



CHEMISTRY

BOOKS - NAVNEET PUBLICATION

CHEMICAL REACTIONS AND EQUATIONS

Solved

1. What are the types of molecules of elements and compounds?

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2. What is meant by valency of elements?

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3. What is the requirement for writing molecular formulae of different compounds? How are the molecular formulae of the compounds written?

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4. Is it possible to produce hydrogen by decomposition of water by means of heat, electricity or light?

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Exercise

1. Fill in the blanks

Organic waste is decomposed by microorganism and as a result manure and..... are formed.

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2. Fill in the blanks

..... is formed on mixing yeast in glucose solution under proper condition.



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3. Fill in the blanks

The chemical reaction during which H_2 is lost is termed as.....



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4. Fill in the blanks

Corrosion can be prevented by using.....



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5. Fill in the blanks

The chemical reactions in which heat is liberated are called

reactions.

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6. The chemical formula of the rust is

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7. The chemical reaction in which heat is absorbed is called ___ reaction.

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8. Fill in the blanks

The process of rusting of iron is process.

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9. Fill in the blanks

When oil and fats are oxidised or even allowed to stand in air for a long time, they become



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10. Fill in the blanks

..... are used to prevent oxidation of food.



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11. Fill in the blanks

Carbon dioxide is passed through water. The reaction is areaction.



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12. Fill in the blanks

Calcium carbonate is heated . The reaction is areaction.



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13. Fill in the blanks

Zinc strip is dipped in a CuSO_4 solution. The reaction is areaction.



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14. Fill in the blanks

Silver nitrate solution is added to NaCl solution . The reaction is areaction.



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15. Fill in the blanks

The slow process of decay or destruction of a metal due to effect of air, moisture and acids on it is known as.....



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16. Choose the correct alternative and write it along with its allotted alphabet:

The reaction of iron nail with copper sulphate solution is.....reaction.

- A. double displacement
- B. displacement
- C. combination
- D. decomposition

Answer:



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17. Choose the correct alternative and write it along with its allotted alphabet:

Reddish brown deposition formed on iron nails kept in a solution of copper sulphate is:

A. Cu_2O

B. Cu

C. CuO

D. CuS

Answer:



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18. Choose the correct alternative and write it along with its allotted alphabet:

The reaction $\text{CuSO}_4 (\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{ZnSO}_4 (\text{aq}) +$

$\text{Cu}(\text{s})$ is a reaction.

is

a

.....reaction

- A. displacement
- B. double displacement
- C. decomposition
- D. combination

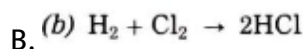
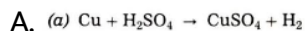
Answer:

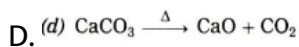
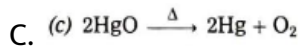


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19. Choose the correct alternative and write it along with its allotted alphabet:

.....is a combination reaction.



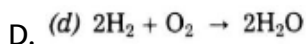
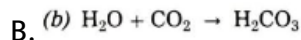
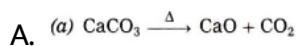


Answer:

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20. Choose the correct alternative and write it along with its allotted alphabet:

.....is a decomposition reaction.



Answer:

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21. Choose the correct alternative and write it along with its allotted alphabet:

In a chemical equation the..... are written on the left hand side.

A. products

B. reactants

C. catalysts

D. elements

Answer:



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22. Choose the correct alternative and write it along with its allotted alphabet:

The delta sign written above the arrow indicates of the reaction

A. reactant

B. product

C. heat

D. direction of the reaction

Answer:



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23. Choose the correct alternative and write it along with its allotted alphabet:



$\text{KNO}_{3(aq)}$ is a/an reaction.

The reaction

A. exothermic

B. endothermic

C. oxidation

D. reduction

Answer:

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24. Choose the correct alternative and write it along with its allotted alphabet:

The reaction $\text{NaOH}_{(s)} + \text{H}_2\text{O}_{(l)} \rightarrow \text{NaOH}_{(aq)}$ is
is a/ anreaction.

- A. exothermic
- B. endothermic
- C. oxidation
- D. reduction

Answer:

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25. Choose the correct alternative and write it along with its allotted alphabet:

A solution of $Al_2(SO_4)_3$ in water is.....

A. blue

B. pink

C. green

D. colourless

Answer:



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26. Choose the correct alternative and write it along with its allotted alphabet:

Carbon dioxide.....

