CHEMISTRY

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PEPIODIC TABLE AND ELECTRONIC CONFIGURATION

Example

1. What is the basis of classification of elements in the periodic table?

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2. Atomic number of sodium is 11 Electronic configuration -2, 8, 1 Group Number -.....

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3. Atomic number of sodium is 11 Electronic configuration -2, 8, 1 Period number -.....

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4. Is the group 1 element a metal or a non metal?

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5. Write the electronic configuration of sodium and argon and complete the Table.

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6. How many electrons are present in the M shell, the outermost shell of argon?

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7. What is the maximum number of electrons that can be accommodated in the M Shell?

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8. The 'K' shell, which is the first shell, has 1 sinshell. The next 'L' shell has 2, and so on. What will be the number of subshells in the 'M' shell and 'N'. M=.....,N=.....

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9. Which sbshell is common to all shells?

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10. What is the maximum number of electrons that can be accommodated in the s subshell?

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11. What is the maximum number of electrons that can be accommodated in the p subshell?

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12. How many electrons are present in the M-shell of an element with atomic number 20?

A) 8

B) 6

C) 18

D) 2

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13. How many electrons are present in helium $(_2He)$?

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14. Complete the subshell electronic configuration of He?

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15. Write the electronic configuration of Lithium $({}_3Li)$

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16. Complete the electronic configuration of beryllium?

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17. Write the electronic configuration of Borom.

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18. Write the electronic configuration of Carbon.

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19. Write the electronic configuration of sodium and argon and complete the Table.

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20. How was the shell wise electronic configu-ration of potassium written?

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21. Compare the energies of 1 s and 2 s subshells. Which one has lower energy?

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22. Among the 3s & 3p subshells which has higher energy?

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23. Among the 3d & 4s sushells which has higher energy?

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24. Write down the subshells in the increasing order of their energies.

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25. Write the subshell wise electronic configu-ration of potassium

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26. The electronic configuration of scandium $(_{21}Sc)$ is

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27. Write the electronic configuration of $22TI$, $23V$, the two elements after Sc.

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28. Which is the nobel gas preceding sodium $(11Na)$/

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29. Write its subshell electronic configuration. (nobel gas preceding sodium)

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30. Subshell electronic configuration of sodium?

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31. Using the symbol of neon, write the subshell electronic configuration of sodium?

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32. Which element has a valency 1?

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33. Which element shows metallic character?

Elements : X, Y

Atomic number: 12, 17 respectively.

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34. Which element has the highest ionisation energy?

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35. Write the chemcial formula of compound formed by the combination of X and Y and label the oxidation states?

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36. Where is the position of d block elements in the periodic table?

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37. From which period onward does the d block begin?

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38. Complete the table 1.16

Compound	Oxidation state of Fe	Symbol of Fe ions
$FeCl_2$		
$FeCl_3$		

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39. How does Fe change to Fe^{2+}?

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40. Write down the subshell electronic config-uration of Fe^{2+}.

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41. If so, which will be the subshel from which iron loses the third electron?

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42. Write the electronic configuratoin of Fe^{3+}

Watch Video Solution

43. Write the subshell electronic configuration of Manganese(Mn).

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44. Complete the table 1.17

Compound	Oxidation state of Mn	Subshell electronic configuration of Mn ions
$MnCl_2$		
MnO_2		
Mn_2O_3		
Mn_2O_7		

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45. Examine these compounds availible. Find more coloured compounds and extend the list.

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46. Which elements shows +2 oxidaion state?

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47. Which elements contains 5 electrons in the outermost shell?

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48. Which is the element that has 5 p electrons in the outermost shell?

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49. Which are the elements in which the last electron enters the d subshell?

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50. Which element has the highest ionisation energy?

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51. Which is the highly reactive non metal?

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52. Which elements shows -2 oxidaion state?

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53. [Ar] 3d5 4s2 is the sub shell electronic configuration of an element.

Which is the element?

