



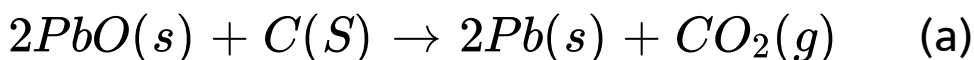
CHEMISTRY

BOOKS - PSEB

Chemical Reactions and Equations

Exercise

1. Which of the statements about reactions below are incorrect?



Lead is getting reduced. (b) Carbon dioxide is getting oxidised. (c) Carbon is getting oxidised. (d) Lead oxide is getting reduced.

A. a)and b)

B. a) and c)

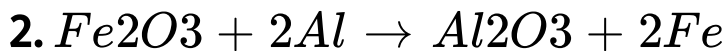
C. a), b)and c)

D. all

Answer:



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The above reaction is an example

- A. combination reaction
- B. double displacement reaction
- C. decomposition reaction
- D. displacement reaction

Answer:



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3. What happens when dilute hydrochloric acid is added to iron filings? Tick the correct answer

A. Hydrogen gas and iron chloride are produced

B. chlorine gas and iron hydroxide are produced

C. No reaction takes place

D. Iron salt and water are produced

Answer:



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4. What is a balanced chemical equation? Why should chemical equations be balanced ?



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5. Translate the following statements into chemical equations and then balance them:



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6. Translate the following statements into chemical equations and then balance them:

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7. Translate the following statements into chemical equations and then balance them:

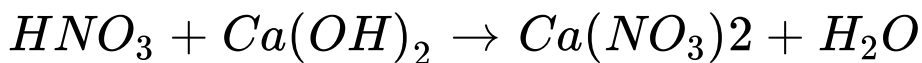
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8. Translate the following statements into chemical equations and then balance them:



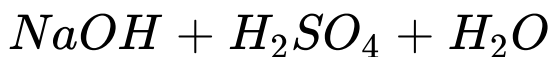
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9. Balance the following chemical equations:



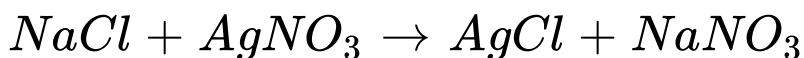
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10. Balance the following chemical equations:



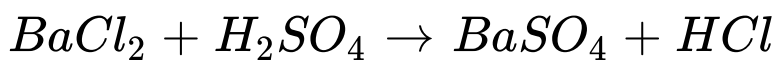
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11. Balance the following chemical equations:



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12. Balance the following chemical equations:



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13. Write balanced chemical equations for the following reaction- a) calcium hydroxide + carbon dioxide \rightarrow calcium carbonate + water



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14. Write balanced chemical equations for the following reaction- zinc+ silver nitrate \rightarrow zinc nitrate+silver



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15. Write balanced chemical equations for the following reaction- aluminium+ copper chloride \rightarrow aluminium chloride + copper



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16. Write balanced chemical equations for the following reaction- barium chloride+ potassium sulphate \rightarrow barium sulphate +potassium sulphate



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17. Write the balanced chemical equation for the following reaction and identify the type of reaction

Potassium bromide (aq) +Barium iodide (aq)

→ Potassium iodide (aq)+Barium bromide

(s)



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18. Write the balanced chemical equation for the following reaction and identify the type of reaction . Zinc carbonate (s) → Zinc oxide (s)+Carbon dioxide



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19. Write the balanced chemical equation for the following reaction and identify the type of reaction.

Hydrogen (g)+Chlorine(g) →
Hydrogen chloride(g)



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20. Write the balanced chemical equation for the following reaction and identify the type of reaction.

Magnesium (s)+Hydrochloric

acid(aq) → Magnesium chloride
(aq)+Hydrogen(g)



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21. What is meant by exothermic and endothermic reaction? Give examples



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22. Why is respiration considered as an exothermic reaction?explain



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23. Why are decomposition reactions called opposite of combination reactions? write equations for these reactions



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24. Write one equation each for decomposition reactions where energy is supplied in the form of heat, light or electricity





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25. What is the difference between the displacement and double displacement reaction write equation for these reactions



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26. In the refining of silver, the recovery of silver from silver nitrate solution involved displacement by copper metal. Write down the reaction involved



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27. What do you mean by precipitation reaction? explain by giving examples



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28. Explain the following in terms of gain or loss of oxygen with two examples each: oxidation and reduction



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29. Explain the following in terms of gain or loss of oxygen with two examples each: oxidation and reduction



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30. A shiny brown coloured element x on heating in air becomes black in colour name the element x and the blacked coloured compound formed



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31. Why do we apply paint on iron articles?



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32. Oil and fat containing food items are flushed with nitrogen why?



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33. Explain the following term with one example : Corrosion



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34. Explain the following term with one example : Rancidity.



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