



CHEMISTRY

BOOKS - U-LIKE CHEMISTRY (HINGLISH)

SAMPLE QUESTIONS PAPER

Section A Read The Given Passage And Answer The

1. A lead storage battery is the most important type of secondary cell having a lead anode and a grid of lead packed with PbO_2 as cathode. A 38%

solution of sulphuric acid is used as electrolyte. (Density = 1.294 g mL^{-1}) The battery holds 35L of the acid. During the discharge of the battery, the density of H_2SO_4 falls to 1.138 g mL^{-1} . (20% H_2SO_4 by mass)

Write the reaction taking place at the cathode when the battery is in use.



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2. A lead storage battery is the most important type of secondary cell having a lead anode and a grid of lead packed with PbO_2 as cathode. A 38%

solution of sulphuric acid is used as electrolyte. (Density = 1.294 g mL^{-1}) The battery holds 35L of the acid. During the discharge of the battery, the density of H_2SO_4 falls to 1.138 g mL^{-1} . (20% H_2SO_4 by mass)

How much electricity in terms of faraday is required to carry out the reduction of one mole of PbO_2 ?



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3. A lead storage battery is the most important type of secondary cell having a lead anode and a

grid of lead packed with PbO_2 as cathode. A 38% solution of sulphuric acid is used as electrolyte. (Density = 1.294 g mL^{-1}) The battery holds 35L of the acid. During the discharge of the battery, the density of H_2SO_4 falls to 1.138 g mL^{-1} . (20% H_2SO_4 by mass)

What is the molarity of sulphuric acid before discharge ?



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4. A lead storage battery is the most important type of secondary cell having a lead anode and a

grid of lead packed with PbO_2 as cathode. A 38% solution of sulphuric acid is used as electrolyte.

(Density = 1.294 g mL^{-1}) The battery holds 35L of the acid. During the discharge of the battery, the density of H_2SO_4 falls to 1.138 g mL^{-1} . (20% H_2SO_4 by mass)

Lead storage battery is considered a secondary cell. Why ? Write the product battery is considered a secondary cell. Why?



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5. A lead storage battery is the most important type of secondary cell having a lead anode and a grid of lead packed with PbO_2 as cathode. A 38% solution of sulphuric acid is used as electrolyte. (Density = 1.294 g mL^{-1}) The battery holds 35L of the acid. During the discharge of the battery, the density of H_2SO_4 falls to 1.138 g mL^{-1} . (20% H_2SO_4 by mass)

write the product of electrolysis when dilute sulphuric acid is electrolysed using platinum electrodes .



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Section A Questions Number One Word Answers

1. Name the substance used as depressant in the sepressant in the saparation of two sulphide ores in Froth floatation method .



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2. Name the unit formed by the attachment of a base to 1' Position of sugar in a nucleoside .



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3. Name the species formed when an aqueous solution of amino acid is dissolved in water .



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4. What type of reaction occurs in the formation of nylon 6, 6 polymer ?



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5. Which of the following compounds would undergo cannizzaro reaction ?

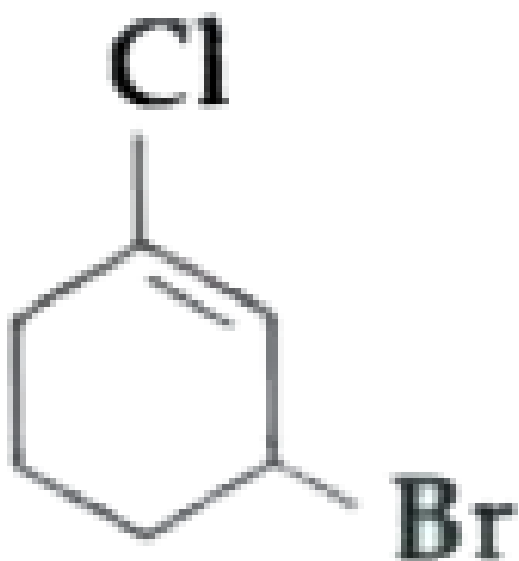
Benzaldehyde, Cyclohexanone , 2-mehtylpentanal.