A. turns lime water milky

B. is odourless

C. is colourless

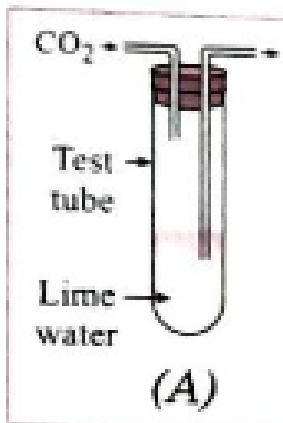
D. All the three (a),(b) and (c) are correct

Answer:

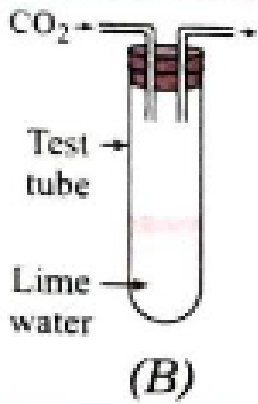
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27. Choose the correct alternative and write it along with its allotted alphabet:

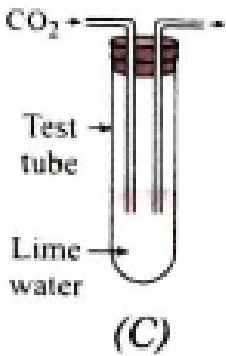
..... is the correct set up to pass CO_2 through lime water.



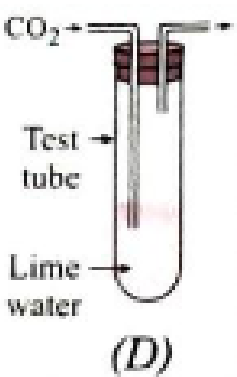
A.



B. _____



C. _____



D. _____

Answer:

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28. Choose the correct alternative and write it along with its allotted alphabet:

When is passed through fresh lime water, it turns milky.

A. H_2

B. CO

C. CO_2

D. SO_2

Answer:



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29. Choose the correct alternative and write it along with its allotted alphabet:

Magnesium reacts with con. HCl to form..... salt

A. copper chloride

B. ferrous chloride

C. calcium chloride

D. magnesium chloride

Answer:

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30. Choose the correct alternative and write it along with its allotted alphabet:

Zinc reacts with hydrochloric acid. The reaction is areaction

A. combination

B. decomposition

C. displacement

D. double decomposition

Answer:



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31. Choose the correct alternative and write it along with its allotted alphabet:

In a double displacement reaction.....

- A. ions remain at rest
- B. ions get liberated
- C. ions are exchanged
- D. ions are not created

Answer:



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32. Choose the correct alternative and write it along with its allotted alphabet:

Combustion of coal in air is a..... reaction.

- A. combination
- B. displacement
- C. decomposition
- D. double decomposition

Answer:



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33. Choose the correct alternative and write it along with its allotted alphabet:

The crystals of ferrous sulphate are.....

- A. blue in colour

B. pink in colour

C. pale green in colour

D. colourless

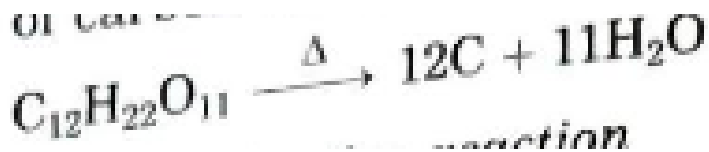
Answer:



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34. Choose the correct alternative and write it along with its allotted alphabet:

Identify the type of following chemical reaction of carbon compound.



A. combination reaction

B. displacement reaction

C. decomposition reaction

D. double displacement

Answer:



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35. Choose the correct alternative and write it along with its allotted alphabet:

The conversion of ferrous sulphate into ferric sulphate is.....reaction

- A. oxidation
- B. displacement
- C. electrolysis
- D. reduction

Answer:



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36. Choose the correct option from the bracket and explain the statement giving reason:

(oxidation, displacement, electrolysis, reduction, zinc, copper, double, displacement, decomposition)

To prevent rusting, a layer of ___ metal is applied on iron sheets

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37. Choose the correct option from the bracket and explain the statement giving reason:

(oxidation, displacement, electrolysis, reduction, zinc, copper, double, displacement, decomposition)

The conversion of ferrous sulphate to ferric sulphate is _____ reaction.

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38. Choose the correct option from the bracket and explain the statement giving reason:

(oxidation, displacement, electrolysis, reduction, zinc, copper, double, displacement, decomposition)

When electric current is passed through acidulated water _ of water takes place.

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39. Choose the correct option from the bracket and explain the statement giving reason:

(oxidation, displacement, electrolysis, reduction, zinc, copper, double, displacement, decomposition)

Addition of an aqueous solution of $ZnSO_4$ to an aqueous solution of $BaCl_2$ is an example of __ reaction

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40. State whether the following statements are true or false.