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54. [Ar] 3d5 4s2 is the sub shell electronic configuration of an element.

Write down the complete subshell electronic configuration?

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55. [Ar] 3d5 4s2 is the sub shell electronic configuration of an element.

Write any two characteristics of this element?

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56. When the last electron of an atom was filled in the 3d subshell, the subshell electronic configuration was recorded as $3d^8$. Answer the questions related to this

atom. Complete subshell electronic config-uration, Atomic number, Block, Period number, Group number.

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57. Pick out the wrong ones from the subshell electronic configuration given below: $1s^2 2s^2 2p^7$

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58. Pick out the wrong ones from the subshell electronic configuration given below: $1s^2 2s^2 2p^2$

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59. Pick out the wrong ones from the subshell electronic configuration given below: $1s^2 2s^2 2p^5 3s^1$

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60. Pick out the wrong ones from the subshell electronic configuration given below: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^1$

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61. Pick out the wrong ones from the subshell electronic configuration given below:

$1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^2$

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62. The element X in group 17 has 3 shells. If so: Write the subshell electronic configura-tion of the element.

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63. The element X in group 17 has 3 shells. If so: Write the period number.

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64. The element X in group 17 has 3 shells. If so: What will be the chemical formula of the compound formed if the element X reacts with element Y of the third period which contains one electron in the p subshell?

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65. The element Cu with atomic number 29 undergoes chemcial reaction to formation with oxidation number +2: write down the subshell electronic con-figuration of the ion.

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66. The element Cu with atomic number 29 undergoes chemcial reaction to formation with oxidation number +2: Can this element show variable valency? Why?

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67. The element Cu with atomic number 29 undergoes chemcial reaction to formation with oxidation number +2: Write downthe chemical formula of one compound formed when this element re-acts with chlorine $(17CI)$.

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68. Certain subshells of an atom are given below. 2s,2d,3f,3d,5s,3p: Which are the subshells that are not possible? Give the reason.

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69. Certain subshells of an atom are given below. 2s,3d,3f,4d,5s,3p: Which are the subshells that are not possible? Give the reason.

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70. Prepare the comprehensive table which indicates the name, symbol, electron configuration, subshell configuration of elements having atomic number 1 to 36?

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71. Some information related to the elements of the p block in the 17th group of the periodic table are given in the table below. Complete the able and analyse the following questions.

a) What is the family name of element belonging to 17th group?

b) what is their common valency?

c) Which element has the highest electronegativity?

d) Which element has highest ionisation energy?

e) List out the name and chemical formula of the compounds formed by this elements with s block elements.

'(##VPU_TTT_CHE_X_P01_C01_E02_002_Q01##)'

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72. What is the family names of elements be-longing to the 17 th group?

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73. What is the common valency of elements belonging to the 17 th group?

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74. Which element has the highest electronega-tivity in 17th group?

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75. Which element has the highest ionisation energy?

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76. List out the name and chemcial formula of the compound formed by this elements Halogen Family with s block elements?

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77. What are shells and subshells.

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78. No of electrons in KLMN shell.

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79. No of electrons present in in the 3rd shell.

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80. Which sbshell is common to all shells?

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81. Write names of subshells in accordance with increasing energy level.

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82. Identify the incorrect subshell electronic configuration.

A. $-1s^2$

B. $-1s^2 2p^6$

C. $-1s^2 2s^2 2p^6$

D. $-1s^2 2s^2 2p^6 3s^2 3p^2$

Answer:

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83. Atomic number of iron is 26. It exhibits Fe^{2+}, Fe^{3+} oxidation state. Write the subshell elec-tronic configuration.

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84. Magnanese, a d-block elements exhibits different oxidation state.Why?

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85. Write the oxidation number and subshell electronic configuration K, CI and O.

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86. Find out atomic number,group, block period using subshell electronic configuration and then complete the table.

