CHEMISTRY

BOOKS - VGS BRILLIANT CHEMISTRY (TELUGU ENGLISH)

CHEMICAL EQUATIONS

Exercise

1. What changes do you notice generally?

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2. "Coal is burnt", "crackers are burnt" .... Changes

Are they physical (or) chemical changes?

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3. Are they (coal, crackers) temporary changes or permanent changes?
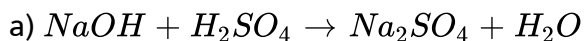
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4. What can be interpreting from a chemical equation ?

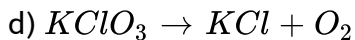> ▶ **Watch Video Solution**

5. Why should we balance a chemical equation?

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6. Balance the following chemical equations.

a) $NaOH + H_2SO_4 \rightarrow Na_2SO_4 + H_2O$

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7. Balance the following chemical equations.

d) $KClO_3 \rightarrow KCl + O_2$

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8. Balance the following chemical equations.

b) $Hg(NO_3)_2 + KI \rightarrow HgI_2 + KNO_3$

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9. Balance the following chemical equations including the physical states.

1) $C_6H_{12}O_6 \rightarrow C_2H_5OH + CO_2$
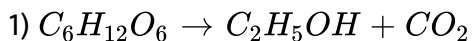
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10. Balance the following chemical equations including the physical states.

3) $NH_3 + Cl_2 \rightarrow N_2H_4 + NH_4Cl$
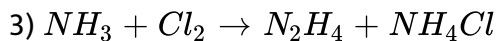
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11. Balance the following chemical equations including the physical states.

4) $Na + H_2O \rightarrow NaOH + H_2$
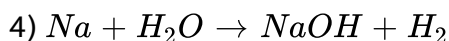
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12. Write the balanced chemical equations for the following and identify the type of reaction in each case.

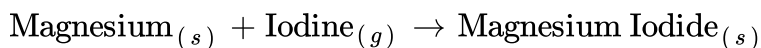$\text{Calcium hydroxide}_{(aq)} + \text{Nitric acid}_{(aq)} \rightarrow \text{water}_{(I)} + \text{Calcium nitrate}_{(}$
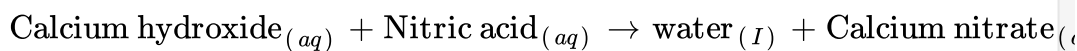
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13. Write the balanced chemical equations for the following and identify the type of reaction in each case.

$\text{Magnesium}_{(s)} + \text{Iodine}_{(g)} \rightarrow \text{Magnesium Iodide}_{(s)}$

14. Balance the chemical equation by including the physical states of the substances for the following reactions.

b) Sodium hydroxide reacts with hydrochloric acid to produce sodium chloride and water.

15. Balance the chemical equation by including the physical states of the substances for the following reactions.

a) Barium chloride and sodium sulphate aqueous solutions react to give insoluble barium sulphate and aqueous solution of sodium chloride.

16. 2 moles of Zinc reacts with a cupric chloride solution containing 6.023×10^{22} formula units of $CuCl_2$. Calculate the moles of copper

obtained. $Zn_{(s)} + CuCl_{2(aq)} \rightarrow ZnCl_{2(aq)} + Cu_{(s)}$

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17. 1 Mole of propane (C_3H_8) on combustion at STP gives 'A' kilo joules of heat energy. Calculate the heat liberated when 2.4 ltrs of propane on combustion at STP.

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18. Calculate the mass and volume of oxygen required at STP to convert 2.4 kg of graphite into carbon dioxide.

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19. How do you test the nature of the solution formed by dissolving CaO in water? what is the nature of the solution?

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20. Explain the reaction between sodium sulphate and Barium chloride.

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21. Explain the reaction of Zinc with HCl and write a balanced equation.

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22. You have brushed the wall with an aqueous suspension of $Ca(OH)_2$. After two days the wall turned to white colour. Write the balanced chemical reactions for the above changes using the appropriate symbols and formulae.