Rusting of iron is a fast reaction.

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41. State whether the following statements are true or false.

Milk is set into curd is a chemical change.

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42. State whether the following statements are true or false.

The reaction between salt and water is an example of exothermic reaction.

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43. State whether the following statements are true or false.

The speed of a chemical reaction depends on the catalyst used in the chemical reaction

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44. State whether the following statements are true or false.

The simple form of representation of a chemical reaction in words is known as word reaction.

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45. State whether the following statements are true or false.

Nascent oxygen is always denoted by showing the symbol of oxygen.

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46. State whether the following statements are true or false.

Antioxidants are used to prevent oxidation of food containing fats and oils.

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47. State whether the following statements are true or false.

When oils and fats are allowed to stand for a long time, they become rancid.

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48. State whether the following statements are true or false.

The chemical formula of rust is $\text{Fe}_3\text{O}_4 \cdot x\text{H}_2\text{O}$.

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49. State whether the following statements are true or false.

Glucose combines with oxygen in our body and provides energy. The reaction is an endothermic reaction.

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50. State whether the following statements are true or false.

Chemical reactions in which reactants gain oxygen are reduction reactions.

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51. State whether the following statements are true or false.

$$\text{CuSO}_{4(aq)} + \text{Zn}_{(s)} \rightarrow \text{ZnSO}_{4(aq)} + \text{Cu}_{(s)}$$
 is an example of decomposition reaction.

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52. State whether the following statements are true or false.

The chemical reactions in which heat is liberated are called endothermic reactions.

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53. State whether the following statements are true or false.

The product or insoluble solid in chemical reactions is indicated by an arrow \uparrow pointing upwards.

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54. State whether the following statements are true or false.

The rate of a reaction increases on increasing the temperature..

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55. State whether the following statements are true or false.

The digestion of food is a chemical decomposition process.

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56. State whether the following statements are true or false.

The reaction between sodium hydroxide and hydrochloric acid is a slow

reaction..

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57. State whether the following statements are true or false.

When calcium carbonate is heated ,it decomposes into calcium oxide and oxygen gas.

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58. State whether the following statements are true or false.

The rate of a chemical reaction changes in presence of catalyst.

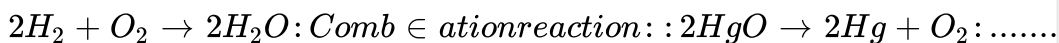
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59. State whether the following statements are true or false.

Chlorine is an oxidant.

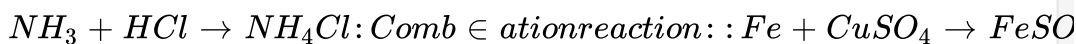
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60. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:



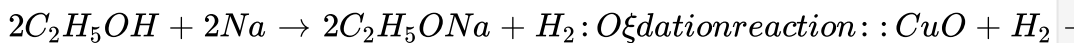
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61. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:



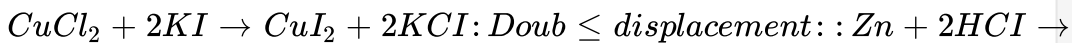
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62. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:



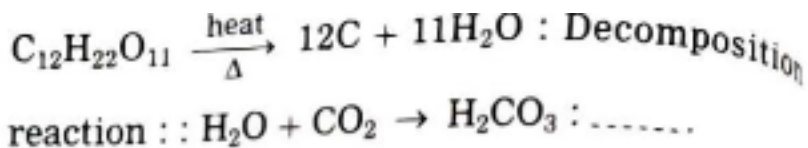
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63. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:



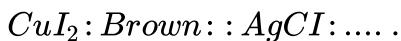
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64. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:



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65. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:



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66. Taking into consideration the relationship in the first pair, complete the second pair. Or complete the following:

Molecular formula of beryllium oxide : BeO : Molecular formula of beryllium chloride :

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67. Match the column in the following table

* (1) Reactants	Products	Type of chemical reaction
$\text{BaCl}_{2(aq)} + \text{ZnSO}_{4(aq)}$	$\text{H}_2\text{CO}_{3(aq)}$	Displacement
$2\text{AgCl}_{(s)}$	$\text{FeSO}_{4(aq)} + \text{Cu}_{(s)}$	Combination
$\text{CuSO}_{4(aq)} + \text{Fe}_{(s)}$	$\text{BaSO}_{4\downarrow} + \text{ZnCl}_{2(aq)}$	Decomposition
$\text{H}_2\text{O}_{(l)} + \text{CO}_{2(g)}$	$2\text{Ag}_{(s)} + \text{Cl}_{2(g)}$	Double displacement