Subshell electronic configuration	Atomic number	Group	Block	Period
$1s^2\ 2s^2 2p^6$				
$1s^2\ 2s^2\ 2p^6\ 3s^1$				
$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^2\ 3d^5\ 4s^1$				
	25			
	28			
	26			

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87. Write down the characteristics of s,d,p,f block elements.

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88. Write down subshell electronic configura-tion of Cu^{1+} and Cu^{2+}

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89. How many 's' subshell electrons are present in $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$

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90. 11,17,10 are the atomic number of elements X,Y and Z: Write down their subshell electronic configuration, group, block, period.

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91. 11,17,10 are the atomic number of elements X,Y and Z: Write the molecular formulac of the compound formed when any two of the above elements are combined.

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92. 11,17,10 are the atomic number of elements X,Y and Z: Write down the oxidation numbers of the elements in the compound formed by any two of the element. Write the subshell electronic configuration of both ions.

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93. Element 'X' is having atomic number 28, it gives two electrons to element 'Y': Write down the electronic configuration of 'X' and its ion.

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94. Element 'X' is having atomic number 28, it gives two electrons to element 'Y': In which block 'X' belongs?

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95. Element 'X' is having atomic number 28, it gives two electrons to element 'Y': Write down the characteristics of that block.

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Sc - [Ar] $3d^1 4s^2$	Ca - [Ar] $4s^2$
Mg - [Ne] $3s^2$	Co - [Ar] $3d^7 4s^2$

96.

Write down the group and period of each element.

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97.

Sc - [Ar] $3d^1 4s^2$	Ca - [Ar] $4s^2$
Mg - [Ne] $3s^2$	Co - [Ar] $3d^7 4s^2$

What are the use of writing electronic configuration this fashion?

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98. $_{24}Cr - [Ar]3d^5 4s^1$ Why chromium exhibits such electronic configuration?

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99. $_{29}Cu - [Ar]3d^{10} 4s^1$ Why chromium and copper exhibits such electronic configuration?

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100. The electronic configuration of the elements are given below.

A-$1s^2 2s^2 2p^6 3s^2 3p^4$

B-$1s^2 2s^2 2p^6 3s^2$

C-$1s^2 2s^2 2p^6 3s^2 3p^5$

D-$1s^2 2s^2 2p^6 3s^1$ Which of these elements show +2 oxida-tion state?

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101. The electronic configuration of the elements, are given below.

A-$1s^2 2s^2 2p^6 3s^2 3p^4$

B-$1s^2 2s^2 2p^6 3s^2$

C-$1s^2 2s^2 2p^6 3s^2 3p^5$

D-$1s^2 2s^2 2p^6 3s^1$ Which metal belongs to 17^{th} group?

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102. The electronic configuration of the elements are given below.

A-$1s^2 2s^2 2p^6 3s^2 3p^4$

B-$1s^2 2s^2 2p^6 3s^2$

C-$1s^2 2s^2 2p^6 3s^2 3p^5$

D-$1s^2 2s^2 2p^6 3s^1$ Which is the period number of the elem-ent A? What is the basis of your find-ings?

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103. The electronic configuration of the eleme-nts, A,B,C,D are given below. $1s^2 2s^2 2p^6 3s^2 3p^4, 1s^2 2s^2 2p^6 3s^2$, $1s^2 2s^2 2p^6 3s^2 3p^5, 1s^2 2s^2 2p^6 3s^1$ Which of these elements can form basic oxides?

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104. Two compounds of iron are given below. $FeSO_4$ $Fe_2(SO_4)_3$ (The oxidation state of sulphate radical is -2): Which of the these compounds show +2 ox-idation state of Fe?

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105. Two compounds of iron are given below. $FeSO_4$ $Fe_2(SO_4)_3$ (The oxidation state of sulphate radical is -2): Which compounds has Fe^{3+} ions?

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106. Two compounds of iron are given below. $FeSO_4$ $Fe_2(SO_4)_3$ (The oxidation state of sulphate radical is -2): Write the subshell electronic confi-gruration of Fe^{3+} ion.