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23. How do we know a chemical reaction has taken place?

24. Count the number of atoms of each elements on left and right of the arrow in the equation $CaO + H_2O \rightarrow Ca(OH)_2$.

25. Na_2SO_4 + $BaCl_2$ $\rightarrow$ $BaSO_4$ + NaCl. Do the atoms of each element on left side equal to the atoms of the element on the right side of the equation?

26. $2C_3H_8 + 10O_2 \rightarrow 6CO_2 + 8H_2O$

Is it a balanced equation as per rules? How do you say?

27. $C_3H_8 + O_2$ rarr $CO_2 + H_2O$ రసాయన సమీకరణాలు తుల్యం చేయండి

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28. If you keep an iron piece in solid state $CuSO_4$ crystals, does it get any reaction? Guess the reason.

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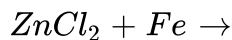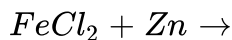
29. Balance the following chemical equation.

i) $Na + H_2O \rightarrow NaOH + H_2$
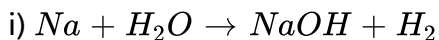
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30. Balance the chemical equation $Fe_2O_3 + Al \rightarrow Fe + Al_2O_3$. Write the steps of balancing.

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31. On adding dilute hydrochloric acid to copper oxide powder, the solution formed is blue green. Write the new compound formed.
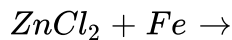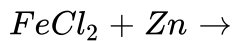
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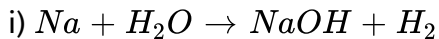**32.** Write the products of given reactions, if any. Give reason.

$FeCl_2 + Zn \rightarrow$

$ZnCl_2 + Fe \rightarrow$

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33. Write the products of given reactions, if any. Give reason.

$FeCl_2 + Zn \rightarrow$

$ZnCl_2 + Fe \rightarrow$

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34. Balance the following chemical equation.

i) $Na + H_2O \rightarrow NaOH + H_2$

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35. Balance the following chemical equation.

ii) $K_2CO_3 + HCl + KCl + H_2O + CO_2$
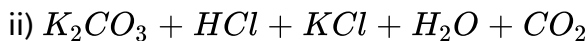
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36. Observe the following balance chemical equations and answer the given question: $C_3H_{8(8)} + 5O_{2(g)} \rightarrow 3CO_{2(g)} + 4H_2O_{(g)}$ How many molecules of oxygen are involved in this chemical reaction?
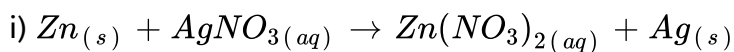
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37. Observe the following balance chemical equations and answer the given question: $C_3H_{8(8)} + 5O_{2(g)} \rightarrow 3CO_{2(g)} + 4H_2O_{(g)}$ How many moles of propane are required to get 20 moles of water?

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38. Why should we balance a chemical equation? Take any one chemical equation and explain the procedure of balancing it.
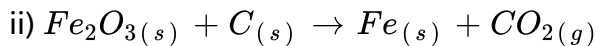
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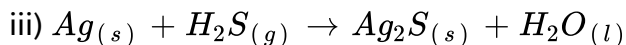**39.** Balance the following chemical equations.

i) $Zn_{(s)} + AgNO_{3(aq)} \rightarrow Zn(NO_3)_{2(aq)} + Ag_{(s)}$

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40. Balance the following chemical equations.

ii) $Fe_2O_{3(s)} + C_{(s)} \rightarrow Fe_{(s)} + CO_{2(g)}$

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41. Balance the following chemical equations.

iii) $Ag_{(s)} + H_2S_{(g)} \rightarrow Ag_2S_{(s)} + H_2O_{(l)}$

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42. Balance the following chemical equations.

iv) $Cu_{(s)} + O_{2(g)} \rightarrow CuO_{(s)}$
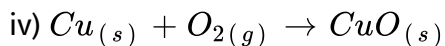
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43. Write the equation for the reaction of zinc with hydrochloric acid and balance the equation. Find, out the number of molecules of hydrogen gas

produced in this reaction, when 1 mole of HCl completely reacts at S.T.P.

