CHEMISTRY

BOOKS - UNITED BOOK HOUSE

MODEL QUESTION PAPER 15

Exercise

1. Hiesenberg uncertainty pinciplestate two parameters for an elctron which can't be measured simultaneously,

A. Momentum & kinetic energy

B. Position & potential energy

C. Kinetic & potential energy

D. Position on and momentum.

Answer:

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2. Which one is non polar

A. CO_2

B. BCl_3

C. NO_2

D. CCl_4

Answer:

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3. Which one does not form Hydorgen bonding

A. Liq. NH_3

B. Phenol

C. H_2O

D. Liq. HCl

Answer:

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4. Why the smell of body spray spreads

A. Diffusion

B. Surface tension

C. Viscosity

D. Density

Answer:

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5. $\Delta G = O$ at equilibrium is satisfied under condition

A. Constant temp. & pressure

B. Constant temp &volume

C. Constant pressure & volume

D. Constant pressure & Density.

Answer:

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6. For endothermic reaction ΔH value is-

A. Positive

B. Negative

C. Zero

D. Constant.

Answer:

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7. The efficiency of a catalyst depends on—

A. Mass of particle

B. Size of particle

C. Magnetice nature of particle

D. Structure of particle.

Answer:

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8. Which one hasleast melting point

A. $CaCl_2$

B. $CaBr_2$

C. Cal_2

D. CaF_2.

Answer:

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9. Which one undergoes thermal decomposition

A. $MgCO_3$

B. K_2CO_3

C. NA_2CO_3

D. $CaCO_3$

Answer:

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10. 1-Chlorobutane on treatment with alcholic KOH gives

A. But-1-ene

B. Butan -1-al

C. But - 2 - ene

D. Butan -2- ol.

Answer:

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11. Which one has highest boiling point—

A. n-hexane

B. n - pentane

C. 2, 2- dimethyl propane

D. 2- Methyl Butane.

Answer:

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12. Which one is not isomeric with diethylether

A. n - propylmethyl ether

B. 1 - butanol

C. 2 - Methyl -2-propanol

D. Butanone

Answer:

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13. No. of isomer possible for C_6H_{14}—

A. 4

B. 6

C. 5

D. 7

Answer:

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14. Which one causes Minamata

A. Cu

B. Fe

C. Hg

D. Pb.

Answer:

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15. Write one defect of law of constant proportion:

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16. Give on example of inner transition metal.

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17. What is an open system?.

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18. Write one' intensive property.

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19. State one limitation of Wartz reaction.

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20. What is 90's benzol?

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21. Establish the relation between normal & vapour density of gas.

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22. State two postulates of Bohr atom model

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23. State with example aufbau principle.

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24. Arrange the lewis acids as per increasing order BCl_3, BF_3, BBr_3, BI_3

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25. Write the structural formula of I, 3 - dimethyl Cyelopentane.

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26. Write IUPAC name of

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27. What is Chemilumisence? State one of its harmful effect.

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28. Write two defects of Bohr Model. What is Quantum number?.

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29. State the drawbacks of Rutherford Model. What is the nature of atomic spectra as per Bohr model.

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30. Why all the 'd' - block elements are not transitional?— explain with proper example .

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31. Write two Characteristics of S block emement. Write general electronic configuration of 'd' block element.

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32. What is dipole moment? if the dipole moment of CS_2be Zero then predict its structure.

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33. Why the boiling point of ethyl alcohol is higher than diethyl ether?

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34. Classify the following intensive and extensive properties:-Internal Energy

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35. Classify the following intensive and extensive properties:-heat Capacity

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36. Classify the following intensive and extensive properties:-enthalpy

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37. Classify the following intensive and extensive properties:-Entorpy.

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38. WHat is thermodynamic equilibrium?

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39. Calculate the work done when 56 gm N_2 gas (ideal) at $28^\circ C$ expands from 10 atm pressure to 2 atm pressure.

