CHEMISTRY

BOOKS - CHETANA CHEMISTRY (MARATHI ENGLISH)

Periodic Classifications of ELements

Exersice

1. Mendeleev arranged__elements known at that time.

A. 30

B. 56

C. 63

D. 92

Answer:

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2. The element eka-boron in Mendeleev's periodic table is known as____in the Modern periodic table

A. Scandium

B. Gallium

C. Germanium

D. polonium

Answer:

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3. In the Modern periodic table,the elements are arranged in accordance with their_____.

A. atomic mass

B. colour

C. atomic number

D. physical state.

Answer:

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4. The____contains the group 1 and 2 elements

A. f-block

B. p-block

C. d-block

D. s-block

Answer:

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5. Elements showing properties of both metals and non-metals are called___.

A. alloys

B. metalloids

C. noble metals

D. mixtures

Answer:

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6. In the family of alkali metals,the number of valence electrons is _____.

A. 2

B. 4

C. 6

D. 1

Answer:

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7. There are seven electrons in the outermost shell of the elements,such elements belong to the family of .

A. noble gases

B. alkali metals

C. halogens

D. alkaline earth metals

Answer:

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8. In the modern periodic table,the elements placed at the bottom of the periodic table are called as _____elements

A. noble

B. normal

C. transition

D. innertransition

Answer:

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9. An element is place in $2^n d$ period,so it has___shells

A. 3

B. 2

C. 4

D. 5

Answer:

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10. _____is the only element in duplet state.

A. Neon

B. Argon

C. Helium

D. krypton

Answer:

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11. First period contains _____elements.

A. 3

B. 4

C. 5

D. 2

Answer:

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12. The formula of chloride of a metal is MCl_2 the metal M belongs to _____group.

A. $1^s t$

B. $2^n d$

C. 13^th

D. 14^th

Answer:

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13. The number of electrons in the outermost shell of alkali metals is _____

A. 1

B. 2

C. 3

D. 7

Answer:

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14. Alkali earth metals have valency 2.this means that their position in the Modern periodic table is in____

A. Group 2

B. Group 16

C. Period 2

D. d-block

Answer:

15. Molecular formula of the chloride of an element X is XCl. This compound is a solid having high melting point. Which of the

following element be present in the same group as X.

A. Na

B. Mg

C. Al

D. Si

Answer:

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16. In which block of the modern periodic table are the non-metlas present?

A. s-block

B. p-block

C. d-block

D. f-block

Answer:

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17. _____resembles alkali metals as well as halogens

A. Lithium

B. sodium

C. hydrogen

D. silicon

Answer:

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18. First period consist of _elements .

A. 1

B. 8

C. 2

D. 4

Answer:

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19. Electronic configuration of Mg is ____

A. (2,8,4)

B. (2,8,1)

C. (2,8,3)

D. (2,8,2)

Answer:

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20. ____is in liquid at room temperature .

A. Fluorine

B. chlorine

C. bromine

D. iodine

Answer:

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21. Elements showing properties of boht metals and non-metals are called___

A. alloys

B. metalloids

C. noble metals

D. mixtures

Answer:

Watch Video Solution

22. Electronic configuration of Mg is ____

A. (2,8,4)

B. (2,8,1)

C. (2,8,3)

D. (2,8,2)

Answer:

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23. Three elements having a single electron in their outermost shell .

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24. Find the odd one word out: Chlorine, Bromine, Arsenic, Fluorine

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25. State whether the following statement is True or False :

The d-block elements are called transition elements

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26. Atomic radius goes on decreasing while going from left to right in a period

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27. Sodium is more metallic than Aluminium .

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28. ${}_{19}K,{}_{3}Li,{}_{11}Na,{}_{4}Be$ .Which of these atoms has smallest atomic radius?

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29. Write down the electronic configuration of the following elements from the given atomic

numbers. Answer the following questions with explanation. `"_6C,_3Li,_9F,_7N,_8O. Which of the above elements has the highest non-metallic character?

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30. What are the demerits of Mendeleev's periodic table?

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31. State the merits of Mendeleev's periodic table.

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32. How could the modern periodic Table remove various anomallies of Mendleev's table?

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33. The following table shows the position of six elements A,B,C,D,E and F in the periodic table

Periods \ Groups	1	2	3 to 12	13	14	15	16	17	18
2	A					B			C
3		D			E				F

Using the above table , answer the following questions:

Which element is a metal with valency 2?

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34. The following table shows the position of six elements A,B,C,D,E and F in the periodic table

Periods \ Groups	1	2	3 to 12	13	14	15	16	17	18
2	A					B			C
3		D			E				F

Using the above table , answer the following questions:

Which element is a non-metal with valency 3?