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68. Match the column in the following table

(2) Reactants	Products	Type of chemical reaction
Fe + S	NaCl + H ₂ O	Oxidation
CuSO ₄ + Zn	2CuO	Neutralization
2Cu + O ₂	ZnSO ₄ + Cu	Displacement
HCl + NaOH	FeS	Combination

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69. Rewrite the second column so as to match the item from first column or match the following :

(1) Column I	Column II
(1) Reduction	(a) Type of chemical reaction
(2) Oxidation	(b) Combination with hydrogen
	(c) Losing hydrogen
	(d) Exchange of ions

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70. Rewrite the second column so as to match the item from first column

or match the following :

(2) Column I	Column II
(1) Oils and fats are allowed to stand in air for a long time	(a) Slow reaction
(2) NaOH dissolves in water	(b) Rancid
	(c) Exothermic reaction
	(d) Colourless solution



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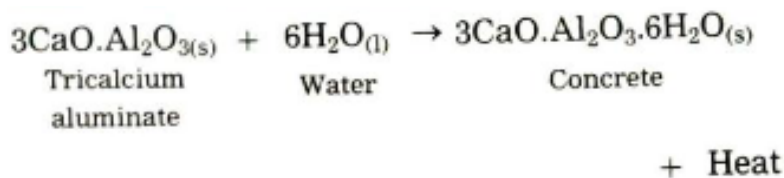
71. Rewrite the second column so as to match the item from first column

or match the following :

(3) Column I	Column II
(1) Combination reaction	(a) $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2 \uparrow$
(2) Double displacement reaction	(b) $\text{C}_{12}\text{H}_{22}\text{O}_{11(s)} \xrightarrow{\Delta} 12\text{C}_{(s)} + 11\text{H}_2\text{O}_{(g)}$
	(c) $2\text{Cu} + \text{O}_2 \rightarrow 2\text{CuO}$
	(d) $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} \downarrow + \text{NaNO}_3$

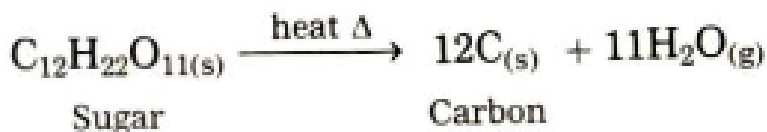
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72. Classify each of the following reactions as combination, decomposition, displacement or double displacement reactions :



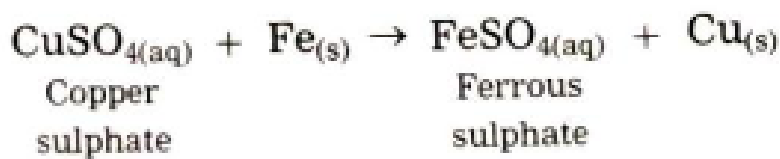
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73. Classify each of the following reactions as combination, decomposition, displacement or double displacement reactions :



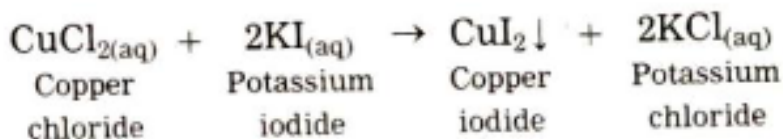
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74. Classify each of the following reactions as combination, decomposition, displacement or double displacement reactions :



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75. Classify each of the following reactions as combination, decomposition, displacement or double displacement reactions :



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78. Name the following

The product formed in the thermal decomposition of sugar.

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79. Name the following

The gas evolved when sodium metal reacts with ethanol.

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80. Name the following

The precipitate formed when barium sulphide reacts with Zinc sulphate.

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81. Name the following

The reducing agent used for the reduction of copper oxide.





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82. Name the following

The catalyst used to accelerate the rate of decomposition of hydrogen peroxide.



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83. Name the following

Which oxidising agent is used to oxidise ferrous sulphate.



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84. Name the following

The product formed in the oxidation of ethyl alcohol.



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85. Name the following

How are the blackened silver utensils and patinated (greenish) brass utensils cleaned ?

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86. Answer the following questions :

What do you understand by a physical change ?

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87. Answer the following questions :

Define physical change.

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88. Answer the following questions :

Explain giving two examples of physical change.

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89. Answer the following questions :

What do you understand by a chemical change ?

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90. Answer the following questions :

Define chemical change ?

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91. Answer the following questions :

Explain giving two examples of chemical change.

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92. What is a chemical reaction ?

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93. Explain the term reactant and product giving examples.

