar is immersed in it. Observe the changes. What are the products formed? Which property of sulphuric acid is revealed here?



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79. What will be the effect of temperature on a rate of a reaction?

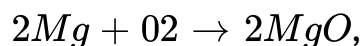
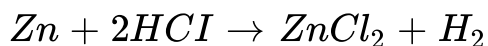


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80. Why rate of a reaction increase when temperature increases?

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81. Some chemical reactions are given below.



$NH_4Cl \leftrightarrow NH_3 + HCl$: What are the peculiarities of first two reactions.

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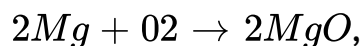
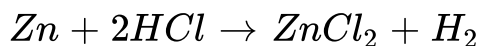
82. Some chemical reactions are given below.



Conduct an experiment for viewing the dissociation and association taking place in the third equation.

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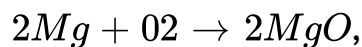
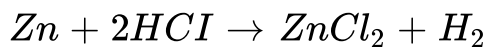
83. Some chemical reactions are given below.



In the three reactions reactants turned into product and products are converted into reactants, is it true?

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84. Some chemical reactions are given below.



What type of reactions are they all represent?

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85. Some chemical reactions are given below.



$NH_4Cl \leftrightarrow NH_3 + HCl$: What are the peculiarities of first two reactions.



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86. $Fe(NO_3)_3 + 3KCNS \rightarrow Fe(CNS)_3 + 3KNO_3$, This balanced chemical equation is written on the black board, when the teacher is going to conduct an experiment on chemical equilibrium: In the above reaction, which chemical has red colour.



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87. $Fe(NO_3)_3 + 3KCNS \rightarrow Fe(CNS)_3 + 3KNO_3$, This balanced chemical equation is wrote on the black board,when the teacher is going to conduct an experiment on chemical equilibrium: In the above reaction, which chemical has red colour.



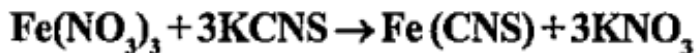
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88. $Fe(NO_3)_3 + 3KCNS \rightarrow Fe(CNS)_3 + 3KNO_3$, This balanced chemical equation is wrote on the black board,when the teacher is going to conduct an experiment on chemical equilibrium: In the above reaction, which chemical has red colour.



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89.



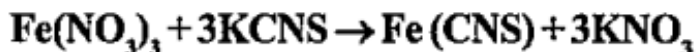
This balanced chemical equation is wrote on the black board, when the teacher is going to conduct an experiment on chemical equilibrium

Point out the characteristics of equilibrium based on the experiment done.



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90.



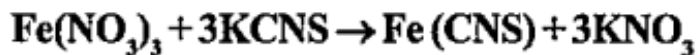
This balanced chemical equation is wrote on the black board, when the teacher is going to conduct an experiment on chemical equilibrium

e. In minute level chemical equilibrium is Kinetic energy why?



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91.



This balanced chemical equation is wrote on the black board, when the teacher is going to conduct an experiment on chemical equilibrium

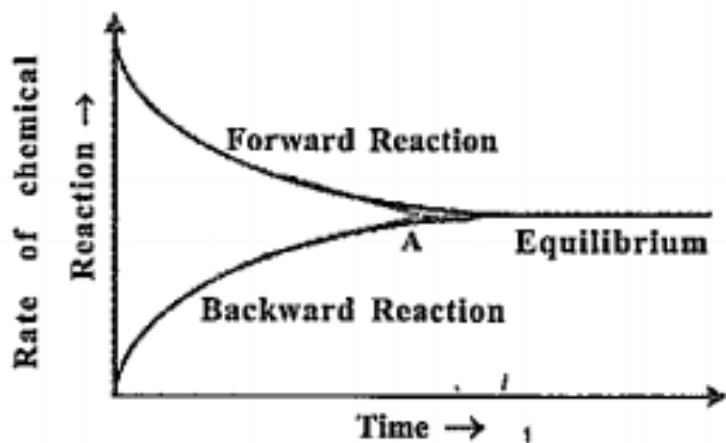
f. How and when a reversible reaction attain chemical equilibrium



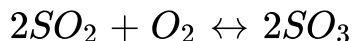
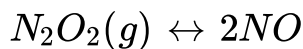
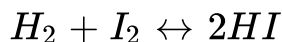
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92. Convert the solution reared into four be: In the graph given below, when the react-ent and product attain the level

A? What are the characteristics of the point A?