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35. The following table shows the position of six elements A,B,C,D,E and F in the periodic table

Periods \ Groups	1	2	3 to 12	13	14	15	16	17	18
2	A					B			C
3		D			E				F

Using the above table , answer the following questions:

Out of D and E, which one has a bigger atomic radius and why?

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36. The following table shows the position of six elements A,B,C,D,E and F in the periodic table

Periods \ Groups	1	2	3 to 12	13	14	15	16	17	18
2	A					B			C
3		D			E				F

Using the above table , answer the following questions:

Write a common name for the family of elements C and F?

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37. The following table shows the position of six elements A,B,C,D,E and F in the periodic table

Groups / Periods	1	2	3 to 12	13	14	15	16	17	18
2	A					B			C
3		D			E				F

Using the above table , answer the following questions:

What is the formula of compound formed between A and B?

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Example

1. Find the odd man out.

Lithium,beryllium,boron,Chlorine

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2. Find the odd man out.

Sodium,Lithium,Copper,Beryllium

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3. Find the odd man out.

Dalton,Dobereiner,Moseley,Newlands

4. Find the odd man out.

boron,Silicon,potassium,antimony

5. Find the odd man out.

Aluminium,Argon,Xenon,Sodium

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6. Find the odd man out.

boron,Silicon,potassium,antimony

Watch Video Solution

7. Find the odd man out.

Lithium,beryllium,boron,Chlorine

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8. Complete the anology

Dobereiner:Traid::Newlands:___

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9. Complete the anology

Mendeleev's periodic Table: Atomic mass::

Modern periodic table:_____

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10. Complete the anology

Hydrogen:First period ::Lithium:__

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11. Complete the anology

Fluorine:2,7:: chlorine:____

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12. Complete the anology

Group 1:Alkali metals :: _____: alkaline earth metals

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13. Complete the anology

Transition elements :d-block:: inner transition elements:_____

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14. Complete the anology

Tellurium:--------:: Radium:Metal

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15. Complete the anology

Transition elements:------Inner transition elements:Three incomplete outermost shells

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16. Complete the anology

Lanthanides:Ce to Lu::Actinides:____

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17. Complete the anology

Ca:Alkaline earth metal:: Cs:_____

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18. Complete the anology

Fe:Electropositive :: CI:____

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19. Complete the anology

(Li,Na,K):____::(F,CI,Br): Group 17

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20. Complete the anology

Valency of Na(2,8,1):1(One)::Valency of P(2,8,5):_____

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21. Match the columns:

Column A	Column B	Column C
(i) Triad	(a) Lightest and negatively charged particle in all the atoms.	(1) Mendeleev
(ii) Octave	(b) Concentrated mass and positive charge	(2) Thomson
(iii) Atomic number	(c) Average of the atomic mass of the first and the third elements	(3) Newlands
(iv) Period	(d) Properties of the eighth element similar to the first	(4) Rutherford
(v) Nucleus	(e) Positive charge on the nucleus	(5) Dobereiner
(vi) Electron	(f) Sequential change in molecular formulae	(6) Moseley

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22. Match the columns:

	Column A		Column B
(1)	Sodium	(a)	Non-metal
(2)	Sulphur	(b)	Lanthanide
(3)	Manganese	(c)	Metal
(4)	Cerium	(d)	Transition metal

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23. Match the columns:

	Column A		Column B
(1)	Alkali metals	(a)	Valency 4
(2)	Alkaline earth metals	(b)	Valency 0
(3)	Argon	(c)	Divalent
(4)	Carbon	(d)	Monovalent

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24. Match the columns:

	Column A		Column B
(1)	Atomic size in a period from left to right	(a)	Atomic mass
(2)	Atomic size in a group from top to bottom	(b)	Increases
(3)	Modern Periodic Law	(c)	Atomic number
(4)	Mendeleev's Periodic Law	(d)	Decreases

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25. Match the columns:

Column A		Column B	
(1)	Alkali metals	(a)	Group 2
(2)	Alkaline earth metals	(b)	Group 1
(3)	Halogens	(c)	Group 18
(4)	Noble gases	(d)	Group 17

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26. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Lithium,Potassium and Sodium are elements forming Dobereiner's traid .

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27. State whether the following statement is True' or False'. If false, write the correct statement for the same.

According to Mendeleev's periodic law,the properties of elements are periodic function of their atomic numbers

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28. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Periods are the horizontal rows of elements

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29. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Group 17 elements are known as Noble gases

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30. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Silicon is a metalloid

Watch Video Solution

31. State whether the following statement is True' or False'. If false, write the correct

statement for the same.

As we move from left to right in a period of Moderns periodic table,atomic size of the elements gradually increases .