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94. Answer the following questions :

What is meant by a chemical equation ?

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95. Answer the following questions :

What is meant by a word equation?

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96. Write down the physical states of reactants and products in following reaction. $SO_2 + 2H_2S \rightarrow 3S + 2H_2O$

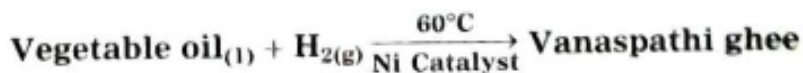
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97. Write down the physical states of reactants and products in following reaction. $2Ag_{(s)} + 2HCl \rightarrow 2AgCl + H_2$

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98. Answer the following questions :

Identify the reactants and products of the following equation:



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99. Answer the following questions :

What is the importance of a chemical equation ?

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100. Answer the following questions :

What are the conventions used in writing a chemical equation ?

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101. Answer the following questions :

Explain the term balanced equation with example.

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102. Answer the following questions :

Write the balanced equations for the following reactions:



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103. Answer the following questions :

Write the balanced equations for the following reactions:



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104. Answer the following questions :

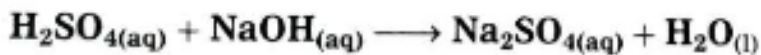
Write the balanced equations for the following reactions:



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105. Answer the following questions :

Write the balanced equations for the following reactions:



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106. Answer the following questions :

Write the balanced equations for the following reactions:



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107. Answer the following questions :

Write the balanced equations for the following reactions:

(vi) Calcium chloride + Sulphuric acid → Calcium sulphate + Hydrogen chloride.



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108. Answer the following questions :

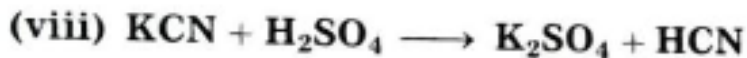
Write the balanced equations for the following reactions:



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109. Answer the following questions :

Write the balanced equations for the following reactions:



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110. Answer the following questions :

Write the balanced equations for the following reactions:



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111. What are the other uses of silver nitrate in everyday life?

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112. Answer the following questions :

What are the different types of chemical reaction ?

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113. Answer the following questions :

Define: combination reaction.

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114. Answer the following questions :

Define: combination reaction.

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115. Answer the following questions :

Give two examples of combination reaction.

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116. Answer the following questions :

Observe the following reaction carefully and answer the sub-questions:



What are the salt and acid in the above reaction ?

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117. Answer the following questions :

Observe the following reaction carefully and answer the sub-questions:



Which is the base in the above reaction? Is it weak or strong ?

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118. Answer the following questions :

Observe the following reaction carefully and answer the sub-questions:



Write a reaction showing dissociation of this base in water.

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119. Answer the following questions :

Explain the term combination reaction with examples.

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120. Define the following/write notes:

Decomposition Reaction.

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121. Answer the following questions :

What is meant by a thermal decomposition ?

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122. Answer the following questions :

Give an example of electrolytic decomposition .

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123. Answer the following questions :

Give two examples of thermal decomposition.



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124. Answer the following questions :

Give an example of electrolytic decomposition .

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125. Answer the following questions :

What are the different types of chemical reaction ?

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126. Answer the following questions :

Study the following reaction and answer the question asked.



Acidic water

Define this reaction.

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127. Answer the following questions :

Study the following reaction and answer the question asked.



Acidic water

Define this reaction.

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128. Answer the following questions :

What is meant by a displacement reaction ?

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129. Answer the following questions :

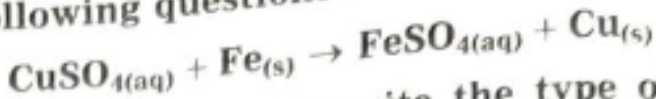
Give an example of displacement reaction.

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130. Answer the following questions :

Observe the reaction and answer the following questions.

following questions.



..... and write the type of chem

Identify and write the type of chemical reaction.



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131. Answer the following questions :

Complete the given chemical reaction:

$\text{CuSO}_4(\text{aq}) + \text{Fe}_\text{s} \rightarrow \text{.....} + \text{.....}$ Name the type of reaction.



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132. Answer the following questions :

Explain the term displacement reaction with examples.



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133. Answer the following questions :

What is meant by a double displacement reaction?

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134. Answer the following questions:

Give two examples of double displacement reaction ?