[Gram molar volume is 22.4 liters at S.T.P., Avogadro's number is 6.023×10^{23}]

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44. How many grams of O_2 is required for combustion of 480 grams of m Mg? Find the mass of 'MgO' formed in this reaction.

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45. What are the reactants and products in a chemical reaction?

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46. What is meant by balanced equation ?

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47. What do you meant by skeleton equation ? Give one example.

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48. Why the white washig gives shiny finish to the walls?

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49. Balance the following equations.

1) $Na + O_2 \rightarrow Na_2O$
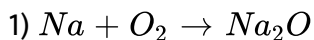
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50. Balance the following equations: $H_2O_2 \rightarrow H_2O + O_2$

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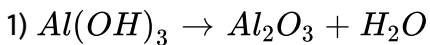
51. Balance the following equations.

3) $Mg(OH)_2 + HCl \rightarrow MgCl_2 + H_2O$

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52. Balance the following equations.

4) $Fe + O_2 \rightarrow Fe_2O_3$
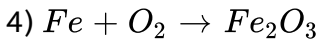
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53. Balance the follwowing equations.

1) $Al(OH)_3 \rightarrow Al_2O_3 + H_2O$

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54. Balance the follwowing equations.

2) $NH_3 + CuO \rightarrow Cu + N_2 + H_2O$
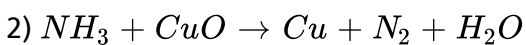
55. Balance the follwowing equations.

3) $Al_2(SO_4)_3 + NaOH \rightarrow Al(OH)_3 + Na_2SO_4$
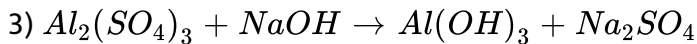
56. Balance the follwowing equations.

4) $HNO_3 + Ca(OH)_2 \rightarrow Ca(NO_3)_2 + H_2O$
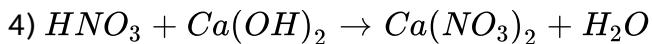
57. Balance the follwowing equations.

5) $NaOH + H_2SO_4 \rightarrow Na_2SO_4 + H_2O$
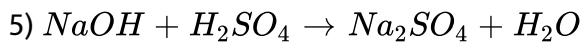
58. Balance the follwowing equations.

6) $BaCl_2 + H_2SO_4 \rightarrow BaSO_4 + HCl$
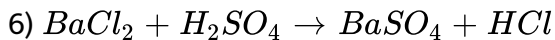
59. Balance the following chemical equation. $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$

60. What do you mean by precipitation reaction ?

61. Why does respiration considered as an exothermic reaction ? Explain.

62. How do you write a chemical equation?

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63. What are the important characteristics of chemical reactions?

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64. Write some chemical equation occurring in our daily life.

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65. How do you express the physical state of substances in the chemical equation ?

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66. $C_{(s)} + O_{2(g)} \rightarrow CO_{2(g)} + Q$ is ....
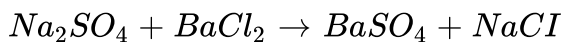
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67. $N_{2(g)} + O_{2(g)} \rightarrow 2NO_{2(g)} - Q$ is ....

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68. $2Cu + O_2 \rightarrow 2CuO$ what information do you get from this equation?

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69. Name the reactants and products in the following chemical equations.

$Na_2SO_4 + BaCl_2 \rightarrow BaSO_4 + NaCI$

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70. Balance the following chemical equation and follow the steps involved in balancing a chemical equation. $Cu_2S + O_2 \rightarrow Cu_2O + SO_2$
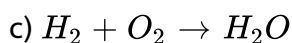
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71. If 40 gm of methane is burnt, then how much amount of CO_2 is released ?