 Watch Video Solution

32. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Group 1 elements in Modern periodic table are refered as"alkali metals"

Watch Video Solution

33. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Argon is not an inert gas.

Watch Video Solution

34. State whether the following statement is True' or False'. If false, write the correct statement for the same.

The d-block elemnts are called transition elements .

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35. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Upto calcium,the law of octaves was found to be applicable

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36. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Metals are electronegative and non-metals are electropositive .

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37. Eka-aluminium was discovered are and named as gallium(Ga)

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38. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Atomic number is the number of protons or electrons present in the nucleus .

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39. State whether the following statement is True' or False'. If false, write the correct

statement for the same.

Eka brorn is known as germanium.

40. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Telluium,Polonium are metalloids

41. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Isotopes have same atomic masses

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42. State whether the following statement is True' or False'. If false, write the correct statement for the same.

Isotopes have similar chemical properties .

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43. State whether the following statement is True' or False'. If false, write the correct statement for the same.

La stands for lutetium

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44. State whether the following statement is True' or False'. If false, write the correct

statement for the same.

F-block elements are metalloids

Watch Video Solution

45. State whether the following statement is True' or False'. If false, write the correct statement for the same.

In an atom of an element,extra-nuclear electrons take part in the chemical reaction

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46. The atom having the smallest size.


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47. The atom having the smallest atomic mass

.


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48. The most electronegative atom


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49. The noble gas with the smallest atomic radius

Watch Video Solution

50. The most reactive nonmetal

Watch Video Solution

51. The group with valency zero

Watch Video Solution

52. The family of non-metals having valency one .

Watch Video Solution

53. The family of metals having valency one .

Watch Video Solution

54. The family of metals having valency two

Watch Video Solution

55. The metalloids in the second and third periods

Watch Video Solution

56. Non-metals in the third period .

Watch Video Solution

57. Two elements having valency 4.

Watch Video Solution

58. Three elements having a single electron in their outermost shell .

Watch Video Solution

59. Three elements with filled outermost shell

 Watch Video Solution

60. Three elements having 7 electrons in their outermost shell

 Watch Video Solution

61. Three elements which are metalloids

 Watch Video Solution

62. Three alkaline earth metals with electronic configuration

Watch Video Solution

63. Two pairs of Dobereiner's traid.

Watch Video Solution

64. The scientist who classified elements on the basis of atomic mass.

Watch Video Solution

65. The scientist who classified elements on the basis of atomic number

Watch Video Solution

66. The 3 elements which were predicted by Mendeleev for which the left blank spaces in the periodic table.

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67. Series of 14 elements placed below the periodic table having atomic no.from 58 to 71(Ce to lu)

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68. X,Y and Z are the elements of a Dobereiner's Traid. If the atomic mass X is 7 and that of Z is 39,what should be the atomic mass of Y?

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69. Two elements X and Y have atomic number 12 and 16 respectively.Write the electronic configuration for these elements.To which period of the Modern periodic table do these two elements belong?

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70. Identify dobereiner's traids from the following group of elements having similar

chemical properties : `Mg(24.3),Ca(40.1),Sr(87.6)

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71. Identify dobereiner's traids from the following group of elements having similar chemical properties : `S(32.1),Se(79.0),T(127.6)

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72. Identify dobereiner's traids from the following group of elements having similar

chemical properties :

$Be(9.0), Mg(24.3), Ca(40.1)$

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73. State the laws/ Define

Newlands's Law of octaves:

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74. State the laws/ Define

Mendeleev's periodic law,

Watch Video Solution

75. State the laws/ Define

Moderns periodic Law,

Watch Video Solution

76. State the laws/ Define

Dobereiner's Traid

Watch Video Solution

77. State the laws/ Define

Periods are the horizontal rows of elements

 Watch Video Solution

78. State the laws/ Define

Groups

 Watch Video Solution

79. State the laws/ Define

Normal elements

Watch Video Solution

80. State the laws/ Define

Lanthanide Series

Watch Video Solution

81. State the laws/ Define

Actinde Series

82. State the laws/ Define

Transition elements :d-block:: inner transition elements:____

83. State the laws/ Define

Inner Transition elements

Watch Video Solution

84. State the laws/ Define

Atomic radius of Lithium is _____pm.

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85. State the laws/ Define

Atomic radius

Watch Video Solution

86. State the laws/ Define

valency

Watch Video Solution

87. State the laws/ Define

Inert elements

Watch Video Solution

88. State the laws/ Define

Isotopes have same atomic masses

Watch Video Solution

89. State the laws/ Define

Metalloids

Watch Video Solution

90. What is the meant by periodicity ?

Watch Video Solution

91. How will the tendency to gain electrons change as we go from left to right across a

period ?why?