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135. Answer the following questions :

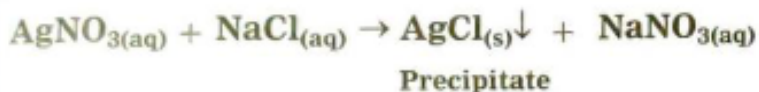
Write down what you understand from the following chemical reaction:



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136. Answer the following questions :

Study the following chemical reaction and answer the questions given below:

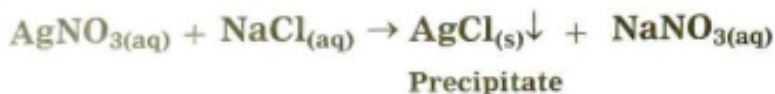


Identify and write the type of chemical reaction.

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137. Answer the following questions :

Study the following chemical reaction and answer the questions given below:

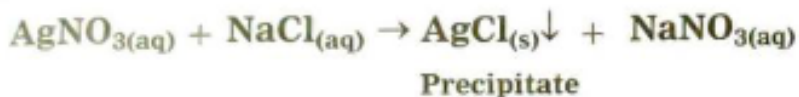


Write the definition of above type of chemical reaction.

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138. Answer the following questions :

Study the following chemical reaction and answer the questions given below:



Write the names of reactants and products of above reaction.

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139. Answer the following questions :

When sodium carbonate solution is mixed with barium sulphate solution, a precipitate is formed.

What is the colour of the precipitate formed ?

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140. Answer the following questions :

When sodium carbonate solution is mixed with barium sulphate solution,

a precipitate is formed.

Name the precipitate.

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141. Answer the following questions :

When sodium carbonate solution is mixed with barium sulphate solution, a precipitate is formed.

What is the type of chemical reaction ?

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142. Answer the following questions :

Explain the terms with examples:

Endothermic reaction.

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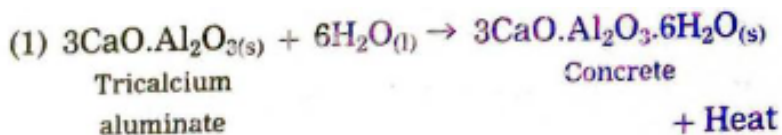
143. Define the following/write notes:

Exothermic Reaction

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144. Answer the following questions :

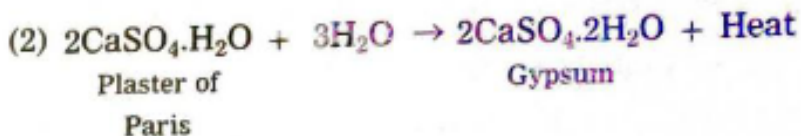
State whether the following reactions are exothermic or endothermic:



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145. Answer the following questions :

State whether the following reactions are exothermic or endothermic:



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146. Answer the following questions :

State whether the following reactions are exothermic or endothermic:



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147. Answer the following questions :

State whether the following reactions are exothermic or endothermic:



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148. Answer the following questions :

State whether the following reactions are exothermic or endothermic:

Transformation of ice into water.

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149. Answer the following questions :

State whether the following reactions are exothermic or endothermic:

Water turns into ice.

 [Watch Video Solution](#)

150. Answer the following questions :

State whether the following reactions are exothermic or endothermic:

Cooking of food

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151. Answer the following questions :

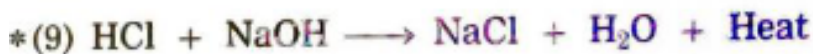
State whether the following reactions are exothermic or endothermic:

Burning candle.

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152. Answer the following questions :

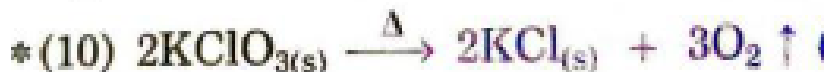
State whether the following reactions are exothermic or endothermic:



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153. Answer the following questions :

State whether the following reactions are exothermic or endothermic:



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154. Answer the following questions :

State whether the following reactions are exothermic or endothermic:



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155. Answer the following questions :

State whether the following reactions are exothermic or endothermic:



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156. Explain the similarity and difference in two events, namely adding NaOH to water and adding CaO to water

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157. Answer the following questions :

What do you mean by slow speed reaction?

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158. Answer the following questions :

Define Slow speed reaction..



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159. Answer the following questions :

What do you mean by fast speed reaction?



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160. Answer the following questions :

Define Fast speed reaction.



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161. Answer the following questions :

Give wo examples of slow speed reactions..





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162. Answer the following questions :

Give two examples of fast speed reactions..



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163. Answer the following questions :

Write a short note on slow speed and fast speed reactions.



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164. Answer the following questions :

State the factors which affect the speed (or rate) of a reaction.



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165. Answer the following questions :

How does the rate of reaction depend on the nature of the reactants?

Illustrate with suitable example.