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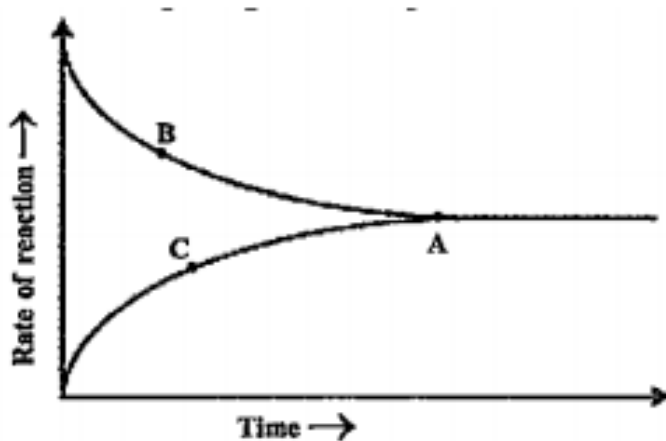
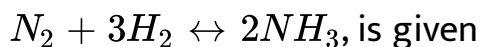
How amount of the products increases in the above reactions
(based on Le chateliers principle) Hints: reference must be

given on each of the following: concentration, pressure, temperature and catalyst



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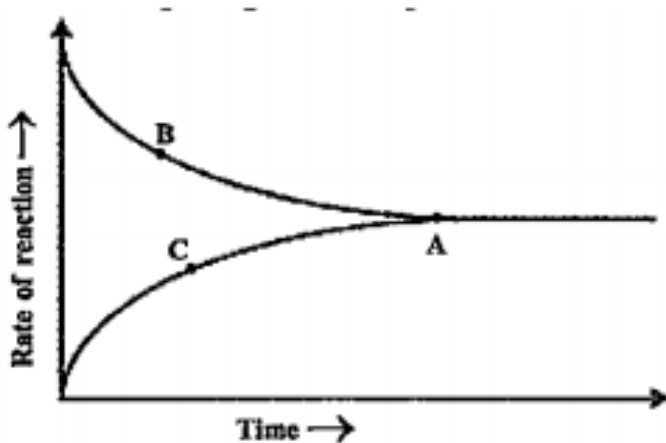
94. The graph showing the progress of the reaction



: Identify

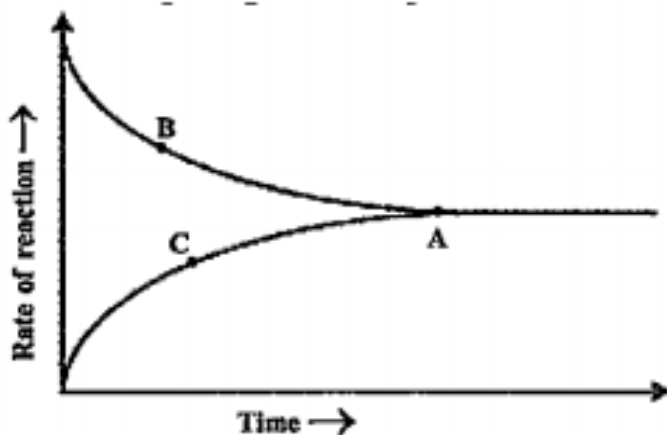
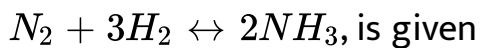
the reactions represented by B and C?

95. The graph showing the progress of the reaction $N_2 + 3H_2 \leftrightarrow 2NH_3$, is given



What is the significance of the state A?

96. The graph showing the progress of the reaction



Is there any change in the concentration, as time passes after attaining the stage A? Explain.



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97. Cold water is taken in one test tube and hot water in another one. Mg ribbon with same size is dropped in each of the test tube: In which test tube, hydrogen is formed with greater speed?



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98. Cold water is taken in one test tube and hot water in another one. Mg ribbon with same size is dropped in each of the test tube: In which test tube, hydrogen is formed with greater speed?



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99. The chemical equation of the industrial preparation of ammonia is given below. $N_2 + 3H_2 \leftrightarrow 2NH_3 + Heat$

Suggest the methods to get more NH_3



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100. $H_2(g) + I_2(g) \leftrightarrow 2HI(g) + Heat$ How do the following circumstances influence the reaction: Increase the concentration of H_2 .



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101. $H_2(g) + I_2 \rightleftharpoons 2HI(g)$ Which of the following has no effect on the given system at equilibrium. (Temperature, Pressure, Concentration).