Watch Video Solution

92. Why was Dobereiner's classification of elements not useful?

Watch Video Solution

93. In the periodic table where are the metalloids placed?

Watch Video Solution

94. Elements of which group are called as alkali metals?

Watch Video Solution

95. Which is the incomplete period in the Moders periodic table?

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96. which law was modified into modern periodic law?

Watch Video Solution

97. what are periodic properties?

Watch Video Solution

98. What are normal elements?

Watch Video Solution

99. On what basis is potassium (Z=19) placed in 4^th period and group 1?

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100. How will you differentiate between metals and non-metals by the number of valence electrons?

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101. What are types of matter?

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102. What are the types of elements ?

Watch Video Solution

103. What are the smallest paritcles of matter called?

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104. What is the difference between the molecules of elements and compounds?

Watch Video Solution

105. What are thevalues of 'n' for the shells K,L and M?

Watch Video Solution

106. What is the maximum number of electrons that can be accommodated in a shell?write the formula

Watch Video Solution

107. Write a short note on:

Mendeleev's periodic law

Watch Video Solution

108. Deduce the maximum electron capacity of the shells K, L and M?

Watch Video Solution

109. Write a short note on:

Structure of the Modern periodic table

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110. Write a short note on:

Position of isotpes in the Mendeleev's and the

Modern periodic table

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111. Write a short note on:

Halogens or group 17 elements

Watch Video Solution

112. Write a short note on:

Transition elements.

113. Write a short note on:

Inner Transition elements

114. Write a short note on:

Metallic and Non-metallic properties

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115. Distinguish between:

Mendeleev's periodic table and Modern periodic table

Watch Video Solution

116. Distinguish between:

Transition elements and Inner transition elements

Watch Video Solution

117. Distinguish between:

Insert gases and normal Elements

Watch Video Solution

118. Distinguish between:

Groups and periods of Modern periodic Table

Watch Video Solution

119. Distinguish between:

Metallic character and non-metallic character

Watch Video Solution

120. Distinguish between:

s-block elements and p-block elements

Watch Video Solution

121. Distinguish between:

Alkali metals and Alkaline earth metals

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122. Atomic radius goes on decreasing while going from left to right in a period

Watch Video Solution

123. Give scientific reasons :

Metallic character goes on decreassing while going from left to right in a period

Watch Video Solution

124. Give scientific reasons :

Atomic radius goes on increasing down a group

Watch Video Solution

125. Give scientific reasons :

Elements belonging to the same group have the same valency

Watch Video Solution

126. Give scientific reasons :

The third period contains only eight elements even though the electorn capacity of the third shell is 18.

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127. Give scientific reasons :

Atomic number is a more fundamental property of an element than its atomic mass.

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128. Give scientific reasons :

Alkali metals are placed in Group 1.

Watch Video Solution

129. Give scientific reasons :

Inert gases exist in the form of free atoms.

Watch Video Solution

130. Give scientific reasons :

Inert gases (zero group elements) are called noble gases.

Watch Video Solution

131. Give scientific reasons :

Fluroine is the most reactive in Group 17.

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132. Write down the electronic configuration of the following elements from the given atomic numbers. Answer the following question with explanation.$_3Li$, _14Si, _2He, _11Na, _15P`. Which of these elements belong to period 3?

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133. Write down the electronic configuration of the following elements from the given

atomic numbers. Answer the following questions with explanation. ${}_{1}H,{}_{7}N,{}_{20}Ca,{}_{16}S,{}_{4}Be,{}_{18}Ar$. Which of these elements belong to the second group ?

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134. Write down the electronic configuration of the following elements from the given atomic numbers. Answer the following questions with explanation. ${}_{7}N,{}_{6}C,{}_{8}O,{}_{5}B,$

Al.which is the the most electronegative element among these?

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135. Write down the electronic configuration of the following elements from the given atomic numbers. Answer the following questions with explanation. `"_4Be,_6C,_8O,_5B,_13Al.which is the most electropostive element among these ?

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136. Write down the electronic configuration of the following elements from the given atomic numbers. Answer the following questions with explanation. `"_11Na,_15P,_17Cl,_14Si,_12Mg.Which of these has the largest atom?

Watch Video Solution

137. Write down the electronic configuration of the following elements from the given atomic

numbers. Answer the following questions with explanation. `"_19K,_3Li,_11Na,_4Be. Which of these atoms has smallest atomic radius?

Watch Video Solution

138. Write down the electronic configuration of the following elements from the given atomic numbers. Answer the following questions with explanation. `"_13Al,_14Si,_11Na,_12Mg,_16S. Which of the

above elements has the highest metallic character?

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139. Write down the electronic configuration of the following elements from the given atomic numbers. Answer the following questions with explanation. `"_6C,_3Li,_9F,_7N,_8O. Which of the above elements has the highest non-metallic character?