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72. What does one mean by exothermic and endothermic reactions? Give examples.

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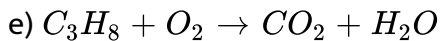**73.** Balance the following chemical equations.

c) $H_2 + O_2 \rightarrow H_2O$

74. Balance the following chemical equations.

e) $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$

75. Balance the chemical equation by including the physical states of the substances for the following reactions.

c) Zinc pieces react with dilute hydrochloric acid to liberate hydrogen gas and forms zinc chloride.

76. Potassium nitrate and sodium nitrate reacts separately with copper sulphate solution. Write balanced chemical equations for the above reactions.

77. Observe the following equation which shows the action of heat on Calcium Nitrate : $2Ca(NO_3)_2 \rightarrow 2CaO + 4NO_2 + O_2$ How many moles of NO_2 are formed when 2 moles of $Ca(NO_3)_2$ is decomposed?

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78. Observe the following equation which shows the action of heat on Calcium Nitrate : $2Ca(NO_3)_2 \rightarrow 2CaO + 4NO_2 + O_2$ What is the volume of NO_2 produced when 164 gm of $Ca(NO_3)_2$ is heated at constant temperature and pressure ?

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79. Observe the following equation which shows the action of heat on Calcium Nitrate : $2Ca(NO_3)_2 \rightarrow 2CaO + 4NO_2 + O_2$ Calculate the mass of Calcium Oxide formed when 82 gm of $Ca(NO_3)_2$ is heated.

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80. Observe the following equation which shows the action of heat on Calcium Nitrate : $2Ca(NO_3)_2 \rightarrow 2CaO + 4NO_2 + O_2$ What is the quantity of $Ca(NO_3)_2$, is required to produce 5 moles of gaseous products?

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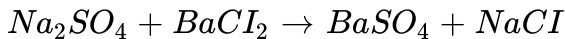81. Write the steps involved in the balancing a chemical reaction. Give an examples balancing the chemical equation.

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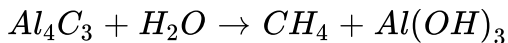82. How to make a chemical equation more informative?

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83. Balance the following chemical equations:

$$Na_2SO_4 + BaCl_2 \rightarrow BaSO_4 + NaCl$$

84. balance the following chemical equations:

$$Al_4C_3 + H_2O \rightarrow CH_4 + Al(OH)_3$$

85. balance the chemical equations: $Pb(NO_3)_2 \rightarrow PbO + NO_2 + O_2$

86. Balance the chemical equation $Fe_2O_3 + Al \rightarrow Fe + Al_2O_3$. Write the steps of balancing.

87. Write the balanced chemical equations for the following reactions.

$\text{Zinc} + \text{Silver nitrate} \rightarrow \text{Zinc nitrate} + \text{Silver}$
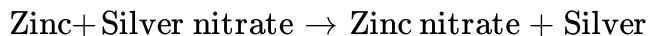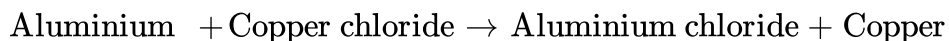
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88. Which gas is produced when zinc granules react with Hydrochloric acid ?

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89. Write the balanced chemical equations for the following reactions.

$\text{Aluminium} + \text{Copper chloride} \rightarrow \text{Aluminium chloride} + \text{Copper}$

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90. Write the balanced chemical equations for the following reactions.

Hydrogen + Chlorine → Hydrogen chloride

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91. Write the balanced chemical equations for the following reactions.
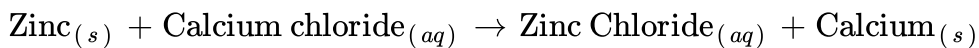
Ammonium nitrate → Nitrous oxide + Water

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92. Write the balanced chemical equations for the following and identify the type of reaction in each case.
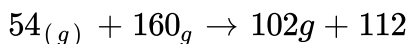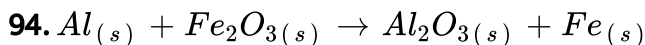
$Magnesium_{(s)} + Hydrochloric\ acid_{(aq)} \rightarrow Magnesium\ chloride_{(aq)} + Hyd$

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93. Write the balanced chemical equations for the following and identify the type of reaction in each case.