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102. $H_2(g) + I_2(g) \leftrightarrow 2HI(g) + \text{Heat}$ How do the following circumstances influence the reaction: Increase the temperature.



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103. The formation of SO_3 in the industrial preparation of sulphuric acid is given below. $2SO_2 + O_2 \leftrightarrow 2SO_3 + \text{Heat}$: Explain the effect of concentration of O_2 to get maximum yield of SO_3 ? State reason.



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104. $H_2(g) + I_2(g) \leftrightarrow 2HI(g) + Heat$ How certain circumstances like P,T,V,C influence the reaction: Identify the law related to it. State it.



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105. The chemical equation of a stage in the industrial preparation of sulphuric acid is given below.

$2SO_2 + O_2 \leftrightarrow 2SO_3 + Heat$: Which is the catalyst used in this reaction?



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106. The chemical equation of a stage in the industrial preparation of sulphuric acid is given below.

$2SO_2 + O_2 \leftrightarrow 2SO_3 + \text{Heat}$: What is the influence of the catalyst in equilibrium?



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107. The chemical equation of the industrial preparation of ammonia is given below. $N_2 + 3H_2 \leftrightarrow 2NH_3 + \text{Heat}$: Temperature is to be decreased to get maximum yield of ammonia, according to the Le Chatelier principle. Why?



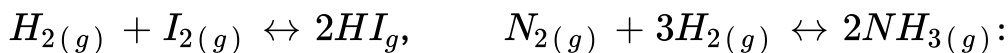
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108. The chemical equation of the industrial preparation of ammonia is given below. $N_2 + 3H_2 \leftrightarrow 2NH_3 + \text{Heat}$: What is the reason for taking an optimum temperature in this reaction?



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109. Analyse the following equations and answer the questions



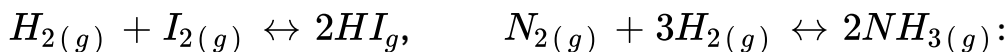
Which of these reactions are affected by change in pressure?

What are the reasons?



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110. Analyse the following equations and answer the questions



Which of these reactions are affected by change in pressure?

What are the reasons?



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111. Catalysts are substances which influence the rate of chemical reactions. Explain how the catalysts influence the rate of reversible reaction?



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112. Some features of a reversible reaction are given below:

Product formation increases when the temperature is increased.

Explain the reason for above inference.



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113. Some features of a reversible reaction are given below:

There is no effect, when the pressure is increased. Explain the reason for above inferences.



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114. $A + B + \text{Heat} \leftrightarrow 2C + D$ This reversible reaction is in equilibrium. What happens to the amount of products under the following conditions: C is removed from the system.



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115. $A + B + \text{Heat} \leftrightarrow 2C + D$ This reversible reaction is in equilibrium. What happens to the amount of products under the following conditions: B is added in excess



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116. $A + B + \text{Heat} \leftrightarrow 2C + D$ This reversible reaction is in equilibrium. What happens to the amount of products under the following conditions: Temperature is increased



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117. $A + B + \text{Heat} \leftrightarrow 2C + D$ This reversible reaction is in equilibrium. What happens to the amount of products under the following conditions: A suitable catalyst is added.



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118. Select the correct statements which are related to the influence of catalyst in a reversible reaction: Forward reaction takes place when a catalyst is used in a reversible reaction.



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119. Select the correct statements which are related to the influence of catalyst in a reversible reaction: Attains equilibrium faster.



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120. Select the correct statements which are related to the influence of catalyst in a reversible reaction: Does not help to produce more product.



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121. Select the correct statements which are related to the influence of catalyst in a reversible reaction: The catalyst increases the rates of both the forward and the backward reactions to the same extent.



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122. Select the correct statements which are related to the influence of catalyst in a reversible reaction: Increases the speed of backward reaction.



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123. Select the correct statements which are related to the influence of catalyst in a reversible reaction: Does not help to produce more product.



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124. $2X(g) + heat \leftrightarrow Y(g) + Z(g)$ Explain the influence of pressure and temperature in this reaction.



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125. Observe the equation of the chemical reactions given below.