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140. Which element is a metal with valency2?

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141. Which element is non-metal with valency 3?

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142. Write a common name for the family of elements C and F?

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143. Chlorine has two isotopes, viz. Cl-35 and Cl-37 Their atomic masses are 35 and 37. Their chemical properties are same. Where should these be placed in Mendeleev's periodic table? In different places or in the same place?

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144. A metal M forms an oxide having the formula M_2O_3. It belongs to 3rd period in the Modern Periodic table. Write the atomic number and valency of the metal.

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145. What is the relationship between the electronic configuration of an element and its valency?

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146. The atomic number of beryllium is 4 while that of oxygen is 8. Write down the electronic configuration of the two and deduce their valency from the same.

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147. Which of the above elements have the biggest and the smallest atom?

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148. What is the periodic trend observed in the variation of atomic radii down a group?

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149. Classify the elements of the third period into metals and non-metals.

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150. On which side of the period did you find the Non-metals?

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151. What is the cause of non-metallic character of element?

Watch Video Solution

152. What is the expected trend in the variation of non-metallic character of element form left to right in a period?

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153. What would be the expected trend in the variation of non-metallic character of elements down a group?

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154. Go through the Modern periodic table (fig2.1) and write the names one below the other of the elements of group 1.

Watch Video Solution

155. Go through the Modern periodic table (fig2.1) and write the names one below the other of the elements of group 1. What similarity do you find in their electronic configuration?

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156. Go through the Modern periodic table (fig2.1) and write the names one below the other of the elements of group 1. How many valence electrons are there in each of these elements ?

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157. On going through the Modern periodic table (fig2.1) it is seen that the elements

Li,Be,B,C,N,O,F and Ne belong to the period -2. Write down the electronic configuration.

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158. On going through the Modern periodic table (fig. 2.1) it is seen that the elements Li,Be,B,C,N,O,F and Ne belong to the period - 2. Write down the electronic configuration. Is the number of valence electrons same for all these elements ?

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159. On going through the Modern periodic table (fig. 2.1) it is seen that the elements Li,Be,B,C,N,O,F and Ne belong to the period - 2. Write down the electronic configuration. Is the number of shells the same in these?

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160. How is the problem regarding the position of cobalt("_59CO) and nickel ("_NI)` in

Mendeleev's periodic table resolved in Modern periodic table ?

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161. How did the position of $\frac{35}{17}Cl$ and $\frac{37}{17}Cl$ get fixed in the Modern periodic table?

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162. Can there be an element with atomic mas 53 or 54 in between the two elements

,Chromium $\frac{52}{(24)Cr}$ and Mangense $\frac{55}{25}$Mn?

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163. How are isotopes of different elements placed in the Modern peridic Table?

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164. What should be the position of Hydrogen in the Modern periodic Table?Why?

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165. How do you calculate valency of an element from its electronic configuration?

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166. What is the valency of elements with atomic number `8,14,17 and 20 ?

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167. What was the basis of Mendeleev's classification ?

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168. What type of relationship of elements was examined by mendeleev?

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169. How many elements were known when Mendeleev started his work?

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170. What is meant by periodicity according to Mendeleev?

Watch Video Solution

171. What is Mendeleev's Periodic law?

Watch Video Solution

172. State the law on which modern periodic table is based

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173. How many groups are there in the modern periodic table?

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174. What is the number of valence electrons in an element of group 1 and group 18 respectively?

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175. What is the trend in the variation of valency while going down a group ?

Watch Video Solution

176. Which pair of elements do you think will have similar properties? : Sodium and potassium

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177. State the merits of Mendeleev's periodic table.

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178. What are the demerits of Mendeleev's periodic table?

Watch Video Solution

179. Merits of Modern periodic table over Mendeleev's periodic table

Watch Video Solution

180. What is the relationship between the electronic configuration of an element and its valency?

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181. The atomic number of beryllium is 4 while that of oxygen is 8.Write down the electronic configuration of the two and deuce their valency from the same

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182. Write the electronic configuration and valency of first 20 elements.

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183. What is the periodic trend in the variation of valency while going from left to right within a period.Explain the answer with reference to period 2 and period 3.

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184. What is the periodic trend in the variation of valency while going down a goup?Explain the

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185. Comparative study of all the four-blocks of Modern periodic table

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186. How could the modern periodic Table remove various anomallies of Mendleev's table?

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187. Answer the equation based on the given data

Elements	K	Na	Rb	Cs	Li
Atomic radius (pm)	231	186	244	262	152

To which group do the elements belong?What is the family called ?