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Section A Multiple Choice Questions

1. The IUPAC name of the compound shown below is :



- A. 2-bromo-6-chlorocyclohex-1-ene
- B. 6-bromo-2-chlorocyclohexene
- C. 3-bromo-1-chlorocyclohexene
- D. 1-bromo-3-chlorocyclohexene

Answer: C

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2. When one mole of $CoCl_3 \cdot 5NH_3$ was treated with excess of silver nitrate solution, 2 mol of $AgCl$ was precipitated. The formula of the compound is :

- A. $[Co(NH_3)_5Cl_2]Cl$
- B. $[Co(NH_3)_5Cl]Cl_2$
- C. $[Co(NH_3)_4Cl_2](NH_3)Cl$
- D. $[Co(NH_3)_3](NH_3)_2$

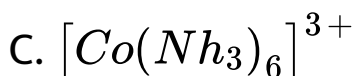
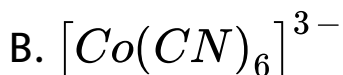
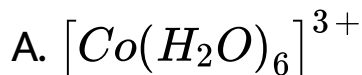
Answer: B



3. The absorption maxima of several octahedral complex ions are as follows:

S.No.	Compound	$\lambda_{\text{max}}/\text{nm}$
1.	$[\text{Co}(\text{NH}_3)_6]^{3+}$	475
2.	$[\text{Co}(\text{CN})_6]^{3-}$	310
3.	$[\text{Co}(\text{H}_2\text{O})_6]^{3+}$	490

The crystal field splitting is maximum for :



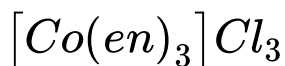
D. All the complex ions have the same splitting

Answer: B



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4. Predict the number of ions produced per formula unit in an aqueous solution of



A. 4

B. 3

C. 6

D. 2

Answer: A



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5. The incorrect statement about LDP is :

- A. It is obtained through the free radical addition of ethene .
- B. It consists of liner molecules .
- C. It is obtained by the H-atom abstraction
- D. Peroxide is used as an initialtor.

Answer: B



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Section A Assertion Reason Questions

1. Assertion (A) : The two strands in double strand helix structure of DNA are complementary to each other

Reason (R) : Disulphide bonds are formed between specific pairs of bases.

A. Both Assertion (a) And Reason (R) are correct Statement . And Reason (R) is the correct explanation of the Assertion (A)

B. Both Assertion (A) and Reason (R) are correct statement , bbut Reason (R) is not the correct expanation of the Assertion (A)

C. Assertion (A) is correct , but Reason (R) is correct statement .

D. Assertion (A) is incorrect but Reason (R) Is correct statement.

Answer: C



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2. Assertion (A) : Glucose reacts with hydroxylamine to form an axime and also adds a molecule of hydrogen cyanide to give cyanohydrin .

Reason (R) : The carbonyl group is present in the open chain structure of glucose .

A. Both Assertion (a) And Reason (R) are correct Statement . And Reason (R) is the

correct explanation of the Assertion (A)

B. Both Assertion (A) and Reason (R) are

correct statement , bbut Reason (R) is not

the correct expansion of the Assertion (A)

C. Assertion (A) is correct , but Reason (R) is

correct statement .

D. Assertion (A) is incorrect but Reason (R) Is

correct statement.

Answer: A

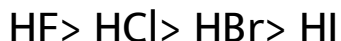


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3. Assertion (A) : The acidic strength of halogen acids varies in the order



Reason (R) : The bond dissociation enthalpy of halogen acids decreases in the order :



A. Both Assertion (a) And Reason (R) are correct Statement . And Reason (R) is the correct explanation of the Assertion (A)

B. Both Assertion (A) and Reason (R) are correct statement , bbut Reason (R) is not

the correct expansion of the Assertion (A)

C. Assertion (A) is correct , but Reason (R) is correct statement .

D. Assertion (A) is incorrect but Reason (R) is correct statement.

Answer: D



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4. Assertion (A) : C_2H_5OH is a weaker base than phenol but is a stronger nucleophile than phenol

Reason (R) : In phenol the lone pair of electrons on oxygen is withdrawn towards the ring due to resonance .

A. Both Assertion (a) And Reason (R) are correct Statement . And Reason (R) is the correct explanation of the Assertion (A)

B. Both Assertion (A) and Reason (R) are correct statement , bbut Reason (R) is not the correct expansion of the Assertion (A)

C. Assertion (A) is correct , but Reason (R) is correct statement .

D. Assertion (A) is incorrect but Reason (R) Is correct statement.

Answer: D



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5. Assertion (A) : Aryl halides undergo nucleophilic substitution reaction with ease.

