

CHEMISTRY

BOOKS - JNAN PUBLICATION

CHEMICAL CALCULATIONS

Example

1. Write an example of direct combination rection.



2. Write an example of decomposition reaction.



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3. Write an example of substitution reaction.



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4. Write an example of acid-base reaction.



5. Write an example of oxidation reaction.



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6. Write an example of addition reaction.



7. Write an example of rearrangement reaction.



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8. Write the equation of rusting of iron.



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9. Who put forward the law of conservation of mass?



10. What is the value of Avogadro's Number?



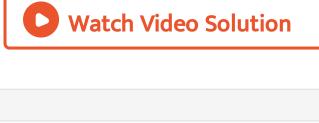
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11. Define Solid.



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12. Define liquid



13. Define Gas



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14. Define vapur density



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15. What is the relative density?



16. Write the relation between vapour density and mass of volume of hydrogen.



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17. What is gram atomic mass?



18. What do you mean by gram molecular weight?



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19. Define-atomic mass.



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20. Write the value of molecular weight of oxygen.



21. Write the value of gram atomic mass of copper.



22. Write the value of gram atomic mass of aluminium.



23. Define-atomicity.



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24. Write three types of atomicity.



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25. Write two examples of monoatomic molecules.



26. Write two examples of diatomic molecules.



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27. Write two examples of polyatomic molecules.



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28. Define-molecular weight





29. What is atom?



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30. What do you mean by gram molecular weight?



31. Write one example of isobar.



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32. Write the value of gram molecular weight of hydrogen gas.



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33. Write the value of gram molecular weight of Nitrogen gas.



34. Write the value of gram molecular weight of chlorine gas.



35. Write the value of gram molecular weight of ammonia gas.



36. Write the value of gram molecular weight of carbon dioxide gas.



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37. Write the value of gram molecular weight of sulphur dioxide gas.



38. Write the value of gram molecular weight of nitrogen dioxide gas.



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39. What is mole?



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40. Define a catalyst.



41. What do you mean by double displacement reactions?



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42. What is the value of atomic mass of hydrogen gas?



43. What is the value of atomic mass of carbon?



44. What is the value of atomic mass of potassium?



45. What is the value of atomic mass of zinc?

46. Write the relation between molecular weight and vapour density.



47. What is meant by 'The vapour density of SO_2 is 32'?



48. Prove, vapour density = $1/2 \times Molecular$ Weight



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49. Why is a chemical equation called as balanced equation?



50. Write the significance of a chemical equation.



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51. Define-chemical equation.



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52. Write an example of balanced equation.



53. What is the substitution reaction? Write an example.



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54. Short Note-Direct Combinatin reaction



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55. Short Note- Double decomposition.



56. Short Noted-Decomposition reaction.



57. Short Note-Addition reaction



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58. Write the limitatins of chemical equation.



59. Short note- Chemical equations or equations for chemical reactions.



60. Write the characteristics of a word equation. In a word equation-



61. Short Note-Unbalanced chemical equation



62. Short Note-Balanced equation.



63. Short Note-Moleculor model of solid



64. Short Note-Molecular model of gas.



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65. Writes the properties of gases



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66. Discuss about the mass concept of mole.



67. Discuss about the particle concept of mole



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68. Discuss about the volume concept of mole



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69. Write two differences between displacement reaction and double displacement reactions.



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70. Why do we store silver chloride in dark coloured bottles?



71. A brown substance 'X' on heating in air forms a substance 'Y' when hydrogen gas is passed over heated 'Y', it again changes back into 'X'



72. The marble statutes often slowly get corroded when kept in open for a long time. Assign a suitable explanation.



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73. Calculate the weight of copper sulphate formed when 128 gm of copper are added to concentrated sulphuric acid. What is the

volume of sulphur dioxide liberated at the same time?



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74. How much mercuric sulphide will be produced when 2.004 gm of meercury reacts completely with 0.32 gm of sulphur.



75. Calculate the number of moles of 16 gm of Hydrogen gas.



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76. Calculate the number of moles of 32 gm of oxygen gas.



77. Calculate the number of moles in 1.5 gm of amonia.