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188. Answer the equation based on the given data

Elements	K	Na	Rb	Cs	Li
Atomic radius (pm)	231	186	244	262	152

Arrange the above elements in an increasing order of atomic radii.Does this arrangement match with the pattern of the group in the above answer?

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189. Answer the equation based on the given data

Elements	K	Na	Rb	Cs	Li
Atomic radius (pm)	231	186	244	262	152

Which of the above elements have the biggest and the smallest atom?

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190. Answer the equation based on the given data

Elements	K	Na	Rb	Cs	Li
Atomic radius (pm)	231	186	244	262	152

What is the periodic trend observed in the variation of atomic radii down a group?

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191. In the following table, seven elements P,Q,R,S,T,U and V (here letters are not the usual symbol of the elements)of the modern periodic table with their atomic numbers are given

3	4	5	6	7	8	9	10
P					T		V
11	12	13	14	15	16	17	18
Q	R		S			U	

Which of these is an inert gas? Name it.

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192. In the following table, seven elements P,Q,R,S,T,U and V (here letters are not the usual symbol of the elements)of the modern periodic table with their atomic numbers are given

3	4	5	6	7	8	9	10
P					T		V
11	12	13	14	15	16	17	18
Q	R		S			U	

Which of these is a halogen? Name it.

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193. In the following table, seven elements P,Q,R,S,T,U and V (here letters are not the usual symbol of the elements)of the modern periodic table with their atomic numbers are given

3	4	5	6	7	8	9	10
P					T		V
11	12	13	14	15	16	17	18
Q	R		S			U	

Which of the these are metals? Name them

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194. In the following table, seven elements P,Q,R,S,T,U and V (here letters are not the usual symbol of the elements)of the modern periodic table with their atomic numbers are given

3	4	5	6	7	8	9	10
P					T		V
11	12	13	14	15	16	17	18
Q	R		S			U	

If Q combines with U, what would be the formula of the compound formed? If Q and U are replaced by their respective metals what will be the formula of the compound formed.

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195. In the following table, seven elements P,Q,R,S,T,U and V (here letters are not the usual symbol of the elements)of the modern

periodic table with their atomic numbers are given

3	4	5	6	7	8	9	10
P					T		V
11	12	13	14	15	16	17	18
Q	R		S			U	

Write the electronic configuration of R and T and the type of bond formed by their combination.

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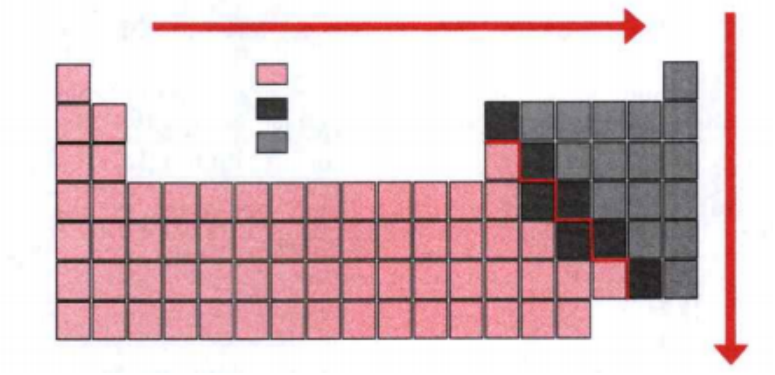
196. Write the answers of the question with refernce to the structure of the periodic table.



which points are considered for the arrangement of the Modern period table?

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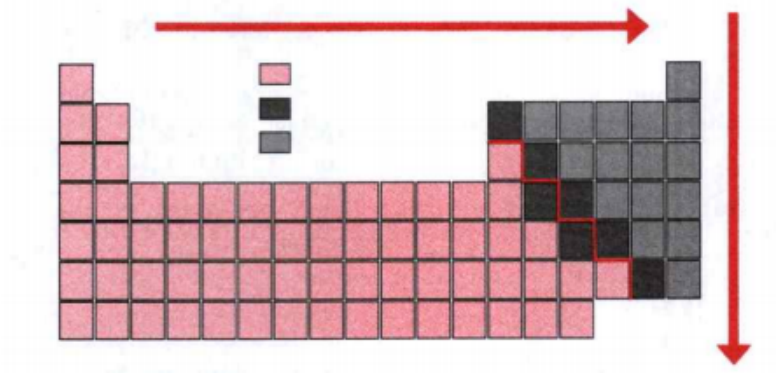
197. Write the answers of the question with reference to the structure of the periodic table.



How are blocks indicated?

