



CHEMISTRY

BOOKS - BEYOND PUBLICATION

ACIDS, BASES AND SALTS

Example

1. Which property do you think of while suggesting the remedy from a problem of an acidity?

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2. What is a neutralization reaction? Give two examples

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3. What is a neutralization reaction? Give two examples

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4. An acid or abase is mixed with water is the process exothermic or endothermic one?

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5. Solid barium oxide has ions, but it does not conduct electricity. Why ?

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6. Dry hydrogen chloride gas does not turn blue litmus to red whereas hydrochloric acid does. Why?

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7. Draw a neat diagram showing acid solution in water conducts electricity.

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8. How does the flow of acid rain into a river make the survival of aquatic life in a river difficult ?

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9. What are the harmful effects of burning fossil fuels? How it effect the environment?

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10. What is Baking powder ?

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11. Give two important uses of washing soda and baking soda.

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12. Five solutions A, B, C, D and E when tested with universal indicator showed pH as 4, 1, 11, 7 and 9 respectively, which solution is:

(a) neutral (b) strongly alkaline (c) strongly acidic (d) weakly acidic (e) weakly alkaline Arrange the pH in increasing order of hydrogen ion concentration.

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13. Five solutions A, B, C, D and E when tested with universal indicator showed pH as 4, 1, 11, 7 and 9 respectively, which solution is:

(a) neutral (b) strongly alkaline (c) strongly acidic (d) weakly acidic (e) weakly alkaline Arrange the pH in increasing order of hydrogen ion concentration.

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14. Five solutions A, B, C, D and E when tested with universal indicator showed pH as 4, 1, 11, 7 and 9 respectively, which solution is:

(a) neutral (b) strongly alkaline (c) strongly acidic (d) weakly acidic (e) weakly alkaline Arrange the pH in increasing order of hydrogen ion concentration.

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15. Five solutions A, B, C, D and E when tested with universal indicator showed pH as 4, 1, 11, 7 and 9 respectively, which solution is:

(a) neutral (b) strongly alkaline (c) strongly acidic (d) weakly acidic (e) weakly alkaline Arrange the pH in increasing order of hydrogen ion concentration.

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16. Five solutions A, B, C, D and E when tested with universal indicator showed pH as 4, 1, 11, 7 and 9 respectively, which solution is:

(a) neutral (b) strongly alkaline (c) strongly acidic (d) weakly acidic (e) weakly alkaline Arrange the pH in increasing order of hydrogen ion concentration.

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17. Why does tooth decay start when the pH of mouth is lower than 5.5?

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18. Is the pH changes tooth decay? Explain.

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19. A milkman adds a very small amount of baking soda to fresh milk.

Why does he shift the pH of the fresh milk from 6 to slightly alkaline

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20. Why does this milk take a long time to set as curd?

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21. Plaster of Paris should be stored in a moisture - proof container.

Explain why ?

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22. Equal lengths of magnesium ribbons are taken in test tubes A and B.

Hydrochloric acid is added to test tube A. while acetic acid is added to

test tube B. Amount and concentration of both the acids are same. In which test tube will the fizzing occur more vigorously and why?

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23. Fresh milk has a pH of 6. Explain why the pH changes as it turns into curd.

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24. How do you prepare your own indicator using beetroot? Explain.

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25. Compounds such as alcohols and glucose contain hydrogen but are not categorized as acids. Describe an activity to prove it

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26. What is meant by "water of crystallization" of a substance? Describe an activity to show the water of crystallization.

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27. How do you select a good place to plant a tree in your school/ at home. Test the soil and investigate and write a report on it.

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28. Do all vegetables are acids? To find this investigate with pH paper and tabulate the values and write a report on it?

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29. Collect Information about importance of the pH value in daily life to human beings as well as plants.





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30. What do you notice from the angle of deviation ?



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31. What do you understand from this activity ?



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32. Give example for use of olfactory indicator in your daily life.



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33. Assertion (A) : Pickles and sour substances are not stored in brass and copper vessels.

Reason: Acids reacts with metals.

A) 'A' and 'R' are correct. R is a correct reason for A.

B) 'A' and 'R' are correct. R is not a correct reason for A.

C) 'A' is correct, but 'R' is wrong.

D) 'A' is wrong, but 'R' is correct.

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34. What do you observe by adding dilute HCl to sodium carbonate and sodium hydrogen carbonate ?

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35. Why did the colour of the phenolphthalein in NaOH solution change after adding the HCl solution ?

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36. What does the pink colour indicate chart ?

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37. Do you guess the reason for reappearance of pink colour of phenolphthalein ?

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38. What do acids have in common?

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39. When bulbs are connected (resistors) in series, What do you notice ?

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40. Now, cover the torch glass with oily paper. What do you observe?

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41. Does the bulb glow in acids ?

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42. What do bases have in common?

 [Watch Video Solution](#)

43. Repeat the above activity using alkalis such as sodium hydroxide, calcium hydroxide solutions etc., instead of acid solutions.

Does the bulb glow ?

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44. If magnesium oxide reacts with an acid, predict the product that formed. What can you conclude from the above activity ?

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45. Do acids produce ions only in aqueous solution ? Explain an activity to observe this

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46. When conc sulphuric acid is added to solid sodium chloride taken in a test tube What do you observe ? Is there a gas coming out of the delivery tube ?

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47. Can you guess, why the leaf is green in colour?

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48. Teacher asked Ramesh to observe small depressions on surface of a potato. What do you infer in this observation ?



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49. The reaction of HCl with NaOH, Is it an exothermic or endothermic process ?



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50. Describe an activity to observe the reaction of metal oxides with acids.
What do you observe ?



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51. What is the nature of each substance on the basis of pH observations ?



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52. What are the strengths of conjugate acids of a strong base and weak base?

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53. What is the reason for the change in colour of copper sulphate when iron nail is kept on it ?

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54. Can you write the chemical equation of the reaction taking place zinc with dilute sulphuric acid?

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55. What can you conclude about the ideal soil pH for the growth of plants in your region?





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56. Under what soil conditions a farmer would treat the soil of his fields with quick lime or calcium carbonate?



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57. Write the formulae of the following salts

Sodium sulphate



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58. Potassium sulphate, sodium sulphate, calcium sulphate, copper sulphate. How many families can you identify among the salts given above ?