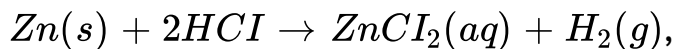
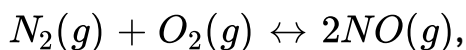
$$NaOH(aq) + HCl(aq) \leftrightarrow 2NaCl(aq) + H_2O(l),$$
$$N_2(g) + O_2(g) \leftrightarrow 2NO(g),$$
$$Zn(s) + 2HCl \rightarrow ZnCl_2(aq) + H_2(g),$$

$2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g)$: In which chemical reaction pressure can influence its speed?



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126. Observe the equation of the chemical reactions given below. $NaOH(aq) + HCl(aq) \leftrightarrow 2NaCl(aq) + H_2O(l)$,

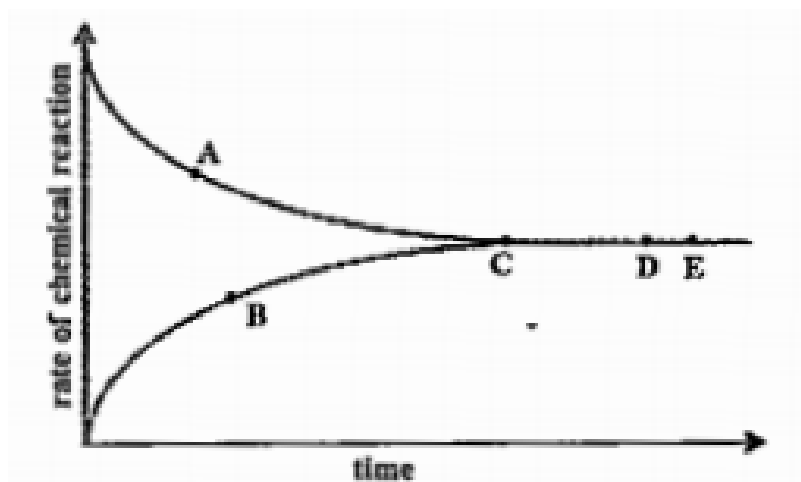


$2SO_2(g) + O_2(g) \leftrightarrow 2SO_3(g)$: In this reaction what changes will be made in pressure for increasing forward reaction? Why?



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127. A graph given below deals with the reversible reaction.

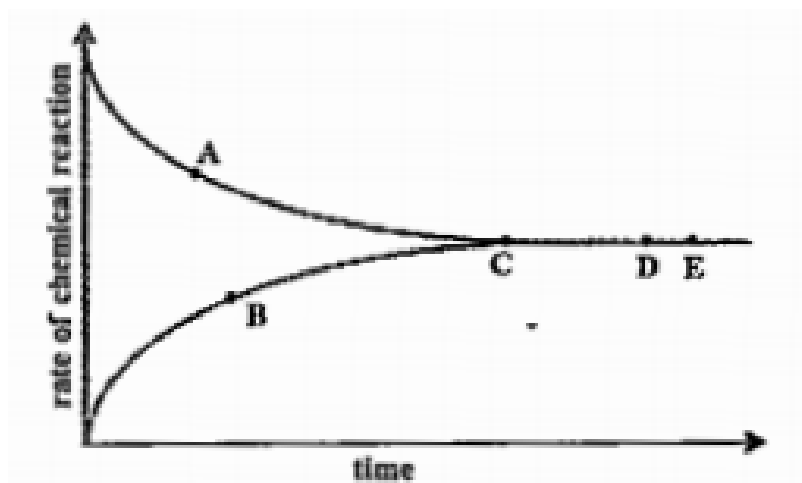


What does A,B,C indicate?



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128. A graph given below deals with the reversible reaction.



What inference can be drawn about the concentration of the reactants and products at the point D and E?

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129. Some chemicals are given below. Sodium chloride, Ammonium hydroxide, Nitric acid, cone. Sulphuric acid, Sodium hydroxide: Which are the substances needed to produce hydrogen chloride?



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130. Some chemicals are given below. Sodium chloride, Ammonium hydroxide, Nitric acid, conc. Sulphuric acid, Sodium hydroxide: Which are the substances needed to produce hydrogen chloride?



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131. When ammonia was leaked to solutions were arising: Spray water, Spray HCl Which method you adopt? Justify your answer.