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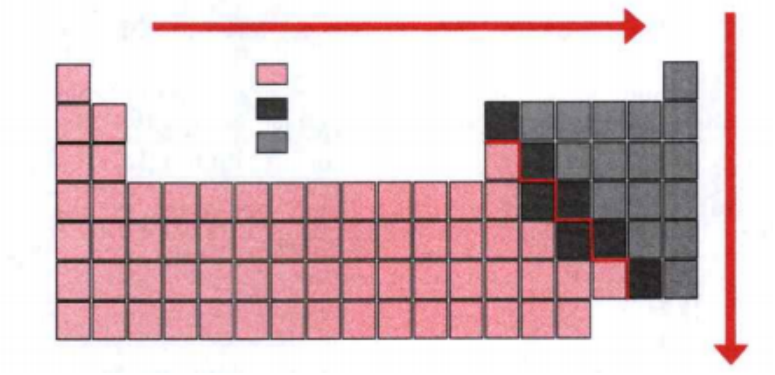
198. Write the answers of the question with refernce to the structure of the periodic table.



Which elements are present near the zig zag line?

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199. Draw the electronic configuration of the period 2 element of first group in the periodic table

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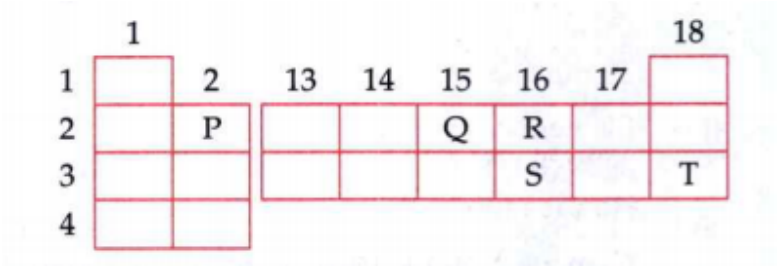
200. A part of periodic table is shown in the following figure



Write the symbol of the element Q'.

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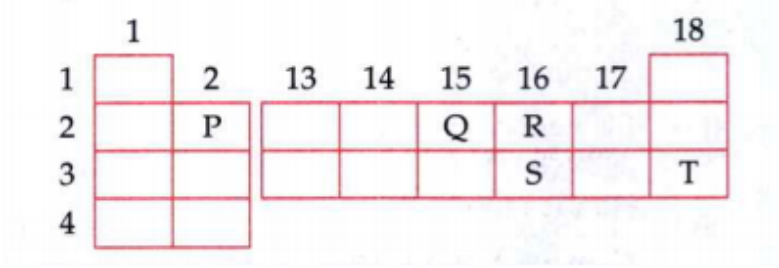
201. A part of periodic table is shown in the following figure



Will elements 'R' and 'S' have same number of valence electrons ?

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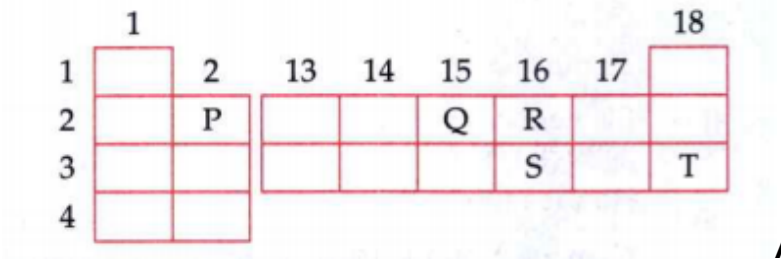
202. A part of periodic table is shown in the following figure



Arrange

elements 'P','Q' and 'R' in increasing order of their metallic character.

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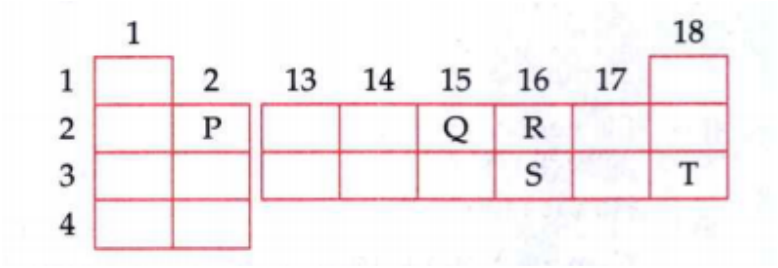
203. A part of periodic table is shown in the following figure



What is the number of electrons in L shell of element T?