$Zinc_{(s)} + Calcium\ chloride_{(aq)} \rightarrow Zinc\ Chloride_{(aq)} + Calcium_{(s)}$

94. $Al_{(s)} + Fe_2O_{3(s)} \rightarrow Al_2O_{3(s)} + Fe_{(s)}$

$54_{(g)} + 160_g \rightarrow 102g + 112$

was given, Then, the amount of aluminium required to get 1120 grams of iron is

95. Calculate the volume, mass and number of molecules of hydrogen liberated when 230 g of sodium reacts with excess of water at STP (atomic masses of Na = 23U, O = 16 U and H = 1U).

96. Calculate the volume, mass and number of molecules of hydrogen liberated when 230 g of sodium reacts with excess of water at STP. (atomic masses of Na = 23U, O = 16U, and H = 1U)

the balanced equation for the above reaction is

$$2Na_{(s)} + 2H_2O_{(l)} \rightarrow 2NaOH_{(aq)} + H_{2(g)} + H_{2(g)} \uparrow$$
$$(2 \times 23)U + 2(2 \times 1 + 1 \times 16)U \rightarrow 2(23 + 16 + 1)U + (2 \times 1)U$$
$$46U + 36U \rightarrow 80U + 2U$$
$$\text{or } 46g + 36g \rightarrow 80g + 2g$$
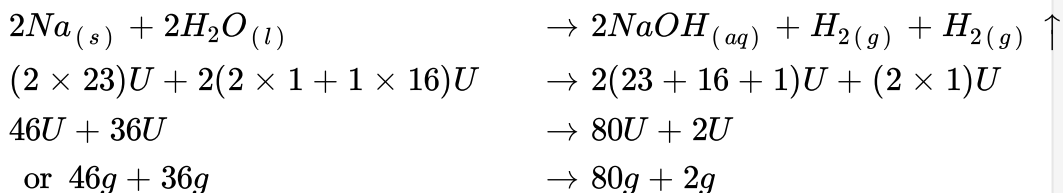
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97. Explain the reaction between sodium sulphate and Barium chloride.

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98. What are the precautions to be observed while mixing Zn and HCl?

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99. What are your observations after adding Zinc granules to HCI?

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100. Which gas is produced when zinc granules react with Hydrochloric acid ?

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101. How do you test the gas evolved while Zn granules are added to HCI ?

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102. Write the balanced chemical reaction for the following and identify the type of reaction in each case.

B)

$$\text{Zinc}_{(s)} + \text{Hydrochloric acid}_{(aq)} \rightarrow \text{Zinc chloride}_{(aq)} + \text{Hydrogen}_{(g)}$$
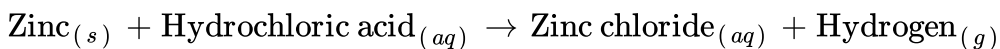
103. How do you test the nature of the solution formed by dissolving CaO in water? what is the nature of the solution?

104. How do you test the nature of the solution formed by dissolving CaO in water? what is the nature of the solution?

105. Water is added to quick lime in a beaker and you touch the bottom of the beaker what do you notice ? What is the reason ?

106. Write the reactants and products in the following reaction. Calcium oxide + Water $\rightarrow$ Calcium hydroxide.

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107. Calculate the volume and the number of molecules of CO_2 liberated at STP if 50 grams of $CaCO_3$ is treated with dilute hydrochloric acid which contains 7.3 grams of dissolved HCl gas.

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108. What is a limiting reagent?