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107. Two compounds of iron are given below. $FeSO_4$ $Fe_2(SO_4)_3$ (The oxidation state of sulphate radical is -2): Why do transition elements show vari-able oxidation states?

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108. Identify the incorrect eletronic configur-ations and correct them. $1s^2 2s^2 2p^3$

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109. Identify the incorrect eletronic configur-ations and correct them. $1s^2 2s^2 2p^6 3s^1$

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110. Identify the incorrect eletronic configur-ations and correct them. $1s^2 2s^2 2p^6 2d^7$

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111. Identify the incorrect eletronic configur-ations and correct them. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4$

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112. Completethetable.

Outermost electronic configuration	Group Number	Oxidation state
$3s^2\,3p^4$	16	a. .....
$3s^1$	b.	+1
$2s^2\,2p^3$	c.	d. .....
$3d^{10}\,4s^2$	e.	f. .....

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113. Two compounds XY_2, XZ_4 are given. The oxidation state of Z is 1. What will be the oxidation state of Y?

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114. Two compounds XY_2, XZ_4 are given. The oxidation state of Z is 1. Write the molecular formula of the compound formed by Y when it combines with aluminum (AI) having oxidation state +3.

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115. Pick out the statements which suits to f block elements. a. All of them are naturally occurring elements. b. Uranium and Thorium are f block elements. c. Last electrons is filled in the shell preceding the outermost shell. d. last electrons are filled up in the antepenultimate shell. e. Includes some

radioactive elements. f. Some of them are used as catalyst in petroleum industry.

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116. The atomic number of four elements are given below.

(The symbols are not real)

A - 8 B - 10

C - 12 D - 18

Write the sub-shell electronic configu-rationof the elements.

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117. The atomic number of four elements are given below.

(The symbols are not real)

A - 8	B - 10
C - 12	D - 18

Which of them are inert gases ?

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118. The atomic number of four elements are given below.

Write the chemical formula of the compound formed by two elements other than inert gases.

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119. The subshell electronic configuration of two elementsends asfollows.(Symbols are not real) $P-3s^2$ $Q-3p^4$: Write the complete subshell electronic configuration.

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120. The subshell electronic configuration of two elementsends asfollows.(Symbols are not real) $P-3s^2$ $Q-3p^4$: Find out the oxidation state of each element.

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121. The subshell electronic configuration of two elementsends asfollows.(Symbols are not real) $P-3s^2$ $Q-3p^4$: The chemical formula of the compound formed by these elements is PQ.Is this statement correct? Justify your answer.

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122. Match the following.

Block	Outer most electronic configuration	Characteristics
s	$3p^5$	Most of the compo-unds are coloured.
p	$3d^14s^2$	Includes Lanthanoids (6^{th}period)
d	$4f^15d^16s^2$	Highest atomic radius in the respective period.
f	$3s^1$	Highelectro negativity

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123. The atomic number of two elements are given below. $Si-14$ $Ni-28$: Write the subshell electronic configu- I ration of these elements.

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124. The atomic number of two elements are given below. $Si-14$ $Ni-28$: Find out the group and period of each I element.

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125. The element 'X' has 4 shells, and its3d sub-shell has6 electrons.(Symbol is not real): Write the complete electronic configuration of the element

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126. The element‘X’ has 4 shells, and its3d sub-shell has6 electrons.(Symbol is not real): What is its group number? Which is the block?

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127. The element‘X’ has 4 shells, and its3d sub-shell has6 electrons.(Symbol is not real): Write any two characteristics of the bl-ock to which element X belongs to.

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128. The element‘X’ has 4 shells, and its3d sub-shell has6 electrons.(Symbol is not real): From which subshell the electrons are lost when the element X shows +2 oxi-dationstate

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129. The outermost electronic configuration of the element A is $2s^2 2p^2$. (Symbolis not real) c. Writ: Find out thegroup number and block of the element.