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40. What is the oxidation number of * market element in Ca(O*Cl) Cl?

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41. Balance the reaction by ion-electron method

$KMnO_4 + FeSO_4 + H_2SO_4 \rightarrow K_2SO_4 + MnSO_4 + Fe_2(So_4)_3$

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42. Calculate the hardness of Water When 1gm $FeCl_3$ dissolved in 1 liter water.

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43. What is the magneitude of Compresibility factor for real gas at very low pressure. Plot long PVS log V for an ideal gas. Write the unit of vanderwall's constant 'a'

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44. What is Boyl's temperature? Define Critical temperature and pressure .

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45. What happens when lithium hybride is neated with Aluminium Chloride?

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46. Write unit of hardness?

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47. Write the IUPAC name for the following Compounds :- `CH_3-CH(NO_2)-CH_2-CH (CH_3)- COOH

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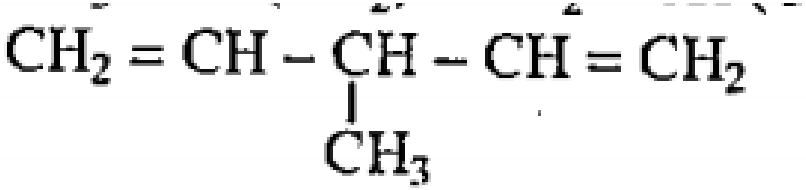
48. Write the IUPAC name for the following Compounds :-



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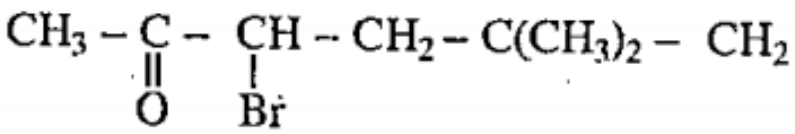
49. Write the IUPAC name for the following Compounds :-



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50. Calculate Kc for the reaction at

$500 \circ^{C} N_2(g) + 3H_2 g = 2NH_3(g) kp. = 1.44 \times 10^{-5} establishtherelationb$

4NH_3 + 5O_2 = 4NO + 6H_2O`

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51. State La-Chatelier priciple for equilibrium of a chemical reaction.

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52. Explain why $\mathbb{C}l_4$ is not hydrolysed while $SiCl_4$ is hydroglysed

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53. State with equation effect of temperature on Boric acid. Why CO can't be dried using concentrated H_2S0_4. Wht is diborane?

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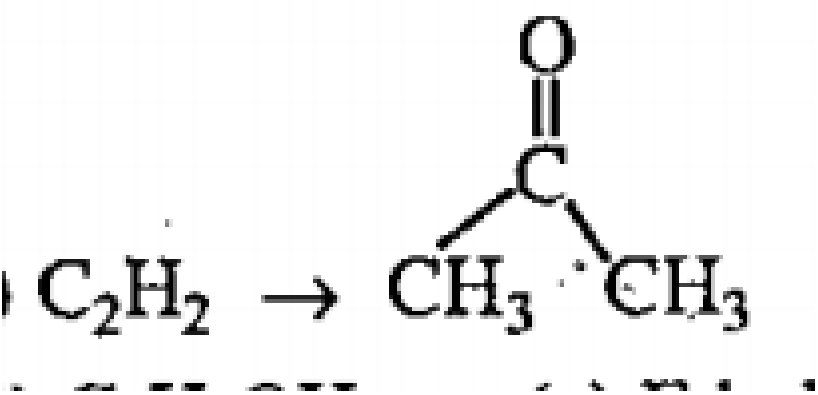
54. How would you carryout the following Conversion :-



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55. How would you Carryout the following Conversion :-$C_2H_5OH \rightarrow$ Ethylene`.

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56. How would you Carryout the following Conversion :-$C_2H_5OH \rightarrow$ $Diethylether$.

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