



# CHEMISTRY

## BOOKS - PSEB

# PERIODIC CLASSIFICATION OF ELEMENTS

### Exercise

1. Which of the following statements is not a correct statement about the trends when

going from left to right across the periods of periodic table.

A. The elements become less metallic in nature.

B. The number of valence electron increases.

C. The atoms lose their electrons more easily.

D. The oxides become more acidic.

**Answer:**



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2. Element 'X' forms a chloride with the formula  $XCl_2$  Which is a solid with a high melting point, X would most likely be in the same group of the periodic table as:

Na, Mg, Al, Si,

A. Na

B. Mg

C. Al

D. Si

**Answer:**



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**3.** Which element has: Two shells both of which are completely filled with electrons?



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4. Which element has: The electronic configuration 2, 8, 2?



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5. Which element has: A total of three shells, with four electrons in its valence shell?



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6. Which element has: A total of two shells with three electrons in its valence shell?



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7. Which element has: Twice as many electrons in its second shell as in its first shell?



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8. What property do all elements in the same column of the periodic table as boron have in common?



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9. What property do all elements in the same column of the periodic table as fluorine have in common?



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10. An atom has electronic configuration 2,8,7.

What is the atomic number of this elements?



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11. The position of three elements A, B and C in the periodic table are as shown below:

Group 16	Group 17
.....	.....
.....	A
.....	.....
B	C

State

whether a is a metal or non-metal



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