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130. The outermost electronic configuration of the element A is $2s^2 2p^2$. (Symbolis not real) c. Writ: Write the chemical formula of the com-pound formed by A when it combines with chlorine.

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131. The outermost electronic configuration of the element A is $2s^2 2p^2$. (Symbolis not real) c. Writ: Write the complete electronic configura-tion of the element just below 'A' in the periodic table.

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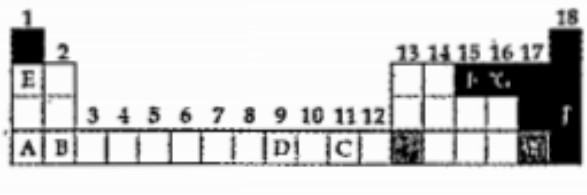
132. The figure of an incomplete periodic table is given below.



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Which one of these elements show -2 oxidation state?

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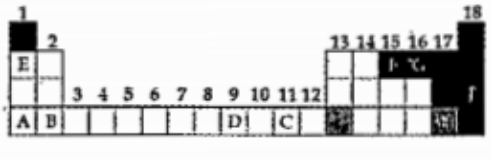
133. The figure of an incomplete periodic table is given below.



Which of these elements have 3 electrons in their outermost p subshell ?

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134. The figure of an incomplete periodic table is given below.

Which element has the highest atomic radius ? Which one has the least ?

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135. The figure of an incomplete periodic table is given below.

Which of these elements show variable oxidation state?

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136. The figure of an incomplete periodic table is given below.

Which of these element has the highest ionization energy ?

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137. Examine the given electronic configurations.

A- $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$

B- $1s^2 2s^2 2p^6 3s^1$

C-$1s^2 2s^2 2p^6 3s^2 3p^6$

D-$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$: Which of these elements belongs to 4^{th} period ?

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138. Examine the given electronic configurations.

A- $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$

B- $1s^2 2s^2 2p^6 3s^1$

C-$1s^2 2s^2 2p^6 3s^2 3p^6$

D-$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$: Which elements belongs to the same group?

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139. Examine the given electronic configurations.

A-$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$

B-$1s^2 2s^2 2p^6 3s^1$

C-$1s^2 2s^2 2p^6 3s^2 3p^6$

D-$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$: Which element doesn't participate in che-mical reactions generally ?

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140. Examine the given electronic configurations.

A-$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$

B-$1s^2 2s^2 2p^6 3s^1$

C-$1s^2 2s^2 2p^6 3s^2 3p^6$

D-$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$: Which element has highest metallic ch-aracter ?

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141. The atomic number of the elements X and Y are 20, 26 respectively. When these ele-ments combines with chlorine, three com-pounds XCI_2, YCI_2, YCI_3 areformed: What is the speciality of the oxidation number of Y,compared to thatof X?

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142. The atomic number of the elements X and Y are 20, 26 respectively. When these ele-ments combines with chlorine, three com-pounds XCI_2, YCI_2, YCI_3 areformed: Explain the reasonfor this,on the basis of the subshellbased electronic configu- ration.

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143. Arrange the following sub-shells in the in-creasing order of energy 5p, 2s, 4f, 3s, 4s, 3d, 6s.

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144. Last electron in f-block elements goes to: Which shell? Outer shell/Penultimate shell /Antepenultimate shell

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145. Last electron in f-block elements goes to: Which sub-shell? Outer f-subshell Penultimate f-subshell/Anitepenulti mate f-subskell.

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146. Sub-shell electronic configuration of X is given below.$1s^2, 2s^2, 2p^5$: The element Y is coming just below the element in same group. Then write the sub-shell electronic configuration of Y.

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147. Sub-shell electronic configuration of X is given below.$1s^2, 2s^2, 2p^5$: Write the sub-shell electronic configuration of the element next to X in same period.

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148. A compound of vanadium pentoxide (V_2O_5) is used as catalyst: What is the oxidation state of vanadium in this compound?

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149. A compound of vanadium pentoxide (V_2O_5) is used as catalyst: How vanadium ion is represented?

