



CHEMISTRY

NCERT - FULL MARKS

CHEMISTRY(TAMIL)

GROUP 2 S - BLOCK ELEMENTS

Question A Choose The Best Answer

1. Among the following, which is known as `alkaline earth metal..

A. Sodium

B. Calcium

C. Lithium

D. Potassium

Answer:



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2. Alkaline earth metals are

A. monovalent

B. trivalent

C. divalent

D. zerovalent

Answer:



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3. Among alkaline earth metals _____ is having the highest ionization energy.

A. Beryllium

B. magnesium

C. Calcium

D. Barium

Answer:



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4. The colour given by barium in flame is

A. Brick red

B. Apple Green

C. Red

D. Blue

Answer:



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5. The third most abundant dissolved ion in the ocean is

A. Beryllium

B. Barium

C. Calcium

D. Magnesium

Answer:



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6. Quick lime is

A. Calcium oxide

B. Calcium hydroxide

C. Calcium nitrate

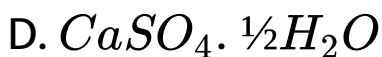
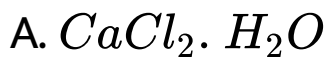
D. Calcium sulphate

Answer:



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7. The formula of bleaching powder is

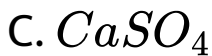


Answer:



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8. Plaster of Paris is _____



Answer:



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9. The compound used in making moulds for statues is

- A. Epsom salt
- B. Calcium sulphide
- C. Plaster of paris
- D. Gypsum

Answer:



10. The element used in pyrotechnics is

A. Magnesium

B. Barium

C. Calcium

D. Beryllium

Answer:



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11. The general electronic configuration of alkaline earth metals is _____ .



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Question B Fill In The Blanks

1. The ionic radius _____ on moving down the group 2.



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2. In flame, calcium gives _____ colour.



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3. Beryllium resembles more with an element in 13th group _____.



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4. Magnesium comes from the name of the mineral _____.





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5. _____ is present chlorophyll.



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6. Magnesium is prepared by the electrolysis of fused _____.



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7. With air, Magnesium forms _____ and _____.



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8. The formula of epsom salt is _____.



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9. Epsom salt is used as _____.



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Question C Match The Following

1. Match the following

1. Magnesite $CaSO_4 \cdot 2H_2O$
2. Dolomite $MgCl_2 \cdot KCl \cdot 6H_2O$
3. Epsom salt $MgCO_3$
4. Carnallite $MgCO_3 \cdot CaCO_3$
5. Gypsum $MgSO_4 \cdot 7H_2O$



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Question Problem

1. An element occupies group number 2 and period number 3. This element reacts with oxygen and nitrogen to form compound A and B. It is a strong electropositive metal so it displaces Ag from $AgNO_3$ solution. With concentrated nitric acid, it forms compound C. Identify the element, compound A, B and C.



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Question D Write In One Or Two Sentence

1. Why the oxides of Group 2 metals have high melting points?



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2. Why there is increase in the ionisation potential for forming M^{3+} is not very much greater than M^{+} ion for group metals?



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3. Why the ionization potential of M^{2+} is not very much greater than M^+ ?



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4. List the carbonates and hydroxide of alkaline earth metals in order of their increasing stability and their solution.



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5. Why do beryllium halides fume in air?



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6. Why group 2 elements are harder than alkali metals?



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7. Why do beryllium halides fume in air?



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8. Why are monoxides of alkaline earth metals are very stable?



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9. Why the oxides of Group 2 metals have high melting points?



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Question D Explain Briefly On The Following

1. What are alkaline earth metals? Why are they called so?



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2. In what respects Be and Mg differ from all the other metals of group 2.



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3. How can you explain the anomalous behaviour of beryllium.



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4. How does magnesium occur in nature? How is the metal extracted from its Ore?



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5. Why the first ionization energy of alkaline earth metals higher than that of 1st group.



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6. Mention the uses of plaster of paris



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7. How is plaster of paris prepared ?



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8. How is $MgSO_4$ prepared?



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9. Mention the uses of Magnesium?



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Question

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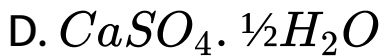
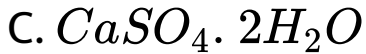
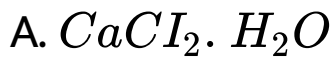
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- C. Calcium nitrate
- D. Calcium sulphate

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7. The formula of bleaching powder is



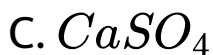
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13. In flame, calcium gives _____ colour.



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14. Beryllium resembles more with an element in 13th group _____.



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15. Magnesium comes from the name of the mineral _____.



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16. _____ is present chlorophyll.



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17. Magnesium is prepared by the electrolysis of fused _____.



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19. The formula of epsom salt is _____.



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20. Epsom salt is used as _____.



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21. Match the following

- | | |
|---------------|--------------------------------|
| 1. Magnesite | $CaSO_4 \cdot 2H_2O$ |
| 2. Dolomite | $MgCl_2 \cdot KCl \cdot 6H_2O$ |
| 3. Epsom salt | $MgCO_3$ |
| 4. Carnallite | $MgCO_3 \cdot CaCO_3$ |
| 5. Gypsum | $MgSO_4 \cdot 7H_2O$ |



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22. An element occupies group number 2 and period number 3. This element reacts with oxygen and nitrogen to form compound A and B. It is a strong electropositive metal so it displaces Ag from $AgNO_3$ solution. With concentrated nitric acid, it forms compound C. Identify the element, compound A, B and C.



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23. Why the oxides of Group 2 metals have high melting points?



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24. Why there is increase in the ionisation potential for forming M^{3+} is not very much greater than M^{+} ion for group metals?



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25. Why there is increase in the ionisation potential for forming M^{3+} is not very much greater than M^{+} ion for group metals?



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26. List the carbonates and hydroxide of alkaline earth metals in order of their increasing stability and their solution.



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27. Why do beryllium halides fume in air?



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28. Why group 2 elements are harder than alkali metals?



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29. Beryllium halides are covalent whereas magnesium halides are ionic. Why?





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30. Why are monoxides of alkaline earth metals are very stable?



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31. The basic strength of the oxides of group 2 elements increases from Be to Ba. Why?



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32. Alkaline earth metals are



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33. In what respects Be and Mg differ from all the other metals of group 2.



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34. How can you explain the anomalous behaviour of beryllium.



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35. How does magnesium occur in nature?

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36. Why the first ionization energy of alkaline earth metals higher than that of 1st group.



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39. How is $MgSO_4$ prepared?



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40. Mention the uses of Magnesium?



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