CHEMISTRY

BOOKS - VGS PUBLICATION-BRILLIANT

MODEL PAPER 12

Section A

1. What are S.T.P conditions?

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2. Calculate the weight of 0.1 mole of sodium carbonate.

3. What are the sign conventions of the work done on the system and work done by the system?

4. What are the $'\Delta H'$ sign conventions for exothermic and endothermic reactions?

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5. No work is done on the system, but heat (q) is taken out from the system by the surroundings. What type of wall does the system have?

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6. Work is done by the system and heat (q) is supplied to the system. What type of system would it be?

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7. All Lewis acids are not Bronsted acids. Why?

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8. SiO_2 is a solid while CO_2 is a gas - explain.

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9. What happens when the following are heated ?

$CaCO_3$ and SiO_2

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10. Which oxides cause acid rain ?

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11. Name the green house gases.

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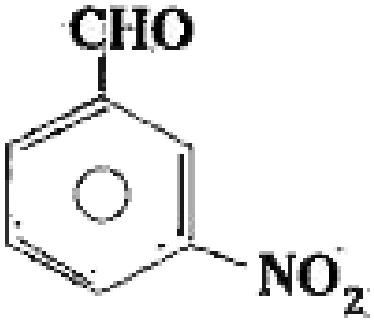
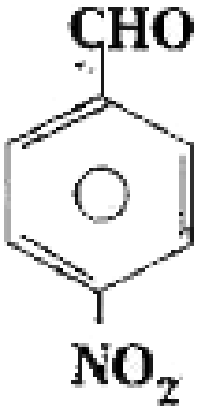
12. Write the IUPAC names of :





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Section B

1. Deduce (a) Boyle's law and (b) Charles law from Kinetic gas equation.

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2. Balance the following redox reactions by ion-electron method :

$MnO_4^-(aq) + SO_2(g) \rightarrow Mn^{2+}(aq) + HSO_4(aq)$

(in acidic solution)

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3. Which buffer solution has maximum pH?

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4. Discuss the principle and the method of softening of hard water by synthetic, ionexchange resins.

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5. What is Plaster of Paris? Write a short note on it.

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6. Explain borax bead test with a suitable example

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7. Explain the following :

b) Thin layer chromatography

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8. In paper chromatography

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9. What is dehydrologenation? Write the equation for the formation of alkene from alkyl halide.

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Section C

1. Write a short notes on Fluorine

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2. Given the outer electronic configuration of s,p,d and f-block elements.

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3. What do you understand by Hybridisation ? Explain different types of hybridisation involving s and p orbitals.

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