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59. Nature of aqueous solution of salts formed by weak acid and weak base is

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60. What does $10H_2O$ signify in the formula : $Na_2CO_{3.10}H_2O$?

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61. Why does Na_2CO_3 solution turn into a suspension, when saturated with CO_2 gas ?

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62. The colour of copper sulphate solution-

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63. The colour of copper sulphate solution-

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64. How can you get half a water molecule ?

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65. Sample solution of pH value 3. Identify the above sample as acidic or basic solution.

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66. You have two solutions, A and B. The pH of solution A is 6 and pH of solution B is 8. Which solution has more hydrogen ion concentration?
Which sample solutions are acidic and which are basic ?

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67. What are OHactory indicators? Write an activity to prove them.

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68. What are OHactory indicators? Write an activity to prove them.

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69. Write an activity to know the reaction of bases with metals.

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70. What are the material / substances required to produce hydrogen gas in your lab ? Write the process

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71. Write an activity to know the reaction of bases with metals.

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72. Write an activity which proves that bases produce hydrogen gas when they react with metals.

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73. Show that the reaction of carbonates and metal hydrogen carbonates with acids produces carbon dioxide gas.

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74. Write an activity to find the change of colour in the reaction of an acid with a base (Neutralization) reaction ?

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75. What is neutralization ? Explain an activity to demonstrate neutralization.

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76. Describe an activity to observe the reaction of metal oxides with acids.
What do you observe ?

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77. Show that the metal oxides are basic in nature through an activity.

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78. Write an activity to show that non-metallic oxide reacts with base is a neutralization.

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79. Write an activity to show that whether all compounds containing hydrogen are acids or not.

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80. Write an activity which proves acids are good conductors of electricity.

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81. Repeat the above activity using alkalis such as sodium hydroxide, calcium hydroxide solutions etc., instead of acid solutions.

Does the bulb glow ?

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82. Repeat the above activity using alkalis such as sodium hydroxide, calcium hydroxide solutions etc., instead of acid solutions.

Does the bulb glow ?

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83. Repeat the above activity using alkalis such as sodium hydroxide, calcium hydroxide solutions etc., instead of acid solutions.

Does the bulb glow ?

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84. Repeat the above activity using alkalis such as sodium hydroxide, calcium hydroxide solutions etc., instead of acid solutions.

Does the bulb glow ?

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85. Do acids produce ions only in aqueous solution ? Explain an activity to observe this

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86. Do acids produce ions only in aqueous solution ? Explain an activity to observe this

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87. What happens when an acid or base is mixed with water?

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88. Write an activity to show that dissolving of an acid in water is an exothermic process (or) endothermic process.

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89. Explain a test to know whether the acid (or base) is strong or weak.

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90. Write an activity to check the colour change in dilute HCl and antacid acid solution in addition of methyl orange.

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91. How can we test the pH value of the soil?

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92. Write the formulae of the following salts and classify them as families based on radicals. Potassium sulphate, sodium sulphate, calcium sulphate, magnesium sulphate, copper sulphate, sodium chloride, sodium nitrate, sodium carbonate and ammonium Chloride.

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93. Write the formulae of the following salts

Ammonium chloride

Identify the acids and bases for which the above salts are obtained also write chemical equations for the reactions between such acids and bases which type of chemical reactions they are.

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94. Collect the salt samples like sodium chloride, aluminium chloride, copper sulphate, sodium acetate, ammonium chloride, sodium hydrogen carbonate and sodium carbonate. Dissolve them in distilled water. Check the action of these solutions with litmus papers. Find the pH using pH paper. Classify them into acidic, basic or neutral salts. Identify the acid and base used to form the above salts. Record your observations in table

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95. Are the crystals of salts really dry ? Write an activity to find it.

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96. Is the substance present in antacid tablet acidic or basic?

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97. What type of reaction takes place in stomach when an antacid tablet is consumed ?

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98. You are provided with three test tubes containing distilled water, an acid and a base solution respectively. If you are given only blue litmus paper, how do you identify the contents of each test tube ?

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99. Which gas is usually liberated when an acid reacts with a metal ? How will you test for the presence of this gas ?

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100. Metal compound A reacts with dilute hydrochloric acid to produce effervescence. The gas evolved extinguishes a burning candle. Write a balanced chemical equation for the reaction if one of the compounds formed is calcium chloride.

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101. Why do HCl, HNO_3 , etc. show acidic characters in aqueous solutions while solutions of compounds like alcohol and glucose do not show acidic character ?

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102. While diluting an acid, why is it recommended that the acid should be added to water and not water to the acid ?

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103. What will happen if the pH value of chemicals in our body increases ?

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104. Why do living organisms have narrow pH range ?

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105. What are olfactory indicators ?

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106. what is plaster of Paris. give its chemical formula?



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107. What are alkalis ?



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108. What is a pH scale ?



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109. Write a short note about the pH of the soil ?



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110. Give pH of neutral, acid and base.



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111. Is the substance present in antacid tablet acidic or basic?



[Watch Video Solution](#)

112. What type of reaction takes place in stomach when an antacid tablet is consumed ?



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113. What is the name of aqueous sodium chloride ?



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114. write the common name of sodium hydrogen carbonate?



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115. Which salt is used in the manufacture of borax?

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116. Give example of the salt that possess water of crystallization?

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117. The chemical formula of plaster of paris is,

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118. Which gas is evolved when an acid reacts with metal ?

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119. reaction in which acid reacts with the base forming salt and water is called as?

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120. Bases which are soluble in water are called as?

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121. Who introduced pH ?

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122. Classify the following examples as acid, base (or) salt.

$Mg(OH)_2$, H_3PO_4 , KNO_2 , $Ba(OH)_2$, KCl , HBr , $NaCl$, HFO_4 , HCl , Al

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123. What is the change you observe in litmus paper with acid ?

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124. What is a litmus solution ?

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125. What is the change you observe in litmus paper with base ?

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126. What is the action of acids and bases with metals ? Give examples.

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127. What is the action of acids with carbonates and metal hydrogen carbonates?.

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128. What is neutralization reaction?

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129. What is the reaction of metal oxides with acids ?

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130. Metal oxide reacts with acid and gives salt and water. What is the nature of metal oxides?