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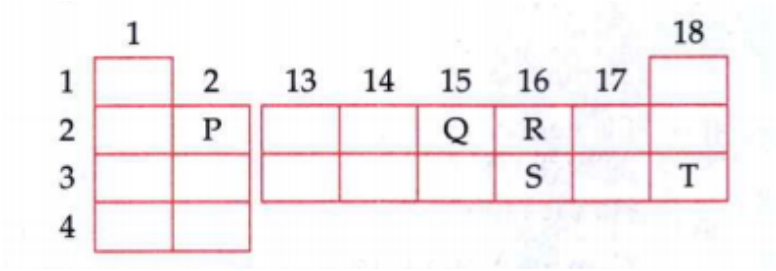
204. A part of periodic table is shown in the following figure



Name any two elements that will have properties similar to that of element 'P'

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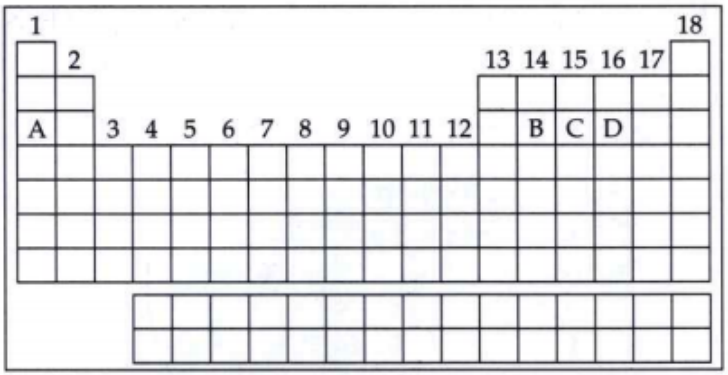
205. Study the below given periodic table in which four elements are indicated by alphabets A,B,C and D



Which element is a metalloid?Name this elemnt and also mention the metalloids in modern periodic table

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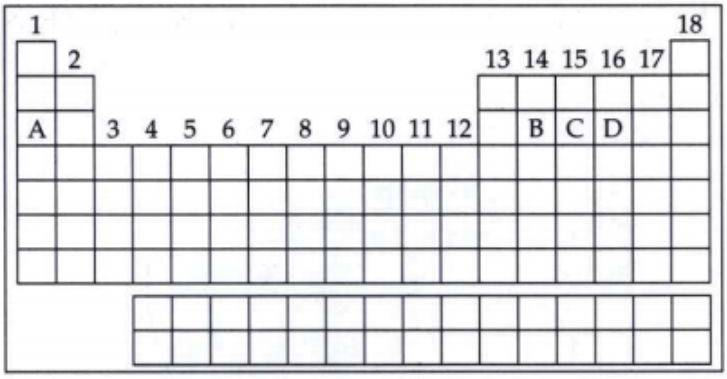
206. Study the below given periodic table in which four elements are indicated by alphabets A,B,C and D



Among 'A' and 'C' which element has larger atomic radius?why?

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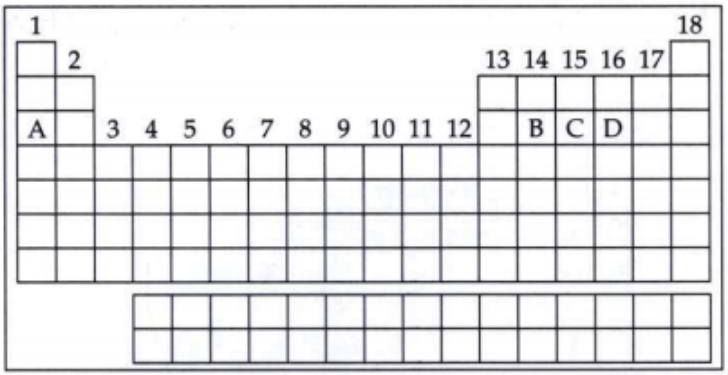
207. Study the below given periodic table in which four elements are indicated by alphabets A,B,C and D



Identify element 'D' and wrtie its electronic configuration Also write the electronic configuration of the elements above and below 'D' in the same group

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208. A scientist studying reactions of metals and non-metals. He knew group 1 and 2 elements are metals while group 17 elements are non-metals. So, he chooses different elements from group 1 and 2 and group17 what is the valency of magnesium?

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209. A scientist studying reactions of metals and non-metals. He knew group 1 and 2

elements are metals while group 17 elements are non-metals. So, he chooses different elements from group 1 and 2 and group17

Name the element in group 17 which forms a diatomic molecule and exists in solid state at room temperature .

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210. A scientist studying reactions of metals and non-metals. He knew group 1 and 2 elements are metals while group 17 elements

are non-metals. So, he chooses different elements from group 1 and 2 and group17

Name the element in group 17 which belongs to the same period as sodium.

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211. A scientist studying reactions of metals and non-metals. He knew group 1 and 2 elements are metals while group 17 elements are non-metals. So, he chooses different elements from group 1 and 2 and group17

Write the formula of compound formed in the reaction between lithium and bromine.

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212. A scientist studying reactions of metals and non-metals. He knew group 1 and 2 elements are metals while group 17 elements are non-metals. So, he chooses different elements from group 1 and 2 and group17 write the formula of compound formed in the reaction between calcium and fluorine.

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