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166. Answer the following questions :

How does the rate of a reaction depend on the size of the particles of reactants?

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167. Answer the following questions :

How does the rate of reaction depend upon the concentration of the reactants? Give suitable example.

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168. Answer the following questions :

How does the rate of a reaction depend upon the temperature of reactants? Give suitable example.

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169. Answer the following questions :

How does the rate of a reaction depend upon the catalyst? Give suitable example.

 [Watch Video Solution](#)

170. Answer the following questions :

State the importance of rate in a chemical reaction.

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171. Answer the following questions :

Define: Oxidation reaction.

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172. Answer the following questions :

Give examples of oxidation.

 [Watch Video Solution](#)

173. Answer the following questions :

What do you mean by oxidant? Explain with suitable example.

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174. Answer the following questions :

Name the various oxidants . How nascent oxygen is liberated from these

oxidants ?

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175. Answer the following questions :

Why is potassium permanganate used during cleaning water tanks ?

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176. Answer the following questions :

Which is the oxidant used for purification of drinking water ?

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177. Answer the following questions :

Acidified potassium permanganate (KMnO_4) is a chemical oxidant and explain, how acidified potassium permanganate oxidise ferrous sulphate (FeSO_4) . Accordingly write a new definition of oxidation and reduction.

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178. Answer the following questions :

Define: reduction reaction

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179. Answer the following questions :

Give two examples of reduction.

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180. Answer the following questions :

What do you mean by reductant? Explain with suitable example.

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181. Explain the types of reaction with reference to oxygen and hydrogen. Illustrate with example

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182. What is the reaction called when oxidation and reduction take place simultaneously? Explain with one example.

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183. Answer the following questions :

Explain redox reaction with suitable example.

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184. Answer the following questions :

Write note on redox reaction..





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185. Answer the following questions :

What are redox reactions? Identify the substances that are oxidised and the substances that are reduced in the following reactions.



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186. Answer the following questions :

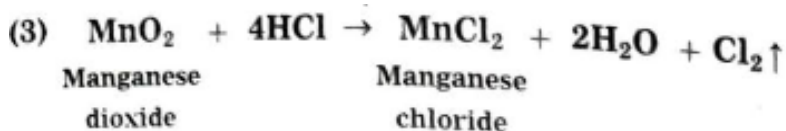
What are redox reactions? Identify the substances that are oxidised and the substances that are reduced in the following reactions.



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187. Answer the following questions :

What are redox reactions? Identify the substances that are oxidised and the substances that are reduced in the following reactions.



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188. Answer the following questions :

Observe the following reaction and answer the questions given below:

~~answer the questions given below~~
answer the questions given below



What type of reaction is it? Justify.

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189. Answer the following questions :

Observe the following reaction and answer the questions given below:



Give one more example.

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190. Answer the following questions :

Identify the following reactions the reactants that undergo oxidation and reduction.

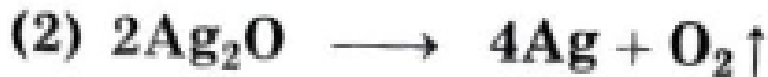


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191. Answer the following questions :

Identify the following reactions the reactants that undergo oxidation and

reduction.



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192. Answer the following questions :

Identify the following reactions the reactants that undergo oxidation and reduction.



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193. Answer the following questions :

Identify the following reactions the reactants that undergo oxidation and reduction.





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194. Answer the following questions :

Identify oxidant and reductant from the following reaction:

following reaction :



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195. Answer the following questions :

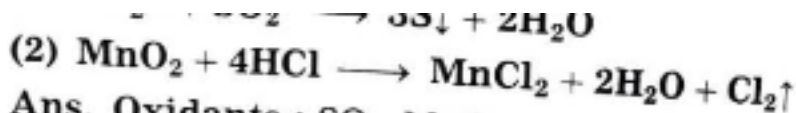
Some more examples of redox reaction are as follows. Identify the reductants and oxidants from them.



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196. Answer the following questions :

Some more examples of redox reaction are as follows. Identify the reductants and oxidants from them.



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197. Answer the following questions :

If oxidation means losing electrons. What is meant by reduction.

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198. Write the reaction of formation of Fe^{2+} by the reduction Fe^{3+} by making use of the symbol(e)?

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199. Define the following :

Corrosion:

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200. Write two conditions necessary for rusting of iron.

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201. How can corrosion of metals be prevented?

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202. What is corrosion? Do gold ornaments corrode? Justify.

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203. The chemical formula of the rust is



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204. Answer the following questions :

What are edible oils?



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205. Answer the following questions :

State the use of antioxidants in food containing fats and oils.