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109. How much minimum volume of CO at STP is needed to react completely with 0.112 L of O_2 at 1.5 atm. Pressure and $127°C$ to give CO_2.

110. How much of lime (CaO) can be obtained by the calcinations of 300 g of lime stone?

111. What volume of H_2 at STP is required to reduce 0.795 g of CuO to give Cu and H_2O.

112. Calculate the volume of O_2 at STP required to completely burn 100 ml. of acetylene.

113. If the gas liberated in an experiment allows the burning splinter to continue burning more brightly in its presence, the gas is ....

    A. oxygen

    B. nitrogen

    C. hydrogen

    D. carbon dioxide

Answer:

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114. Number of moles of Oxygen needed to produce 4 moles of water on reacting with 4 moles of Hydrogen gas is ....

    A. 1 mole

    B. 2 moles

    C. 3 moles

D. 4 moles

Answer:

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115. The reaction that takes place when quicklime is added to water is

....

A. displacement reaction

B. gas liberating reaction

C. heat liberating reaction

D. combustion reaction

Answer:

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116. The balanced chemical equation, among the following, is ....

A. $NaOH + Zn \rightarrow NaZnO_2 + H_2$

B. $2NaOH + Zn \rightarrow Na_2ZnO_2 + H_2$

C. $2NaOH + 2Zn \rightarrow 2NaZnO_2 + H_2$

D. $NaOH + 2Zn \rightarrow NaZn_2o_2 + H_2$
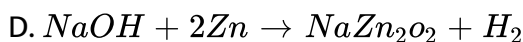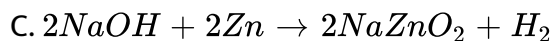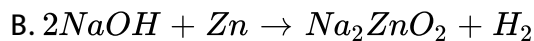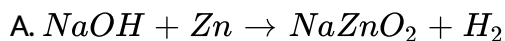
Answer:

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117. Solutions of Copper sulphate, Iron sulphate and Sodium sulphate are marked as X, Y and Z respectively. Few pieces of Aluminium are added to each solution. In which solutions, change will be observed?

A. X and Y

B. Y and Z

C. X and Z

D. X. Y and Z

Answer:

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118. The balanced chemical equation of $AgCl \xrightarrow{Sunlight} Ag + Cl_2$ is___

A. $AgCl_2 \xrightarrow{Sunlight} Ag + Cl_2$

B. $AgCl_2 \xrightarrow{Sunlight} 2Ag + Cl_2$

C. $2AgCl \xrightarrow{Sunlight} Ag + Cl_2$

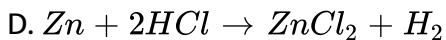D. $2AgCl \xrightarrow{Sunlight} 2Ag + Cl_2$
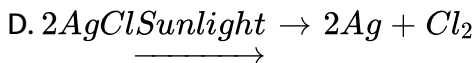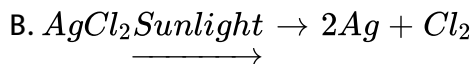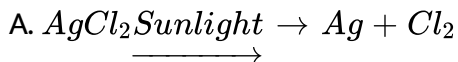
Answer:

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119. Which of the following is a balanced equation?

A. $Mg + O_2 \rightarrow MgO$

B. $C + O_2 \rightarrow CO$

C. $H_2 + O_2 \rightarrow H_2O$

D. $CaCo_3 \rightarrow Cao + co_2$
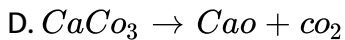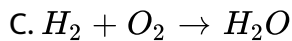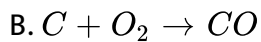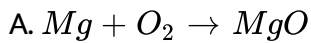
Answer:

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120. Which of the following is correct