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132. $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$ How do the following factors influence the forward reaction: One of the products is removed



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133. $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$ How do the following factors influence the forward reaction: Increase in pressure



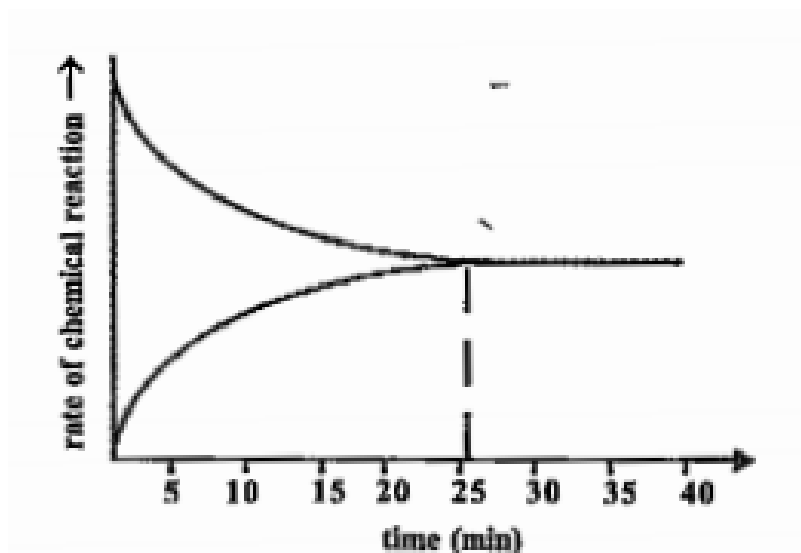
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134. $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g) + \text{heat}$ How do the following factors influence the forward reaction: More N_2 is added.



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135. A graph given below deals with the reversible reaction.



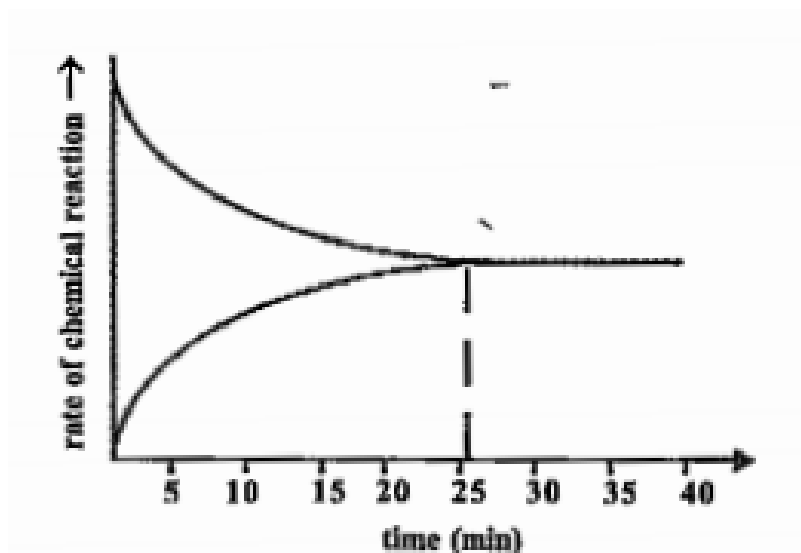
What

happened to the forward and back-ward reactions as time passes?



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136. A graph given below deals with the revers-ible reaction.



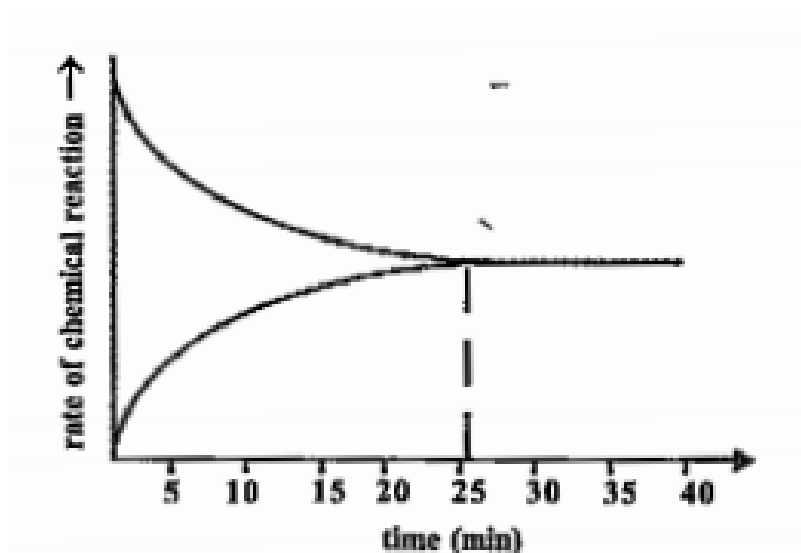
In which

minute does the system attain equilibrium?



