



CHEMISTRY

BOOKS - UNITED BOOK HOUSE

PHYSICAL AND CHEMICAL PROPERTIES OF MATTER

Exercise

1. Saturated hydrocarbon is-

A. ethane

B. ethalene

C. acetylene

D. none of these

Answer:



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2. – $COOH$ group is called-

A. carboxyl

B. carbonyl

C. aldehyde

D. none of these

Answer:



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3. Which one is not a green house gas -

A. acetylene

B. ethylene

C. methane

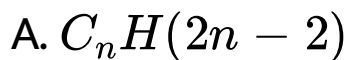
D. none of these

Answer:



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4. General formula of alkene is-





Answer:



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5. A covalent compound is-

A. CaO

B. KCl

C. H_2O

D. NaCl

Answer:



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6. In CH_4 valency of carbon is-

A. 1

B. 2

C. 3

D. 4

Answer:



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7. Arrange the increasing order of atomic radii

Li, Na, K.

A. KltLiltNa

B. LiltNaltK

C. NaltLiltK

D. KltNaltLi

Answer:



8. In the modern periodic table the coinage metals belong to group-

A. 1

B. 12

C. 11

D. 10

Answer:



9. Hydrogen is called-

A. rogue element

B. noble element

C. a cation

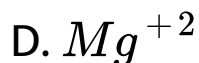
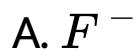
D. none of these

Answer:



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10. Choose the ion with the largest ion size-



Answer:



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11. Elements A, B, C and D have atomic number 9, 17, 19, 35 respectively. Choose the odd element-

A. A

B. B

C. C

D. D

Answer:



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12. Elements belonging to the same period have-

A. same valency electrons

B. same number of shells

C. same atomic size

D. same electron affinity

Answer:



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13. Which group of the periodic table contains maximum elements?

A. Third

B. Eighth

C. Seventeenth

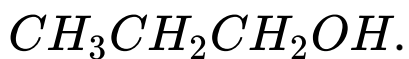
D. Eighteenth

Answer:



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14. Write down the IUPAC name of



- A. propan-1-ol
- B. n-propanol
- C. isopropanol
- D. propyl alcohol

Answer:



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15. Which does not react with metallic sodium-

A. Methyl alcohol

B. Ethyl alcohol

C. Acetic acid

D. Dimethyl ether

Answer:



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16. Organic compounds are-

A. high melting point

B. low melting point

C. soluble in water

D. conducts electricity in molten state

Answer:



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17. Percentage of silver in german silver is-

A. 55

B. 0

C. 0.5

D. 24

Answer:



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18. Which is extensively used for making aeroplane body-

A. Zn

B. Cu

C. Fe

D. Al

Answer:



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19. The shape of NaCl crystal is-

A. tetrahedral

B. octahedral

C. hexagonal

D. icosohedral

Answer:



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20. Long form of periodic table based on-

A. atomic number

B. atomic mass

C. number of neutrons

D. none of these

Answer:



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21. The number of elements discovered till date is-

A. 63

B. 109

C. 60

D. 118

Answer:



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22. At the time of publication of mendeleevs periodic tables. The number of inert gas elimants discovered are-

A. 1

B. 3

C. 3

D. 0

Answer:



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23. How many periods are there in the mendellevs periodic table-

A. 7

B. 10

C. 6

D. 14

Answer:



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24. Long from of periodic table based on-

A. atomic number

B. atomic mass

C. number of neutrons

D. none of these

Answer:



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25. The similarity between the subgroup A and B element is with respect to their-

A. Physical properties

B. Chemical properties

C. Valency

D. None

Answer:



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26. Why should copper sulphate solution not be kept in an iron vessel?



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27. All ores are minerals but all minerals are not ores'-Explain it.



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28. Will 'o' the wisp is seen in marshy land. Is it a supernatural phenomenon or something else? Explain ?



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29. Methan is a greenhouse gas-Explain.



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30. State whether pickle should be taken in aluminium foil or not . Give reasons for your answer.



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31. Why cannot fire caused due to burning of magnesium be extinguished by CO_2 gas?



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32. Why does sugar turn black when concentrated H_2SO_4 is added to it?



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33. Name two gases which react with each other to form a solid. Give the equation and write down the name of the solid formed.



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34. What is fuming nitric acid?



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35. What is aqua regia? Write down its uses.



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36. A solution of common salt in water can conduct electricity but that of sugar in water can not. Explain why.



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37. What is electroplanting.



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38. What is cation and anion.



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39. Define covalency. Explain with an example.



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40. Define electronegativity.



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41. State Mendeleev's periodic law.



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42. Define ionization energy.



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43. Define ionic bond with example.



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44. Define electrolytes.



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45. What is Electroplating.



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46. What is oleum?



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47. How NH_3 is detected using Nessler's reagent?



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48. Describe the manufacture of nitric acid by Ostwald's process.



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49. Justify the position of hydrogen in the periodic table on the basis of its electronic configuration.



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50. Discuss the drawbacks of Mendeleev's periodic table.



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51. Identify the following as electrovalent or covalent compounds :



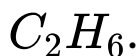
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52. Identify the following as electrovalent or covalent compounds :



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53. Identify the following as electrovalent or covalent compounds :



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54. Point out the differences between the properties of NaCl and Cl_4 .



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55. Give the differences between metallic conductor and electrolyte.



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56. How is impure copper refined by the process of electrolyte.



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57. Prove that hydrochloric acid contains hydrogen and chlorine.



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58. Which one is used to identify nitrogen in Lassign test -



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59. Show that nitric acid is an oxidising acid.



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60. Prove that nitric acid contains nitrogen, oxygen and hydrogen.



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61. Write down the principle of manufacture of sulphuric acid by contact process.



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62. State with equations, what happens when concentrate HNO_3 is added drop. By drop to hot pumice stone.



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63. State with equations, what happens when Hot and conc. HNO_3 is kept in an iron vessel.



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64. Give differences between organic and inorganic compounds.



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65. Write down the addition reactions of acetylene with hydrogen and bromine.



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66. Write down the substitution reaction of methane with chlorine.



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67. What is teflon? Write down the name of its monomer. Write down its uses.



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68. Why alkali metals are placed in group 1 A, why they are named so?



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69. Describe with reason the variation of ionic radii along a period.



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70. What is atomic radius mention they type of atomic radius?



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71. Describe with reason how oxidising and reducing properties of element vary along a period?



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72. What are transuranic elements give example.



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Example

1. Write down the formula of bauxite.



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2. An atom has electronic configuration 2, 8, 7.

What is the atomic number of this element?



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3. Write down the full form of DNA.



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4. Give an example of an artificial polymer.

Name its monomer.



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5. Name the gas used for welding.



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6. Name the simplest hydrocarbon with structural formula.



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7. Which gas is responsible for the formation of will the wisp?



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8. Name a metal which does not react with alkalis.



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9. Which metal is used in the preparation of grignard's reagent?



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