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131. What do acids have in common?

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132. What is responsible for acidic property of acids?

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133. What do bases have in common?

 [Watch Video Solution](#)

134. What do the given symbol represents?

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135. When acid is added to water, what type of reaction is it ?

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136. How do you decide the strength of acid or base ?

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137. How do the universal indicator help us to know the strength of acid or base ?

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138. what is the pH value of solution?

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139. What is the range of a pH scale?

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140. How do you use pH scale to know whether a solution is acid or base ?

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141. What is the chemical name and formula of table salt?

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142. What is the nature of the salt $CaSO_4$ formed by the reaction between calcium hydroxide and sulphuric acid?

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143. What are the uses of Hydrochloric acid (HCl) ?

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144. What is bleaching powder? Write its formula.

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145. Write the chemical equation for preparation of Baking soda.

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146. What is Baking powder ?

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147. What is water of crystallization?

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148. How do Plaster of Paris obtain from gypsum?

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149. What is the reaction of Plaster of Paris with water ?

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150. What are the salts obtained from common salt ?

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151. What is a rock salt ?

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152. What does $10H_2O$ signify in the formula $:Na_2CO_{3.10}H_2O$?

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153. how do acids said neutralize bases ?



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154. how do acids and bases react with each other?



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155. how strong are acids and base solutions?



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156. Write the chemical formulae of the following: Bleaching powder



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157. Write the chemical formulae of the following: sodium chloride



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158. Write the chemical formulae of the following: slaked lime



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159. Write the chemical formulae of the following: baking soda



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160. Write the chemical formulae of the following: washing soda



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161. Write the chemical formulae of the following: gypsum



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162. Write the chemical formulae of the following: plaster of paris

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163. Write the chemical formulae of the following: acetic acid

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164. Write the chemical formulae of the following: sodium hydroxide

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165. Why do acids not show acidic behaviour in the absence of water?

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166. If someone in the family is suffering from a problem of acidity, which of the following would you suggest as a remedy : lemon juice, vinegar or baking soda solution ? Which property do you think of while suggesting the remedy ?

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167. What happens when an acid or base is mixed with water?

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168. Why pura acetic acid does not turn blue litmus to red?

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169. Why does the soil of agricultural lands get tested for pH?

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170. How does the clean cloth act, when it is kept with finely chopped onion in plastic bag.

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171. The pH of rain water collected from two cities A and B were found to be 6 and 5 respectively. The water of which city is more acidic ?

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172. What precaution to be taken while diluting the con. Acid?

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173. What happens if the copper sulphate crystals taken into dry test tube are heated ?

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174. How do you know the nature of salt formed due to the reaction between acids and bases?



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175. how washing soda is obtained?



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176. Write a short note on pH scale.



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177. What is the role of pH in our digestive system ?



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178. explain the self-defence by animal and plants through a chemical warfare?

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179. Write about universal indicator.

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180. What do acids have in common?

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181. What are antacids? Give example.

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182. What do bases have in common?

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183. Categorize the following as acids, bases, and salts :

lemon juice, salt water, soap water, tamarind juice, surf water, lime water.

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184. How can you prepare turmeric indicator ? What is the use of it ?

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185. Name two salts and water their formulae which possesses water of crystallization .

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186. What is neutralization reaction?

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187. Fresh milk has a pH of 6. Explain why the pH changes as it turns into curd.

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188. Why do curd and sour substances not be kept in copper vessels ?

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189. Acid should be added to water but not water to the acid. Why ?

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190. What value of pH in the mouth leads to tooth decay ? Why ?

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191. Name the four chemicals that are obtained from common salt and write their molecular formulae.

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192. Show that the reaction of carbonates and metal hydrogen carbonates with acids produces carbondio:dde gas.

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193. Dry hydrogen chloride gas does not turn blue litmus to red whereas hydrochloric acid does. Why?

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194. Equal lengths of magnesium ribbons are taken in test tubes A and B. Hydrochloric acid is added to test tube A. while acetic acid is added to test tube B. Amount and concentration of both the acids are same. In which test tube will the fizzing occur more vigorously and why?

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195. Give two important uses of washing soda and baking soda.

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196. Which gas is liberated when, acids react with metals ? Give one example?

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197. Write the chemical formulae of the following ?

Gypsum , Plaster of Paris.

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198. Describe how sodium hydroxide is obtained from common salt.

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199. Describe process of preparation of bleaching powder ? Write its uses.

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200. Write the chemical equation of preparation of baking soda. What are the uses of baking soda ?

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201. How do you prepare washing soda ? What are its uses ?

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202. Draw a neat diagram showing variation of pH with the change in concentration of H^+ ions and OH^- (aq) ions.

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203. Additional Information can we get from a chemical equation?

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204. Describe the uses of Acids, Bases and Salts.

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205. Distinguish between acids and bases.

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206. Define pH. Calculate the pH of 0.001 M of HCl

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207. Define the following. Give one example for each.

Strong acid

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208. Define the following. Give one example for each.

Strong base

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209. Define the following. Give one example for each.

Weak acid



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210. Define the following. Give one example for each.

Weak acid



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211. Write any four chemical properties of acids.



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212. Write the formulae of the following salts

Ammonium chloride

Identify the acids and bases for which the above salts are obtained also

write chemical equations for the reactions between such acids and bases which type of chemical reactions they are.

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213. Five solutions A, B, C, D, E, F as 5, 2, 1, 3, 7 and 9 respectively which solution is

- (a) Neutral
- (b) Strongly alkaline
- (c) Strongly acid
- (d) Weakly acidic

Arrange the pH in increasing order of Hydrogen ion concentration.

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214. Five solutions A, B, C, D, E, F as 5, 2, 1, 3, 7 and 9 respectively which solution is

- (a) Neutral
- (b) Strongly alkaline

(c) Strongly acid

(d) Weakly acidic

Arrange the pH in increasing order of Hydrogen ion concentration.

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215. Five solutions A, B, C, D, E, F as 5, 2, 1, 3, 7 and 9 respectively which solution is

(a) Neutral

(b) Strongly alkaline

(c) Strongly acid

(d) Weakly acidic

Arrange the pH in increasing order of Hydrogen ion concentration.

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216. Five solutions A, B, C, D, E, F as 5, 2, 1, 3, 7 and 9 respectively which solution is

(a) Neutral

(b) Strongly alkaline

(c) Strongly acid

(d) Weakly acidic

Arrange the pH in increasing order of Hydrogen ion concentration.

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217. Name some natural indicators.

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218. Give some examples of synthetic indicators.

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219. Observe the following table.

If the aqueous solutions of A and B are mixed, what will be that colour of the solution so formed in the presence of methyl orange indicator ? Why

?

Substance	Test with indicator			
	Methyl orange	Phenolphthalene	Red litmus	Blue litmus
A	Red	—	—	Red
B	Yellow	Pink	Blue	—
C	No change	No change	No change	No change

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220. Observe the following table.

If red litmus paper is dipped in solution A, what is the change that you observed? Why?

Substance	Test with indicator			
	Methyl orange	Phenolphthalene	Red litmus	Blue litmus
A	Red	—	—	Red
B	Yellow	Pink	Blue	—
C	No change	No change	No change	No change

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221. What is the smell of onion when treated with an acid and a base?

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222. What is the smell of vanilla extract when treated with an acid and a base ?

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223. Which gas is evolved when an acid reacts with metal ?

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224. What is the action of acids with carbonates and metal hydrogen carbonates?.

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225. Who am I?

I can roughly measure pH value from 0-14.

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226. Who am I?

I am called antichlor and am used to remove excess chlorine from clothes when treated with bleaching powder.



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227. Who am I?

I am a product of gypsum and am used to making chalks and fire proof materials.



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228. Who am I?

I am a compound of calcium and can be used for disinfecting drinking water as well as for decolourisation.



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229. Who am I?

I give different smell in acid and base solution

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230. Who am I?

I am an oxide capable of showing properties for both acids and bases.

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231. Who am I?

I am a covalent compound and conducts electricity in aqueous medium.

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232. Who am I?

I am a salt of potassium hydroxide and nitric acid.



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233. Who am I?

I am the term used when a solid becomes liquid when exposed to moist air.

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234. Who am I?

I am derived from tomato and turn blue litmus into red.

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235. If the pH values of solutions X, Y and Z are 13, 6 and 2 respectively then
which solution is a strong acid ? Why ?

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236. If the pH values of solutions X, Y and Z are 13, 6 and 2 respectively then which solution contains ions along with molecules of solution '?

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237. If the pH values of solutions X, Y and Z are 13, 6 and 2 respectively then which solution is a strong base? Why?

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238. If the pH values of solutions X, Y and Z are 13, 6 and 2 respectively then does the pH value of a solution increase or decrease when a base is added to it? Why?

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239. Discuss briefly the examples showing the importance of pH in daily life.

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240. Read the information given in the table and answer the following questions.

List out the acids in the above table.

Solution	pH value	Reaction with Phenolphthalein solution	Reaction with Methyl orange solution
HCl	1	No colour change.	Turns into red colour.
Distilled water	7	No colour change.	No colour change.

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241. Read the information given in the table and answer the following questions.

Name the strongest acid and the strongest base among the given solutions.

Solution	pH value	Reaction with Phenolphthalein solution	Reaction with Methyl orange solution
HCl	1	No colour change.	Turns into red colour.
Distilled water	7	No colour change.	No colour change.

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242. Name the four chemicals that are obtained from common salt and write their molecular formulae.

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243. Based on the properties of acids, bases and neutral solutions, fill the following table.

Indicators	Acidic solution	Basic solution	Neutral solution
Red litmus			No change in colour
Blue litmus	Red		
Phenolphthalein	No change in colour		
Methyl orange		yellow	
Universal			Parrot green

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244. Fill the following of results of reactions between some substances (acids, bases, neutral substances) and indicators.

Indicator Substance	Litmus blue paper	Litmus Red paper	Methyl orange Indicator	Phenolphthalein solution
HCl		No reaction		
NaOH			Turned into yellow	
Tomato juice				No reaction
Normal		Normal		

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1. The colour of methyl orange indicator in acidic medium is

- A. yellow
- B. green
- C. orange
- D. red

Answer:



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2. The colour of phenolphthalein indicator in basic solution is

- A. yellow
- B. green
- C. pink

D. orange

Answer:



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3. Colour of methyl orange in alkali conditions

A. orange

B. yellow

C. red

D. blue

Answer:



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4. A solution turns red litmus blue, its pH is likely to be

A. 1

B. 4

C. 5

D. 10

Answer:



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5. One of the following solutions reacts with crushed egg shells to give a gas that turns lime- water milky, the solution is of ____

A. NaCl

B. HCl

C. LiCl

D. KCl

Answer:

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6. If a base dissolves in water, by what name is it better known?

A. neutral

B. base

C. acid

D. alkali

Answer:

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7. Which of the following substances when mixed together will produce table salt ?

A. Sodium thiosulphate and sulphar dioxide

B. Hydrochloric acid and sodium hydorxide

C. Chlorine and oxygen

D. Nitric acid and sodium hydrogen carbonate

Answer:

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8. What colour would hydrochloric acid (pH = 1) turn universal indicator '?

A. orange

B. purple

C. yellow

D. red

Answer:

 [Watch Video Solution](#)

9. Which one of the following types of medicines is used for treating indigestion ?

- A. antibiotic
- B. analgesic
- C. antacid
- D. antiseptic

Answer:



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10. What gas is produced when magnesium is made to react with hydrochloric acid ?

- A. Hydrogen
- B. Oxygen
- C. Carbon dioxide

D. No gas is produced

Answer:

 [Watch Video Solution](#)

11. Does the bulb glow in acids ?

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12. Which measurements are able to construct a triangle?

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13. What is the chemical name of baking Soda ?

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14. What is a neutralization reaction? Give two examples

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15. How does the flow of acid rain into a river make the survival of aquatic life in a river difficult ?

 [Watch Video Solution](#)

16. Why does tooth decay start when the pH of mouth is lower than 5.5?

 [Watch Video Solution](#)

17. Which of the following is the most accurate way of showing neutralization ?

A. Acid + Base \rightarrow Acid - base solution

B. Acid + Base \rightarrow Salt + Water

C. Acid + Base \rightarrow Sodium chloride + Hydrogen

D. Acid + Base \rightarrow Neutral solution

Answer:

 [Watch Video Solution](#)

18. Why is universal indicator a better one Than litmus paper ?

A. Litmus paper can only by used for acids.

B. Litmus paper can only be used for alkalis.

C. Universal indicator goes green if some thing is neutral.

D. Universal indicator is useful for all ranges of pH of the solution.

Answer:

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19. Which of the following is used for making toys ?

- A. Gypsum
- B. Calcium carbonate
- C. Plaster of Paris
- D. Bleachng power

Answer:



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20. Brine solution is called

- A. sodium chloride
- B. potassium chloride
- C. copper chloride
- D. calcium chloride

Answer:



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21. Tooth decay start when the pH of the mouth is lower than-

A. 5.5

B. greater than 5.5

C. lower than 5.5

D. all of these

Answer:



[Watch Video Solution](#)

22. The chemical formula of hydrated copper sulphate is-

A. $CuSO_{4.6}H_2O$



Answer:



[Watch Video Solution](#)

23. If the pH value of a solution is 12, then the character of solution of the solution is

A. Base

B. Strong acid

C. Acid

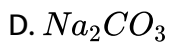
D. Strong base

Answer:



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24. Which of the following is used in cakes or pastries to making them light and fluffy ?

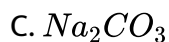
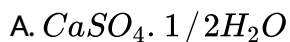


Answer:



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25. Which salt is used in the manufacture of borax?



D. NaHCO_3

Answer:

 [Watch Video Solution](#)

26. Which of the following is the reason for the bluish green colour solution, when copper oxide reacts with dilute HCl ?

- A. copper(II)chloride
- B. copper (IV) chloride
- C. copper(III) chloride
- D. copper (I) chloride

Answer:

 [Watch Video Solution](#)

27. Excess of sodium hydroxide reacts with zinc to form

A. H_2

B. N_2

C. O_2

D. None of these

Answer:



[Watch Video Solution](#)

28. __ is a mild non-corrosive base.

A. Washing soda

B. Baking soda

C. Baking soda

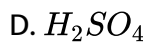
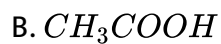
D. Bleaching powder

Answer:



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29. Which of the following is the weakest acid ?

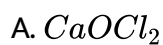


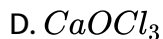
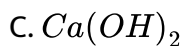
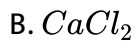
Answer:



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30. Formula of bleaching powder





Answer:

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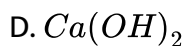
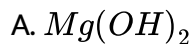
31. Which of the following gas is liberated, when sodium carbonate is reacted with Hydrochloric acid ?



Answer:

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32. Which one of the following is the weakest base?

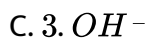
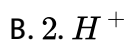
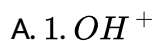


Answer:



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33. Because which of the absence of following ions, in the glucose and alcohol solution the bulb did not glow ?



D. 4. H^-

Answer:

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34. reaction in which acid reacts with the base forming salt and water is called as?

A. Neutralization

B. Reduction

C. Oxidation

D. Crystallisation

Answer:

 [Watch Video Solution](#)

35. Which of the following mild base is used on the stung area gives relief ?

- A. Calcium hydroxide
- B. Baking soda
- C. Sodium hydroxide
- D. Washing soda

Answer:



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36. Bases which are soluble in water are called as?

- A. Acidic
- B. Basic
- C. Alkali
- D. Neutral

Answer:

 [Watch Video Solution](#)

37. Which of the following is used to enhance the taste of food ?

A. sodium carbonate

B. sodium sulphate

C. sodium chloride

D. sodium bicarbon

Answer:

 [Watch Video Solution](#)

38. ___ is slightly soluble in water.

A. $Ca(OH)_2$

B. KOH

C. $Mg(OH)_2$

D. $Be(OH)_2$

Answer:

 [Watch Video Solution](#)

39. Stinging hair of leaves of nettle plant, Inject _____ acid causing burning pain

A. Methanoic acid

B. Hydrochloric acid

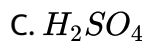
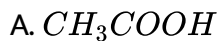
C. Nitric acid

D. Sodium thiosulphate

Answer:

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40. Which of the following acid is produced by human's stomach ?



Answer:



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41. Which of the following precipitate is formed on mixing calcium hydroxide with carbon dioxide ?

A. Red precipitate of calcium carbonate

B. Blue precipitate of calcium carbonate

C. White precipitate of calcium carbonate

D. Black precipitate of calcium carbonate

Answer:

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42. Which of the following is obtained recrystallization of sodium carbonate ?

- A. Washing soda
- B. Baking soda
- C. Bleaching soda
- D. None of these

Answer:

 [Watch Video Solution](#)

43. __ is produced by action of chlorine on dry slaked lime.

- A. Washing soda
- B. Bleaching soda
- C. Plaster of Paris
- D. Baking soda

Answer:



[Watch Video Solution](#)

44. Acidic substances contains ions.

- A. OH^-
- B. H^+
- C. Na^+
- D. Cl^-

Answer:



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45. pH for base is more than

A. 4

B. 3

C. 7

D. 5

Answer:



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46. What is neutralization reaction?

A. acid

B. salt

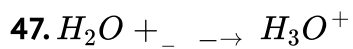
C. base

D. ice

Answer:



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A. H^+

B. OH^-

C. H_2O

D. H_3O

Answer:



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48. The aqueous solution of ----- conducts electricity.

- A. Ethyl alcohol
- B. Acetic acid
- C. Acetone
- D. Ether

Answer:



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49. Which of the following substance is used as an antichlor ?

- A. $CaOCl_2$
- B. $Na_2S_2O_3$
- C. Na_2SO_4
- D. $CuSO_4$

Answer:



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50. Which of the following substance has the lowest pH value?

- A. Sugar
- B. Tomato juice
- C. Vinegar
- D. Washing soda

Answer:



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51. the acid formed in stomach is-

- A. HCl

B. H_2SO_4

C. CH_3COOH

D. HNO_3

Answer:



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52. Many salts absorb water from atmosphere this property is called

A. Crystallisation

B. Hydration

C. Deliquescence

D. Efflorescence

Answer:



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53. Antacid medicine is used for indigestion because

- A. it neutralizes the acid produced
- B. it neutralizes digested food material
- C. it oxidizes food material
- D. it helps to produce digestive juices

Answer:



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54. Antacids are used.....