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150. A compound of vanadium pentoxide (V_2O_5) is used as catalyst: Write thesub-shell electronic

configur-ation of this ion (V- 23)

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151. Find the wrong electronic configurations fromthefollowing. Whatiswrongin these? $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 3d^9, 4s^2$

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152. Find the wrong electronic configurations fromthefollowing. What is wrong in these? $1s^1$

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153. Find the wrong electronic configurations fromthefollowing. Whatiswrongin these? $1s^2, 2s^1, 2p^6$.

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154. Find the wrong electronic configurations fromthefollowing. Whatiswrongin them? $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$.

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155. Group and period number of two elements are given. P-group17,period-3, Q-group2,period-3: Write the

sub-shell electronic configu-ration of each.

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156. Group and period number of two elements are given. P-group17,period-3, Q-group2,period-3: Write the chemical formula of the co- I mpound formed by their combination.

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157. Write the reason for the statement given below: d-block elements in the same period shows similarity.

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158. Write the reason for the statement given below: Transition elements shows variable oxidation state.

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159. Write the sub-shell electronic configuration of following elements. Predict the block, groupand period. (Symbolsare not real): M-27

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160. Write the sub-shell electronic configuration of following elements. Predict the block, groupand period. (Symbolsare not real): N-19

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161. Write the sub-shell electronic configuration of following elements. Predict the block, groupand period. (Symbolsare not real): P-15

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162. Which element is having 5 electrons in valence shell?

A. N

B. O

C. F

D. C

Answer:

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163. Which elements are having 2electrons in valencesub-shell?

A. N

B. C

C. F

D. O

Answer:

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164. Which element is having last electron in 3p?

A. Na

B. Al

C. Mg

D. F

Answer:

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165. Atomic number of someelementsaregiven. $A-15, B-8, C-11, D-18, E-20. F-34, G-10$

: Which are the elementsin sameperiod?

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166. Atomic number of someelementsaregiven. $A-15, B-8, C-11, D-18, E-20. F-34, G-10$: Which are the elements in same group?

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167. Electronic configuration of some elements are given. Write answers to the following questions. $[Ne]3s^1$, $[Ar]3d^2$, $4s^2$, $[Xe]6s^1$, $[Ne]3s^2$: Which metal is having high reactivity ?

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168. Pick thewrongstatement fromthe follow-ing: Elements with atomic number 5 belong to group 15.

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169. Pick thewrongstatement fromthe follow-ing: Electronic configuration of scandium (Atomic number 21) is 2,8,8,3.

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170. Pick thewrongstatement fromthe follow-ing: d-block elements are known as transi-tion elements.

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171. Pick thewrongstatement fromthe follow-ing: All s-blockelementsaremetals.

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172. Write the electronic configuraiton of Mg 2+ ion

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173. Fill in the blanks: the maximum number of electrons that can be accomondated in the L shell is.....

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174. Fill in the blanks: s and p block elecments together are known as .....

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175. The oxidation state of Cromium in the compound Cr_2O_3 is +3: What is the oxidation state of Cromium in the compound CrO_3?

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176. The oxidation state of Cromium in the compound Cr_2O_3 is +3: Write the subshell electronic configuration fo Cromium in Cr_2O_3 and CrO_3. (Atomic numbr of `Cr-24).

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177. The Atomic number of Copper is 29: Write the subshell electronic configuration of copper.

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178. The Atomic number of Copper is 29: Cu_2O and `CuO are 2 compounds of copper. Find the oxidation state of copper in there compounds.

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179. The Atomic number of Copper is 29: Write down the subshell configuration of copper ions in there compounds.

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180. The atom of a compound in the 4^{th} period has 3 electrons in the d subshell just inside the outer most shell: Write down the complete electronic configuration of this compound.

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181. The atom of a compound in the 4^{th} period has 3 electrons in the d subshell just inside the outer most shell: To which group does this compound belong?

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