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137. A graph given below deals with the revers-ible reaction.



What

change occurs to the equilibrium, when a catalyst is used?



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138. Ammonia is an industrially useful compound of nitrogen:

Name the industrial production of ammonia.



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139. Ammonia is an industrially useful compound of nitrogen:

Write the equation of the reaction.



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140. Ammonia is an industrially useful compound of nitrogen:

Write any two uses of ammonia.



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141. Ammonia is a pungent smelling gas: How does ammonia convert into: liquor ammonia.



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142. Ammonia is a pungent smelling gas: How does ammonia convert into: liquid ammonia



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143. Ammonia is a pungent smelling gas: Write the colour, smell, solubility in water and density of ammonia.



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144. Sulphuric acid is industrially manufactured by using contact process. Write the equation of this process.



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145. Write equations for the following chemicals to prepare from sulphuric acid: Hydrogen chloride



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146. Write equations for the following chemicals to prepare from sulphuric acid: Oleum



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147. Write equations for the following chemicals to prepare from sulphuric acid: Sodium sulphate



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148. Observe the following chemical equation

$Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$: Which one has oxidation state as 0?



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149. Observe the following chemical equation

$Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$: What is its oxidation state after the chemical reaction?



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150. Observe the following chemical equation

$Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$: Is the change oxidation or reduction?



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151. Observe the following chemical equation

$Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$: Which chemical nature of sulphuric acid is found here?



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152. Litmus has coloured in the solution. When $BaCl_2$ solution was added to his. Solution, a white precipitate insoluble in HCl is produced: Write the chemical name and chemical formula of the white precipitate produced when it was added with $BaCl_2$.



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153. Litmus has coloured in the solution. When $BaCl_2$ solution was added to his. Solution, a white precipitate insoluble in HCl is produced: Write the chemical name and chemical formula of the white precipitate produced when it was added with $BaCl_2$.



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154. Litmus has coloured in the solution. When $BaCl_2$ solution was added to his. Solution, a white precipitate insoluble in HCl is produced: Write the chemical name and chemical formula of the white precipitate produced when it was added with $BaCl_2$.



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155. Sulphuric acid is a drying agent as well as a dehydrating agent: Find out the difference between these two activities.



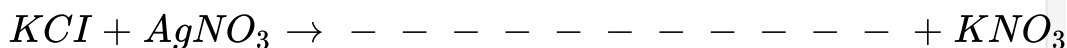
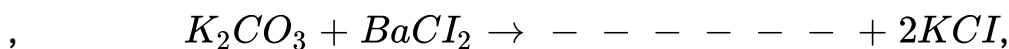
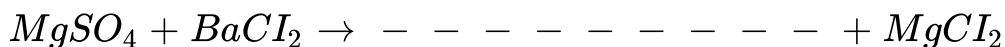
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156. Sulphuric acid is a drying agent as well as a dehydrating agent: Write one example each for these nature of sulphuric acid.



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157. Some incomplete equations are given below:

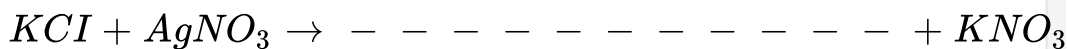
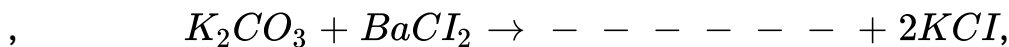
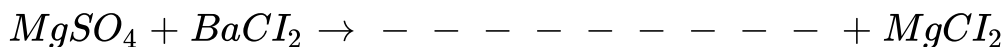


: Complete the equations.



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158. Some incomplete equations are given below:



: Find out the white precipitate in each of the reactions.



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Exercies

1. Fill in the blanks: Liquor ammonia: concentrated aqueous solution of ammonia; Liquid ammonia:.....



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2. Sulphuric acid is not prepared by dissolving SO_3 in water.

Why?



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3. Ammonium chloride is used in the laboratory for the manufacture of ammonia : Which chemical is reacted with ammonium chloride to produce ammonia?



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4. Ammonium chloride is used in the laboratory for the manufacture of ammonia : Write the equation for the reaction.



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5. Ammonium chloride is used in the laboratory for the manufacture of ammonia : Which substance is used to remove the water content in ammonia. Which property is utilized here?



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6. State the Le-Chatelier's principle.



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7. On the basis of Le-Chatelier's Principle explain what will be the effect of removing the product from the system.



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