A. Zn+ HCl $\rightarrow$ $ZnCl_2 + H_2$

B. 2 Zn+HCl $\rightarrow$ $2ZnCl_2 + H_2$

C. $Zn + 2HCl \rightarrow ZnCl_2 + H_2$

D. $Zn + 2HCl \rightarrow ZnCl_2 + H_2$
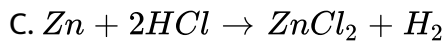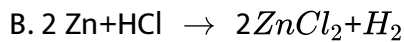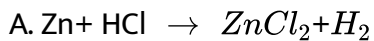
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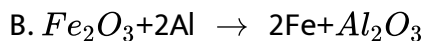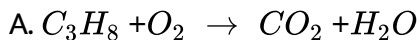**121.** Downward arrow in chemical equation indicates-

A. Direction

B. Gas

C. Precipitate

D. No reaction

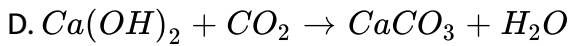**Answer:**

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122. which of the following is a skeleton reaction?

A. $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$

B. $Fe_2O_3 + 2Al \rightarrow 2Fe + Al_2O_3$

C. $AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$

D. $Ca(OH)_2 + CO_2 \rightarrow CaCO_3 + H_2O$
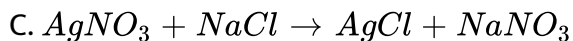
Answer:

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123. The chemical reaction in which energy is absorbed to form a new compound is called. ....

    A. Exothermic reaction

    B. Endothermic reaction

    C. Thermal reaction

    D. Photochemical reaction

Answer:

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124. The substances that are present on left side of a chemical equation are called ....

    A. Reactants

    B. Products

    C. Pracipitates

    D. Gases

Answer:

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125. Chemical equation is balanced according to the law of

    A. Constant proportions

    B. Conservation of mass

    C. Law of equality

    D. Law of balance

Answer:

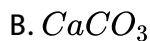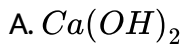**126.** $C+O_2 \rightarrow CO_2 + Q$. This is ___ reaction.

    A. Endothermic

    B. Chemical
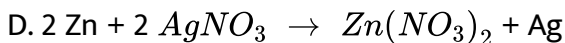
    C. Exothermic

    D. Photochemical

Answer:

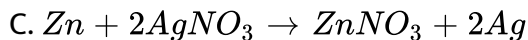
127. $Ca(OH)_2 + CO_2 \rightarrow CaCO_3 + H_2O$. In this reaction shiny finish to

walls is due to _____

A. $Ca(OH)_2$

B. $CaCO_3$

C. CO_2

D. H_2O

Answer:

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128. Balance the following chemical equations.

i) $Zn_{(s)} + AgNO_{3(aq)} \rightarrow Zn(NO_3)_{2(aq)} + Ag_{(s)}$

A. 2 Zn+ $AgNO_3 \rightarrow$ Zn(NO_3)_2`+2 Ag

B. Zn + $2AgNO_3 \rightarrow Zn(NO_3)_2$+2Ag

C. $Zn + 2AgNO_3 \rightarrow ZnNO_3 + 2Ag$

D. 2 Zn + 2 $AgNO_3 \rightarrow Zn(NO_3)_2$ + Ag
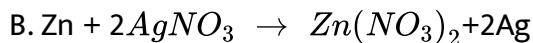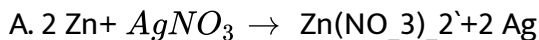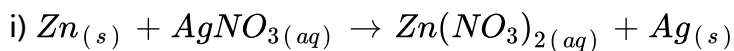
Answer:

129. If some amount of energy is released in a chemical reaction, then it is called ___ reaction.

A. Exothermic

B. Endothermic

C. Oxidation

D. Reduction

Answer:

130. Crackers burnt is a

A. Physical change

B. Chemical change

C. Both A & B

D. None

Answer:

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131. Skeleton equation means

A. A balanced chemical equation

B. An unbalanced chemical equation

C. Chemical equation without reactants

D. Chemical equation without products

Answer:

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132. How do you express the physical state of substances in the chemical equation ?