- A. to produce acid in the stomach.
- B. to produce water in the stomach.
- C. to neutralise the excess base in the stomach.
- D. to neutralise the excess acid in the stomach.

Answer:



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55. Tooth paste is.....in nature.

- A. Acidic
- B. Base
- C. Neutral
- D. Amphoteric

Answer:



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56. Which one of the following types of medicines is used for treating indigestion ?

A. Antibiotic

B. Analgesic

C. Antacid

D. Antiseptic

Answer:

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57. Metal oxide + Acid \rightarrow salt + Water. To perform this reaction in your lab. You should collect

A. Salt + metal

B. Salt + water

C. Base + water

D. Non-metallic oxide + Base

Answer:

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58. Test tube 'P' contain $NaHCO_3$ solution. Test tube 'Q' contain lemon juice. On introducing pH paper strips on both of them it is observed that the pH paper turns.

- A. Blue in P and Red in Q
- B. Red in P and Pink in Q
- C. Red in P and Blue in Q
- D. Blue in both

Answer:

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59. Match the following Set A with Set B.

<u>Set A</u>	<u>Set B</u>
i) POP	P) NaHCO_3
ii) Bleaching powder	Q) CaOCl_2
iii) Baking Soda	R) $\text{CaSO}_4 \cdot \frac{1}{2} \text{H}_2\text{O}$
iv) Washing Soda	S) Na_2CO_3

A. i-R,ii-Q,iii-P,iv-S

B. i-R,ii-P,iii-Q,iv-S

C. i-P,ii-R,iii-Q,iv-S

D. i-p,ii-R,iii-S,iv-Q

Answer:



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60. Acid should be added to water but not water to the acid. Why ?

A. Both Naoushad and Sreenu are correct

B. Naoushad correct, Sreenu incorrect

C. Naoushad is incorrect and Sreenu is correct

D. Both Naoushad and Sreenu are incorrect.

Answer:

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61. Match the following.

<u>A</u>	<u>B</u>
P) Plaster of Paris	i) NaHCO_3
Q) Gypsum	ii) $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
R) Baking Soda	iii) Na_2CO_3
S) Washing Soda	iv) $\text{CaSO}_4 \cdot \frac{1}{2} \text{H}_2\text{O}$

A. P-iv,Q-ii,R-I,S-iii

B. P-iv,Q-ii,R-iii,S-i

C. P-ii,Q-iv,R-I,S-iii

D. P-ii,Q-iv,R-iii,S-i

Answer:



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62. Identify the pair of pH values of strong acid and strong base in the following.

A. (6,14)

B. (1,8)

C. (7,7)

D. (2,14)

Answer:



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63. The acid which enters the body by the sting of bee is

- A. Acetic acid
- B. Methanoic acid
- C. Sulphuric acid
- D. Fatty acid

Answer:



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64. Which one of the following metals reacts both with acid and base and release hydrogen gas ?

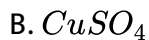
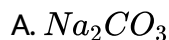
- A. Na
- B. Fe
- C. Cu
- D. Zn

Answer:



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65. The gas that turns lime water to milky is _____ .



Answer:



[Watch Video Solution](#)

66. Which one of the following is given to a person who suffers from acidity to get relief from it ?

A. Carbonated water

B. Baking soda

C. Vinegar

D. Lime juice

Answer:

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67. An aqueous solution of the salt is acidic which of the following acids and bases react to give this salt ?

A. Strong acid and strong base

B. Strong acid and weak base

C. Weak acid and strong base

D. Weak acid and weak base

Answer:

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68. Select a pair of basic salts among the following.

- A. Sodium chloride and Sodium acetate
- B. Sodium acetate and Sodium bicarbonate
- C. Sodium carbonate and Sodium sulphate
- D. Sodium nitrate and Sodium oxalate

Answer:

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69. Which of the following is oxide mineral

- A. Oxalic acid
- B. Citric acid
- C. Acetic acid

D. Phosphoric acid

Answer:

 [Watch Video Solution](#)

70. An acid is regarded as strong if it :

- A. turns blue litmus red
- B. has a strong action on skin has
- C. is completely dissociated in aqueous solution
- D. is highly soluble in water

Answer:

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71. Which property is mostly related to the acids ?

- A. Sour taste
- B. Bitter taste
- C. Soapy touch
- D. Pleasant smell

Answer:

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72. Bases on ionisation release

- A. chloride ions
- B. hydroxide ions
- C. hydrogen ions
- D. sodium ions

Answer:

 [Watch Video Solution](#)

73. The two colours seen at the extreme end of the pH chart are :

- A. green and blue
- B. orange and green
- C. red and green
- D. red and violet

Answer:



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74. Which of the following has $\text{pH} = 0$?

- A. pure water
- B. $1\text{MCH}_3\text{COOH}$ solution in water
- C. 1M HCl solution in water

D. 1M NaOH solution in water

Answer:

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75. Which is not required to find the pH of a solution ?

A. Universal indicator

B. pH paper

C. Standard pH chart

D. Litmus paper

Answer:

 [Watch Video Solution](#)

76. If the pH of a solution is 13, this means that it is :

A. strongly acidic

B. strongly basic

C. weakly acidic

D. weakly basic

Answer:

 [Watch Video Solution](#)

77. How do Plaster of Paris obtain from gypsum?

A. Slaked lime

B. Gypsum

C. Lime stone

D. Quick lime

Answer:

 [Watch Video Solution](#)

78. Bleaching powder is obtained by the reaction chlorine with

- A. washing soda with chlorine
- B. caustic soda with chlorine
- C. slaked lime with chlorine
- D. baking soda with chlorine

Answer:



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79. The difference of water molecules in one molecule of each gypsum and plaster of paris is

- A. $3/2$
- B. $1/2$
- C. 2

Answer:



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80. A few drops of methyl orange are added to a soap solution the colour of the solution becomes-

A. Orange

B. Yellow

C. Pink

D. Remains colourless

Answer:



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81. A few drops of methyl orange are added to a soap solution the colour of the solution becomes-

- A. Red
- B. Pink
- C. Yellow
- D. No specific

Answer:



[Watch Video Solution](#)

82. A few drops of aqueous NaOH are dropped on a pH paper. The colour of the pH paper will turn to :

- A. violet
- B. yellow
- C. green

D. blue

Answer:



[Watch Video Solution](#)

83. Which of the following metal does not produces dihydrogen gas with dilute hydrochloric acid?

A. Cooper

B. Zinc

C. Iron

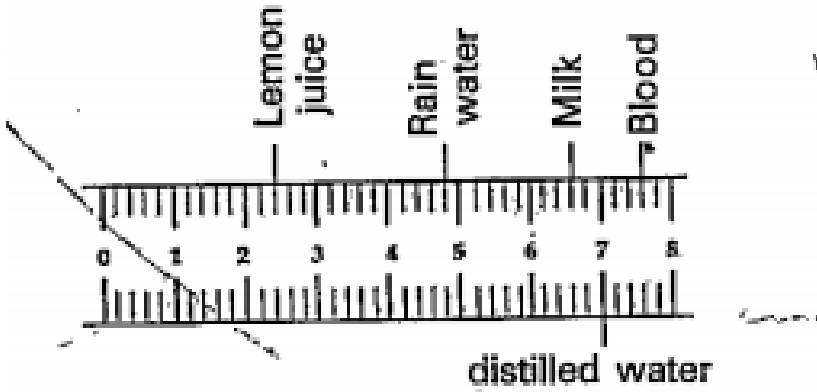
D. Aluminium

Answer:



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84. Observe the figure. Which is base among them ?



A. Lemon Juice

B. Rainwater

C. Blood

D. Milk

Answer:

[Watch Video Solution](#)

85. Observe the box. Which is the weak acid among them ?

Substance	pH
Gastric	1.2
Washing Soda	12.8
Sea water	8.0
Beer	4.3
Milk	6.6

- A. Washing Soda
- B. Milk
- C. Gastric
- D. Beer

Answer:



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86. Four students were given colourless liquids A,B,C of water, lemon juice and a mixture of water and lemon juice respectively. After testing these liquids with pH paper, following sequences in colour change of pH paper were reported.

Blue, Red and Green

Orange, Green and Green

Green, Red and Red

Red, Red and Green

A. c

B. a,b

C. d,c

D. b,d

Answer:



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87. What happens when a solution of an acid is mixed with a solution of a base in a test tube ?

The temperature of the solution increases

The temperature of the solution decreases

The temperature of the solution remains constant

Salt formation takes place

A. b,c

B. a,d

C. a

D. a,c

Answer:



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88. Sodium hydrogen carbonate when added to acetic acid evolves a gas.

Which of the following statements are true about the gas evolved ?

It turns lime water milky

It extinguishes burning splinter

It dissolves in a solution of sodium hydroxide

It has pungent odour

A. b,c,d

B. a,b,c

C. a,d

D. a,b

Answer:



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89. Common salt used in kitchen as well as be used as the raw material for making :

Washing soda

Bleaching powder

Baking soda

Slaked lime

A. a,b,d

B. a,c,d

C. a,c

D. a,b

Answer:



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90. Bases turn phenolphthalein into colour.

A. Red

B. Yellow

C. Pink

D. No change

Answer:



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91. __ are sour to taste and turn blue litmus to _

- A. Bases, red
- B. Bases, blue
- C. Acids, red
- D. Acids blue

Answer:



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92. The substances that are soapy to touch

- A. Bases, red

B. Bases, blue

C. Acids, red

D. Acids, blue

Answer:



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93. reaction in which acid reacts with the base forming salt and water is called as?

A. salt & water

B. only water

C. concentrated acid

D. concentrated base

Answer:



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94. __ is example of synthetic indicator.

A. Acid

B. Red litmus

C. Blue litmus

D. Phenolphthalein

Answer:



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95. Which gas is usually liberated when an acid reacts with a metal ? How will you test for the presence of this gas ?

A. O_2

B. H_2

C. Cl_2

D. CO_2

Answer:



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96. When base react with metal ___ gas is evolved.

A. O_2

B. H_2

C. Cl_2

D. CO_2

Answer:



[Watch Video Solution](#)

97. reaction in which acid reacts with the base forming salt and water is called as?

- A. Crystallisation
- B. oxidation
- C. reduction
- D. neutralisation

Answer:



[Watch Video Solution](#)

98. mixing an acid or base with water is process called

- A. dilution
- B. concentration
- C. addition
- D. substitution

Answer:



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99. The p in pH stands for _____.

- A. potenz
- B. positive
- C. presence
- D. potato

Answer:



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100. Higher the hydronium ion concentration lower is the _____ value.

- A. OH

B. pH

C. pOH

D. concentration

Answer:



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101. The pH value in acidic solution is

A. 7

B. gt 7

C. lt 7

D. No value

Answer:



[Watch Video Solution](#)

102. The pH value in basic solution is _____.

- A. 7
- B. gt7
- C. lt7
- D. No value

Answer:



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103. Acids that gives less H_3O^+ ions are said to be

- A. acids
- B. weak acids
- C. strong acids
- D. strong bases

Answer:

 [Watch Video Solution](#)

104. Acids that gives less H_3O^+ ions are said to be

- A. strong acid
- B. strong base
- C. weak base
- D. weak acid

Answer:

 [Watch Video Solution](#)

105. Our body works within the pH range of.....to.....

- A. 1.0, 2.3

B. 7.0, 7.8

C. 8.5, 9.5

D. 6.3 , 7.0

Answer:

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106. Among the following elements most acidic oxide is given by

A. Mg

B. Na

C. S

D. Zn

Answer:

 [Watch Video Solution](#)

107. pH was introduced by

- A. Sorensen
- B. Bohr
- C. Lewis
- D. Rutherford

Answer:



[Watch Video Solution](#)

108. pH value of Lemon juice is

- A. 4.2
- B. 7.4
- C. 2.5
- D. 5.8

Answer:



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109. Indicators used in acid-base filtrations are

- A. strong organic acids
- B. strong organic bases
- C. weak organic acids (or) bases
- D. non electrolytes

Answer:



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110. A universal indicator

- A. can be used in all acid-base filtration

B. is a mixture of several indicators

C. is used in the filtration of a weak and against weak base

D. none

Answer:

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111. Which of the following is the weakest base ?