The carbon halogen bond in aryl halides has partial double bond character.

A. Both Assertion (A) and Reason (R) are correct statements. Reason (R) is the correct explanation of the Assertion (A).

B. Both Assertion (A) and Reason (R) are correct statements, but Reason (R) is not the correct explanation of the Assertion (A).

C. Assertion (A) is correct, but Reason (R) is not a correct statement.

D. Assertion (A) is incorrect but Reason (R) Is correct statement.

Answer: D



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Section B

1. Calculate the number of lone pairs on central atom in the following molecule and predict the geometry





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2. The rate of a reaction depends upon the temperature and is quantitatively expressed as

$$K = Ae^{-E_a/RT}$$

(i) If a graph between $\log K$ and $1/t$ write the expression for the slope of the reaction ?

(ii) If at under different condition E_{a1} and E_{a2} are the activation energy of two reactions .

If $E_{a1} = 40 \frac{j}{m} \text{ol}$ and $E_{a2} = 80 \frac{j}{m} \text{ol}$. Which of the

two has larger value of the rate constasnt ?



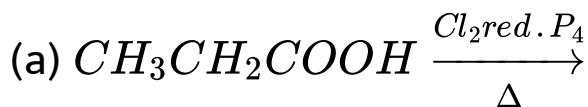
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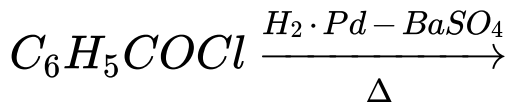
3. The experimentally determined molar mass for what type of substances is always lower than the true value when water is used as solvent. Explain. Give one example of such a substance and one example of a substance which does not show a large variation from the true value.



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4. Write structure of the product formed :





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5. Draw one of the geometrical isomers of the complex $[Pt(en)_2Cl_2]^{2+}$ which is optically inactive . Also write the name of this entity according to the IUPAC nomenclature .



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6. Discuss the bonding in the coordination entity $[CO(Nh_3)_6]^{3+}$ on the basis of valence bond

theory . Also , comment on the geometry and spin of the given entity .



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7. What is meant by vapour phase refining ? Write any one example of the process which illustrates this technique, giving the chemical equation involved .



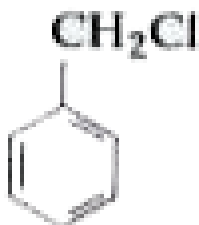
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8. Write and explain the reaction involved in the extraction of gold .



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9. Which one of the following compound will undergo hydrolysis at a faster rate by S_N1 mechanism ? Justify .



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Section C

1. Calculate the freezing point of a solution containing 0.5g KCl (Molar mass = 74.5g/mol) dissolved in 100 g water , assuming KCl to be 92% ionised . K_f of water = 1.86Kkg/mol.



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2. For the reaction $A + B \rightarrow \text{product}$, the following initial rates were obtained at various

given initial concentrations:

S.No.	[A] mol/L	[B] mol/L	Initial rate M/s
1.	0.1	0.1	0.05
2.	0.2	0.1	0.10
3.	0.1	0.2	0.05

Determine the half - life period.



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3. A first order reaction is 50% complete in 50 minutes at 300K and the same reaction is again 50% complete in 25 minutes at 350K . Calculate activation energy of the reaction .



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Section C Answer The Following Questions

1. which of the following electrolytes is most effective for the coagulation of AgI / Ag^+ Sol?

$MgCl_2$, K_2SO_4 , $K_4[Fe(CN)_6]$



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2. What happens when a freshly precipitated $Fe(OH)_3$ is shaken with a little amount of dilute solution of $FeCl_3$?



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Section C Account For The Following

1. Out of sulphur sol and proteins , which one forms macromolecular colloids?



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2. Moist SO_2 decolourises $KMnO_4$ solution .



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3. In general interhalogen compound are more reactive than halogens (expect fluorine)



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4. Ozone acts as a powerful oxidising agent .



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5. identify the product formed when propan -1-ol is treated with conc. H_2SO_4 at 413K. Write the mechanism involved for the above reaction .



