



CHEMISTRY

BOOKS - MAXIMUM PUBLICATION

QUESTION PAPER MARCH 2019

Example

1. Which of the following molecule can undergo addition reaction?

A. methane

B. ethane

C. propene

D. butane

Answer: C



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2. The glass used to make lenses and prisms
is.....



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3. Atomic mass of Nitrogen is 14. which of the following sample contain 6.022×10^{23} Nitrogen atoms?

(7g Nitrogen, 14g Nitrogen, 28g Nitrogen, 1g Nitrogen)



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4. The ore of a metal is lighter than the impurities. Which method is suitable for its

concentration?



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5. A fresh piece of Mg ribbon loses its luster after a few days. This is due to the formation of the compound.....



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6. The last subshell of an element is 3p and there are 3 electrons in it.

Write the complete electronic configuration of the element.



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7. The last subshell of an element is $3p$ and there are 3 electrons in it.

Identify its period and group.



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8. An iron nail is dipped in $CuSO_4$, solution.

(Reactivity order $Fe > Cu$)

What is the change that can be noticed on the iron nail after a while?



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9. An iron nail is dipped in $CuSO_4$, solution.

(Reactivity order $Fe > Cu$)

Write down the CHEical equation of the oxidation reaction occurs here.





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10. 4g of NaOH is dissolved in water and the volume is made upto 1 L.

(1 mole of NaOH = 40 g)

Calculate the molarity of the resultant solution.



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11. 4g of NaOH is dissolved in water and the volume is made upto 1 L.

(1 mole of NaOH = 40 g)

How will you make 1 M solution of NaOH using the same amount (4g) of NaOH?



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12. Concentrated Cu_2S is converted into oxide by roasting.

Write the process of roasting.



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13. Concentrated Cu_2S is converted into oxide by roasting.

How impurities like sulphur and phosphorus are removed in this process?



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14. Ethanoic acid is an organic compound having industrial values.

How ethanoic acid is manufactured industrially?





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15. Ethanoic acid is an organic compound having industrial values.

Give any one use of ethanoic acid.



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16. The CHEical equation for the manufacture of ammonia is $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

Complete the following:

1 mol N_2 + H_2 \rightarrow NH_3 .



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17. The chemical equation for the manufacture of ammonia is $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

Calculate the amount of H_2 required to react with 28 g of N_2 , completely.

[Hint Molecular mass of $N_2 = 28$, $H_2 = 2$]



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18. The chemical equation for the manufacture of ammonia is $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

What will be the volume of NH_3 , formed at STP, if 22.4L of N_2 is completely reacted?



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19. Alumina is mixed with cryolite and subjected to electrolysis to extract aluminium.

Why cryolite is added to alumina?



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20. Alumina is mixed with cryolite and subjected to electrolysis to extract aluminium.

Which are the ions present in alumina?



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21. Alumina is mixed with cryolite and subjected to electrolysis to extract aluminium.

Write the equation of the reduction reaction taking place at negative electrode.



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22. Zinc piece and zinc powder are taken in two test tubes and equal amount of dil. HCl is added.

In which test tube does the reaction proceed faster?



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23. Zinc piece and zinc powder are taken in two test tubes and equal amount of dil. HCl is

added.

Give the reason.



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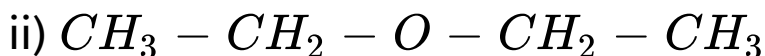
24. Zinc piece and zinc powder are taken in two test tubes and equal amount of dil. HCl is added.

Give an instance from daily life, where such condition is made use.



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25. The structure of two organic compounds are given below:



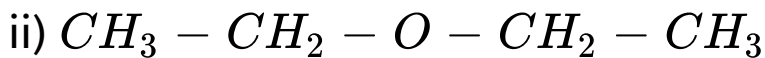
Write the molecular formula of these compounds.



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26. The structure of two organic compounds are given below:



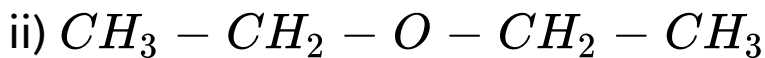


Which type of isomerism do they exhibit?



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27. The structure of two organic compounds are given below:

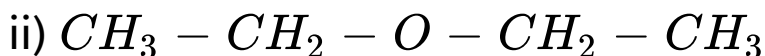


Explain this isomerism.



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28. The structure of two organic compounds are given below:



Write the structural formula of a position isomer of compound



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29. The atomic number of an element is 19.

Write the subshell electronic configuration.



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30. The atomic number of an element is 19.

Identify its group, period, block and oxidation state.



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31. The atomic number of an element is 19.

Write any one characteristic of the block to which the element belongs.



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32. • Aspirin is an antipyretic.

• Amoxicillin is an antibiotic

Give the functions of antipyretics and antibiotics.



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33. • Aspirin is an antipyretic.

• Amoxicillin is an antibiotic

Write any two unhealthy practices among people in using medicines.



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