

CHEMISTRY

BOOKS - JNAN PUBLICATION

CHEMICAL CALCULATIONS

Example

1. Write an example of direct combination reaction.



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2. Write an example of decomposition reaction.



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3. Write an example of substitution reaction.



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4. Write an example of acid-base reaction.



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5. Write an example of oxidation reaction.



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6. Write an example of addition reaction.



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7. Write an example of rearrangement reaction.



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8. Write the equation of rusting of iron.



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9. Who put forward the law of conservation of mass?



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10. What is the value of Avogadro's Number?



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11. Define Solid.



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12. Define liquid



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13. Define Gas



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14. Define vapour density



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15. What is the relative density?



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16. Write the relation between vapour density and mass of volume of hydrogen.



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17. What is gram atomic mass?



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18. What do you mean by gram molecular weight?



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19. Define-atomic mass.



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20. Write the value of molecular weight of oxygen.



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21. Write the value of gram atomic mass of copper.



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22. Write the value of gram atomic mass of aluminium.



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23. Define-atomicity.



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24. Write three types of atomicity.



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25. Write two examples of monoatomic molecules.



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26. Write two examples of diatomic molecules.



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27. Write two examples of polyatomic molecules.



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28. Define-molecular weight





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29. What is atom?



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30. What do you mean by gram molecular weight?



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31. Write one example of isobar.



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32. Write the value of gram molecular weight of hydrogen gas.



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33. Write the value of gram molecular weight of Nitrogen gas.



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34. Write the value of gram molecular weight of chlorine gas.



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35. Write the value of gram molecular weight of ammonia gas.



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36. Write the value of gram molecular weight of carbon dioxide gas.



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37. Write the value of gram molecular weight of sulphur dioxide gas.



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38. Write the value of gram molecular weight of nitrogen dioxide gas.



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39. What is mole?



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40. Define a catalyst.



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41. What do you mean by double displacement reactions?



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42. What is the value of atomic mass of hydrogen gas?



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43. What is the value of atomic mass of carbon?



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44. What is the value of atomic mass of potassium?



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45. What is the value of atomic mass of zinc?



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46. Write the relation between molecular weight and vapour density.



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47. What is meant by 'The vapour density of SO_2 is 32'?



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48. Prove, vapour density = $\frac{1}{2} \times$ Molecular Weight



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49. Why is a chemical equation called as balanced equation?



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50. Write the significance of a chemical equation.



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51. Define-chemical equation.



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52. Write an example of balanced equation.



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53. What is the substitution reaction? Write an example.



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54. Short Note- Direct Combinatin reaction



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55. Short Note- Double decomposition.



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56. Short Noted-Decomposition reaction.



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57. Short Note-Addition reaction



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58. Write the limitatins of chemical equation.



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59. Short note- Chemical equations or equations for chemical reactions.



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60. Write the characteristics of a word equation. In a word equation-



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61. Short Note-Unbalanced chemical equation



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62. Short Note-Balanced equation.



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63. Short Note-Molecular model of solid



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64. Short Note-Molecular model of gas.



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65. Writes the properties of gases



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66. Discuss about the mass concept of mole.



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67. Discuss about the particle concept of mole



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68. Discuss about the volume concept of mole



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69. Write two differences between displacement reaction and double displacement reactions.





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70. Why do we store silver chloride in dark coloured bottles?



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71. A brown substance 'X' on heating in air forms a substance 'Y' when hydrogen gas is passed over heated 'Y', it again changes back into 'X'



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72. The marble statues often slowly get corroded when kept in open for a long time. Assign a suitable explanation.



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73. Calculate the weight of copper sulphate formed when 128 gm of copper are added to concentrated sulphuric acid. What is the

volume of sulphur dioxide liberated at the same time?



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74. How much mercuric sulphide will be produced when 2.004 gm of meercury reacts completely with 0.32 gm of sulphur.



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75. Calculate the number of moles of 16 gm of Hydrogen gas.



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76. Calculate the number of moles of 32 gm of oxygen gas.



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77. Calculate the number of moles in 1.5 gm of ammonia.



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