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6. Give chemical tests to distinguish between the following pairs of compounds :

(i) Ethanal and propanone .

(ii) pentan -2-one and pentan -3-one .

Arrange the following compounds in increasing order of their acid strength : Benzoic acid , 4-Nitrobenzoic acid , 3, 4- Dinitrobenzoic acid , 4-Methoxybenzoic acid .



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7. Compare the reactivity of benzaldehyde and ethanal towards nucleophilic addition reactions.

Write cross aldol condensation product between benzaldehyde and ethanal.



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8. Define and write an example for the following :

(a) Broad spectrum antibiotics .

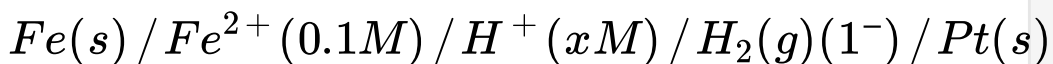
(b) Analgesics.



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Section D

1. The e.m.f of the following cell at 298K is 0.1745V,



Calculate the H^+ ions concentration of the solution at the electrode where hydrogen is being produced .

(b) Aqueous solution of copper sulphate and silver nitrate are electrolysed by 1 ampere current for 10 minutes in separate electrolytic cells. Will the mass of copper and silver deposited on the cathode be same or different ? explain your answer.



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2. (a) Calculate the degree of dissociation 'A' and 'B' are dilute . The limiting molar conductivity of 'B' increases to a smaller extent while that 'A' increase to a much larger extent comparatively .

Which of the two is a strong electrolyte ? Justify your answer .

(b) Solution of two electrolytes 'A' and 'B' are dilute . The limiting molar conductivity of 'B' increases to a smaller extent while that of 'A' increase to a much larger comparatively .

Which of the two is a strong electrolyte ? justify your answer.



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3. An organic compound 'A' with molecular formula C_7H_7NO reacts with $Br_2/aqKOH$ to give compound 'B', which upon reaction with $NaNO_2$ and HCl at $0^\circ C$ gives 'C'. Compound 'C' on heating with CH_3CH_2OH gives a hydrocarbon 'D'. Compound 'B' on further reaction with Br_2 water gives white precipitate of compound 'E'. Identify the compound A, B, C, D and E, also justify

your answer by giving relevant chemical equations.



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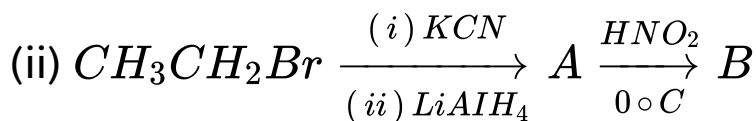
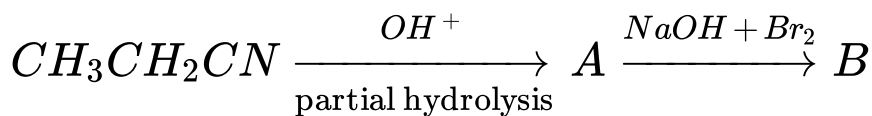
4. (a) How will you convert:

(i) Aniline into fluorobenzene.

(ii) Benzamide into Benzylamine .

(iii) Ethanamine to N, N-Diethylethanamine.

(b) Write the structures of A and B in the following :





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5. (a) when a chromite ore (A) is fused with an aqueous solution of sodium carbonate in free excess of air, a yellow solution of compound (B) is obtained. This solution is filtered and acidified with sulphuric acid to form compound (C). Compound (C) on treatment with solution of KCl gives orange crystals of compound (D). Write the chemical and actinoids:

(i) Greater range of oxidation states of actinoid as compared to lanthanoids.

(ii) Greater actinoid contraction as compared to lanthanoid contraction .

Lower ionisation enthalpy of early actinoids as compared to the early lanthanoids.



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6. what happens when :

(i) Manganate ions (MnO_4^{2-}) undergoes disproportionation reaction in acidic medium ?

(ii) Lanthanum is heated with sulphur ?

(b) explain the following trends in the properties of the members of the first series of transition

elements:

(i) $E^\circ (M^{2+} / M)$ value for copper is positive (+ 0.34V) in contrast to the other members of the series .

(ii) Cr^{2+} is reducing while Mn^{3+} is oxidising power in the series increases in the order $VO_2^{(+)}$



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