A. NaOH

B. $Ca(OH)_2$

C. NH_4OH

D. KOH

Answer:

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112. Acetic acid is a weak acid. List in order of descending concentration all of the ionic and molecular species present in 1M aqueous solution of acetic acid.

- A. its molecular mass is high
- B. it is a covalent compound
- C. it is highly unstable
- D. it doesnot dissociate much (or) its ionisation is very small.

Answer:

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113. NH_4^+ ion in an aqueous solution will behave as

- A. a base
- B. an acid
- C. both acid and base

D. neutral

Answer:



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114. Which of the following is an acidic salt ?

A. $NaHSO_4$

B. Na_2SO_4

C. Na_2SO_3

D. Na_2SO_4

Answer:



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115. When pH of 0.0001 M HCl solution is

A. 1.0

B. 2.0

C. 13,1

D. 1, 13

Answer:

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116. Aqueous solution of copper sulphate

A. turns blue litmus red

B. turns red litmus blue

C. does not affect litmus

D. affects both red and blue litmus

Answer:

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117. Which Is most basic?

- A. pH = 8
- B. pH = 7 - 8
- C. pH = 12
- D. pH = 6

Answer:

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118. Which apparatus is required to test hydrogen gas?

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119. What do you know by glowing bulb ?

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120. What are the reason for the flow of current through the solution ?

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121. Due to which ions, acids allow current through them ?

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122. What does it mean that pH of a solution is 1 ?

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123. Does it tell only about acidic nature or also about basic nature ?

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124. When pH value increase from 1 to 7, which ion concentration decrease ?

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125. What value may a strong base has ?

A. 1. 1 to 4

B. 2. 4 to 5

C. 3. 9 to 12

D. 4. 12 to 14

Answer:

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126. Who introduced pH ?

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127. For what purpose, baking soda is used in cooking ?



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128. How do baking powder is made from baking soda ?



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129. Why do it is used in antacid tablets ?



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130. Why do it is used in antacid tablets ?



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131. When a honeybee stings you, what is the role of baking soda to give relief from the pain ?

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132. If magnesium oxide reacts with an acid, predict the product that formed. What can you conclude from the above activity ?

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133. Ramesh took two beakers A and B full of two various solutions when connected them to an electric circuit. In beaker A the bulb glows, in beaker B bulb doesn't glow why ? What are the solution may be ?

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134. Suresh's skin has burned while diluting conc. H_2SO_4 . Why is it so ?

How will he dilute conc.

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135. In rocky areas, a curdy substance formed when a soap is added to water. To change its nature, which chemical is to be added to water ?

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136. The reagent used in the preparation of chloroform is ___

A. $CaOCl_2$

B. $CaSO_4$

C. $NaHCO_3$

D. Na_2CO_3

Answer:



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137. Bases turn phenolphthalein into colour.

- A. Red
- B. Yellow
- C. Pink
- D. No change

Answer:



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138. Non-metallic oxides are in ___ nature

- A. acidic
- B. basic
- C. neutral

D. effervescence

Answer:



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139. In glucose and alcohol solution _____ are absent.

A. acids

B. bases

C. ions

D. reagents

Answer:



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140. To get rid of stomach pain caused by indigestion, people use bases called _____

- A. antacid
- B. paracetamol
- C. aspirin
- D. glucose

Answer:



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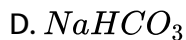
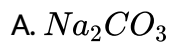
141. Tooth decay start when the pH of the mouth is lower than-

- A. 6.3
- B. 7.2
- C. 8.5
- D. 5.5

Answer:

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142. Which is the formula of washing soda ?



Answer:

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143. $pH = 0$ for a solution. It contains

A. M or number of H^+ ions

B. *M* or *evmberof* $OH^{-i}ons$

C. *Equalvumberof* H^{+} and $OH^{-i}ons$

D. *It does \neg conta \in* H^{+} and $OH^{-i}ons$

Answer:

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144. HCl , H_2SO_4 , CH_3COOH are some acids. Which of the following is not true for the above acids ?

A. All are produce H^{+} ions in water

B. All are turns blue litmus to red

C. pH of the compounds is less than 7

D. All are strong acids

Answer:

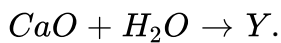
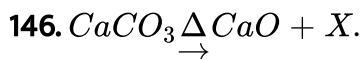
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145. Which of the following is correct ?

- A. Only synthetic indicators are used to identify the nature of solution
- B. Indicators are strong acids or strong bases
- C. Indicators are weak acids or weak bases
- D. Universal indicator gives the same colour at different pH values.

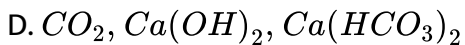
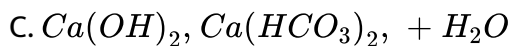
Answer:

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X, Y, Z are:





Answer:



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147. Which of the following is not true about pH ?

A. It indicates the concentration of H^+ ions in solutions of less than 1 molar.

B. At pH = 0, the hydronium ion concentration is 1 molar.

C. At pH = 0, the hydroxyl ion concentration is 1 molar.

D. All the above

Answer:



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148. $NaHCO_3 + HCl \xrightarrow{\Delta} NaCl + H_2O + X$. The gas 'X' is passed through aqueous Na_2CO_3 solution to form compound 'Y'. Then X and Y are

- A. $X = CO_2, Y = Na_2CO_3$
- B. $X = CO_2, Y = NaHCO_3$
- C. $X = H_2, Y = NaHCO_3$
- D. $X = Cl_2, Y = NaOH$

Answer:

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149. A substance 'A' on heating gives a Colourless gas. The residue is dissolved in water to form 'B'. When excess of CO_2 is bubbled through a solution of 'B', 'C' is formed. 'C' on gently heating forms 'A'. Then the substance 'A' is

A. Calcium carbonate

B. Sodium carbonate

C. Calcium nitrate

D. Sodium bicarbonate

Answer:

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150. Which type of bases allow electricity through them ?

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151. Some bases do not allow electricity through them. Why ?

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152. Yet sodium ethoxide is a salt, it has basic nature. Why ?

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153. When a soap added to water, white curdy substance will form why ?

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