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206. Answer the following questions :

Is rancidity a phenomenon of oxidation or reduction?



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207. Answer the following questions :

What is the type of this reaction , in which vanaspathi ghee is formed from vegetable oil?

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208. Answer the following questions :

Define Rancidity.

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209. Give scientific reasons

It takes time for pieces of shahabad tiles to disappear in HCl , but its powder disappears rapidly.

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210. Give scientific reasons

Grills of doors and windows are always painted before they are used.

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211. Give scientific reasons:

Physical states of reactants and products are mentioned while writing a chemical equation.

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212. Give scientific reasons

When the gas formed on heating limestone is passed through freshly prepared lime water, the lime water turns milky.

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213. Give scientific reasons

While preparing dilute sulphuric acid from concentrated sulphuric acid in the laboratory, the concentrated sulphuric acid is added slowly to water with constant stirring.

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214. Give scientific reasons

It is recommended to use air tight container for storing oil for long time.

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215. Give scientific reasons:

Iron articles rust readily whereas steel which is also mainly made of iron does not undergo corrosion.

 [Watch Video Solution](#)

216. Give scientific reasons

Concentrated hydrochloric acid reacts rapidly with $CaCO_3$ than dilute hydrochloric acid.

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217. Give scientific reasons:

Zinc powder reacts faster than zinc granules when added to copper sulphate ($CuSO_4$) solution.

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218. Give scientific reasons:

When copper articles exposed to air for a long time, gets corroded.

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219. Give scientific reasons:

When silver vessels exposed to air turns blackish after sometime.

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220. Explain the following reactions giving their balanced chemical equations:

Calcium carbonate (Lime stone) is heated.

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221. Explain the following reactions giving their balanced chemical equations:

Copper reacts with dil. nitric acid.

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222. Explain the following reactions giving their balanced chemical equations:

Copper reacts with conc. nitric acid..

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223. Explain the following reactions giving their balanced chemical equations:

Ammonia gas reacts with hydrogen chloride gas.

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224. Explain the following reactions giving their balanced chemical equations:

Magnesium strip is burnt in air.

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225. Explain the following reactions giving their balanced chemical equations:

Calcium oxide is mixed with water.

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226. Explain the following reactions giving their balanced chemical equations:

Sugar is heated.

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227. Explain the following reactions giving their balanced chemical equations:

Electric current is passed through acidulated water.

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228. Explain what happens when following reactions take place and give the balanced chemical equations. Zinc dust is added to copper sulphate solution.

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229. The reaction of iron nail with copper sulphate solution is _____ reaction

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230. Explain the following reactions giving their balanced chemical equations:

Lead is added to copper sulphate solution.

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231. Explain the following reactions giving their balanced chemical equations:

Potassium chromate solution is added to barium sulphate solution.

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232. Explain the following reactions giving their balanced chemical equations:

Calcium chloride solution is added to sodium carbonate solution..

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233. Explain the following reactions giving their balanced chemical equations:

Sodium chloride solution is mixed with silver nitrate solution..

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234. Explain the following reactions giving their balanced chemical equations:

Dilute sulphuric acid is added to barium chloride solution.

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235. Explain the following reactions giving their balanced chemical equations:

Calcium carbonate (Lime stone) is treated with dil.hydrochloric acid..

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236. Explain the following reactions giving their balanced chemical equations:

Aluminium is treated with dil. hydrochloric acid.

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237. Explain the following reactions giving their balanced chemical equations:

Magnesium is treated with hydrochloric acid.

 [Watch Video Solution](#)

238. Explain the following reactions giving their balanced chemical equations:

Hydrogen peroxide is decomposed in the presence of manganese dioxide (MnO_2).

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239. Explain the following reactions giving their balanced chemical equations:

ethyl alcohol is treated with acidified potassium dichromate..

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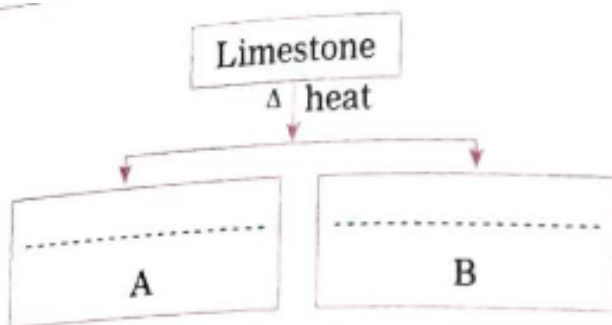
240. Explain the following reactions giving their balanced chemical equations:

Hydrogen gas is passed over black copper oxide..

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241. Identify A and B from the following table and complete the table .

Write the chemical equation.



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