CHEMISTRY

BOOKS - MAXIMUM PUBLICATION

QUESTION PAPER FEB 2020

Example

1. Which of the following subshells is not possible in an atom ?

A. 2p

B. 3f

C. 1s

D. 4d

Answer: B

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2. In which of the following, oxidation number of chlorine is +5?

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3. The relationship between volume and number of molecules of a gas at constant temperature and pressure is known as .....

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4. When molten sodium chloride (NaCl) is electrolysed, the gas liberated at the anode is.....

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5. Find the relation and fill up suitably

Iron Haematite

Aluminium: .....

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6. The molecular mass of water (H_2O) is 18.

Find the mass of 1GMM of H_2O .

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7. The molecular mass of water (H_2O) is 18.

How many moles of molecules are there in 180g of H_2O ?

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8. . Hematite is converted into iron by using the blast fumace .

Which are the substances fed into the blast furnace along with the ore of iron ?

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9. . Hematite is converted into iron by using the blast fumace .

Which compound acts as the reducing agent in the blast furnace ?

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10. $FeCl_2$ and $FeCl_3$ are two different chlorides of iron. [Hint : Atomic Number of $Fe = 26$

Oxidation state of CI = (-1)

Find the oxidations state of Fe in $FeCl_2$.

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11. $FeCl_2$ and $FeCl_3$ are two different chlorides of iron. [Hint : Atomic Number of $Fe = 26$

Oxidation state of CI = (-1)

Write down the subshell wise electronic configuration of Fe^{3+}

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12. Which homologous series do the hydrocarbons with general formula C_n , H_{2n} belong to?

[alkane, alkene, alkyne]

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13. Write down the structural formula of a member of the same homologous series having 3 carbon atoms .

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14. Soaps and detergents are cleansing agents. Name the byproduct in the industrial production of soap.

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15. Soaps and detergents are cleansing agents. How does excessive use of detergents destroy aquatic life ?

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16. Electroplating is one of the practical utilities of electrolysis .Copper can be coated on an iron bangle by this process.

Which metal is connected to negative terminal of the battery in this process ?

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17. Electroplating is one of the practical utilities of electrolysis .Copper can be coated

on an iron bangle by this process.

Which is the electrolyte used here ?

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18. Electroplating is one of the practical utilities of electrolysis .Copper can be coated on an iron bangle by this process.

Write down any other practical utility of electrolysis.

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19. Which are the chemicals required for the preparation of ammonia in the laboratory?

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20. Which is the drying agent used to remove moisture from ammonia ?

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21. The gas jar used for collecting ammonia is kept inverted .Why?

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22. The atomic number of element X is 12. [Symbol is not real]

Write the subshell wise electronic configuration of X.

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23. The atomic number of element X is 12. [Symbol is not real]

Which period does this element belong to ?

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24. The atomic number of element X is 12. [Symbol is not real]

Which block does it belong to ?

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25. The atomic number of element X is 12. [Symbol is not real]

Write down the molecular formula of the chloride of X.

[Hint Valency of CI=1]

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26. $2SO_2 + O_2 \rightarrow 2SO_3$

represents an important stage in the Industrial Preparation of Sulphuric acid.

By what name is the Industrial Preparation of Sulphuric acid known as?

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27. $2SO_2 + O_2 \rightarrow 2SO_3$

represents an important stage in the Industrial Preparation of Sulphuric acid.

Which is the catalyst used in this process ?

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28. $2SO_2 + O_2 \rightarrow 2SO_3$

represents an important stage in the Industrial Preparation of Sulphuric acid.

How do the following changes influence the forward reaction ?

i) More oxygen (O_2) is added.

ii) Pressure is decreased.

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29. The structural formula of an organic compound is given below:

$CH_3 - CH_2 - O - CH_3$

Identify the functional group present in this compound?

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30. The structural formula of an organic compound is given below:

$CH_3 - CH_2 - O - CH_3$

What are the compounds with the given functional group commonly called ?

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31. The structural formula of an organic compound is given below:

$CH_3 - CH_2 - O - CH_3$

Write down the structural formula of its functional isomer and its IUPAC name.

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