A. (l),(g) and (s)

B. (S),(g) and (l)

C. (S),(l) and (g)

D. (g),(l), and (s)

Answer:

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133. Aqueous means

A. Substance in alcohol

B. Substance in water

C. Substance in mercury

D. Substance in KOH

Answer:

---

134. Aqueous indicates in chemical equation as

A. (q)

B. Q

C. (aq)

D. (W)

Answer:

---

135. The reaction in which heat release

A. Exothermic reaction

B. Endothermic reaction

C. Both A & B

D. We can't say

Answer:

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136. 1 Gram molar volume of gas occupies

A. 2.24 litres

B. 22.4 litres

C. 4.22 litres

D. 42 .2 litres

Answer:

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137. STP conditions
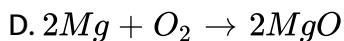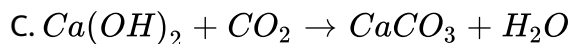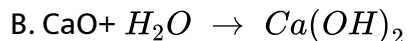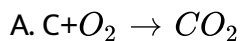
A. 273K,1bar

B. $30° C, 1bar$

C. 273 K, 76 bar

D. $30° C, 76bar$

Answer:

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138. Not an exothermic reaction

A. C+$O_2 \rightarrow CO_2$

B. CaO+ $H_2O \rightarrow Ca(OH)_2$

C. $Ca(OH)_2 + CO_2 \rightarrow CaCO_3 + H_2O$

D. $2Mg + O_2 \rightarrow 2MgO$

Answer:

---

139. The number of molecules present in 1g of hydrogen

A. 6×10^{23}

B. 3×10^{23}

C. 1×10^{23}

D. 2×10^{23}

Answer:

---

140. $xH_2 + yO_2 \rightarrow zH_2O$. The values of x, y, z are

A. X=1,y=1,z=1

B. X=2,y=1,z=2

C. X=2,y=2,z=2

D. X=2,y=2,z=1

Answer:

▶ **Watch Video Solution**

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141. Which one is necessary to test for a hydrogen gas ?

A. $Ca(OH)_2$

B. Battery

C. Burning stick

D. Litmus paper

Answer:

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142. A student added dilute HCL to a test tube containing Zn granules and made following observations. 1)The Zn surface became dull and black. 2) A gas evolved which burnt with a pop sound 3) the solution remained colorless
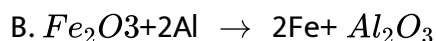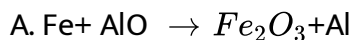
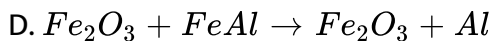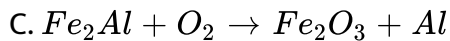A. 1,2

B. 2,3

C. 1,3

D. All the above

Answer:

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143. The chemical reaction is

A. Fe+ AlO $\rightarrow Fe_2O_3$+Al

B. Fe_2O3+2Al $\rightarrow$ 2Fe+ Al_2O_3

C. $Fe_2Al + O_2 \rightarrow Fe_2O_3 + Al$

D. $Fe_2O_3 + FeAl \rightarrow Fe_2O_3 + Al$

Answer:

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144. $Al_{(s)} + Fe_2O_{3(s)} \rightarrow Al_2O_{3(s)} + Fe_{(s)}$

$54_{(g)} + 160_g \rightarrow 102g + 112$
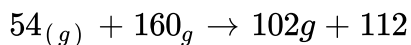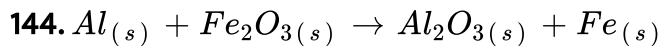
was given, Then, the amount of aluminium required to get 1120 grams of iron is

A. 540 gr

B. 540 kg

C. 112 gr

D. 1120 gr

Answer:

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145. 10 g of hydrogen contain .... molecules. If 2g of hydrogen contain 6.02×10^{23} molecules.

A. 3.01×10^{12}

B. 3.01×10^{24}

C. 6.02×10^{24}

D. 6.02×10^{23}

Answer: